PPSC Chemistry Part VI Applied/Industrial Chemistry Online Test

Sr	Questions	Answers Choice
1	The rate constant of a reaction depends on	A. Concentration of reactants B. Concentration of products C. Temperature D. Time
2	In which pair of species, the Lewis formulae contain same number of ion pairs and bond pairs but they are not isoelectronic.	A. O2, N2 B. SO2, O3 C. PCl3 , BF3 D. SOCl2, COCl3
3	Which of the following is not a property of Ni.	A. it is a soft silvery white metal B. It is malleable and ductile C. It is highly magnetic D. It has high electrical and thermal conductivities
4	DDT is	A. Biodegradable pollutant B. Nodegradable contaminant C. Air pollutant D. An antibiotic
5	Which of the following disposal method is used for agriculture wastes.	A. Dump B. Landfill C. Incineration D. Open burning E. All above
6	Fluorine finds considerble use of DDFT which is used as.	A. herbicide B. Fungacide C. Insecticide D. Nomatocides
7	In plant noise control, which of the following method is used for reducing noise	A. Plant planning B. Control at the source C. Control of radiated noise D. All above
8	A catalyst	A. Actually participates in the reaction B. Changes the equilibrium concentration of the products C. Does not affect a reaction energy path D. Always decreases the rate for a reaction
9	Which of the following is most basic.	A. Aniline B. Benzylamine C. Diphenylamine D. N-methylaniline
10	The unit cell having dimensions, a = b=c, alpha = beta= gama not equal 90 ^{o} is known.	A. Cubic B. Trigonal C. Tetragonal D. Monoclinic
11	Plane polarized light is affected by	A. Identical molecules B. All polymers C. Chiral molecules D. All biomolecules
12	A molecule is said to be chiral	A. If it contains plane of symmetry B. If it contains centre of symmetry C. If it can be superimposed on its mirror image D. None of the above
13	The alkaline hydrolysis of fat is know as	A. Condensation B. Esterification C. Saponification D. Emulsification
14	Select the correct IUPAC nae for [Co(NH3)6]2+	A. Hexammoniacobaltate (II) ion B. Hexaamminecobaltate (II) ion C. Hexammoniacobalt (II) ion

		D. Hexaamminecobalt (II) ion
15	According to systematic nomenclature which hydrogen compound is sulphane.	A. HF B. Si H4 C. SF4 D. H2S
16	A trend which is common to elements of both the group IA and group VII A ongoing from top to bottom.	A. Boiling point increases B. Electron affinity increases C. Oxidizing power increases D. Ionization energy decrease
17	Which of the following will have the largest pH?	A. O.1 N HCl B. 0.1 N CH3COOH C. 0.1 N NaOH D. 0.01 N NaOH
18	The flow of solvent into a solution when two are separated by a semi -permeable membrane is called.	A. Mixing B. Effusion C. Diffusion D. Osmosis
19	What refers to the removal of zinc from brasses?	A. Dezincification<o:p></o:p> B. Graphitization<o:p></o:p> C. Stabilization<o:p></o:p> D. Denitration<o:p></o:p>
20	Most Hazardous metal pollutant of automobile exhaust is.	A. Tin B. Mercury C. Cadmium D. Lead
21	Ozone hole refers to.	A. Black hole B. Decrease to thickness of ozone layer in stratosphere C. Decrease of thickness of ozone in troposphere D. Increase concentration of ozone in the atmosphere
22	Lactic acid is a molecule which shows	A. Epimersim B. Tautomerism C. Opical isomerism D. Metamerism
23	Rectified spirit obtained by formentation contains 5% of water .So in order to remove it, rectified spirit is mixed with suitable quantity of benzene and heated Benzene helps because.	A. It is dehydreating agent and so removes water B. It forms the loweer layer which retains all the water so that alcohol can be distilled off C. It form an azeotropic mixture having high boiling point and thus allows the alcohol to distall over D. It forms low boiling azeotropic mixture which distille over leaving behind pure alcohol which can than be distilled.
24	Which of the following does NOT react with sodium hydroxide solution.	A. Fat B. Vinegar C. Ethanol D. Water
25	The principal ores of copper are	A. Copper sulphides B. Copper oxides C. Both sulphides and oxides D. Copper carbonate
26	The shape of SO_4^{2-} ion is.	A. Tetrahedral B. Trigonal planar C. Square planar D. Octahedral
27	When metal orbital are rotated in octahedral field the following representation obtained.	A. t2 g + lg B. a1 g C. t 1 u D. All above
		A. Smoke

28	Which of the following cause water pollution.	A. Smoke B. Automobile exhausts C. Aeroplanes D. Silt and posticides
29	What refers to the deterioration of material by oscillatory relative motion of small amplitude between two solid surfaces in a corrosive environment?	A. Stray current corrosion<o:p></o:p> B. Microbiologicla corrosion<o:p></o:p> C. Fretting corrosion<o:p></o:p> D. None of these<o:p></o:p>
30	Which of the following is an important aspect of industrial ecology.	A. Minimising air emissions B. Minimising liquid waste C. Recycling after use D. All above
31	Which of the following statement is not true with respect to hydrocarbons.	A. They are gaseous and liquids B. They can be saturated or unsaturated C. They in air by themselves alone cause harmful effects D. They form photochemical oxidants
32	The metallic character of group 14 elements	A. Decreases from top to bottom B. Increases from top to bottom C. Does not change gradully D. Has no significance
33	Hybridization involves.	A. Orbitals of same atom with slightly different energies. B. Orbitals of different atoms, but with equal energies. C. Orbitals of the same atom but with widely different energies. D. Orbitals of different atoms with different energies.
34	With which one of the following configurations, the lowest value of first IE is associated.	A. ls2, 2s2, 2p6, 3s1 B. 1s2, 2s2, 2p5 C. 1s2, 2s2, 2p6 D. 1s2, 2s2, 2p6, 3s2, 3p2
35	The most widely used method of extracting metal ions is the formation of a chelate molecule with an organic chelating agent The chelating agents are.	A. Strong acids B. Strong bases C. Weak bases D. Weak acids
36	The magnetic quantum number (m) specifies the individual orbital in a Sub shell for a given l, m can be.	A. l,l- l.....-1 B. l.....2, l -3.....-2l C. l-l-2,.....-l D. l-2,l-4, ...-4l
37	The unit of sodium chloride structure is.	A. Linear B. Cubic C. Tetrahedral D. Square planner
38	The internal resistance to flow possessed by a liquid is called its.	A. Fluidity B. Viscosity C. Surface tension D. Turbidity
39	The hydrogen bond is strongest in.	A. O - H .....S B. S - H.....O C. F- H.....F D. F - H.....O
40	Which of the following is not an ore of iron.	A. Haematite B. Magnetite C. Siderite D. Monazite
41	In the kinetic study of a reaction A _____ products. A straight line was observed when a graph between time and 1/C2 was plotted. the reaction is.	A. Second order B. First order C. Third order D. Zero order
		A. C1 B. C2

42	Linear molecules have _____ axis of rotation	B. C2 C. C D. C3
43	Elements of group 14	A. Exhibit oxidation state of -4 B. Exhibit oxidation state of +4 C. From M3+ and M4+ ions D. Form M4- and M4+ iona
44	Purpose of sizing is.	A. To increase the strength B. To improve formation C. To increase resistance toward water D. To remove wastes
45	In the metallurgy of iron, when limestone is added to the blast furnace, the calcium ion ends up in.	A. Slag B. Gangue C. Metallic calcium D. Calcium carbonate
46	The rate constant of a reaction has same units as the rate of the reaction The reaction is of.	A. Second order B. First order C. Three order D. Zero order
47	Which process of adsorption of hydrogen on palladium is known as.	A. Syneresis B. Occlusion C. Diffusion D. Erosion
48	Hydrolytic reaction of fat with caustic soda is known as _____	A. Esterification B. Saponification C. Acetylation D. Carboxylation
49	The configuration of valence shell of certain atom X is 3s2 , 3p5, which valences can it exhibit.	A. 1,3 only B. 1,5 only C. 1,3,5,7 D. 1,3,4
50	The melting of nearly all glass is done in a continuous tank furnace. which operates steadily over periods of up to.	A. a day B. a month C. a year D. None of these
51	The value of Kw increase with temperature because the ionization of water.	A. Positive B. Negative C. Endothermic D. Exothermic
52	The binding site on ribosome t-RNA and m-RNA is provided by	A. Polysome B. Ribosomal RNA C. Codone D. DNA
53	The particle motion in solids is	A. Only vibratory B. Only translator C. Vibratory and rotatory D. Only translatory
54	The following statements are true except one which one.	A. Carburizing does not harden a steel<o:p></o:p> B. Flame and induction hardening require the use of hard enable steels.<o:p></o:p> C. Quench –hardened steel does not require tempering to present brittleness<o:p></o:p> D. None of these
55	In order to understand the nature of H , bond the theory has been suggested.	A. Electrostatic approach B. Molecular orbital approach C. Valance bond approach D. All the above approaches
56	The light absorbed in UV and visible region causes.	A. Vibrational energy changes B. Rotational energy changes C. Electronic excitation D. All of these
57	Gold dissolves in aqua regia forming	A. AuCl B. Au(NO3)3 C. AuCl3 D. HAuCl4

58	A closed system is one which can exchange with surrounding.	A. Matter but not energy B. Energy but not matter C. Both matter and energy D. Neither matter nor energy
59	How pig iron is usually obtained from	A. iron pyrite B. Limonite C. Hematite D. Siderite
60	Which property is not exhibited by carbon in its compounds.	A. Forming bounds to other carbon atoms B. Formation multiple forms C. Exhibiting allotropic forms D. Forming compounds with coordination number beyond four
61	Which of the following metal acts as pollutant.	A. Hg B. Pb C. Zn D. Ni E. All above
62	Non localised bonds are referred as	A. Metallic bond B. Long range bonds C. Ionic bond D. Covalent bonds
63	Which of the following technique is the application of voltammetry at a fixed potential to detect changes int he currents as a function of the concentrating of the analyte	A. Amperometry B. Coulometry C. Polarography D. Potentiomertry
64	The element with the highest first ionization potential is.	A. Boron B. Carbon C. Nitrogen D. Oxygen
65	Which of the following statement is correct.	A. The wavelength of phosphorescence is less than the wavelength absorbed B. Teh transition from T₁ to S₀ without the emission of light is called phosphorescence C. The combination CO2 and water in plants, in the presence of chlorophyll, is an example of bioluminescence. D. Population inversion is a necessary condition for laser action
66	When two H atoms approach each other then forces operates.	A. Attractive forces B. Repulaive forces C. Attractive and repulsive D. None of above
67	Which of the following is the strongest oxidant.	A. F2 B. Cl2 C. br2 D. I2
68	According to Henry's Law. the mole fraction of a gas (x) dissolved in a solvent is related to the pressure of the gas.	A. x = k/p B. x = p/k C. x = k D. p = k/x
69	The inert gases Ar, Kr and Ke form compounds with water at low temperature and high pressure. These compounds are called.	A. Halides B. Hydrates C. Clathrates D. All of above
70	The 'shape' of molecule ." Xe F6 is.	A. Pentagonal bipyramidal B. Regular octahedral C. Distorted octahedral D. Square planar
71	Which of the following is not a pyrimidine base.	A. Uracil B. Thymine C. Cytosine D. Guanine
72	An element with atomic number 20 is placed in which period of the periodic table.	A. 1 B. 2 C. 3 D. 4
73	Which of the following techniques is capable of seperating minute quantities of the substances in a relatively short times with high resolutions.	A. Gel electrophoresis B. Capillary electrophoresis C. GC D. HPLC

74	Which is incorrect statement for Xe F2.	A. It has linear structure. B. It is hydrolyzed rapidly in aqueous solution of a base C. It oxidizes Cl and I to Cl2 and I 2 respectively D. It cannot act as F donor
75	Which type of polymer the Nylon -06 is	A. Polyamide B. Polyester C. Addition D. Homopolymer
76	Ammonia is utilized for	A. Manufacture of urea B. Oxidation to nitric acid C. Manufacture of ammonium sulphate D. All above
77	The sugar present in RNA is	A. D- ribose B. D-Arabinose C. D-Glucose D. Deoxyribose
78	Halogens are coloured because.	A. They are strong oxidant B. Their molecules are held together by weak van der Waals forces C. Their atoms absorb radiations form visible range causing the excitation of valence electrons to higher energy of levels D. Their molecules absorb light radiation forming the oxcited state.
79	in monel metal copper is alloyed wiht which metal.	A. Fe B. Mn C. Ni D. All
80	Which of the following physical properties is employed int he analytical methods.	A. Electric current B. Transition temperature C. Surface tension D. All above
81	Which of the following is not a colligative property.	A. Elevation of B.P. B. Depresaion in F.P C. Viscosity D. Osmotic pressure
82	Chemical compounds which are added to reduce to reactivity of glass are called.	A. Formers B. Modifiers C. Stabilizers D. None of these
83	What nickel alloy has high electrical and corrosion resistance and high strength at red heat temperature and contain 15 to 20% chromium.	A. Alnico B. Nichrome C. Invar D. None of above
84	The most stable oxidation state of chromium is.	A. +6 B. +3 C. +4 D. +2
85	Which of the following molecule does not contain the covalent bond between similar atoms.	A. N2H4 B. F2O2 C. H2F2 D. H2O2
86	Which of the following acid radical gives chromyl chorate test.	A. F- B. I- C. Cl- D. Br-
87	Enzymes are	A. Complex non living compunds B. Laving organisms C. Complex protein molecules D. Bacterial colonies
88	The size of quantum dot is _____m	A. 5 B. 5 x 10⁻⁹ C. 5 x10⁻¹⁰ D. 5 x10⁻¹¹
89	Hydrometallurgy of copper involves extraction of copper from poor ores by which process.	A. Dry process B. Wet process C. Both dry and wet process D. None of these
		A. Zn salt

90	Which of the following salt is colourless.	B. Co salt C. Ni salt D. Mn salt
91	Which is the correct order of wave number of the following redistions.	A. X-rays > uv > Infrared > visible> radio waves B. X-rays > uv > visible > Infrared > radio waves C. X-rays > radio waves > uv > visible > Infrared D. X-rays > Infrared > uv> visible > radio waves
92	Which of the following elements does not impart any characteristic colour to the flame.	A. Ca B. Mg C. Ba D. Sr
93	The number of electrons involved in bonding in Lewis structure of oxalate ion is	A. 20 B. 14 C. 22 D. 18
94	In C4-axis of rotation, an object in rotated through an angle of.	A. 120^o B. 180^o C. 100 ^o D. 90^o
95	H-Bond has more energy than the van der Waals forces i.e.	A. 1.0 kcal/mole B. 2.0 k cal/mole C. 10.0 kcal/mole D. 20. 0 kcal mole
96	The branch of chemistry which deals with the rate of reaction as well as mechanism is known as	A. Wave mechanism B. Classical thermodynamcis C. Chemical kinetics D. Phtochemistry
97	At the some temperature 0.1 M solution of urea is isotonic with.	A. 0.1 M glucose solution B. 0.1 M NaCl solution C. 0.1 M urea solution D. 0.1 M BaCl2 solution
98	In the long form of periodic table, elements are arranged according to.	A. Increasing atomic number B. Decreasing atomic number C. Increasing atomic mass D. Decreasing atomic mass
99	Brass is an alloy of	A. Copper and tin B. Copper and zinc C. Aluminium and nickel D. Leed and tin
100	Which of the following level is an indicator of hearing loss.	A. > 25 dB B. < 25 dB C. <20 dB D. None of these
101	The atomic and ionic radii value on moving from left to right in the sereis.	A. Increase B. Decrease C. Does not change D. None of above
102	In the process of electrosmosis	A. Colloidial particles move towards the electrodes B. Both colloidal particles and dis persons medium move C. Only dispersion medium moves to carry the current D. Positively charged colloidal particles move, but negatively charged particles remain stationary
103	Which of the following reactions does not take place with light radiation.	A. Oxidation B. Reduction C. Polymerization D. Double displacement
104	Which of the following is not a component of hollow cathode lamp.	A. Anode B. Cathode C. Filter gas D. Atomic vapour
105	Carbides because of their hardness are	A. Ionic carbides B. Interstitial carbides C. covalent carbides D. Any of above
		A. -5

106	The pKa of an acid having ionization constant 1 x 10^{-5} is	B. 5 C. 9 D. -9
107	Among the elements A,B,C and D having atomic numbers 9,10,11, and 12 respectively, the correct order of ionization energies is.	A. A > B > C &gtD B. B > A > D > C C. B > A > C > D D. D > C > B > A
108	_____ remove the remaining color producing a water white sugar syrup	A. Carbon filters<o:p></o:p> B. Centrifuge<o:p></o:p> C. Annealing<o:p></o:p> D. Refining<o:p></o:p>
109	in the system of designating wrought aluminum alloys the letter F that follows the number indicates what condition of the alloy.	A. As fabricated B. Calcined C. Annealed D. Strain hardened
110	Which of the following carbonates decomposes at the highest temperature.	A. Mg CO3 B. CaCO3 C. Sr CO3 D. Ba CO3
111	How many planes of symmetry are present in benzene.	A. 1 plane B. 3 planes C. 5 planes D. 7 planes
112	The polarity of bonds can lead to polarity of molecules and affect	A. Melting point B. Boiling point C. Solubility D. All of above
113	The particles of about 1 nm need _____ activation energy to enter either aggregation processes or reactions to give to new chemicals.	A. Higher B. Lesser C. N o D. All above
114	The temperature at which the vapour pressure becomes equal to external pressure is called.	A. Saturation point B. Critical temperature C. Consolute temperature D. Boiling point
115	Which of the following technique is used to separate substance based on their charge to mass ratio.	A. HPLC B. HPTLC C. GC D. Electrophoresis
116	The oxidation number of Mn in KMnO4	A. +5 B. +7 C. +4 D. +3
117	The increase in boiling points of noble gases from He to Xe is due to the	A. Decreases in ionization energy B. Increases in polarizability C. Increase in electron affinity D. Increase in atomic volume
118	The maximum degree of freedom for a pure substance under equilibrium constitutions is	A. 1 B. 2 C. 3 D. zero
119	In Nano synthesis new unsual chemical reactions are due to.	A. Non equilibrium system B. Equilibrium system C. Isothermal system D. Adiabatic process
120	The carbonate of which of the following will have highest lattice energy.	A. Barium B. Magnesium C. Calcium D. Strontium
121	What element is added to copper to increase its strength and fatigue propertioes	A. Silicon B. Aluminium

121	What element is added to copper to increase its strength and fatigue properties.	C. Beryllium D. Copper
122	In radial direction the thermal conductivity of a nano tube is _____ watt/(m.k)	A. 3500 B. 385 C. 0 D. 350
123	Which of the following is the active ingredient in ordinary household bleach.	A. HCl B. Cl2 C. NaCl D. NaClO
124	Bitumen is used in	A. Electric generators B. Road surfacing C. Coal tar D. All of above
125	A compound with an congruent melting point decomposes on heating into.	A. A liquid of the same composition as the solid B. A new solid phase and a solution with a compositional from that of the solid phase C. A new solid phase and a solution with the same composition as that of the solid phase D. A solution of fixed composition
126	The sugar present in DNA is	A. D- Ribose B. D-Glucose C. 2- Doxy D-Ribose D. 3-Deoxy D-ribose
127	Which of the following is NOT a hardware requirement for die casting.	A. Water cooled metal cavities<o:p></o:p> B. Machined metal holding blocks<o:p></o:p> C. Ejection mechanism<o:p></o:p> D. Metal mold<o:p></o:p>
128	Sanger's reagent is	A. Carbobenzyloxy chlride B. Dimethyl amino sulphonyl chloride C. I-Fluoro -2,4-dinitrobnzene D. 2,4- Dinitrophenyl hydrazine
129	Metallic magnesium is obtained by	A. Reduction of MgO with Coke B. Electrolysis of an aqueous solution of MgCl2 C. Electrolysis of molten MgCl2 D. Displacement of magnesium by iron form MgCl2 solution.
130	The expected specific wastes of textile industry is	A. Cloth residue B. Fibre residue C. Dyes D. All above
131	Which property is used in volumetric methods of analysis.	A. Density B. Viscosity C. Volume D. Molar volume
132	Which of the following statement is not correct n respect of Arrhenius concept.	A. The concept is applicable only for aqueous systems. B. Neutralization takes place in aqueous mediam only C. TH+ ion cannot remain as such in water D. This concept is applicable for non aqueous system only.
133	Which of the following anionic species is not separated by gravimetric analysis.	A. Cl- B. SO4 C. CH3COO- D. PO4
134	Chemical and physical properties of metal nano particles of atoms were observed to	A. Number of atoms in a particle B. Shape of particle

