

## PPSC Chemistry Part VI Applied/Industrial Chemistry Online Test

Sr	Questions	Answers Choice
1	The Langmuir adsorption isotherm shows that the amount of adsorbed gas per gram of the solid is equal to.	A. $\frac{ap}{1+bp}$ B. $\frac{ap+1}{1-bp}$ C. $\frac{1+ap}{1-bp}$ D. $a(1+bp)$
2	Which of the following statements is not a part of Bohr's theory of the hydrogen atom.	A. An electron in an atom revolves around the nucleus only in circular paths. B. An electron does not absorb energy in the stationary orbit C. An electron does not emit energy in the stationary orbit D. Energy is emitted or absorbed in a discrete amount from the stationary orbit
3	Which of the following quantity is correct for micro analysis.	A. 1-10 mg or <math>\leq 50\text{ ml}</math> B. 10-20 mg or <math>\leq 50\text{ mL}</math> C. 50-100 mg or <math>\leq 100\text{ mL}</math> D. None of above
4	The common temperature detecting device in DTA are.	A. Thermocouples B. Thermopiles C. Thermistors D. All
5	Which one of following is paramagnetic and has the bond order equal to 0.57	A. $\text{N}_2$ B. $\text{H}_2^+$ C. $\text{O}_2$ D. $\text{F}_2$
6	Which of the following $\alpha$ -amino acid is not capable of exhibiting optical isomerism.	A. Glycine B. Leucine C. Arginine D. Alanine
7	Buffer solution are used to.	A. Increase the pH B. Resist the pH C. Decrease the pH D. None of above
8	Rutherford proposed the nuclear model of the atom to account for the result of experiments in which the alpha particles are scattered from metal foils. Which of the following statements is not related to Rutherford observation.	A. An atom consists of central core or nucleus around which the protons exist. B. The nucleus has most of the mass of the atom C. the nucleus consists of protons and neutrons. D. Each distinct atom has a specific number of protons.
9	The reagent which can react with 1-chlorobutane to give substitution product is	A. $\text{AlCl}_3$ B. $\text{KOH}-\text{CH}_3\text{OH}$ C. $\text{NaCN}$ D. $\text{Mg/ether}$
10	VBT is unable to explain the nature of some of the complexes of.	A. Cobalt B. Copper C. Nickel D. Manganese
11	Iron is said to be abundant in nature. About how many percent of the earth's crust is iron.	A. 10% B. 5% C. 20% D. 8%
12	The half life period of any first order reaction.	A. Is half the specific rate constant B. Is independent of the initial concentration C. Is always the same whatever the reaction D. Is directly proportional to the initial concentration of the reactant

13	The titration involving oxidation reduction reactions is called.	A. Complex titration B. Simplex titration C. Redox titration D. Acid base titration
14	Glass obtained by placing a layer of butyral plastic with a suitable adhesive between two layers of glass and cementing them by heat and pressure is called.	A. Glass wool B. Safety glass C. Optical glass D. Jena glass
15	Albumin is classified as	A. Simple protein B. Conjugated protein C. Lipoprotein D. Derived protein
16	Which of the following is not chemical characteristics of water.	A. pH B. COD C. BOD D. Colour
17	Ziegler -Natta catalysta is	A. $(C_2H_3)_3 Al$ B. $TiCl_4$ C. $(C_2H_5)_3 Al/TiCl_4$ D. $(C_2H_3)_3 B/TiCl_2$
18	The possible sub levels in the $n = 4$ energy level are.	A. s,p,d B. s,p,d,f C. s D. s,p
19	The osmotic pressure of a solution with definite composition.	A. Varies directly as the volume and temperature. B. Various inversely as the temperature. C. Varies inversely as the volume and directly as the temperature. D. None of the above
20	The concept of telluric helix was developed by	A. Lothar meyer B. A.E. de Chancourtois C. New lands D. Doberieiner