134	Chemical and physical properties of metal nano particles or atoms were observed to change periodically depending upon	B. Shape of particle C. type of organization D. All of the above
135	Amorphous boron on burning in air form	A. B(OH)3 B. Only B2O3 C. Only BN D. Mixture of B2O3 and BN
136	An sp3 hybrid orbital contains	A. 1/4 a character B. 1/2 a character C. 2/3 a character D. 3/4 a character
137	Which of the following is strong adheaive.	A. Epoxy resin B. Melamine -formadehyde resin C. Alkyd resins D. Bakelite
138	Which statement is false.	A. If a reaction is thermodynamically spontaneous it may occur rapidly B. If a reaction is thermodynamically spontaneous it may occur slowly. C. Activation energy is a kinetic quantity rather than a thermodynamic quantity. D. If a reaction is thermodynamics spontaneous, it must have a low activation energy.
139	Which of the following analytical method is based on the rotation of light ratiation	A. Refractomerty B. Polarimetry C. Interformetry D. Polarography
140	Colloids can be purified by	A. Peptization B. Coagulation C. The Breeding are method D. Dialysis
141	Which of the following symmetry element leaves the molecule or an object unchanged.	A. Proper rotation B. Improper rotation C. Inversion axis D. Identity
142	Which of the following case of acid or base strength is not explained by inductive effect.	A. Formic acid> acetic acid B. Dimethyl amine > trimethyl amine C. Dimethyl amine > methyl amine D. Chloroacetic acid > acetic acid
143	Reaction in which molecules absorbing light do not themselves react but induce other moleculaes to react are called.	A. Chain reactions B. Photosenaitized reactions C. Reversible reactions D. Free radical reactions
144	Which of the following statements is not related with flame photometric analysis.	A. Vaporization of the solvent leaving back the residue B. Conversion of solid slat to the gaseous state C. Dissociation of gaseous molecules into free atoms D. Measurement of the intensity of absorbed tradition
145	Rutherford proposed the nuclear model of the atom to account for the result of experiments in which the alpha particles are scattered from metal foils.Which of the following statements in not related to Rutherford observation.	A. An atom consists of central core or nucleus around which the protons exist. B. The nucleus has most of the mass of the atom C. the nucleus consists of protons and neutrons. D. Each distinct atom has a specific number of protons.
146	Which of the following process involves the use fo organic compound as an electron acceptor.	A. Aerobic respiration B. Anaerobic respiration C. Fermentation D. Glycolsis
147	The common ligands can be arranged in order of their increasing splitting power to cause d-orbitals splitting. This series is called as.	A. Electro-chemical B. Spectro -chemical C. Physico-chemical D. Spectro -electrical
148	Which type of organic compounds does fat belong to.	A. Alkene B. Ester C. Alkanol D. Alkanoic acid
		A. C,N, Si, P B. N, Si, C,P

149	The electronegativity of the following elements increases in the order.	C. Si, P, C, N D. P, Si, N, C
150	In which pair of species, the Lewis formula contain same number of Lone pairs and bond pairs but they are not iso electronci.	A. O2 B2 B. SO2, O3 C. PCl3, BF3 D. SOCl2, COCl2
151	Formula of orthophosphoric acid.	A. H2PO4 B. H3PO3 C. H3PO2 D. H4P2O5
152	If a chemical reaction in equilibrium is subjected to a change the reaction tends to more in such a direction that the effect of the change would be neutralized This is a statement of.	A. Law of mass action B. Le Chatlier's principle C. Henery's law D. Correspondence principle
153	Which of the following elements has the highest value of IE.	A. Na B. K C. Mg D. Ca
154	The temperature at which two conjugate solutions change into one homogeneous solution is called.	A. Azeotrope B. Conjugate temperature C. Consolute temperature D. Transition temperature
155	What does 'F' stand for in AFM.	A. Fine B. Front C. Force D. Flux
156	Which of the following statements are correct for Linear polymers.	A. Linear polymers may be condensation as well as addition polymers B. Structure is well packed in nature C. Linear polymers have higher density higher melting point and higher tensile strength D. All are correct
157	The reduction of an alkyne to alkene using Lindlar's catalyst results into	A. Syn addition of hydrogen atoms B. Anti addition of hydrogen atoms C. A mixture obtained by ayn and anti addition of hydrogen which are equilibrium with each other D. A mixture obtained by syn and anti addition of hydrogen which are not in equilibrium with each other.
158	Final paper wound in the form of a real having final moisture of about.	A. 6-8% B. 9 - 12 % C. 13-15 % D. 4 - 10%
159	An example of nitro dyes is.	A. Martius yellow B. Auramine O C. Malachite green D. Methyl red
160	The maximum covalence of an element equal to.	A. The number of unpaired d electrons B. The number of paired p electrons C. The number of unpaired a and P electors D. The actual number of a and P electrons in the outermost shell
161	_____ surfactants perform well over a wide range of water hardness and pH.	A. Anionic B. Cationic C. Nonionic D. Neutral
162	The vitamin which is related to monossaccharides is.	A. Vitamin A B. Vitamin C C. Vitamin D D. Vitamin E
163	All the strong acids have very close pKas value and they appear to have nearly equal strengths in aqueous solutions. The phenomenon is called as.	A. Levelling effect B. Differnetiating effect C. Levelling solvent D. Differnetiating solvent
164	Which of the following is not correct.	A. Rusting of iron can be stopped by increasing the concentration of CO2 in water B. Rusting of iron is electrochemical in nature. C. Rusting of iron takes place in moist air D. Rusting of iron produces hydrated

		D. Rusting of iron produces hydrated iron (III) oxide
165	The electrolytic method super passes all other methods due to.	A. Furity B. Cheapness C. Easy available D. All above
166	Sulphate ores of aluminium	A. Alumite B. Cryolite C. Fekdsper D. Kaolin
167	Valences bond theory was put forward by	A. Pauling and Slatter B. Heitler and London C. Lewis D. Pauli
168	Which of the following test is not shown by proteins.	A. Xanthoproten test B. Ninhydrin test C. Hopkin cole test D. Muliken Barker test
169	Drying agent which react with CO2 and removes water vapours is.	A. CaO B. CaCl2 C. CaCO3 D. Ca(NO3)2
170	In biological ecosystem which of the following substance is used by organisms.	A. Water B. Sunlight C. Minerals D. All above
171	The ease of hydrohalogenation of alkyl halide with alcoholic KOH is.	A. 3^o > 2^o > 1^o B. 3^o < 2_o <1^o C. 3^o > 2^o < 1^o D. 3^o < 2 ^o > 1^o
172	Phosphorus normally exhibit a covalency of.	A. +1 and +2 B. +2 and +3 C. +3 and +4 D. +4 and +5
173	Which of the following cast irons is a high carbon silicon alloy.	A. Gray iron B. White iron C. Malleable iron D. Alloy iron
174	Permanent hardness of water is due to.	A. Sulphate of Ca B. Chloride of Ca C. Sulphate of Mg D. All above
175	In graph of atomic volume versus atomic weight the elements corresponding to peaks in the curve belong to.	A. Group 1 B. Group 18 C. Group 4 D. Group 14
176	Complexing reactions are useful for which of the following method of analysis	A. Gravimetry B. Spectrophotometry C. Interfering ions masking D. All of the above
177	What is a measure of rigidity?	A. Stiffiness<o:p></o:p> B. Jardmess<o:p></o:p> C. Strength<o:p></o:p> D. Modulus of elasticity<o:p></o:p>
178	Gutta percha is	A. Cis poly imprene B. Trans -polyisoprene C. Polyethylene D. Polyisobutylene

179	An Ideal gas is one which obeys all the gas law at.	A. Low pressure B. High Pressure C. Low and High temperature D. All condition of pressure and temperate re
180	Lime water is an aqueous solution of.	A. MgSO4 B. Ca (OH)2 C. CaCO3 D. CaSO4
181	Which of the following process is a source of nuclear pollution.	A. Uranium mining B. Uranium processing C. Reactor waste D. All above
182	In nature nickel is found in the form of.	A. Sulphides B. Silicates C. Arsenides D. All
183	Which of the following statements is worng.	A. Covalent compounds are generally soluble is polar solvents B. Covalent compounds have low melting and boiling point. C. Ionic solids do not conduct electricity is solid state D. Ionic compounds conduct electricity in the fused state.
184	Gases and dust particles are removed from H2SO4 by	A. Tydal effect B. Drying tower C. Absorption tower D. Contact converter
185	Which of the following statements false about transition metals.	A. They form complexes B. They show variable valency C. All transiting metal compounds are paramagnetic D. They form coloured ions
186	In diborane (B2H6)	A. The structure is similar to that of C2H6 B. All the atoms are in one plane C. The born atoms are linked through hydrogen bridges D. There is a direct boron boron bond
187	If Principal quantum number n = 4 the quantum number l can have calue.	A. 1,2,3 and 4 B. 0,1,2 and 3 C. 1,2 and 3 only D. None of the abvoe
188	The spectral line obtained when an electron jumps from n = 6 to n= 3 belongs to.	A. Balmer series B. Layman series C. Paschen series D. Bracket series
189	Commercial incinerators produce.	A. Smoke B. CO C. NOx D. All above
190	The substance added to the soil to provide one or more nutrient elements essential for plants growth are called.	A. Growth hormones<o:p></o:p> B. Fertilizers<o:p></o:p> C. Salts<o:p></o:p> D. Minerals
191	Which treatment is done with pulp before delivering it to paper making machine.	A. Pulp is disperse din water to make slurry B. Mechanical refining or heating of the fibers C. Addition of chemical additives and recycled fibres from the waste paper plant D. All above
		A. Metallury<o:p></o:p> B. <p class="MsoNormal" style="margin

192	What field of study encompasses the procurement and production of metals.	B. Geology<o:p></o:p> C. Metagraphy<o:p></o:p> D. Nanochemistry<o:p></o:p> <o:p></o:p>
193	The formula of sulphur sequioxide	A. SO4 B. S2O7 C. S2O3 D. SO3
194	What is the ASTM tension testing designation for standard method for steel products.	A. A 370 B. E 345 C. E8 D. E 9
195	What combination of elements has high electrical resistance high corrosion resistance, and high strength at red hear temperatures, making it useful in resistance heating.	A. Aluminium bronze B. Nichrome C. Hastelloy D. None of above
196	Which of the following property of liquids concern with the interval resistance to its flow.	A. Refractive inded B. Viscosity C. Optical activity D. Dipole moment
197	The hybridization of S in SO2 is.	A. sp B. sp2 C. sp3 D. dsp2
198	Major principle underlying the sustainability of natural ecosystems is that they run on.	A. Electric energy B. Solar energy C. Wind energy D. None of the above
199	Washing soap can be prepared by saponification with alkali of _____ of the following oil.	A. Rose oil B. Paraffin oil C. Groundnut oil D. Coconut oil
200	Which of the following gas does not exist free on earth.	A. N2 B. H2 C. O2 D. CH4
201	Potentiometry is based on the measurement of which physical property.	A. Electrical conductance B. Electrical potential C. Thermal conductance D. Current
202	Which of the following is the weakest base.	A. KOH B. NaOH C. LiOH D. RbOH
203	In each period the element with lest electron affinity belongs to.	A. Group 1 B. Group 14 C. Group 17 D. Group 18
204	The percentage of nitrogen in ammonia is _____%	A. 32 B. 82 C. 25 D. 55
205	The correct order of electron affinities of Sl, P, and Cl is.	A. P > Si > Cl B. Cl > P > Si C. Cl > Si, > P D. Si > P . Cl
206	Which of the following techniques are used for minimizing water pollution.	A. Stabilization of ecosystem B. Recharge of the waste C. Waste treatment D. All above
207	Considering the elements B, Al, Mg and K , the correct order of their metallic character is.	A. B > Al > Mg > K B. Al > Mg > B > K C. Mg > Al > K > B D. K > Mg > Al > B

		D. K > Mg > Al > B
208	Which of the following is not a characteristics of terpenoids.	A. They are pleasant smelling liquids B. They are steam volatile C. They are nitrogenous bases D. they are insoluble in water
209	The terpenoid present in oil of lemon grass is	A. Citral B. Geranial C. Nerol D. a- terpineol
210	Select an acidic amino acid	A. Lysine B. Cystine C. Aspartic acid D. Aminoacetic acid
211	Monomer of neoprene rubber to	A. 1-chloro 1,3- butadiene B. 2- chlro, 1,3-butadiene C. 2-Bromo -1,3- butadiene D. 2-Methyl 1,3-butadiene
212	The bond order for BO molecule is.	A. 2.5 B. 3.0 C. 2.0 D. 3.5
213	The size of iso electronic species - F- , Ne, and Na+ is affected by	A. Nuclear charge (Z) B. Valence principal quantum number (n) C. Electron electron interaction in the outer orbital D. None of the factors because their size to the same.
214	When of the following steps is involved in structure determination of an organic compound.	A. Purification of compund. B. Qualitative and quantitative analysis of elements present C. Determination of molar mass D. All above steps
215	Which of the following has highest ionization energy.	A. Oxygen B. Argon C. Barium D. Caealum
216	Which are not considered member of d-block elements.	A. Zn B. Cd C. Hg D. All above
217	In a system , when the chemical potential of each component is the same for all phases. the equilibrium is said to be in	A. Metastable equilibrium B. Thermal equilibrium C. Composition equilibrium D. Mechanical equilibrium
218	Used in filling luminous tubes.	A. Xenon B. Krypton C. Radon D. Helium
219	An auxochrome is a group which	A. Absorbs in UV region B. Absorbs in visible region C. Absorbs in IR region D. Increase absorption wavelength of chromophore
220	Which of the following salt is not used in salt bridge to minimize liquids junction potential.	A. KCl B. NH4Cl C. KNO3 D. CaCl2
221	The number of degrees of freedom and number of components for a system of containing undissolved salt , in equilibrium with water vapor are.	A. 2,2 B. 3 , 2 C. 1 , 1 D. 1 , 2
222	Which of the following statement represent disadvantages of sanitary landfill	A. Public opposition B. Uneconomical C. Health hazard D. All above
223	The inert gasses AR, Ka, and Xe form solid compounds with certain organic molecules under pressure..	A. Halides B. Hydrates C. Clathrates D. All of above
		A. The rate of nitration of benzene is almost the same as that of hexadeutero benzene

224	Among the following statements in the nitration of aromatic compounds, the false one is.	B. The rate of nitration of toluene is greater than that of benzene C. The rate of nitrationof benzen is greater than that of hexadeutero benzene. D. Nitration in an electrophite substitution reaction.
225	Putrefaction is	A. Hydrolysis of proteins B. Reduction of proteins C. Bacterial oxidation of proteins D. All of these
226	The height to which a liquid will rise in an open capillary tube is inversely proportional to.	A. Temperature of the liquid B. Surface tension C. Density of the liquid D. Air pressure
227	Shows a regular increase on moving down the group form carbon to lead	A. Atomic volume B. Atomic radius C. Density D. All above
228	Isotopes are atoms whom nuclei have the same atomic number but different mass numbers. A specific isotope has an atomic number of 18 and a mass number of 35. How many electrons are there in the neutral atom.	A. 17 B. 18 C. 34 D. 35
229	Which of the following steps is involved in quantitative analysis.	A. Sampling B. Conversion of the desired constituent into a suitable form per analysis. C. Measurement of some physical or chemicals property, on which the determination is based. D. All above steps
230	Each fat or oil in made up of	A. A distinctive mixture of several different triglycorides B. A distinctive mixture of several aldehydes C. Mixture of above both D. None of above
231	The conductance of 1 cm3 of an electrolytes solution is called its.	A. Specific resistance B. Specific conductance C. Molar conductance D. Equivalent conductance
232	The brown colour of the pulp obtained from chemical pulping is due to the present of	A. Chlorine B. Residual lignin<o:p></o:p> C. Sodium hydochlorite<o:p></o:p> D. All above
233	Which of the following has hexagonal structure.	A. Sodium chloride B. Potassium choride C. Diamond D. Graphite
234	Which element among the following cannot exhibit variable electronvalency	A. ₂₉Cu B. ₅₀Sn C. ₂₅ Mn D. ₃₈ Sr
235	The electronic configuration of sodium (Z=11)	A. 1s2, 2s2, 2p4 B. 1s2, 2s2, 2p6, 3s2 , 2p5 C. 1s2, 2s2, 2p6, 3s1 D. 1s2, 2s2, 2p6, 3s2
236	Organic substance responsible for the smell of flowers etc. are grouped together in chemistry as.	A. Perfumes B. Terpenoids C. Flavonoids D. Alkaloids
237	Complete hydrolysis of nucleotide result in the formation of.	A. Heterocyclic bases B. A pentose C. A phosphate ion D. All of these
238	Which of the following statements do not represent Lewis idea of acids and base?	A. Compounds which have completely filled orbitals B. Compounds which have incompletely filled orbitals C. Compounds in which the central atom

		can expand its octel D. All simple metal ions like Ag+, Al3+ etc.
239	The number of formula weight of the solute dissolved per dm3 of the solution is called.	A. Mole fraction B. Normality C. Formality D. Molaiity
240	_____ is preferred for horticultural crops and for tobacco and potatoes.	A. Potassium Chloride<o:p></o:p> B. Potassium Sulphate C. Potassium Nltrate D. None of these
241	Molecule of oxygen is	A. Diamagnetic B. Paramagnetic C. Both A and B D. None of above
242	Catenation is a process of.	A. Formaton of cations B. Deposition of cations C. Formation of long chain of identical atoms D. Formation of covalent bond
243	An example of acrylic monoterpenoid is	A. Dipentene B. Myocene C. a- terpineol D. Limonene
244	How many sigma and pi bonds are there in a CO2 molecule.	A. 2 sigma B. 2 sigma and 4 pi C. 2 sigma and 2 pi D. 4 sigma and no pi
245	Which of the following statement is not related to MOT	A. Atomic orbitals lose their identities B. MOTgives as idea of denationalization C. MOT uses all the orbitals and elections D. It treated bond as purely covalent
246	Which of the following elements would have the lowest first ionization energy	A. Mg B. Rb C. Li D. Ca
247	Stereotsomers not related to each other as object and minor image are called.	A. Enantiomers B. Diastereolsomers C. Conformations D. Antipodes
248	Which of the following are neutral ligands.	A. NH3 B. H2O C. CO & NO D. All of above
249	In Glass of vitreous state solid the atoms are arranged in.	A. Regular fashion B. Random fashion C. linear fashion D. All of these
250	Which of the following is atmospheric pollutant.	A. CO2 B. CO C. O2 D. N2
251	The noble gases are found in the atmosphere to the extent of about some percent by volume.	A. 0.5% B. 1.0% C. 1.5% D. 2.0%
252	Calander stock is a process in paper making in which.	A. Thickness of the paper is reduced B. Surface of paper is made smooth C. Moisture is removed D. Both A and B
253	Group IV A consist of elements	A. 3 B. 4 C. 5 D. 6
254	The geometry of the molecule is primarily decided by	A. Bond pairs around the central atom B. No of k bond around the central atom C. No of bond pairs as well as lone pairs around the central atom D. No. of lone pairs on central atom

		D. No. of lone pairs on central atom
255	The compound insoluble in acetic acid is.	A. Calcium oxide B. Calcium carbonate C. Calcium oxalate D. Calcium hydroxide
256	Which of the following hydrocarbon cannot be obtained on reacting chloomethane with sodium metal in the presence of dry ether.	A. C4 H10 B. C2H6 C. C2H4 D. C3H8
257	Which of the following acids acts as acid waste from coal mines.	A. HCl B. HNO3 C. CH3COOH D. H3PO4
258	Suppose a sample is analyzed for a particular constituent by two different method One can tell whether the two average values are significantly different by applying which of the following test.	A. Student's test B. F test C. Chi square test D. Variance
259	Sterols are steroids having the functional gruup.	A. Ketonic B. Al;coholic C. Phenolic D. Aldehydic
260	Aviation Fuel contains.	A. Light Naphtha B. Medium Naphtha C. Kerosene D. Diesel
261	The equilibrium constants Kp and Kc are related as	A. Kp = Kx Pdelta n B. K p= Ks/P delta n C. Ks = Kp (RT)delta n D. Kx = Kp (P/RT) delta n
262	Major ingredients of traditional ceramics	A. Silica B. Clay C. Feldspar D. All
263	Which of the following colligative properties can be used to characterize colloidal particles.	A. Lowering in vapour pressure B. Elevation in boiling point C. Depression in freezing point D. Osmotic pressure
264	Which of the following pair on aldol condensation followed by dehydration gives methyl vinyl ketone.	A. HCHO and CH3COCH3 B. HCHO and CH3CHO C. CH3CHO and CH3CHO D. CH3COCH3 and CH3COCH3
265	Which of the following statement is not correct with reference to cell constant.	A. The dimensions of cell constant is cm-1 B. It is used to determine the specific conductance C. It is measured with KCl solution D. Specific conductance does not vary with concentration.
266	Vitamin which contains cobalt is.	A. Vitamin B1 B. Vitamin B2 C. Vitamin B6 D. Vitamin B12
267	Among LiCl, BeCl2, BCl3, and CCl4 the covalent bond character follows the order.	A. LiCl < Becl2< Bcl3 < CCl4 B. LiCl > BeCl2 < BCl3<CCl4 C. LiCl< Becl2 < BCl3 < CCl4 D. LiCl > Becl2 > BCl3 < CCl4
268	The number of significance figures in the number 80.7 is.	A. 1 B. 2 C. 3 D. 4
269	Which of the following potassium fertilizers are more useful for horticultural crops tobacco and potatatoes.	A. KNO3 B. KCl C. HNO3 D. H2SO4
270	Which is the correct configuration of Fe3+ (Z= 26)?	A. [Ar] 4s² , 3d⁶ B. [Ar] 4s² , 4d⁵ C. [Ar] 3d⁵ D. None of these
		A. 10 B. 25

271	The tensile strength of a carbon nanotube is _____ times that of steel.	C. 100 D. 1000
272	What is the purpose of molybdenum in steel alloying.	A. To increase dynamic and high temperature strength and hardness B. To increase brittleness C. To increase corrosion and resistance D. All above
273	Which of the following statement in not correct with respect to inductive effect.	A. Bond length decrease with increase in inductive effect. B. Inductive effect generates polar character in bonds C. Variation in strength of aliphatic acids can be explained D. It alone cna explain the basically of triphenylamine
274	The aluminium salt commonly used to stop bleeding is	A. Aluminium sulphate B. Potash Alum C. Aluminium chloride D. Aluminium fluroide
275	Which of the following is not an ore of Cr.	A. Chrome iron B. Nicollite C. Crocisite D. Chrome ochre
276	Which one of the following statements is not true.	A. Transition metals form alloys B. Transition metals form complexes C. Zn, Cd and Hg are transition metals D. K2[PtCl6] is a well known compound but corresponding nikel compound is not knonwn
277	Which of the following steps is not involved in chemical analysis.	A. Separation of sample in pure form B. Separation of the sample in the mixture form C. Preparation of sample for the analyisis D. Validity of experimental results
278	WIhich of the following can act as a protective colloid	A. Gelatin B. Silica gel C. Oil in water emulsion D. All three
279	Which of the following term refers to nearness between several measurements of the same quantity.	A. Accuracy B. Precision C. Standard error D. Standard error of mean
280	Phenol on reaction with ethanoic anhydrides in the presence of sodium ethanoate gives.	A. Phenyl benzoate B. Ethyl benzoate C. Phenyl ethanoate D. Phenyl methyl ether
281	What is the raw material of sugar industry.	A. Sugar cane B. Potato C. Carrot D. Sugar heat E. Both A and C
282	According to R, S system the correct order of priority of the following groups is .	A. -CH2OH > - CHO > ----- COOH B. -COOH > ------CHO > ------ CH2OH C. ---- CH2OH -----> - COOH > -----CHO D. --- COOH > - CH2OH >----- CHO
283	Which of the following statement is false about resonance.	A. It increase the stability of a molecule B. It leads to similar type of bonds C. It increase the reactivity of the molecule D. It decrease the reactivity of the molecule.
284	The three isotopes of hydrogen differ from one another in	A. Atomic number B. Number of protons C. Nuclear charge D. Nuclear mass
285	Nitric acid has the property	A. Nitrating B. Reducing C. Redoxing D. None of above
		A. 0.15 to 0.30

286	The chrome vanadium steels contain how many percent of vanadium.	B. 0.05 to 0.15 C. 0.30 to 0.45 D. 0.45 to 0.60
287	Which of the following organic molecule is not aromatic.	A. Benzene B. Naphthalene C. Anthracene D. Cyclo-octatetraene
288	What is a coal that has been previously burned in an oxygen poor environment?	A. Tuyere<o:p></o:p> B. Coke C. Silver D. Diamond
289	A^{o} or 10 Dq is called crystal field.	A. Energy B. Splitting energy C. Stabilization energy D. None of above
290	Which of the following chloride is soluble in hot water.	A. Hg2Cl2 B. AgCl C. PbCl2 D. All above
291	The attraction which exists between carbon dioxide molecules in solid carbon dioxide is due to.	A. Van der Waal's forces B. Molecule ion forces C. ionic bonds D. hydrogen bonds
292	Which of the following statements is not correct regarding the structure of DNA.	A. It has a double helix structure. B. There are hydrogen bonds in its structure C. Unlike RNA there is no fixed ratio of bases in DNA D. The code for protein synthesis is given by the sequence of bases in DNA
293	Tetra halides do not undergo hydrolysis	A. C B. Si C. Sn D. Pb
294	The formula of sulphur sesquioxide	A. SO4 B. S2O7 C. S2O3 D. SO3
295	Which of the following cast iron is heat treated for ductility.	A. Gray iron B. Malleable iron C. White iron D. None of these
296	An aromatic compound has a molecules formula C7H8O. How many isomers are possible for this compound.	A. 3 B. 4 C. 5 D. 6
297	The base which in not present in DNA is	A. Adenine B. Guanine C. Thymine D. Cytosine
298	Which of the following sulphide is yellow in colour.	A. HgS B. PbS C. CdS D. SnS
299	The formation of daughter DNA's from parent DNA is called.	A. Transalation B. Transcription C. Reproduction D. Replication
300	For a given mass of a gas, if pressure in reduced to half and temperature in doubled, then volume.	A. 2V² B. 4V C. 8 V D. V²
301	What types of bonding occurs in d-block elements.	A. Ionic B. Covalent C. Metallic D. Both B and C
302	Which of the following statement is not true with respect to photo chemical reactions.	A. These take place in the presence of light B. Free energy of these reactions may be positive or negative

		C. Light intensity affect these reactions D. Temperature has significant affect n rate of these reactions
303	Which of the following reaction cannot be used for the synthesis of a amino acids.	A. Gabriel phthalimide B. Streckers synthesis C. Sorensen synthesis D. Schmidt synthesis
304	Are used as water repellents	A. Carbides B. Silicon C. Silicones D. Silicates
305	The first ionization energy in electron volts of nitrogen and oxygen atoms are respectively given by.	A. 14.6, 13.6 B. 13.6, 14.6 C. 13.6, 13.6 D. 14.6, 14.6
306	Beryllium shows diagonal relationship with.	A. Mg B. Al C. Na D. B
307	While compacting the concrete by a mechanical vibrator, the slump should not exceed.	A. 2.5 cm B. 10 cm C. 3.1 cm D. 5.0 cm
308	In a system of designating wrought aluminum alloys a sour digit number is used what does the first digit indicate.	A. The purity of aluminum B. The identity of the alloy C. The alloy group D. All of above
309	When rain is accompanied by a thunderstorm, the collected rain water will have pH	A. Slightly lower than that of rain water without thunderatorm B. Slightly higher than that of rain water without thunderastorm C. Uninfluenced by occurrence of thunderatorm D. Which depends on amount of dust in air
310	Hydrogen bons holding the strand to nucleic acids are formed between	A. Sugar and base units B. Base unit C. Sugar ane phosphate units D. Sugar units
311	The sample characteristics affecting the weight loss curve include.	A. Amount of sample B. Sample particle site C. Heat of decomposition reactions D. All
312	The expression of specific conductance is given by	A. Ls= l/R , l/A B. Ls = L l/A C. Ls = l/L , A/l D. LS = r l/A
313	The most suitable method of separation in mixture of o-and p- nitrophenol is.	A. Steam distillation B. Chromatography C. Ion-exchange D. Sublimation
314	Which of the following is most acidic.	A. Phenol B. p-nitrophenol C. o-Nitrophenol D. m-Nitrophenol
315	Magnalium is alloy of Aluminium which is used in	A. Scientific apparatus B. Aircraft parts C. Rail road care D. Boat machinery
316	Which one of the following ions is colourless.	A. Cu+ B. Co2+ C. Ni2+ D. Fe3+
317	The H2SO4 obtained by the contact process having purity	A. 70% B. 74% C. 78% D. 82%
318	Pi bond is formed	A. By the overlapping of atomic orbitals on internuclear axis B. By transference of electrons C. By sidewise overlapping to half filled p orbitals D. By overlapping of s-orbitals with p

		D. By overlapping of s orbitals with p orbitals
319	Photochemical among is related to pollution of	A. Air B. Water C. Soil D. All of the above
320	The silicate chains are present in	A. Silica B. asbestos C. Beryl D. Clays
321	Which of the following steps are involved in the extraction of copper.	A. Roasting B. Smelting C. Refining D. All
322	Increasing oxygen contents in oxyacids leads to.	A. An increase in thermal stability B. An increase in acid strength C. A decrease in oxidizing power D. All above
323	Which of the following does not belong in the group of herocyclic dyes.	A. Acridine B. Cyanine C. Methylene blue D. Amido black
324	The most common oxidation state of alkaline earth metals is.	A. +1 B. +2 C. -2 D. -1
325	Molecules have zero dipole moment	A. CO2 B. BCl3 C. CH4 & CCl4 D. All above
326	Which of the following compounds would you use in order to obtain a crystalline derivative of an aromatic amine.	A. 2,4 Dinitrophenyl hydrazine B. Nitrous acid C. Benzoyl chlride D. None of these
327	Granulated sugar also known as.	A. Brown sugar B. Refined sugar C. White sugar D. None of these
328	Which of the following statements is not related with entropy.	A. It is a measure of disorder B. It is a measure of unavailable energy C. It is a function of thermodynamics probability D. It is a path function
329	What element is the most abundant by mass in the Earth's crust.	A. Fe B. H C. O D. K
330	The electronic configuration of some elements are given below. The element with highest electron affinity is	A. 1s2, 2s2, 2p3 B. 1s2, 2s2, 2p4 C. 1s2, 2s2, 2p5 D. 1s2, 2s2, 2p6
331	The different types of glass are.	A. A- glass, C-Glass, E-Glass and S-Glass B. A-Glass, B-Glass, E-Glass, S-Glass C. AR -Glass, C-Glass, E-Glass and S -Glass D. A-B Glass
332	Aluminium does not corrode as does iron because.	A. Al does not react with O2 B. a-protective layer of Al2O3 forms on the metal surface C. Al is harder to oxidize than is Fe D. Fe gives chathodic protection to Al
333	Which one of the following pairs are chemically dissimilar.	A. Na and K B. Ba and Sr C. Zr and HF D. Ca and Zn
334	Casionic polymerization is initiated by	A. BF3 B. NaNH2 C. BuI D. Both b and c
335	Which of the following is the third most abundant element in the nature	A. Oxygen B. Sulphur

335	Which of the following is the third most abundant element in the nature.	C. Aluminum D. Hydrogen
336	The electrolysis of molten metal hydride will produce dihydrogen gas.	A. At cathode B. At anode C. At both the electrodes D. At none of the electrodes
337	The second order rate constant can have units.	A. dm-6 mol 2 s-1 B. dm3 mol s-1 C. dm³ mol⁻¹ s⁻¹ D. dm6 mol-1 s-1
338	A chemical reaction resulting in a change in the electric charge on the reacting particles may be called as.	A. Add ion reaction B. Redox reaction C. Elimination reaction D. Chain reaction
339	Ammonia when used directly as a fertilizer is to be injected about_____ under the surface to keep it from seeping out.	A. 2 inches<o:p></o:p> B. 4 inches<o:p></o:p> C. 6 inches<o:p></o:p> D. 10 inches<o:p></o:p>
340	Which of the following dyes belongs to the group of acridine dyes.	A. Acriflavin B. Alizarin C. Indigotin D. Cyanine
341	Pick out the incorrect statement about K2Cr2O7	A. It is thermally stable B. It dissolves in alkali to form chromate C. It oxidizes acidified FeSO4 solution to Fe2(SO4)3 D. It is used as cleansing agent for glassware. etc. When mixed with cold con. H2SO4
342	Which of the following regions of the spectrum would be used to determine the structure of the crystalline solids.	A. Microwave B. X-rays C. Visible D. Infrared
343	When orthoboric acid is heated strongly it gives.	A. B2O3 B. H2B3O7 C. HBO2 D. B
344	The gap between occupied and the unoccupied orbitals is not very large and the conduction of electricity is negligible at lower temperature and appreciable at high temperatures then it will be.	A. Good conductor B. Non conductor C. Semi conductor D. None of the above
345	The atomic orbitals are progressively filled in order of increasing energy. This statement is called as	A. Hund's rule B. Aufbau's rule C. (n+1) rule D. Planck's rule
346	Which of the following techniques involve gas as the mobile phase.	A. HPLC B. GLC C. TLC D. Paper chromatography
347	Which of the following analytical technique is based on the emission of light radiation.	A. Flame photometry B. Atomic absorption spectrophotometry C. Raman spectroscopy D. Conductometry
348	The formula of Tetraboric acid is.	A. H2BO3 B. HBO2 C. H2B4O7 D. H6B4O9
349	Pick out incorrect statemtn about K2r2O7	A. It oxidizes acidified solution of H2S to S B. It oxidizes Ki TO i2 C. It oxidizes HCl to Cl2 D. It gives oxygen, when treated with

		cold conc. H2SO4
350	The law of trinds was proposed by	A. Dobereiner B. Newlands C. Lother Mayer D. Chancourtois
351	The maximum oxidation shown by managanese is.	A. +2 B. +7 C. +4 D. +5
352	The number of bonds formed by the central atom is called its.	A. Valence number B. Complex number C. Coordination number D. Avogadro's number
353	Which of the following statement is not correct regarding the constant R . and in ideal gas equation PV= nRT	A. Its value in independent of temperature B. Its value is independent of pressure C. In SI Units its value is 8.314 K⁻¹ mol⁻¹ D. It is called the universal gas constant per molecule.
354	The bond angle between hybrid orbitals in methane is	A. 115.5^o B. 109.5^o C. 105,7^o D. 120^o
355	Which of the following configuration is associated with biggest jump between second and third IE.	A. 1s2, 2s2, 2p2 B. 1s2, 2s2, 2p6, 3s1 C. 1s2,2s2,2p6, 3s2 D. 1s2, 2s2,2p6
356	Which of the following method is based on the solubility difference between the analyte and the unwanted components.	A. Distillation B. Complex formation C. Electrodepostion D. Precipitation
357	Which of the following reactions have small enthalpy change.	A. NaOH with HCl B. NaOH with CH3 COOH C. HCl with NH4 OH D. None of these
358	The element having highest ionization energy and least electron affinity belong to	A. Period 1 , group 18 B. Period 2, group 17 C. Period 2, group 1 D. Period 2, group 2
359	The total number of crystal systems and the number of Bra via is lattices are.	A. 7,7 B. 7,14 C. 14,7 D. 14,28
360	An equilibrium the free energy change delta F for a reaction is.	A. Maximum B. Minimum C. Zero D. Negative
361	Manganese steel usually contains how many percent of manganese.	A. 1 to 5 B. 3 to 10 C. 11 to 14 D. 14 to 18
362	The number of Glass products now manufactured is.	A. 25,000 B. 75,000 C. 50,000 D. All of these
363	Soapy detergents and soapless detergents behave differently in hard water because they	A. Have different hydrophilic heads B. Have different hydrophobic hydrocarbon chains C. Have different pH values D. Above A and C both
364	Which of the following compound will be optically active.	A. Suceinic acid B. Meso tartaric acid C. Acetic acid D. Lactic acid
365	The element with maximum first ionization energy is.	A. B B. N C. O D. C
366	The most stable carbonium ion is	A. See butyl B. n-butyl C. Tert butyl

		D. None of the above
367	The deficiency of which vitamin leads to beri brainteaser	A. Thiamine B. Riboflavin C. Pyridoxine D. Asorbic acid
368	The Lewis formula of SOCl2 the total number of bond pairs and lone pairs of electrons around sulphur are.	A. 2,1 B. 2 ,2 C. 3 ,1 D. 3 ,0
369	Enfleurage process is ued to extract the essential oils from	A. Bark of plant B. Seeds of plant C. Leaves of plant D. Flowers of plant
370	The correct order of ionization energies of alkali metals is.	A. Li > Na > K > Rb B. Na > K > Rb > Li C. Rb > K > Na > Li D. Rb > K > Li > Na
371	Alumina is not used as	A. Refractory material B. A medium in chromatography C. An abrasive D. A White pigment
372	What is defined as an intimate mechanical mixture of two or more phases having a definite composition and a definite temperature of transformation within the solid state.	A. Pearlite<o:p></o:p> B. Eutechtoid<o:p></o:p> C. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri","sans-serif";mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:Arial;mso-bidi-theme-font:minor-bidi;mso-ansi-language: EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">Delta solid solution D. None of these
373	Which of the following is an alloy of copper	A. Brass B. Bronze C. Monel metal D. All
374	Which of the following technique is used to separate substance of high molecular weight of different charges.	A. Dialysis B. Electrophoresis C. Solvent D. None of the abvoe
375	The main active contaminants of uranium processing are.	A. U - 235 B. U-238 C. Pu -234 D. All above
376	Which of the following group will have hyper conjugation effect when attached to benzene.	A. ------ CH3 B. ----C6H5 C. ------C(CH3)3 D. ---------CH(CH3)2
377	B.P of heavy water is	A. equal to that of ordinary water B. greater than that of ordinary water C. Less than that of ordinary water D. equal to that of distilled water
378	Among the elements of second period the element with highest melting point belongs to group.	A. 1 B. 14 C. 17 D. 18
379	The concept is also known as electron pair donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux -Flood D. Usanovich
380	Layer of the C -atom in graphite are hold together by	A. Covalent bonds B. Free electrons C. Ionic bond D. Van Dar Waals forces
		A. Phosphorescence

381	The reverse of photo chemical reaction is called.	A. Phosphorescence B. Chemiluminescence C. Fluorescence D. Photscnaitization
382	Carbylamine reaction proceeds via the intermediate formation of.	A. Alkyl isocyanide B. Chloride ion C. Alkyl carbonion D. Dichloro methylene
383	Which of the following technique involves the bonding of hydrophobic functional group to solid particle, surface and acts as extracting phase	A. Liquid phase extraction B. Solid phase extraction C. Electrophoresis D. Gel electrophoresis
384	What is the effect of aluminum in cast iron.	A. To increase hardness above 0.5 % B. To deoxidize molten cast iron C. To affect machinability, ductility and shrinkage depending on form D. Both A and B
385	Chlorine is used in	A. Sterilizatin of water B. Extractionof gold C. Bleaching of cotton D. All above
386	The volume of a given mass of gas at constant temperature varies inversely with the pressure. This is a statement of.	A. Charlea's law B. Avogadro's law C. Boyle's law D. Dalton 's law
387	Which of the following radical is not a member of III group	A. Al³⁺ B. Fe²⁺ C. Ca²⁺ D. Fe³⁺
388	One ppm solution of NaOH Contain 1000 mg of the solute per how much of the volume of the solution.	A. 1000 mL B. 100 mL C. 10 mL D. 1 mL
389	Which of the following pollutant results from combustion of fossil fuels.	A. SO2 B. NO2 C. CO D. All above
390	If steel is heated to a temperature well below red heated and is then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
391	Which of the following polymers is chlorinated.	A. Orlon B. Neoprene C. Dacron D. None of these
392	Which of the following techniques is involved in purification of organic compound.	A. Distillation B. Sublimation C. Solvent extraction D. All above
393	What is a process of producing a hard surface in a steel having a sufficiently high carbon content to respond to hardening by a rapid cooling of the surface?	A. Cyaniding<o:p></o:p> B. Nitriding<o:p></o:p> C. Flame hardening<o:p></o:p> D. Stability<o:p></o:p>
394	Which is the following is not a buffer.	A. H2CO3/HCO3 B. NH4Cl/NH4OH C. CH3COOH/CH3COONa D. NH3OH/CH3COOH
395	The correct order of acid strength is.	A. HIO4 > HBrO4 > HClO4 B. HClO4> HBrO4 > HIO4 C. HBrO4 > HIO4 > HClO4 D. HBrO4 > HClO4 > HIO4
396	An explosive	A. Nitroglycerine B. Trinitrotoluene C. Fluorine perchlorate

		D. All above
397	How much amount of NaOH is required to prepare 100 mL of 1 N solution.	A. 80 g B. 4 g C. 40 g D. zero
398	Al Cl3 is used in	A. Manufacturing of petrol B. In borax bead test C. Prezervation of food D. All above
399	The main constituents of _____ are boron oxide and silica.	A. Pyrex glass B. Low silica glass C. Soda lime glass D. Super hard glass
400	The bond between two identical non metal atoms has a pair of electrons.	A. Unequally shared between the two B. Transferred fully from one atom to another C. With identical spins D. Equally shared between them
401	Which of the following methods is chemical in nature.	A. Acid bas titration B. Redox titration C. Complexometric titration D. All above methods
402	What is the scaling off of a surface in flakes or layers as the result of corrosion?	A. Expoliation<o:p></o:p> B. Corrosion fatigue<o:p></o:p> C. Scaping<o:p></o:p> D. Fretting<o:p></o:p>
403	Boric acid is added to glass because is.	A. Makes the glass opalescent B. Reduces the coefficient of expansion C. Makes the glass brittle D. Increase refractive index of the glass.
404	The colloidal solution of arsenic sulphide prefers to absorb	A. NO3 B. K+ C. S2- D. H+
405	Which of the following generally increases on going from top to bottom in a group.	A. Metallic charcter B. Electronegativity C. Oxidising behaviour D. Raducing behaviour
406	The basic strength of hydrides of group 15 elements very in the following order.	A. NH3 > PH3 > AsH3 > SbH3 >BiH3 B. PH3 > NH3 > AsH3> SbH3 > BiH3 C. BiH3> NH3 >PH3 > AsH3 > SbH3 D. NH3 > PH3 > SbH3 > AsH3 > BiH3
407	The Langmuir theory of unimolecular adsorption is generally valid at.	A. Low pressures and low temperature B. Low pressures and high temperature C. High pressures and low temperature D. High pressure and high temperature
408	Among alkali metals, the least metallic element is.	A. Li B. Na C. Rb D. Cs
409	Of the following the commonly used n the laboratory desiccator is.	A. Anhyd. Na2Co3 B. Anhyd Ca Cl2 C. Dry NaCl D. None of the above
410	In coordination chemistry the donor atom of a ligand is.	A. A Lewis acid B. The counter ion C. The central metal atom D. The atom in the legend that shares an electron pair with the metal
		A. HF

411	Which is the strongest reducing agent.	A. HF B. HCl C. HBr D. HI
412	Which of the following properties are not related to an atom.	A. An atom consists of two basic parts , a nucleus and one or more electons. B. The nucleus is the central core of an atom C. An electron is a heavy and negatively charged particle. D. The nucleus itself consiste of two particles.
413	Which of the following item is not symmetry element.	A. Pllane of symmetry B. Inversion centre C. Improper rotation D. Optical activity
414	A pH of a neutral solution at 100 ^{o}C when Kw = 1.0 x 10^{-12}	A. 0 B. 7 C. 6 D. 7
415	Which of the following compounds has highest boiling point.	A. HI B. HF C. HBr D. HCl
416	Which of the following hydroxide is getatinous in nature.	A. Fe(OH)3 B. Al(OH)3 C. Ca(OH)3 D. Cr (OH)3
417	Which of the following is a planar molecule.	A. Acetone B. Formic acid C. Acetic acid D. All above
418	Molecular weight of proteins may be determined by	A. Osmotic pressure measurements B. Sedimentation methods C. Light scattering methods D. All of these
419	Which one of the following ionsis colourless.	A. Cu+ B. Ni2+ C. Co2+ D. Fe3+
420	A half cell reaction is one that	A. Occurs at one electrode B. Goes only half way to completion C. Involves a half mole of the concentration of the solution D. Always oxidizes
421	Which ratio decides the efficiency nano substance.	A. Weight /volume B. Surface area/volume C. Volume/weight D. Pressure/volume
422	If diesel has cetane number of 50 then the diesel index will be.	A. 36 B. 46 C. 56 D. 66
423	LPG is use this	A. Vehicles<o:p></o:p> B. Aviation Fuel<o:p></o:p> C. Home D. All above
424	Stablization of particles and their reactivity is affected by.	A. Surface properties B. Bulk properties C. Regardless to the surface properties D. No of particles
425	The force responsible for dissolution of ionic compounds in water are	A. Hydrogen bonds B. Ion dipole forces C. Ionic bonds D. Van Der Waal forces
426	The concentration required to give a signal equal to three times the standard deviation of the baseline is called.	A. Sensitivity B. Detection limit C. Signal to noise ratio

		D. None of the above
427	Which of the following is the major process when neopentyl bromide is dehydrogenate with alcoholic potash.	A. 2- methyl -1- butene B. 2- methyl- 1- butene C. 2,3 -dimethyl butene D. 2- butene
428	Which of the following statements is not relevant to the Plank's quantum Theory.	A. Radiant energy is not absorbed or emitted continuously B. Radiant energy is emitted or absorbed in the form of small packets of energy. C. The quantum oflight energy is called photon D. The energy associated with photon of radiation is directly proportional to the wavelength.
429	In XeF2 molecules, Xe atom undergoes hybridization	A. spd B. sp2 C. sp3 D. sp3d
430	Which of the following responsible for depletion of ozone layer in upper strata of the atmosphere.	A. Polyhalogens B. Ferrocene C. Freons D. Fullerencs
431	A diameter of human hair is approximately _____ m	A. 75000 B. 75 C. 7.5 x 10⁻⁵ D. 7.5 x 10⁻⁹
432	Carbon in wrought iron is present as	A. Silicon carbide B. Iron carbide cementite C. Graphite D. Partly as iron carbide and partly as graphite
433	For quality control of Portland cement, the test essentially done is.	A. Setting time B. Soundness C. Tensile strength D. All
434	Which of the following method is used to separate small molecules form the larges molecules from the larger molecules in diffusing through a membrane.	A. Dialysis B. HPLC C. FPLC D. TLC
435	The yellow colour of chromates changes to orange red on acidification, due to the formation of.	A. Cr3+ B. Cr2O3 C. Cr2O7²⁻ D. Cro3
436	Human hearing is sensitive to frequency in the range of about	A. 10,000 - 20,000 Hz B. 10 - 10,000 Hz C. 16- 20,000 Hz D. None of the above
437	Inorganic acids (HCl, HBr, HNO3 etc) have K value.	A. < 1 B. >1 C. >10 D. <10
438	Which of the following is not strong electrolytes.	A. HCl B. H2SO4 C. HNO3 D. CH3COOH
439	In order to increase the rate of the reaction one should.	A. Increase the concentration of products B. Decrease the concentration of reactants C. Decreases the concentration of products D. Both C and D statement are correct
440	What is the colour of pulp obtained from chemical pulping.	A. Black B. Brown C. Blue D. Red
441	The acetylene molecule contain a	A. Single bond B. Double bond C. Triple bond D. Co ordinate bond
		A. Silicon carbide B. Iron carbide or cementite

442	Carbon in wrought iron is present as	C. Graphite D. Partly as iron carbide and partly as graphite
443	Which among the following is a Talse statement.	A. SO3 is obtained by the catalytic oxidation of SO2 B. SO3 has trigonal planar geometry in gaseous state C. SO3 in nauseous state has all S-O bonds equivalent D. SO3 gas shows more solubility in water than in H2SO4
444	Domestic waste mostly constitutes	A. Non biodegradable pollution B. Biodegradable pollution C. Effluents D. Air pollution
445	Any property whose magnitude is independent of the amount of substance present is called a/an	A. Extensive property B. Colligative property C. Structural propety D. Intensive property
446	For a compound to act as a dye it must have	A. A suitable colour B. Ability to fix to fibre C. Both A and B D. None of these
447	The physical methods of nano roads syntheses involves.	A. Top down approach B. Bottom up approach C. Left right approach D. Right left approach
448	Which group contains elements that exist as monoatomic molecules.	A. 1 B. 2 C. 14 D. 18
449	The first noble gas compound was	A. XeO3 B. XeF4 C. XeF6 D. Xe +[PtF6]
450	If the absorbed light is green the transmitted light will be	A. Purple B. Orange C. Violet D. Black
451	The rate of a chemical reaction is proportional to the product of the active mean of the reactants, This is a statement of.	A. Law of dynamic equilibrium B. Le Chatlier's principle C. Law of mass action D. Solubility product principle
452	Biogical role of nucleic acid doe snot include	A. Genetic continuity B. Protein syathesis C. Hybridization D. Mutation
453	The quantum yield of a Photo chemical reaction in	A. Always less than unity B. Always equal to unity C. Always greater than unity D. Can have any value > 0 depending on the reaction
454	Which of the following source of energy is abundant everlasting and non polluting.	A. Nuclear B. Electric C. Solar D. All above
455	Radon is obtained only in the radioactive decay of	A. Radium B. Thorium C. Actinium D. Any of above
456	Which of the following instruments is used to measure the optical activity.	A. Refractometer B. Conductivity meter C. Polarimeter D. Torsion meter
457	The presence of which of the following in drinking water is responsible for mottling of teach.	A. Mercury B. Iodine C. Chlorine D. Flourine
458	Which of the following liquids has lowest vapour pressure at 25 ^{o}C	A. Benzene B. Chloroform C. Ether D. H2O

459	Homolytic fission of convalent bond results in the formation of.	A. Free redicals B. Carbocations C. Carbonions D. Both B and C
460	The most stable oxidation state shown by lead is.	A. +2 , +4 B. +2 only C. +3 , +4 D. +4 only
461	Excluding H-atom , Hydrogen bond never involves more than atoms.	A. One B. Two C. Three D. Four
462	Vet days are generally applied to the fabric in the form of.	A. Mordants B. Leuco base C. Oxidised base D. Dispersed dyes.
463	In which of the following group, each member given a positive iodoform test.	A. Methanol, ethanol, propanone B. Ethanol, isopropyl alcohol, methanol C. Ethanol, ethanal, isopropyl alcohol D. Propanal 2-propanol, propanone
464	Which of the following molecules can oxhibit geometrical isomerism.	A. CH3CH = CH2 B. CH3CH = CHCH3 C. (CH3)2 C = CH2 D. CH3CH = C(CH3)2
465	Which of the following analytical method is based on scattering of radiation.	A. Emission spectroscopy B. Colorimetry C. Turbidimetry D. Polarimetry
466	Cyclic polymers of ethylene glycol formed by condensation are called.	A. Crown ether B. Brown ether C. Cryptates D. Both A and C
467	The electrical resistance of stainless steels can be as much as_____ time that of carbon steel.	A. 5 B. 6 C. 10- D. 15
468	Increased asthmatic attacks in certain seasons are related to.	A. Inhalation of seasonal pollens B. Eating of seasonal vegetables C. Low temperature D. Wet and dry environment
469	Bromine number is measure of.	A. Paraffins B. Unsaturates C. Saturates D. None of these
470	Oxygen and sulphur exist in state	A. Free B. Combined C. _{Both free & combined} D. None of above
471	Which type of polymers the Vulcanised rubbers is.	A. Linear B. Cross jinked C. Branch chain D. Any one of these
472	Sodium react more vigorously than lithium because.	A. It is a metal B. It has higher atomic mass C. It is more electronegative D. It is more electropositive
473	In normal mode of operations of liquid liquid partition, a polar stationary phase is used with a non polar mobile phase Which of the following solvent is used as mobile phase.	A. Ethanol B. Propanol C. Butanol D. Hexane
474	The temperature of which the compound melts into a liquid to the same composition as the solid is called the	A. Congruent melting point B. Incongruent melting point C. Peritectic temperatures D. Metastable point
475	UV radiation from the sun causes a reaction in the atmosphere that leads to production	A. Fluorides B. Carbon monoxide C. Sulphur dioxide D. Ozone
476	Trimethylamine is a weaker base than dimethylamine is explained by	A. Steric effct B. Resonance effect

476	Trimethylamine is a weaker base than dimethylamine is explained by	C. Inductive effect D. All above
477	Which of the following statements is not correct with the concept of Bronsted concept of acids and bases.	A. An acid can donate a proton B. A base can accept a proton C. This concept has many bases that have OH^- ions D. This concept is more general
478	Which of the following substance is generally not considered an air polutant.	A. CO B. CO_2 C. SO_2 D. NO_2
479	Which of the following test to used to find out whether the observed data differ significantly from the one obtained from theoretical distribution.	A. Chi square test B. F -Test C. Student's test D. Coefficient of variance
480	Which one of the following sets of elements has the strongest tendency to form positive ions in gaseous state.	A. Li, Na, K B. F, Cl, Br C. Be, Mg, Ca D. O, S , Se
481	In glass making the whole combination of ingredients is called a.	A. Gangue B. Batch C. Mixture D. None of these
482	Elements in the same vertical group of the periodical have same	A. Number of electron B. Atomic number C. Number of valence elections D. Electronic configuration
483	To complete transfer of a shared pir of electrons to one of the atoms joined by a double or triple bond at the requirement of an attacking reagent is known as.	A. Inductive effect B. Resonance effect C. Eletromeric effect D. Stark effect
484	30 mL of an acid solution is neutralized by 15 mL of 0.2 N base. The strength of acid solution is.	A. 0.1 N B. 0.15 N C. 0.3 N D. 0.4 N
485	Which of the following compounds is electrovalent in nature.	A. SO_2 B. ICl C. KBr D. CHI_3
486	The condensation between formaldehyde and acetaldehyde in the presence of conc. NaOH and heat gives.	A. Acrolein B. Mixture of CH_3OH and CH_3COO Na. C. Mixture of CH_3CH_2OH and $HCOO^-Na^+$ D. None of these
487	Which of the following technique is not related to instrumental analysis.	A. Optical method B. Colorimetry C. Polarography D. Gravimetric analysis
488	Which of the following elements has the highest ionization energy.	A. Na B. Si C. Ar D. Cl
489	A red color gas, on condensing ti gives a dark blue liquid.	A. NO B. N_2O C. N_2O_3 D. N_2O_4
490	iodine is used as a	A. Photography B. Manufacture of dyes C. Analgesic D. All above
491	In the Lewis structure of H_2SO_4 molecule the total number of unshared electrons in valence shell of various atoms is.	A. 8 B. 16 C. 12 D. 20
492	The one which is not a purine base	A. Cytosine B. Guanine C. None of these D. Adenine
493	Which compound among the following does not contain an ionic bond.	A. NaOH B. HCl C. KaS D. LiH

		D. LiH
494	An element having low IE and low EA is likely to belong to.	A. Group IA B. Group IB C. Group VII A D. Group VIII
495	What element constitutes the major component of most bronzes.	A. Tin B. Zinc C. Carbon D. Aluminum
496	Glass industry requires soda ash with	A. Solids density 1.91 and bulk density 1.0 B. Solids density 1.86 and bulk density 0.6 C. Solid density 1.80 and bulk density 0.58 D. All of above
497	For a given mass of a gas at constant temperature, if the value V becomes a times, the pressure will become.	A. 3P B. P/3 C. 9P D. 3P/T
498	Which of the following metals is the most abundant in the earth's crust.	A. Mg B. Ca C. K D. Na
499	At constant temperature , the decrease in Halmholts free energy is equal to.	A. Decrease in entropy B. Increase in entropy C. Reversible work done by the system D. All types of work done
500	Which of the following compounds does not show dipole moment.	A. CH3OH B. HBr C. CCl4 D. CHCl3
501	The ration of thermal conductivity of silver to that of a carbon nanotube is.	A. 100 : 1 B. 1 :100 C. 10:1 D. 1:10
502	Which of the following phenomena is not explained by the classical mechanics.	A. Black body radiation B. Photoelectric effect C. Atomic and molecular spectra D. Heat capacities of solids E. All of the above
503	What letter suffix steel identification means that it is steel with boron as an alloying elements.	A. xxL xx B. xxBxx C. xxHxx D. xxKxx
504	J.J. Thomson established certain properties about cathode rays. Which of the following is not related to cathode rays.	A. Cathode rays from a gas discharge tube consists of negatively charged particles B. Cathode rays are called electrons. C. The e/m ratio of cathode rays depends on the gas inside D. Cathode rays are affected by electric and magnetic fields.
505	Which of the following trace elements may be present in the particulate materials.	A. Cadmium B. Nickel C. Mercury D. Lead E. All of the above
506	Which of the following techniques is used for cleanup of samples prior to introduction into chromatographic column.	A. Paper chromatography B. TLC C. Solvnent extraction D. Solid phase extraction E. Both C and D
507	Which of the following is not an ore of nickel.	A. Pentalandite B. Siderite C. Garnierite D. Nicollite
508	The number of mole of the solute dissolved per dm3 of the solution is called.	A. Molality B. Formality C. Normality D. Molarity
509	In Pakistan how many units are involved to the production of glass.	A. 20 B. 25 C. 30

		D. None of these
510	What is the oxidation number of the central meal atom in the coordination compound. [Pt(NH#)#Cl]Cl	A. -1 B. 0 C. +2 D. +3
511	Inert pair effect is that	A. When an element shows incertness in chemical combination B. When higher oxidation state is more stable than lower oxidation state C. When an elecrron pair is present on the atom of an element D. When two s -electrons or outermost shell remain paired and do not participate in bonding.
512	The pH of milk is	A. 6.0 B. 6.5 C. 7.0 D. 7.5
513	The pH of water 7 at 25 ^{O}C if water is heated to 70 ^{O}C . Which of the following should be true.	A. pH will decrees B. pH will increase C. pH will remain constant D. None of these
514	Of the molecules, SF4, XeF4, and CF which has square planar geometry.	A. SF4, XeF4 and CF4 B. Sf4 only C. CF4 only D. XeF4
515	Which of the following statement about molecularity is not correct.	A. It cannot be fraction B. It can be obtained from balanced equation C. It may be or may not be equal to the order of the reaction D. it can not be more than 3
516	The addition HCl to 2-pentene give	A. 3-Chloropentane B. 2- Chloropentyne C. 2- Chloropentane D. 2-Chloro-2-methyl butane
517	Indicate false statement about stainless steel	A. The density of stainless steel is about the same as carbon or low alloy steels<o:p></o:p> B. Stainless steels are poor conductors of heat<o:p></o:p> C. Stainless steels are poor conductors of electricity<o:p></o:p> D. Stainless steels have tensile moduli greater than those of carbon and alloy steels. <o:p></o:p>
518	Polyamide jinkage is present in	A. Nylon B. Silk C. Protein D. All of these
519	In terms of number of phases (p) components (C) and degree of freedom (F) the phase rule is expressed as.	A. P + C = F + 2 B. F = P+ C - 2 C. P + F= C + 2 D. P - F = C = 2
520	The lowest K.E. for an electron is three dime national cubic box is given by	A. h²/ 8m^a B. 3h² / 8 ma² C. 9h²/ 8ma² D. 16h² / 8ma²
521	Elements in which differentiating electron enters the (n-1) the d-orbitals of the (n-1) the main shell are called elements.	A. s- block B. p-block C. d-block D. f-block
		A. 3 B. 4

522	Group VA of the periodic table consist of elements.	C. 5 D. 6
523	_____ is heat treatment cycle that prevents glass from harmful stress.	A. Forming B. Annealing C. Batching D. None of these
524	The experimental relationship between rate of the reaction and concentration of the reactants is called.	A. Rate law B. Law of mass action C. Le-Chatelier's principle D. Rate constant
525	Copper is resistant to	A. Air B. Water C. Acid and Alkali D. All of the above
526	The velocity possessed by maximum fraction of molecules at a given temperature is called.	A. Average velocity B. Root mean aquare velocity C. Most probable velocity D. None of the above
527	All steroids on heating wish solenium give	A. phenanthrene B. Cholesterol C. Diels hydrocarbon D. Isoprene
528	At high temperature nitrogen combines with calcium carbide to give	A. Calcium cyanide B. Calcium cyanamide C. Calcium nitride D. Calcium cabonate
529	Which of the following process is not related with cannot cycle.	A. Iso thermal expansion B. Adiabatic expansion C. Isothermal compression D. Isobaric compression
530	Which of the following methods does not give the weight average molecular weight.	A. Sedimentation equilibrium B. Sedimentation velocity C. Light scattering D. Osmotic method
531	Which of the following is not a ligand or complexing agent.	A. NH3 B. CH3COOH C. EDTA D. CN-
532	Which of the following is homopolymer.	A. Starch B. Plexiglas C. Orlon D. All of these
533	Which of the following gas form weakly acidic sulphurous acid	A. SO2 B. SO3 C. NO2 D. NO
534	Which one of the following noble gas is obtained by radioactive disintegrastion	A. Kr B. Br C. Rn D. Xe
535	The terpenoid responsible fo the smell	A. Camphor B. Genenial C. Citral D. Carvone
536	Which of the following reactions is employed to produce ozone in the laboratory.	A. Exposure of air to UV light B. Reaction of F2 with H2O at low temperature C. Reaction SO2 with H2O2 D. Passage of silent electric discharge through oxygen
537	Which of the following represents the correct order of ionic radii	A. La+ < Na+ < K+ < Rb+ B. Li+ > Na + > K+ >Rb+ C. Li = Na + = K+ = Rb+ D. Rb +> Na+> K+ > Li+
538	The product obtained on heating n-heptane with Cr2O3 ____Al2O3 at 600 ^{o}C is.	A. Cycloheptane B. Methyl cyclohexane C. Benzene D. Teluene
539	In group theory the triple degenerate set is denoted by	A. eg B. t2g C. e2g

		D. tg
540	Aluminium hydroxide is.	A. An acid B. An amphoteric hydroxide C. A base D. An explosive hydroexide
541	Which of the following is an allotrops of hydrogen.	A. 0- H2 B. P-H2 C. Both A and B D. None of these
542	Mangalium is an alloy of.	A. Al + Mg B. Mg + Ai + Mn C. Mg + Ai + Cu D. Mg + Ai + Cu + Mn
543	Which of the following statement is not related with the effect of thermal pollution.	A. Decrease in BOD B. Increase in BOD C. Reduction in DO D. Change in algal production
544	PCRA stand for	A. Pollution control research association B. Petroleum conversation Research association C. Petroleum control research association D. All of above
545	In Ostwald's process of manufacturing nitric acid a mixture of ammonia gas with air is mantained with ratio.	A. 1 : 4 B. 1 :3 C. 1: 8 D. 1 : 10
546	During the last two centuries, the atmospheric CO2 contents are increased by	A. 15% B. 25% C. 35% D. 50%
547	The addition of HCl in the presence of poroxule does not follow anti Markovnikov's rule because.	A. HCl bond is too strong to be broken homolytically B. Cl atom is not reative enough to add on to a double bond C. Cl combines with H to give back HCl D. HCl is a reducing agent.
548	A well packed column may hve	A. 100 plates /m B. 1000 plates /m C. 10 plates /m D. 10,000 plates/m
549	Which of the following technique has flame as a source of excitation energy.	A. UV spectroscopy B. I-R spectrocscopy C. Flame photometry D. Raman spectroscopy
550	The maximum number of electron is an atom with l = 2 and n = 3 is	A. 2 B. 6 C. 10 D. 12
551	The element Uuu has atomic numebr	A. 102 B. 111 C. 101 D. 110
552	Which of the following does not apply to nanotechnology.	A. It is a general purpose technology B. It can be called Green technology C. Newtonian mechanics can describe it. D. It involves rearrangement of atoms
553	Primary structure of protein refers to	A. Amino acid sequence B. Arrangement of peptide chains C. Orientation of amino acids D. Whether is has a or b helix in space structure.
554	Which of the following is not an adsorption indicator.	A. Eosin B. Bromocrsol green C. Fluorescein D. Phenolphthalein
555	Which of the following electrode is normally used as reference electrode for a potentiometer.	A. Platinum electrode B. Calomel electrrode C. Silver electrode D. Copper electrode
		A. CO+ NO

556	Which pair of species can undergo chemical reaction with each other.	B. LiH and H2O C. CO2 and HCl D. CaH2 and Sl H4
557	Hot isostatic pressing is not a viable option if the chief criterion is	A. Strength without gram growth B. Lost cost C. Zero porosity D. Make it hard
558	CCl4 has zero dipole moment because of.	A. Planar structure B. Tetrahedral structure C. Similar size of C and Cl atoms D. Similar electrons affinity of C and Cl
559	Stainless steel consists of which elements.	A. Fe only B. Cr only C. Fe and Ni D. Fe ,Ni and Cr
560	Anhydrous AlCl3 cannot be obtained by heating hydrated AlOl3 ,6h2o Because.	A. It decomposes completely to give Al2O3 B. It does not lose water completely C. It undergoes hydrolysis to give Al(OH)3 D. AlCl3 .6H2O is very stable.
561	A colorless gas with pleasant odour and sweet taste.	A. N2O B. N2O3 C. NO D. N2O4
562	Which of the following technique is based on the absorption of light radiation.	A. Spectrophotomerty B. Colorimetry C. NMR D. All the above technique
563	When CH3COOH is titrated against NaOH the pH as the equivalence point is.	A. 7 B. > 7 C. o >7 D. 6.8
564	In which property listed below hydrogen does not resemble alkali metals.	A. Tendency to form cation B. Nature of oxide C. Combination with halogens D. Reducing character.
565	Amino acids are are important in biochemistry which of the following statements is not correct regarding amino acids.	A. These are amphoteric substances tend to undergo internal protein transfer B. In aqueous solutions these substances tend to undergo internal proton transfer C. These for zwitter ion in aqueous medium D. These always contain two amino groups.
566	Heisenberg's uncertainty principle precludes the exact simultaneous measurement of.	A. Velocity and energy B. Velocity and time C. Charge density and probability D. Position and momentum
567	1 meter = _____ nm	A. 10⁹ B. 10⁻⁹ C. 10¹⁰ D. 10⁻¹⁰
568	The ionic product equilibrium constant is.	A. Ka B. Kb C. Kc D. Kw
569	Which of the following enthalpies is always negative.	A. Enthalpy of melting B. Enthalpy of combustion C. Enthalpy of solution D. Enthalpy of formation
570	In the process of preparation of detergents the organic acids produced are neutralized with.	A. Sodium hydroxide B. Sodium sulphate C. Sodium chloride D. Potassium hydroxide
571	Water is often treated with chlorine to	A. Increases oxygen content B. Kill germs C. Cause sedimentation D. Remove insoluble impurities.
572	The osmotic pressure of a solution with definite composition	A. Varies directly as the volume and temperature. B. Various inversely as the temperature.

572	The osmotic pressure of a solution with definite composition.	C. Varies inversely as the volume and directly as the temperature. D. None of the above
573	Among sodium phosphate, sodium sulphate and sodium chloride the solubility in water increases as.	A. Chloride > Phosphate > Sulphate B. Sulphate > Pohosphate> Chloride C. Chloride > Sulphate > Phosphate D. Phosphate > Chloride > Sulphate
574	Which of the following is most soluble in water	A. CaSO4 B. Sr SO4 C. MgSO4 D. BaSO4
575	Consider the coordination compound K2[Cu(CN)4] A coordinate covalent bond exists between	A. K+ and CN- B. Cu2+ and CN- C. K_ and [Cu(CN)4]2+ D. C and N in Cn
576	It has been observed that if one goes on adding KNO3 solution to a precipitate of AgCl the solubility of these precipitates goes on increasing with increasing concentration of K+ and NO3= ions which are not common to AgCl This is due to which effect.	A. Divers ion effect B. Uncommon ion effect C. Activity effect D. All above
577	Which of the microscope techniques is similar to the Atomic Force Microscopy (AFM)	A. Scanning Electron Microscopy B. Scanning Tunneling Microscopy C. Transmission Electron Microscopy D. None of the above
578	What is caustic potash	A. NaOH B. KOH C. NaCl D. KCl
579	The green colour of glass is due to the presence of.	A. Chromium (III) B. Cobalt (II) C. Mn (IV) D. Iron(III)
580	The most electronegative element of the third period is.	A. F B. P C. Br D. Cl
581	The property measured in DTA is	A. Heat effects B. Weight loss C. Rate of change in weight D. Change in temperature
582	The rising world temperature will have serious effect on.	A. Agriculture B. Animal production C. Human being D. All above
583	Which of the following statements is not correct with respect to applications of Hammett equations.	A. It develops a quantitative relationship between structure and reactivity B. This equation on can be used to calculate the value of pK_a C. This equation does not help to calculate the rate of some reactions D. This equation has mechanistic implications
584	Acid rain effects	A. Human being B. Crops C. Aquatic life D. All above
585	Phosphorus has the oxidation state of +3 in	A. Orthophosphoric acid B. Hypophosphoric acid C. Metaphosphoric acid D. Orthophosphorus acid
586	Which of the following process is involved in nitrogen flotation	A. Non symmetric fixation of nitrogen B. Fixation by soil bacteria C. Fixation by yeast D. Fixation by blue green algae E. All above
587	Which one of the following has the highest boiling point.	A. H2O B. H2S C. H2Se D. H2Te
		A. Cu²⁺

588	Which of the following radical is not a member of II group.	B. Cd²⁺ C. Ba³⁺ D. K+
589	Oxidative enzymes are responsible for	A. Biological processes B. Biological oxidation C. Biological hydrolysis D. Biological isomerisation
590	Which of the following is branch chain polymer.	A. Glycogen B. Terylene C. PVC D. Orlen
591	Which of the following compounds cannot be a monomer.	A. CH3-CHOOH -CH2OH B. NH2 -CH2-NH2 C. CH3-CH2-NH3 D. NH2-CH2-CH-CH2 -NH2 CH3
592	CoCl3 ,6NH3 has six NH3 molecules that satisfy the valency of the Cu3+ metal ion	A. Primary B. Secondary C. Both A and B D. None of above
593	Which of the following fuel is used in flame photometry.	A. Hydrogen gas B. Acetylene gas C. Methane D. Propane E. All above
594	The solution of NaOH pH -10.46 contain [OH-]	A. 2.0 X. 10⁻⁴ B. 4.6 X 10⁻⁴ C. 4.6 X 10⁻² D. 4.6 X 10⁻³
595	Which of the following bonds will be non polar.	A. N - H B. O - H C. C - H D. C l - Cl
596	Ionic compounds in general possess both	A. High melting point and non - directional bonds B. High melting points and low boiling poinits C. Directional bonds and low boiling points D. High solubility in polar and non -polar bonds.
597	Which of the following have +3 oxidation states.	A. B & Al B. In & TI C. B & In D. Al & TI
598	The number of gram equivalents of the solute per dm3 of the solution is called.	A. Formality B. Normality C. Molality D. Molarity
599	NH3 has a not dipole moment while BF3 has zero dipole moment Thsi is because.	A. NH3 is not a planar molecule while BF3 is a planar molecule. B. NH3 is a planar molecule, while BF3 is a planner molecule. C. Fluorine is more electronetative than nitrogen D. Born is more electronegative than nitrogen
600	What do you call earth and stone missed with the iron oxide	A. Hematite<o:p></o:p> B. Gangue<o:p></o:p> C. Ore<o:p></o:p> D. Residue<o:p></o:p>
601	A man has to think of alternate sources of energy due to	A. Shortage of vehicles B. Shortage of fossil fuels C. Construction of house D. Running of power plant

602	Any substance which has solidified from the liquid state with crystallization is known as	A. Steel B. Fibre C. Glass D. Asbestos
603	The number of degree freedom at the triple point for the water system in.	A. One B. Two C. Three D. Zero
604	The wire of flash bulb is made up of.	A. Cu B. Ag C. Mg D. Ba
605	In quantum theory, which of the following tells us that the prediction of quantum mechanics must pass smoothly into those of classical mechanics as we progress in a continuous way from microsopic to macroecopic.	A. Uncertainty principle B. Correspondence principle C. Probability distribution D. Aufbau principle
606	The alkali metal with highest melting point is	A. K B. Na C. Li D. Ca
607	Length of semiconductor nanorods are in the range of.	A. 1.50 nm B. 1-50 micro meter C. 100-500 nm D. 50-100 nm
608	The full form of STM is	A. Scanning Tunneling Microscope B. Scientific Technical Microscope C. Systematic Technical Microscope D. SuperTensile Microscope
609	Which of the following technique is used for separation of volatile components.	A. GC B. HPLC C. FPLC D. TLC
610	In the electrolysis of alumina, cryolite is added to.	A. Lower the melting point of alumina B. Increase the electric la conductivity C. Minize anodize affect D. Remove impurites from alumina
611	Which of the following techniques involves the distribution of solute between two immiscible liquid phases.	A. Chromatography B. Electrophoresis C. Solvent extractions D. Solid phase extraction
612	Which of the following statement is not correct regarding dissociation constant (Ka)?	A. It is a measure of the tendency of an acid to split up into ions B. The greater the value of Ka, more is the dissociation C. It is determined by conductimetric method D. It is not a proper parameter for weak acids
613	The emission of light in a biological reaction is known as	A. Fluorescence B. Phosphorescence C. Bioluminescence D. Photolysia
614	Orion is polymer of.	A. Styrene B. CF2 = CF2 C. Vinyl chloride D. Acrylontrile
615	What typical penetrator is used in Brinell hardness test	A. 10 mm ball<o:p></o:p> B. 1.6 mm diameter ball<o:p></o:p> C. 20^o needle<o:p></o:p> D. None of these<o:p></o:p>
616	Which of the following is not an acid radical	A. Cl- B. Br-

	Which of the following is not an acid radical	C. K+ D. I-
617	"There is a plenty of room at the bottom" This was stated by	A. Issac Newton B. Albert Einstein C. Richard Feynman D. Eric Drexler
618	The range of sound pressure for uncomfortable level is.	A. 80 - 90 dB B. 100 - 120 dB C. 130-140 dB D. All above
619	CFT can very well explain	A. Color B. Magnetic properties C. Spectra of transition metal D. All
620	Which of the following extract is used for wet tests of acid radicals.	A. Calcium carbonate extract B. Sodium iodide extract C. Sodium carbonate extract D. Ammonium carbonate extract
621	According to the Debye-Huckel theory of strong electrolytes, and ion moving in an atmosphere of oppositely charged ions experience a drag This effect is known as	A. Aaymmetric effect B. Electrophoretic effect C. Inter ionic effect D. Concentration effect
622	Which of the following contains isoprene unite.	A. Natural rubber B. Nylong -6,6 C. Polyethylene D. Decron
623	Which of the following is not an alkali metal.	A. Potassium B. Francium C. Sodium D. Strontium
624	The oxidation state of HCiO4	A. + 7 B. + 3 C. + 5 D. + 1
625	Electron gas theory is able to explain	A. Metallic lusture and optical properties B. Malleability and ductility C. High electrical and thermal conductivity D. All of the above
626	Which of the following statements regarding phenols is not correct.	A. Phenol ls are stronger acids than water and alcohols. B. Phenol are weaker acids than carboxylic acids C. Phenol are solubel in both aqueous NaOH and aqueous sodium hydrogen carbonate D. Phenoxides ions are more stable than the corresponding phenol
627	Which of the following element is usually determined by flame photometry.	A. Li B. Na C. K D. All above elements
628	Various compound corresponding to molecular formula C1H10 are.	A. Functional isomers B. Position isomers C. Chain isomers D. None of the abvoe
629	The ionization potential of K would be numerically equal to.	A. Electron affinity of Ar B. Electromagnetically of K C. Electron affinity of K+ D. ionization energy of Ca
630	The half life for a first order reactions 32 s, What was the original concertation if after 2.0 minutes, the reactant concentration is 0.062 M.	A. 0.84 M B. 0.069 M C. 0.091 M D. 0.075 M
631	The capacity of normal human eye to see the smallest object is _____ micro meter	A. 10000 B. 1000 C. 100 D. 10
632	The first ionization energy of Mg is lower than	A. Na B. Ca C. Al D. Be

633	Which of the following is an example of molecular solids.	A. MgO B. ZnO C. Ice D. Graphite
634	Electronegativity of oxygen is.	A. 2,5 B. 3,5 C. 2,4 D. 2.1
635	In hydrogen bonding a hydrogen atom is bonded to which of the highly electronegative atoms.	A. N B. O C. F D. N,O,F
636	Which of the following properties does not depend upon the numeb rof solute particles.	A. Elevation in B.P. B. Osmotic pressure C. Depression in F.P. D. Boiling point of the solvent
637	Pick out the incorrect statement for SO2	A. It turns filter paper moistened with acidified K2Cr2O7 B. It turns starch iodate paper blue C. It does not react with chlorine in presence of charcoal D. It decolourises acidified KMnO4 solution.
638	Nitrobenzen can be prepared from benzene by using a mixture of conc. HNO3 and conc. H2SO4 In the nitrating mixture. HNO3 acts as a.	A. Base B. Acid C. Oxidizing agent D. Catalyst
639	When a drop of detergent solution is added onto a clean towel. It spreads instead of existing as a droplet Which of the following statements explains this phenomenon.	A. Detergent acts as an emulsifying agent B. Detergent reduce surface tension of water C. Detergent reduces surface tension of water D. All of above
640	Which of the following methods is used in qualitativ eanalysis.	A. Physical method B. Chemical method C. Instrumental method D. All above
641	The orientation of a crystalline surface is confidently defined in terms of.	A. Lijima Indices B. Miller indices C. Clausen indices D. None
642	The reciprocal of the coefficient of viscosity in called.	A. Density B. Specific gravity C. Fluidity D. Conductance
643	A terpenoid which as an alcoholic group in the molecule is	A. Citral B. Camphor C. Menthol D. Carvone
644	Which of the following compounds would be most ionic to charcter.	A. PbCl4 B. PbCl2 C. SnCl4 D. SnCl2
645	Transition elements, in general exhibit the following properties, except one Name that property.	A. Variable oxidation state B. Natural radioactivity C. Tendency to form complexes D. Formation of alloys
646	Which of the following techniques is used to reduce the need for large volumes of organic solvents.	A. Solid phase extraction B. Gel permeation C. Electrophoresis D. TLC
647	Which of the following substance is most abundant of all components of atmospheric air.	A. O2 B. N2 C. CO2 D. A2
648	Which of the following salt is green in colour	A. Mn salt B. Cr salt C. Co salt D. Ba salt
		A. UGU B. GGC

649	One arm of each t-RNA terminates in the base sequence.	C. ACT D. CCA
650	XeF4 is obtained, when a mixture of Xenon and fluorine in the ratio is heated in a nickle vessel at 400 ^{o}C	A. 1 : 3 B. 5 :1 C. 1 :20 D. 1 :5
651	Which of the following sets of quantum number is possible for an electron in a 4f orbital.	A. n = 4,l = 3, m = 4, s = +1/2 B. n = 4, l= 4, m = +4 , s = +1/2 C. n = 4, l = 3, m = +1, s = -1/2 D. n = 4, l = 4, m = +1 , s = -1/2
652	Which of the following would decompose at lowest temperature.	A. MgCO3 B. SrCO3 C. BaCO3 D. CaCO3
653	Alums are generally used	A. In Dying and water proofing of fabric B. In arrest bleeding C. IN water purification D. All above
654	The formula of copper pyrite is.	A. CuFeS B. CuFeS2 C. Cu2FEs D. Cu Fe2S
655	Which of the following factors does not effect the rate of the reaction.	A. Pressure B. Temperature C. Concentration D. Catalyst E. All of the above
656	Which of the following is not true of ozone.	A. It is a strong electilizing agent B. It attacks organic compounds containing carbon carbon double bond C. Its molecular is linear and has two different O-O bond lengths D. It is more powerful oxidising agent at molecular oxygen
657	Boron and aluminum halides are electron deficient compounds in this respect. they act as.	A. Lewis acid B. Lewis base C. Oxidizing agent D. Reducing agent
658	Which of the following statement is not related with the advantages of TLC.	A. A variety of adsorbents can be used B. The thickness of adsorbent can be varied C. Fluorescence can be introduced D. Different detectors can be used
659	Combination of a -amino acid through which linkages results result in formation of protein	A. Ester linkage B. Glycosidic linkage C. Lactum linkage D. Peptide linkage
660	The electromagevitiy of the following elements increase in the order	A. F > Cl > O > S B. S > Cl > O > F C. F > O > N > C D. C > O > N > F
661	Pesticide residues appear in which of the following foods.	A. Milk B. Fruit C. Fish D. Vegetables E. All above
662	Both the elements are typical non metals.	A. B & Ai B. B & Si C. Al & Si D. Any of above
663	Which of the following pairs of fundamental particles are present in equal numbers in a neutral atom.	A. Proton and neutron B. Proton and positron C. Electron and proton D. Neutron and electron
664	A boy accidently splashes a few drops of conc. H2SO4 on his cotton shirt. A few minutes later, the splashed part blacken and holes appear.This is because the sulphuric acid.	A. Heats up the cotton so that it burns B. Dehydrates the cotton C. Causes cotton to react with oxygen of the air D. Removes the elements of water from cotton
665	The technique which involves the equivalence relation between the quality of electric current passed and quantity of chemical change taking place in the electrochemical cell	A. Voltametry B. Coulometry

	is called.	C. Polarography D. Potentiometry
666	Which of the following statements about anhydrous aluminium chloride is correct.	A. It exist as AlCl3 molecules B. It is not easily hydrolysed C. It sublimes at 100 ^oC under vacuum D. Boron does not form B³⁺ ions
667	The bond length is measured by	A. X-ray diffraction B. Neutron diffraction C. Microwave spectroscopy D. All of above
668	Mostly used solvents for ionic compounds.	A. Liquid ammonia B. Liquid SO2 C. Liquid HF D. All above
669	Al Cl3 acts as a strong Lewis acid, because it is.	A. A covalent compound B. Readily hydrolyzed C. Electron deficient D. An ionic compound
670	The hardest material found in nature is	A. Steel B. Topaz C. Diamond D. Quartz
671	One of the best fluorinating agent is	A. XeF2 B. XeF4 C. XeF6 D. None of above
672	Pick out the incorrect statement	A. The geometry around 'N' atom in trimethylamine is pyramidal B. The geometry around N atom in trisilylamine is planar C. The nitrogen atim is trimethylamine is sp2 hybridized whilst in trisilylamine it is sp2 hybridized D. Trisilylamine has donor properties whilst trimethylamine has no donor properties.
673	Used in producing intense light in cinematography	A. Xenon B. Krypton C. Radon D. Helium
674	Is an instate able colourles gas with a sticky sweet odor and is extremely toxic.	A. B2 H6 B. B4 H10 C. B3 H9 D. B6 H10
675	The energy gap between tag and eg sets in denoted by	A. A- B. 10 Dq C. Both A and B D. None of above
676	When HCl is titrated against NaOH , the pH at the equivalence point is.	A. zero B. > 7 C. < 7 D. 14
677	If the values of standrd deviations for the first and second method differ, then which of the following test helps one to know whether this difference is significant.	A. Student's test B. F-Test C. Chi square test D. Standard deviation
678	in the process of production of soap the soap can be salted out by adding	A. Concentrated sulphuric acid B. Concentrated potassium hydroxide solution C. Concentrated sodium chloride solution D. None of above
679	What type of inter molecular force present in nylon-66^{o}	A. Vander wall B. Hydrogen bond C. Dipole -dipole interactions D. Sulphide linkage
680	Which of the following species is not a basic radical.	A. Ag+ B. Cl- C. Ba2+ D. K+
		A. Labila

681	The types of coordinate compounds.	B. Inert C. Both A and B D. None of above
682	Which of the following solutions of sulphuric acid will exactly neutralize 25 mL. of 0.2 M NaOH	A. 12.5 mL of 0.1 M solution B. 24 mL OF 0.1 m Solution C. 50 mL of 0.1 M solution D. None of the above
683	Principal constituents of noble gases is	A. Argon B. Neon C. Xenon D. Helium
684	The oxidation Number of I in HIO4 is.	A. +6 B. + 7 C. + 3 D. + 14
685	The units of surface tension in SI system are	A. Joule m-1 B. Newton m-1 C. Erg cm-1 D. Dynes cm-2
686	Which of the following expression is correct.	A. C = n/RT B. C = RT/n C. RT = Cn D. Cn = 1/RT
687	The vapours attacks the eyes and mucous membrane of nose and throat	A. F B. Cl C. I D. Br
688	The _____ sphere is enclosed in brackets in formulas for complex species, and it includes the central metal ion plus the coordinated group	A. Ligand B. Donor C. Coordination D. Oxiation
689	Un-like s -block elements d-block elements form which compounds as well	A. ionic compounds B. Co valent compound C. Co ordinate compounds D. None of above
690	What is the minimum tensile strength of gray Cast Iron class 50	A. 25000 ibf /in² B. 50000 ibf/in² C. 100000 ibf/in² D. 900000 ib/in²
691	The width of a typical DNA molecule is _____nm	A. 1 B. 2 C. 5 D. 10
692	A major constituent of materials one whose amount in the materials is	A. 1% or more B. 0.1% C. 0.01% D. 0.001 %
693	Highly dangerous acid and produces severe wounds on the skin.	A. HClO B. HClO2 C. HClO3 D. HClO4
694	Which of the following compounds shows optical activity	A. Lactic acid B. Maltose C. Glucose D. All above
695	The structure of SO2 is	A. Linear B. Angular C. V-shaped D. Planner
696	Group IV A consist of elements.	A. 3 B. 4 C. 5 D. 6
697	CIF is	A. Chlorine monoflouride B. Fluourine C. Monochlorine fluoride D. Monofluorine chloride
698	Codon for amino acid glycinc is not represented by base pair	A. GCA B. GGC C. GGA D. GGU

699	Main constituent of all inorganic matter	A. Carbon B. Silicon C. Tin D. Lead
700	Sodium silicate is ued	A. In fire proofing of wood and textiles B. As a preservative of eggs C. As a furniture polish D. All above
701	In vinyl cyanide, the number of a bonds in	A. 2 B. 3 C. 1 D. 4
702	Form electron deficient compounds	A. B B. Al C. Both B and Al D. None of above
703	What is a method of casehardening involving diffusion in which the steel to be casehardened is machined, heat treated placed in an air tight box and heated to about 1000 ^{o}F	A. Annealing<o:p></o:p> B. Normalizing<o:p></o:p> C. Carburizing<o:p></o:p> D. Decomposition<o:p></o:p>
704	In B2H6 molecule	A. There exists a direct B-B a -bond B. All the atoms are in one plane C. All the B-H bonds are normal covalent bonds D. There exist two bonds between the boron atoms.
705	When Phosphate rock Ca3(PO4) 2 is converted to phosphorus.	A. One of the products of the reaction is water B. Sulphuric acid is added to generate insoluble calcium sulphate C. Hydrogen is used to reduce the phosphate to phosphorus D. Silica is added to form a calcium silicate slag
706	Proteins have characteristics	A. Melting poirn t B. Iso electric point C. Boiling point D. All of these
707	When two bodies have equality of temperature with a 3rd body they in turn have equality of temperature with each other. This is a statement of.	A. First law of thermodynamics B. Zeroth law of thermodynamics C. Second law of thermodynamics D. Third law of thermodynamics
708	The reagent which can be used to distinguish acetophenone from benzophenone is.	A. 2,4 -dinito phenyl hydrazine B. Li AlH4 C. Benedict reagetn D. I2 and Na2CO3
709	The IUPAC name of C2(CN)3 is	A. 2,3-dicvano butanedinitrile B. 2,3 -dicyano -2- butenedinitrile C. 1,1,2,2-tetrcyanoethane D. 1,1,2,2, tetracyanocthenc
710	What is the approximate chromium range of a ferritic stainless steel.	A. 12% to 18% B. 16% to 20% C. 20% to 24% D. 12% to 16%
711	Which of the following has the bighest value.	A. Transnational partition function B. Rotational partition function C. Vibrational partition function D. Electronic partition function
712	The smallest cluster of carbon atoms in Bucky balls known till today consists of _____ carbon atoms.	A. 75 B. 20 C. 60 D. 15
		A. Silver bromide

713	Black and white photographic film contain small grains of.	B. Silver cholride C. Silver iodide D. Any of above
714	Turpentine is obtained from._____	A. Oak tree B. Pine tree C. Birch tree D. Lemon tree
715	Which of the following statement is not true with respect to the role of matter undergoing decomposition.	A. Decomposed matter increase soil fertility B. They provide a texture which is favorable for plant growth C. Organic compounds for complexes with mineral nutrients which enhance uptake by plants. D. In high concentration the decomposition product may increase the photosynthesis
716	a-pinene hydrochloride on warming rearrangements to form bornyl chloride	A. Pinacol pinacolone B. hofmann C. Wagner Meerwein D. Wolf
717	The correct order of second ionization potential of carbon nitrogen, oxygen and fluorine is.	A. C > N > O > F B. O > F > N > C C. O > N > F > C D. F > O > N > C
718	Setting of plaster of Paris volves.	A. Oxidation with atmospheric oxygen B. Combination with atmosphere CO2 C. Dehydration D. Hydration to yield another hydrate
719	The change in the concentration of the reactant of product per units time is called.	A. Order of the reaction B. Melecularity of the reaction C. Rate of reaction D. None of the above
720	The order in O2+ is	A. 1.0 B. 1.5 C. 2.0 D. 2.5
721	Which type of the coal preferred for metallurgical coal.	A. Lignite B. Peat C. Bituminous coal D. None of these
722	A mixture of weak acid and its salt is.	A. Alkaline buffer B. Acidic buffer C. Neutral buffer D. All of above
723	Which of the following statements is not related with joule Thomson effect.	A. Joule Thomson is 0 isenthalpic in nature B. H2 and He show heating effect C. All gases show change in temperature D. The change in temperature depends on initial temperature and nature of the gas. E. Joule Thomson coefficient is defined as meo= (op/oT)H
724	Lead pencil contain	A. Lead B. Lead sulphide C. a mixture of lead and silica D. graphite
725	In the electronic structure of acetic acid,the total number of shared and unshaped pair of electrons are respectively.	A. 16 ,8 B. 8 ,4 C. 12 ,8 D. 8 ,12
726	According to SHAH concept the Lewis bases were classified on the basis os.	A. Charge ion size B. Polarization consideration C. Electron and co coordinating ability D. All of above
727	Buffer solution are used to.	A. Increase the pH B. Resist the pH C. Decrease the pH D. None of above
728	Which of the following is capable of shown g optical isomersm.	A. CH3COCOOH B. CH3CHOHCOOH C. Botha a and b D. All of th

		D. All of these
729	In the presence of dilute alkali monosaccharides undergo reversible isomerisation . The reaction known as.	A. Kiliani reaction B. Weermann rearrrangement C. Lobry do Bruyn Van Ekenstein rearrangement D. Mutarotation
730	During sintering densification is not due to	A. Atomic diffusion B. Surface diffusion C. Bulk diffusion D. Surface tenstion
731	The vapour pressure of a liquid	A. Always increase's with temperature B. Always decreases with temperature C. Is independent of temperature D. Increase up to the boiling point
732	Which of the following statements is incorrect.	A. The elements of group 18 are known as aerogens. B. Group 2 elements are all metals C. Metallic character increase on going down a group D. All the elements belonging to a particular period have same valence shell configuration.
733	The by -product of the process of saponification is.	A. Methanol B. Glycol C. Glycerol D. Absolute alcohol
734	The bond order gives the following valuable information.	A. Stability of the molecules of ions B. Bond dissociation energy and bond length C. Magnetic properties D. All of the above
735	Soft drinks and baby feeding bottles are generally made up	A. Polyeater B. Polyurethens C. Polyamide D. Polyetyrene
736	Hydrolith is the common name of	A. NaH B. CaH2 C. NaF D. CaF2
737	In the Aluminothermite process, aluminium acts as.	A. An oxidizing agent B. A reducing agent C. A flux D. A Solder
738	For assciated liquids, the value of d/M n x 10^8 should be (where d is the density, M is the molar mass and n is the coefficient of viscosity)	A. Zero B. Infinte C. Higher than 70 D. Less than 70
739	Which one of the following is not formed when an electric discharge passes through helium.	A. HeH+ B. HeH2+ C. He2+ D. He2-
740	Equivalent conductance is expressed in the units.	A. S cm-1 eq -1 B. S cm eq-1 C. S cm2 eq-1 D. S cm2 eq
741	Electron affinities of halogens are in the order.	A. F > Cl > Br > I B. Cl > F > Br > I C. Cl > Br > I > F D. Cl > Br > F > I
742	Amino acids have	A. Acidic group B. Basic group C. Both of these D. None of these
743	The expected specific waste fo petroleum industry is.	A. Asphalt and tars B. Paper C. Cloth D. Fibre
744	The process of passing of a precipitate into colloidal solution, on adding an electrolyte is called.	A. Dialysis B. Peptization C. Electrophoresis D. Electromsmosis
		A. Vinyl cyanide

745	Which of the following is used to make non-stick material.	A. Vinyl cyanide B. Tetrafluoroethene C. Vinyl chloride D. Styrene
746	Strong field ligands such as CN	A. Usually produce high spin complexes and small crystal field splitting B. Usually produce low spin complexes and small crystal field splitting C. Usually produce low spin complexes and high crystal field splitting D. Cannot form low spin complexes
747	Which substance has the greatest lattice energy.	A. CuBr B. MgO C. KI D. NaF
748	Chlorination of toluene in the presence of light and heat followed by treatment with aqueous NaOH gives.	A. o - creaol B. p - creaol C. 2,4 -dihydroxy toluene D. Benzoic acid
749	Which one of the following statement is incorrect in relation to ionization enthalpy.	A. Ionization enthalpy increase for each successive electron B. The greatest increase in ionization enthaly is experienced on removal of electron from core noble gas configuration C. End of the valence electron is marked by a big jump in ionization enthalpy D. Removal of electron from orbitals bearing lower value to easier than from orbital having highest n value.
750	The mole of photon is known as	A. Quantum B. Eienstein C. Energy Packet D. None of the above
751	In second group of inorganic qualitative analysis, the S^{2-} ions does not form precipitate with which of the following ions.	A. ^{Hg2+} B. Cu²⁺ C. Ai³⁺ D. Cd³⁺
752	Total pressure exerted by a mistress of two or more than two gases in a definite volume as any given temperature is equal to the sum of partial pressures which each gas would exert, if it occupied the same volume alone, at the same temperature This is a statement.	A. Boyle's law B. Charle's law C. Graham's law D. Dalton's law
753	Which of the following is not alloy of aluminium.	A. Aluminium bronze B. Magnalum C. Duralumin D. Stellite
754	A 2M solution of H2SO4 would have how many moles of H+ ion in one liter	A. 1.0 B. 2.0 C. 4.0 D. 5.0
755	The normality of 2.3 M H2SO4 solution is.	A. 0.46 N B. 0.23 N C. 2.3 N D. 4.6 N
756	Which of the following type of lattice has maximum numb of atoms per unit cell.	A. Simple cubic B. Body centred cubic C. Face centred cubic D. All of them
757	The freezing point of a solvent	A. Will increase on adding a solute B. Will decrease on adding a solute C. Will note change on adding solute D. None of the above
758	In confining and growing nano roade CNTs will act as.	A. Template B. Support C. Source of oxidant D. Sieve
759	Which of the following is not adsorptive separation process.	A. Parex B. Olex C. Penex D. None of these
760	Which halide of cesium will be highly ionic in nature.	A. K+ B. Ag+ C. Rb+ D. Ca+

		D. Cu^+
761	Which of the following is NOT true 7 ceramic materials are.	A. Hard, have high densities high compressive strength and very good thermal resistance and strength at higher temperature Silicon B. Hard, have low densities high compressive strength and very good thermal resistance and strength at higher temperature. C. Hard, have low densities low compressive strength and very good thermal resistance and strength at higher temperature. D. Hard, have low densities' high compressive strength and very good thermal resistance and strength at higher temperature.
762	The maximum absorption in [Ti(OH)2)6 3+ take place at wavelength of.	A. 4000 A^o B. 5000 A^o C. 6000 A^o D. 10000 A^o
763	The splitting of H2O can be carried out through	A. Photolysis B. Electrolysis C. Dialysis D. Hydrogenation
764	Berllium has diagonal relationship with	A. Li B. Al C. B D. Na
765	Each of the following compound react with Grignard's reagent to form alkane exxcept.	A. Ethanal B. Ethanoic acid C. Ethanol D. Ethync
766	The prefiex 'nano' comes from a	A. French word meaning billion B. Greek word meaning dwarf C. Latin word meaning invisible D. Spanish word meaning particle
767	Red brass contain about how many percent of zinc.	A. 20 % B. 15 % C. 30 % D. 25 %
768	Argillaceous material does not include.	A. Vlay<o:p></o:p> B. Marine shells<o:p></o:p> C. Slate<o:p></o:p> D. Blast furnace slag<o:p></o:p>
769	Iron which contains up to 1% carbon is called.	A. Steel B. Cast iron C. Wrought iron D. Pig iron
770	The penultimate shell of carbon contains electrons.	A. s2 B. s2p6 C. s2p6d10 D. s2p6d8
771	The smog is essentially caused by the presence of.	A. O3 and N2 B. O2 and N2 C. Oxides of sulphur and nitrogen D. O2 and O3
772	The structure of SO2	A. Linear B. Angular C. V-shaped D. Planner
773	Strongest inter molecular hydrogen bond is formed in	A. H2O B. NH3 C. HF D. H2S
		A. Dehydrogenations high endothermic

774	Which of the following statement is not true is case of catalytic reforming.	B. Dehydrogenation is exothermic C. Hydrodealkylation reactions are endothermic D. None of these
775	Which of the following electrolytes will be most effective in the coagulation of arsenic sulphide sol.	A. NaNO3 B. Al PO4 C. MgSO4 D. K4[Fe(CN)6]
776	In an adiabatic system, if work in done, the temperature must.	A. Increase B. Decrease C. Remain the same D. Increase than decrease
777	The depolarizer used in dry cell batteries in.	A. NH4Cl B. MnO2 C. KOH D. Na2PO4
778	The one in which the acceptor atom is of low positive charge, Large size and has several outer electrons which can be easily excited is a.	A. Soft base B. Hard Base C. Soft acid D. Hard acid
779	Which of the following method of analysis is based on diffraction of radiation.	A. Mass spectrometry B. Polarography C. Potentiometry D. Raman scattering
780	Which number of halogen family does not show positive oxidation state.	A. Fluorine B. Chlorine C. Bromine D. Iodine
781	The word 'ceramic' meant for.	A. Soft material B. Hard material C. Burnt material D. Dry material
782	The unit of nucleic acid having base sugar combination is called.	A. Nucliec acid B. Nucleoside C. Nucleotide D. None of these
783	For covalent bond to form between two atoms A and B	A. Transference of electrons must take place from A to B B. A pair of electrons of A is shared by both A and B C. A and B contribute equal no. of electrons for mutual sharing by A and B D. One of the atom A or B must already have octet of electrons.
784	The wear-resistance of the material is attributed to its ability on _____ that is the hardness is increased greatly when the steel is cold worked.	A. Cold harden<o:p></o:p> B. Stress harden<o:p></o:p> C. Strain harden<o:p></o:p> D. Cool temperature<o:p></o:p>
785	The state of hybridization of carbon in CO2 is	A. sp2 B. sp C. sp3 D. dsp2
786	The plate height is the length of the column divide by	A. Length of the column B. Width of the column C. Number of theoretical plates D. Number of components of the mixture.
787	Which of the following pollutants does not a leave a residue.	A. Air pollutant B. Chemical pollutant C. Soil pollutant D. Noise pollutant
788	Metal are generally elements	A. Electronegative B. Electropositive C. Neutral

		D. None of the above
789	Oxalic acid when heated withe conc. H2SO4 it gives out.	A. H2O and CO2 B. CO and CO2 C. CO2 and H2S D. Oxalic sulphate
790	The statement that heat cannot flow spontaneously from a colder to a hotter body is the result of.	A. The first law of thermodynamics B. The second law of thermodynamics C. The third law of thermodynamics D. Henry's law
791	Which of the following solution has highest normality.	A. 1 N H2PO4 B. 0.5 N H2SO4 C. 6 g NaOH per 100 cm3 D. 4 g NaOH PER 1000 cm3
792	The simplest formula of a compound containing 50% of element X	A. XY2 B. XY C. X2Y D. None of the abvoe
793	Which of the following materials is not suitable as adsorbent for chromatography.	A. Silica gel B. Activated charocal C. Alumina D. Calciu7m chloride
794	Nitrogen (N2)) is relatively unreactive because.	A. Its electronegativity is high B. Its dissociation energy is large C. Its atomic radius is small D. It is the first element of group 15
795	The chrome molybdenum steels contain how many percent of molybdenum	A. 0.10 B. 0.20 C. 0.30 D. 0.40
796	The unit of specific conductance will be	A. S cm-1 B. Ohm cm C. Ohm cm-1 D. Mho cm
797	PCl5 is an example of hybridization	A. d sp³ B. d2 sp2 C. sp2 D. sp3
798	Which of the following is not a component of a gas chromatography system.	A. Carrier gas B. Capillary column C. Packed column D. Cathode lamp
799	Carbohydrates are characterized by the presence of.	A. Hydroxyl group B. Carbony group C. Asymmetric carbon D. All of these
800	Which of the following dye is used an antiseptic .	A. Methyl orange B. Mercurochrome C. Alizarin D. Bismarck brown
801	A group that causes deepning of the colour is known as	A. Bathchromic B. Hypsochromic C. Hypochromic D. Hyperchromic
802	Solid phase micro extraction is a solvent less extraction technique This technique is used for preparation of samples for analysis by which of the following technique.	A. HPLC B. GC C. TLC D. Electrophoreals
803	In which polymerization branching of chain cannot be possible.	A. Free radical B. Cationic C. Anionic D. Anionic and Zieglar Natta
804	What refers to the tin mill steel, without a coating.	A. White plate B. Black plate C. Tin steel free D. Dichromate tin
805	Petroleum is formed from	A. Domestic animal B. Organisms in sea<o:p></o:p> C. <p class="MsoNormal" style="margin-

		bottom:0in;margin-bottom:.0001pt;line-height: normal">Wild animals<o:p></o:p></p> D. All above
806	Most commercial glasses consist of	A. Lime B. Soda C. Silica D. All
807	The formula of Borax is.	A. Na2B4O7 6H2O B. Na2B4O7 8H2O C. Na2B4O7 10H2O D. Na2B4O7 12H2O
808	Which of the following is renewable resources of energy.	A. Hydropower B. Wind power C. Solar power D. All above
809	Coagulation of protein on treatment with heavy motal salts or heating is called.	A. Decolorisation B. Denaturation C. ^{Sedimentation process} D. Reversible precipitation
810	Which of the following techniques is used to separate a mixture of cations.	A. GC B. FPLC C. Ion exchange chromatography D. Size exchange chromatography
811	Which of the following is not a component of HPLC system.	A. Pumps B. Columns C. Particle collector D. Injection system.
812	It is possible to distinguish between optical isomers.	A. Using chemical tests B. By mass spectrometry C. By IR spectroscopy D. By polarimetry
813	Cement containing higher percentage of gypsum than required.	A. Sets slowly B. Sets repidly C. _{Does not set at all} D. Has no effect
814	The electron gain enthalpy of chlorine is -349 Kj mol-1 ionization energy of Cl would be.	A. -349 kJ mol -1 B. 349 kJ mol-1 C. -698 kJ mol-1 D. 698 kJ mol-1
815	Which element out of the following can ehibit a maximum con valency of seven.	A. Chlorine B. Sulphur C. Fluorine D. both Cl and F
816	Which of the following statements is not related with principal requisites of water for industrial purposes.	A. It should be pure and cool B. It should not contain iron C. It contains less quantity of line D. It must be soft
817	Which of the following substance is colloidal in nature.	A. Clay B. Al2O3 C. Fe2O3 D. All above
818	Which of the following substance acts as gaseous pollutant.	A. NO B. NO2 C. CO D. SO2 E. All above
819	Which of the following solution would exhibit abnormal colligative proportions.	A. 0.1 M NaCl B. 0.1 M urea C. 0.1 M sucrose D. 0.1 M glucose
820	The term 'brass' is very commonly used to designate any alloy primarily of.	A. Copper and zinc B. Aluminum and iron C. Copper and aluminum D. Zinc and nickel
821	The movement of an electric charge produce a magnetic field is known from the	A. Elementary Physics B. Elementary Chemistry C. Both A and B D. None of above
		A. The electrons are shared between atoms.

822	Which of the following statements regarding covalent bond is false.	B. The bond in non -directional C. The strength of the bond depend upon the extent of overlapping D. The bond formed may be polar or non-polar
823	During the preparation of ethane by Kolbe's electrolytic method using inert electrodes the pH of the electrolyte.	A. Increases progressively as the reaction proceeds B. Decreases progressively as the reaction proceeds C. Remains constant throughout the reaction D. May decrease if the concentration of the electrolyte is not very high
824	What is the ratio of the maximum load in a tension test to the original cross sectional area of the test bar.	A. Tensile strength<o:p></o:p> B. Yield strength<o:p></o:p> C. Shear strength<o:p></o:p> D. Torsion<o:p></o:p>
825	IUPAC name of HCONH2 is.	A. Methanamide B. Methanoylamine C. Ammoethanal D. Formanide
826	Ozone filters out radiation below.	A. o 1000 A B. o 2000 A C. o 3000 A D. o 4000 A
827	Which of the flowing operator combination would yield eight value equation	A. d/x (sin x) B. d/dx (cos x) C. d /dx (sin 4x) D. d /dx (cos 4x) E. d/dx (e^x)
828	Stainless steel contains	A. Fe+Cr+Ni B. Fe+Ni+Cu C. Fe + Cr+ Cu D. Cu + C + Ni
829	Which of the following statement is not true regarding Open Hearth process.	A. No iron is lost B. The process is economical and simple C. Steel obtained is of high quality D. Scrap iron cannot be used in this process.
830	The interactions in HF are.	A. dipole dipole interasctions B. Hydrogen bonds C. dipole -dipole and dispersion forces D. Hydrogen bond and diapersin forces
831	Hemimorphite is an example of.	A. Orthosilicate B. Pyrosilicate C. Cyclic silicate D. Meta silicate
832	The rate at which a substance reacts depends on its.	A. Molecular mass B. Active mass C. Equivalent mass D. Molar mass
833	Ozone layer of upper atmosphere is being destroyed by	A. chlorofluorocarbons B. SO2 C. Photochemical oxidants O2 and CO2 D. Smog
834	The range of sound pressure which is painful is as	A. 130-140 dB B. 100 - 120 dB C. 90 - 80 dB D. All above
835	In the reaction RCO2Na + Na OH (CaO) _____ RH, we eliminate carboxylate group as.	A. CO2 B. Na2CO3 C. -CO D. CaCO3
		A. Gravimetric analysis

836	Thermogravimetic analysis has application is which of the following fields	A. Gravimetric analysis B. Discovery of new methods ofseparation C. Determination of purity and thermal stability D. All above
837	Cryolite is used in the electrolytic extraction of aluminium to.	A. Obtain more aluminium B. Reduce alumina C. Protective electrodes D. Dissolve bauxite and increase the electrical conductivity
838	Zeigler Natta catalyst is.	A. Pt/PtO B. TiCl4/Al(C2H5)3 C. Pt/Rh D. Pt
839	The fraction of the total current carried to an ion is called itss.	A. Ionic mobility B. Transport number C. Limiting ionic conductance D. None of these
840	What cast iron has modular or spheroidal graphite?	A. Ductile iron<o:p></o:p> B. Gray iron<o:p></o:p> C. White iron<o:p></o:p> D. Raw iron<o:p></o:p>
841	Hydrolysis of protein gives	A. a -amino acid only B. b-amino acids only C. gama amino acid only D. A mixture of all of these
842	Which of these historical works of art contain nanotechnology.	A. Lycurgus cup B. Medieval stained glass windows in churches C. Damascus steel swords D. All of the above
843	Which of the following reacts with hemoglobin of blood and produce toxic effect.	A. Carbon dioxide B. Carbon monoxide C. Oxygen D. Carbon suboxide
844	The change of chemical potential of any component with temperature an constant P and composition, is euqal to.	A. Partial molar enthalpy of that component B. Partial molar volume C. Partial molar free energy D. Negative of the partial molar entropy
845	The size of nanoparticles is between _____ nm	A. 100 to 1000 B. 1 to 100 C. 0.1 to 10 D. 0.01 to 1
846	LiAlH4 is most useful reducing agent It reduce to alcohol	A. Aldehydes B. Ketone C. Carboxylic acid D. Any of above
847	Which of the following analytical technique is used for the separation of an interfering substance or analyte from the mixture.	A. Precipitation B. Distillation C. Electrode position D. All above these
848	Which of the following steps in involved in the metallurgy of aluminium.	A. Purification of bauxite B. Electrolytic reduction of alumina C. Refining of aluminum D. All above
849	A salt solution is treated with chloroform drops.Then it is shaken with chlorine water, chloroform layer become violet solution contains.	A. NO2 ion B. NO3 ion C. Br ion D. I- ion
		A. Mean value

850	The term accuracy refers to how near the observed value is to.	B. Low value C. True value D. Standard value E. Both C and D
851	Electronegativity of Oygen is.	A. 2.5 B. 3.5 C. 2.4 D. 2.1
852	In Pakistan the total production of glass is over _____ tons per year.	A. 800 B. 8000 C. 80,000 D. None of these
853	Which of the following methods gives the number average molecular weight of a polymer.	A. Light scattering method B. Osmotic method C. Sedimentation equilibrium method D. Viscosity method
854	Which of the following is diamagnetic	A. O2 B. O2+ C. O2- D. O2 ²⁻
855	[Ti(OH2)6]3+ gives colour	A. Green B. Red C. Purple D. Blue
856	Attention should be focused on qualitative changes in particle properties as a function of.	A. Particle numebrs B. Particle mass C. Particle size D. Particle density
857	Temporary hard water is softened on industrial scale by adding.	A. Mg(OH)2 B. Ca(OH)2 C. KOH D. NaOH
858	When propyne is treated with equeous H2SO4 in the presence of HgSO4 the functional isomer of the major product obtained in.	A. Propanal B. Acetone C. Propane 2 -nl D. Propanol
859	The following oxo acids have been arranged in the order decreasir acid strength identify the correct order.	A. III > IV > II > I B. III > II > I >IV C. I > II >III > IV D. IV > III > II > I
860	Select the major product obtained from the addition HBr to 1-methyl cyclohexene.	A. 1- bromo -2- methyl cyclohexane B. 6- bromo-1- methyl cyclohex - 1- ene C. 3- bromo-1- methyl cyclohex -1- ene D. 1- bromo-1- methyl cyclohexane
861	What % if nickel is present in the major ore Pentlandite.	A. 22% B. 18% C. 14% D. 10%
862	The value of an Einstein	A. Is independent of wavelength B. Decrease with increase in wavelength C. Increase with increase in wavelength D. Depends on the temperature of the absorbing system
863	What impurity in steel can cause ted shortness which means the steel becomes unworkable at high temperature.	A. Sulphur B. Silicon C. Magnesium D. Aluminium
864	When 0.01 moles of NaOH are added to a buffer solution, its pH changes from 4.745 to 4.832 WHAT IS ITS.	A. 0.115 B. 0.900 C. 0.015 D. 0.215
865	An emulsifier is an agents which	A. Stabilizes an emulsion B. Homgeneises and emulsion C. Causes coagulation of an emulsion D. Helps in the formation of an emulsion
866	Which cast iron is hard and wear resistant.	A. Grey iron B. White iron C. Melleable iron D. None of these
		A. These travel in a straight lines. B. These are deflected by magnetic and

867	The rays emitted by the cathode in a gas discharge tube under low pressure and high voltage of electricity are called cathode rays. Which of the following properties are not related to cathode rays.	electric field. C. Minerals Fluoreace with a characteristic color when placed in a beam of cathode rays. D. These are dependent of the material used for the electrode.
868	Ammonium nitrate is sold as a mixture with	A. Soda Ash B. Lime stone C. Zinc D. None of above
869	Noble gases are used in discharge tubes to give different colours. Raddish orange glow is due to.	A. Ar B. Ne C. Xe D. Kr
870	Which of the following pollutant result from roasting and heating processes.	A. Dust B. Smoke C. Metal fumes D. All above
871	The important condition for the formation of chemical bond is that.	A. Their electron clouds should not diffuse B. Both atoms should have high electron affinites. C. Both atoms should have same electronegativities D. The process should be accompanied by the lowering in potential energy.
872	When a concentrated solute of an electrolyte is diluted.	A. Its specific conductance increases B. Its equivalent conductance decreases C. The specific conductance decreases and equivalent conductance increases D. Both specific and equivalent conductance increase
873	The percentage of nitrogen in Urea is _____%	A. 46 B. 37 C. 82 D. 50
874	The relative lowering of vapour pressure of a solution on the addition of non -volatile solute.	A. Is equal to the mole fraction of solute B. Is equal to the sum of the mole fraction of the solute and solvent C. Depends upon the nature of the solute D. Depends upon the mole fraction of the solvent
875	Molecule is a diatomic	A. Nitrogen B. Phosphorous C. Arsenic D. Antimony
876	A combination of atomic orbitals produces a large number of closely special energy states brown as.	A. Packet of energy B. Band of energy C. Botha a and b D. None of the above
877	Considering the elements F, Cl ,O and N, the correct order of their chemical reactivity in terms of oxidizing property is.	A. F > Cl > N B. F > O > Cl > N C. Cl > F > O > N D. O > F > N > Cl
878	Which of the following compounds has highest dipole moment.	A. Dichloromethane B. Chloroform C. Chloromathane D. All above
879	In extraction of iron, the furnace charge consists of iron ore, coke and limestone. The function of limestone is to act as.	A. An oxidation agent B. A reducing agent C. Flux D. Slag
880	A steel cannot qualify for stainless prefix until it has at least how many percent of chromium.	A. 10 % B. 20 % C. 5 % D. 30 %
		A. Plastics and ceramics are immune to many forms of corrosion because they are not good conduction of electricity.<o:p></o:p> B. <p class="MsoNormal" style="margin-bottom:0in;margin-bottom:.0001pt;line-

881	Indicate the false statement about corrosion.	height: normal">The corroded member in a corrosion cell is the cathode <o:p></o:p> C. Passivity is a prerequisite for the corrosion protect on many metals<o:p></o:p> D. None of these
882	When alkyl iodidies are decomposed by light then the product obtained is.	A. R - R B. R - H C. RCH2I D. RCHI2
883	Which of the following interaction is the strong.	A. Dipole -dipole B. Ion induced dipole C. Ion -dipole D. Dipole induced diple
884	Which of the following is not an androgen i.e. male sex hormones.	A. Androsterone B. Testosterone C. Oestrone D. All of these are make hormone
885	In their ionic compounds halogens exhinit the oxidation states of.	A. -1 B. -2 C. -3 D. -4
886	Which of the following is not a physical test.	A. Colour test B. Flame test C. Beed test D. Wet test
887	The process requirieng the absorption of energy of.	A. F = F B. Cl = Cl C. H = H D. O = O
888	The use of acids to remove oxides and acids on hot worked steels is known as	A. Tempering B. Picking C. Machining D. Sizing
889	Which of the following statement is false regarding lyphilic sols.	A. The colloidal particles show a linking for the dispersion medium B. These are generally easy to prepare C. These are more stable than lyophobic sols D. The stability of the sols is mainly due to the electrical double layer
890	Water that easily forms a lather of films and frotha when agitated with a soap solution called.	A. Hard water B. Heavy water C. Soft water D. Washing water
891	Relative order of acidity of oxy acid	A. HClO > HClO2 > HClO3 > HClO4 B. HClO4 > HClO3 >HClO2 > HClO C. HClO3 > HClO2 > HClO > HClO4 D. HClO2 > HClO4 > HClO3 > HClO
892	The expected specific waste of food industry is.	A. Meats B. Nuts C. Fats or Oils D. All above
893	The molar mass of an organic acids is determined by	A. Depression of freezing point B. Elevation of boiling point C. Volumetric method D. Victor Myer's method
894	Which of the following elements is most electropositive.	A. C B. N C. O D. Be
895	Co enzyme cna be separated from enzyme by	A. Precipitation B. Dalysis C. Hydrolysis D. Distillation
896	In which of the following compounds does hydrogen bonding occur.	A. CCl4 B. NaH

	In which of the following compounds does hydrogen bonding occur?	C. HI D. NH3
897	Which idea of envisioned the construction of nano robots	A. Building nano materials atom by atom B. Destruction of macromolecules to nano ones C. Bothe of the above D. None of the above
898	Which of the following is a triphenylmethane dye.	A. Auramine G B. Crystal violes C. Fluorescein D. Fast green O
899	The principle former of almost all glasses is	A. (SIO2)n B. (SiO3)n C. (SiO2) D. None of these
900	According to CFT the metal ligand bond is considered to be ionic to presentage.	A. 100% B. 90% C. 50% D. 70%
901	The first ionization energies of the elements of the first transition series. (Ti _____ Cu)	A. Increases as the atomic number increases B. decreases as the atomic number increases C. Do not show any change as the addition of electrons takes place in the inner (n-1) d -orbitals. D. Increases from Ti to Mn and then decreases from Mn to Cu
902	Pauling has suggested that the calculate of energy can be improved by considering.	A. Screening effect B. Polarization effect C. Both A and B D. None of abvoe
903	Which of the following hydroxides is most stable.	A. Mg (OH)2 B. Ca(OH)2 C. Sr (OH)2 D. Ba (OH)2
904	Which among the following is secondary pollutant.	A. CO B. CO2 C. PAN D. Aerosol
905	What ASTM test for compression is designated for plastics.	A. D 638 B. D 695 C. D 790 D. D 732
906	The alkali metal that react with nitrogen directly to form nitrides.	A. Na B. K C. Rb D. Li
907	The number of vibrational degree of freedom for CO2 is	A. 2 B. 3 C. 4 D. 5
908	The juice is allowed to boil at lower temperatures to protect the sugar from	A. Hardening B. Solubility in water C. Caramelization D. Dwatering
909	Opticla tweezers	A. Are used to remove facial hair with miniaturized laser beams B. use light to manipulate particles as small as single atom C. Are a nanotechnology bases tool for stamp collectors D. Don't exist
910	The correct order of electron affinities is.	A. C > Si, > Na > Ar B. Si > Cl > Na > Ar C. C > Na > Si > Ar D. C > Si > Ar > Na
911	Which of the following statement is not correct with respect to radioactive pollutnats.	A. Carcinoma and breast cancer B. Leukemia C. Increases biological immune system D. Somatic and generic disorder
912	Which of the following a -amino acid is not capable of exhibiting optical isomerism.	A. Glycine B. Leucine C. Arginine

		C. Arginine D. Alanine
913	Alnico is an alloy containing how many percent nickel.	A. 10% B. 14% C. 18% D. 22%
914	The solution of the transition metal complexes having one or more unpaired electrons in the d-orbital are.	A. Coloured B. Colourless C. White D. None of above
915	Fullerene or bucky ball is made up of _____ carbon atoms.	A. 100 B. 20 C. 75 D. 60
916	Which of the following is not an organic precipitating agent.	A. Diemethglyoxime B. Cuperon C. Oxime D. Acetate
917	In order to give strength and elasticity natural rubber is heated with.	A. Sulphur B. Oxygen C. Nitrogen D. Chlorine
918	The oxidation state of Pt in Xe+ [Pt F6] is	A. +4 B. +5 C. +6 D. None of these
919	The process of determining amounts of each of the components in a sample of matter is termed as.	A. Gravimetric analysis B. Coulometric analysis C. Quantitative analysis D. Qualitative analysis
920	The suffix '-ene' in the name of fullerene shows the presence of ___ in the molecule.	A. One triple bond B. One double bond C. Two single bonds D. Two triple bonds
921	The reason why phenylamine is a much weaker base tahn ammonia when each is in aqueous solution to that.	A. Teh ion pair of electron on two nitrogen atom of phenylamine is delocalised over the benzene ring. B. The phenylamine molecule is too large to capture hydrogen ion easily C. Phenylamine is much less soluble is water than is ammonia D. The benzene ring has a tendency to increase the acidity of its substituents.
922	By applying an external force the ionic solid can be easily broken to powder form so the ionic solid are highly	A. Hard B. Brittle C. Tough D. Soft
923	DTA is of great importance in which of the following field	A. Ceramic B. Metallurgy C. Mineralogy D. All
924	What is the advantage of quench hardening?	A. Imporoved strength<o:p></o:p> B. Hardness<o:p></o:p> C. Wear characteristics<o:p></o:p> D. All of the choice
925	Stainless steel contains.	A. Fe + Cr+ Ni B. Fe + Ni + Cu C. Fe + Cr+ Cu D. Cu + C + Ni
926	_____ is used for fruits, vegetables and tobacco	A. Potassium Chloride B. Potassium Sulphate C. Potassium nitraite D. All above
		A. pH at which it does not have any charge

927	The isoelectric point of a protein or amino acid to.	B. pH at which it does not have not charge and does not migrate in electric field C. pH at which the concentration of cation is greater than amino D. pH at which the concentration of anion is greater than cation
928	What exactly is quantum dot	A. A semiconductor nanostructure that confines the motion of conduction band electrons, valence band holes or excitation in all three spatial directions B. The sharpest possible tip of an atomic force microscope C. A fictional term used in science fiction for the endpoints of wormholes D. Unexplained spots that appear electron microscopy images of nanostructures smaller than 1 nanometer
929	The unit of sound pressure level is	A. Pascal B. Decibel C. Newton D. Ampere
930	Which of the following compounds is must acdic.	A. H2O B. H2S C. H2Se D. H2Te
931	Presence of nitrogen in organic compound to tested as.	A. Nitrogen gas B. NH3 C. NO D. Amide
932	What element is added to copper to make it extremely hard.	A. Aluminum B. Zinc C. Lead D. Tin
933	_____ is used for Annealing	A. Klin B. Batch C. Converter D. Oven
934	The correct order of ionic radii for the following ions is.	A. S²- < P³⁻ < Cl- < K+ B. Cl- > S2+ > P3- > k+ C. K+ > Cr > S2+ > P3- D. P3+ > S2 > Cl - > K+
935	What of the following is not a Lewis base.	A. CN- B. AlCl3 C. NH3 D. ROH
936	An electron has types of motion	A. Spin motion B. Orbital motion C. Both A and B D. None of above
937	What refers to the application of any process whereby the surface of steel is altered so that it will become hard.	A. Caburizing<o:p></o:p> B. Case harden zing<o:p></o:p> C. Ammealing<o:p></o:p> D. Surface hardening<o:p></o:p>
938	Bromine is soluble in	A. Alcohol B. Water C. Chloroform D. All above
		A. mol-2 s-1

939	The rate constant for 3rd order reaction has the dimentions of.	B. l2 mol-2 s-1 C. mol l -1 s-1 D. l-1 mol -1 s-1
940	The green color of water in a lake is due to	A. Excessive growth of sea weeds B. Algae C. Pollution D. Grass
941	The absorbance is directly proportional to the path length in the flame and to the concentration of atomic vapor in flame is a statement of.	A. Lambert's law B. Beer's law C. Honery's law D. Starke law
942	The rusting of iron is catalyzed by which of the following.	A. Fe B. H+ C. O2 D. Zn
943	The pH of the 0.0032 M H2SO4 is.	A. 3.2 B. 4.0 C. 2.198 D. 1.0
944	The unequal sharing of bonded pair of electrons between the two atoms in a molecule causes.	A. Dipole B. Radical formation C. Decomposition of found D. Covalent found
945	HClO evolves Cl2 and O2 when dissolve	A. Ca B. Ni C. Cu D. Any of above
946	What is the activation energy of a reaction whose rate constant increases by a factor of 100 upon increasing the temperature from 300 K to 360 K.	A. 27 B. 35 C. 42 D. 69
947	Which of the following statements is not correct with respect to second law of thermodynamics.	A. It helps in know the position of chemical equilibrium B. It helps to know the position of chemical equilibrium C. It determines the conversion of heat into work D. It is based on Nerst heat theorem
948	Which of the following biogeochemical cycle is not component of ecosystem.	A. Carbon cycle B. Potassium cycle C. Oxygen cycle D. Nitrogen cycle
949	When the colourless liquid chlorobenzene is shaken with bromine water, the chlorobenzen becomes a yellow-orange colour. Which of the following is the best interpretation of this.	A. An addition compound of chlorobenzene and bromine has been formed. B. The chlorine atom has been replaced by a bromine atom C. The bromine is mor esoluble in chlorobenzene than in water D. A hydrogen atom has been replaced by a bromine atom
950	Ground water is threatened with pollution from which of the following source.	A. Domestic wastes B. Industrial wastes C. Agricultural wastes D. All above
951	The substance added to the soil in very small amounts are called.	A. Macronutrients<o:p></o:p> B. Micronutrients<o:p></o:p> C. Fertilizers<o:p></o:p> D. None of these<o:p></o:p>
952	Among all halogens no oxyacid of the following is known	A. F B. Cl C. Br D. I
		A. Ionization enthalpy

953	A property which gradually increases on moving down group in the periodic table is	A. Ionization enthalpy B. Elemtronegativity C. Electron affinity D. atomic size
954	Bryllium salts on hydrolysis give.	A. Basic solutions B. Acidic solutions C. Neutral solutions D. Amphoteric solutions.
955	The most reactive alkali metal among the following is	A. Li B. Na C. Cs D. Rb
956	The maximum number of electrons in first energy levels are.	A. 1 B. 2 C. 8 D. 10
957	A process in which no heat enters leaves the system is called.	A. Isochoric B. Isobaric C. Adiabatic D. Reveraible
958	The blue colour of CuSO4 disappears on adding Zn granules to it . it is because of .	A. Oxidation of Cu atom B. Oxidation of Zn²⁺ C. Oxidation Cu²⁺ D. Oxidation of Zn²⁺
959	Which of the following equations represent linear free energy relationship.	A. Hammett equation B. Taft equation C. Helmholtz equation D. Differential equation
960	Which of the halogens has lowest bond energy.	A. Cl2 B. Br2 C. F2 D. I2
961	What is the lowest temperature diffusion hardening process and does not require a quench	A. Carburizing<o:p></o:p> B. Tempering<o:p></o:p> C. Nitriding<o:p></o:p> D. Melting<o:p></o:p>
962	Photochemical smog is caused primarily by	A. CO B. CO2 C. NO2 D. O3
963	Acid rain effects	A. Human being B. Crops C. Aquatic life D. All above
964	What is a process for making glass reinforced shapes that can be general by polling resin impregnated glass stands though a die.	A. Continuous pultrusion B. Bulk molding C. Vacuum bag forming D. Computational analysis
965	The Langmuir adsorption iso therrn shows that the amount of adsorbed gas per gram of the solid is equal to.	A. ap/1+bp B. ap+1/1-bp C. 1+ap/1-bp D. a(1+bp)
966	Suppose the activatin energy of a certain reaction is 250 kj/mol, If the rate constant at T1 =300 K is k1 and the rate constant at T2= 320 K is ks, then the reaction is _____ times faster at 320 K than at 300 K	A. 3 x 10⁻²⁹ B. 0.067 C. 525 D. 15.0
967	The percentage of nitrogen in urea is.	A. 36% B. 46% C. 55% D. 65%
		A. H2O molecule in linear B. Sp3 hydrogen bonding is involved in the formation of water

968	The boiling point of water is unexpectedly high because.	the formation of water C. There is hydrogen bonding and consequent association of H2O molecules. D. Oxygen is the first member of the VI group
969	Which name is associated with the rules which help in predicting the portability of anion.	A. Soddy B. Slater C. Fajan D. Linus pauling
970	Coordination compound show	A. Structural isomerism B. Stereo isomerism C. Both A and B D. None of above
971	Chromium is found in nature in the the form of.	A. Oxides B. Silicates C. Borates D. Sulphides
972	How many oxygen atoms lined up in a now would fit in a one nanometer speace.	A. None an oxygen atoms is bigger than 1 nm B. One C. Seven D. None of the above
973	Select the correct IUPAC name for [FeF4(OH)2]-	A. Diaquatrafluoriron (III) ion B. Diaquateratrafluoriferrate (III) ion C. Diaquatertrafluoroiron (I) D. None of these
974	Citral when heated with KHSO4 forms.	A. Isoprene B. p-cymene C. p-menthane D. Dipentene
975	Petroleum is mixture of	A. Petrol B. Diesal C. Petroleum D. All of these
976	Pyrolysis gasoline is obtained from.	A. Catalytic cracking B. Gasification C. Steam cracking D. Reforming
977	The relative populations of gourd state and excited state populations at a given flame temperature can be estimated using.	A. Boltzmann distribution law B. Maxwell law C. Lambertie law D. Beer's law
978	Which of the following radical is a member of VI group.	A. Mg²⁺ B. Na+ C. K+ D. NH₄⁺ E. All above
979	Pick out incorrect statement regarding HF	A. It is used for making chlorofluorocarbon used as refrigerating fluids and as propellants in aerosols B. It is used in making ASIF3 and synthetic cryolite C. Aqueous HF is used for etchine glass D. HF does not react with B2O3 even in presence of conc. H2SO4
980	Which of the following iso -electronic ion would require least energy for the removal of electron.	A. Ca2+ B. Cl- C. Ar- D. K+
981	The most convenient and has nearest approach to a universal pH measurement	A. pH strips B. pH indicator C. The emf method D. The colorimetric
982	Sulphur can exist in	A. One phase B. Two phase C. Three phase D. Four phase
983	carbon monoxide is harmful to human beings as it.	A. Is carcinogenic B. Is antagonistic to CO2 C. Has higher affinity for haemoglobin as compared to oxygen D. Is destructive to O3

984	Which of the following is a pseudohalide.	A. I3- B. IF7 C. CN- D. ICl
985	When FeSO4 is added in the sodium extract the compound formed is.	A. Only Na4[Fe (CN)6] B. Only Fe (OH)2 C. Only Na2So4 D. Mixture of all these
986	An induction of dipole or polarity in non polar bond, and consequent electron shifting along a chain of atoms is known as.	A. Inductive effect B. Resonance effect C. Hyper conjugation D. None of the above
987	A colloidal system in which both the dispersion phase and dispersed phase are liquid is.	A. Smoke B. Emulsion C. Whipped cream D. Mist
988	Thermocouples have been constructed from	A. Chromel ve elumel B. Copper vs platinum C. Both D. None
989	The correct order of thermal stabilities of hydriedes of group 15 is.	A. NH3 > PH3 > AsH3 > BiH3> SbH3 B. NH3 > PH3 >AsH3 > SbH3> BiH3 C. NH3 < PH3 < SbH3 > AsH3 > BiH3 D. BiH3 > Sb H3 > AsH3 > PH3 > NH3
990	Ionic reactions mainly take place in.	A. Aqueous solutions and organic solvents of high polarity B. Non aqueous solvents of low polarity C. Gaseous state D. Solid state
991	Which of the following extractant is used to solid phase extraction	A. Bonding of C18 chains on silica B. Bonding of C20 on paper C. Bonding of C18 on glass D. Bonding of C20 on cellulose
992	Which of the following is not a correct postulate of the kinetic theory of gases.	A. The molecules are in random motion B. The gaseous collisions are perfectly elastic C. The average kinetic energies of different gases are equal at a particular temperature. D. The pressure exerted on the walls of the container is due to inter molecular forces.
993	Each of the following when present at para position decreases the acidic strength of phenol except.	A. -NH2 B. -Cl C. CH3O- D. CH3-
994	Relative order of acidity of HF, HCl, HBr, and HI acids is	A. HCl > HBr > HI > HF B. HF > HCl > HBr > HI C. HI > HBr > HCl >HF D. HF > HI > HCl > HBr
995	The kinetics of the decomposition of ammonia on the tungsten surface follows	A. Zero order B. First order C. Second order D. Third order
996	Among the unit cells given below, which has the highest symmetry	A. Monoclinic B. Cubic C. Hexagonal D. Orthorhombic
997	Which of the following technique is most sanative one.	A. Photomerty B. AAS C. Flame photometry D. Flourimetry
998	Egyptians were using _____ to prepare make up for eyes.	A. Nanoaluminium B. Nanocopper C. Nanosteel D. Nanolead
999	Which of the following elements of group 15 is a typical metal.	A. P B. As C. Bi D. Sb

1000	Consider the coordination compound Na2[Pt(CN)4] the Lewis and is	A. [Pr(CN)4]2 B. Na+ C. Pt D. Pt2+
1001	Which of the following in not a characteristics of crystalline solids.	A. Sharp melting point B. Isotropic C. Long range orderly arrangemnt D. None of above
1002	Which of the following is a false statement.	A. Halogens are strong oxidizing agent B. Halogens show only (-I) Oxidation state C. Hf molecules from intermolecular H-Bonds D. Fluorine is highly reactive
1003	CO belong to which group.	A. C_{2v} B. D_{2h} C. C_{av} D. D_{ah}
1004	A mixture of ethyl iodide and n-propyl iodide is subjected to Wurts reaction. The hydrocarbon that will nto be formed is	A. n-butane B. n-propane C. n-pentane D. n-hexane
1005	Pick out the incorrect statement for transition metals.	A. Cu+ is not a transition metal ion B. Transition metals do not exhibit variable oxidation states C. Transition metal ions are coloured D. Transition metals and majority of their compounds are paramagnetic
1006	Which one of the following is not correct.	A. Ar is used in electric bulbs B. Kr is obtained during radioactive decay C. Boiling point of H2 is lowest among all noble gases. D. Xe forms Xe OF4
1007	Which of the following species has highest bond energy.	A. H2 B. T2 C. D D. Cl
1008	Inductive effect can be used to explain	A. Dipole moment of chemical bonds B. Strength of acids C. Strength of bases D. All above
1009	Which of the following is the second anciently known metal.	A. Nickel B. Copper C. Gold D. Silver
1010	Petrol can be saved by	A. Driving at a constant and moderate speed B. Ensuring correct type pressure C. Switching off the engine at traffic lights D. All of these
1011	Which of the following statement is not related with nitric oxide.	A. It is a colorless and odourless gas B. It is produced largely by fuel combustion C. It is a brown pungent gas D. It is oxidized to NO2
1012	A correct reaction mechanism for a given reaction usually is.	A. The same as the balanced chemical equation B. Obvious if its heat of reaction is known C. Sometimes difficult to prove D. Obvious if the activation energy is known
1013	Al2Cl6 is an example of	A. Ionic bond B. Covalent bond C. Co ordinate bond D. Metallic bond
1014	Greeks and Romans had used nanoparticles in the manufacture of.	A. Cosmetics for eyes B. Medicines C. Metals D. Hair -dye
		A. K2SO4

1015	Which of the following salt is water insoluble.	B. Na2SO4 C. BaSO4 D. None of abvoe
1016	A molecule the cannot be susperimposed on its mirror image is said to exhibit which of the following.	A. Geometrical isomerism B. Optical isomerism C. Linkage isomerism D. Reactive isomerism
1017	An equal volume mixture explodes with violence	A. H2 & N2O B. H2 & NO C. H2 & N2O4 D. H2 & N2O3
1018	For a given mass of a gas if temperature increase	A. Pressure and volume remain Constance B. Volume increases provided pressure is kept constant C. Pressure decreases provided volume is constant D. Both volume and pressure decrease
1019	The correct order of reactivity among I , Ii, and III IS.	A. i > ii >iii B. i > iii > ii C. II > III > I D. III > II > I
1020	The bond length of C = C is	A. 1.20 A^o B. 1.34 A^o C. 1.54 A^o D. 1.68 A^o
1021	Treatment of phenol with cold dilute nitric acid gives.	A. Only o-nitro phenol B. Only p-nitro phenol C. 2,4,6 -Teinitro phenol D. Mixture of o-and p-nitro phenol
1022	The pH of the tears is	A. 7.0 B. 7.4 C. 7.8 D. 8.2
1023	The state of hybridization of Xe in Xe F6 are	A. sp2 B. sp3 C. sp3 d D. dsp3
1024	Of the molecules, SF4 Xe F4 and CF4 which have square planar geometry.	A. SF4 , Xe f4 and CF4 B. SF4 only C. CF4 only D. XeF4 only
1025	Gravimetric method is based on which of the following property.	A. Volume of a liquid B. Volume of gas C. Mass of substance D. Viscosity
1026	Monomers are Teflon is	A. Monochloroethene B. 1,2- Difluoroethene C. 1,1,2- Trifluoroethene D. Tetrafluoroethene
1027	What is the function of Head Box in paper making machine.	A. It dry the paper B. It reduces thickness of paper C. It makes the surface of paper smooth D. It discharge the pulp at the screen of fourdrinier table
1028	Peppermint oil contains.	A. Menthol B. Thymol C. a-pinene D. Comphene
1029	The bond along Sp2 hybridization is.	A. 180^o B. 120^o C. 109.5^o D. 160^o
1030	Which of the following statements is not correct regarding electromagnetic spectra?	A. The frequency of microwave is less than uv B. The velocity of X-rays is more than uv C. Cosmic rays have shorter wave length than radio waves. D. The frequency of uv is greater than visible rays.
1031	All halogens exist as covalent molecules.	A. Monoatomic B. Daitomic C. Triatonic

		D. Tetra atomic
1032	Which of the following quantity is correct for micro analysis.	A. 1 -10 mg or < 50 ml B. 10-20 mg or > 50 mL C. 50-100 mg or < 100 mL D. None of above
1033	Which of the following statement is not correct with respect to harmful effects of ground water pollution.	A. It causes lungs cancer B. It causes jaundice C. It damages crops D. It helps to prevent epidermises
1034	A catalyst increases the rate of a reaction because.	A. It provides the necessary energy to the colliding molecules to cross energy barrier B. It decreases the heat of the reaction C. It decreases the order of the reaction D. It provides a different path of lower activation energy.
1035	Variable electrovalency is due to the following reasons.	A. Unstable configuration of core B. Inset electron pair effect C. All of above D. None of above
1036	Peeling of ozone umbrella is due to.	A. CFCa B. PAN C. CO2 D. Coal burning
1037	What is defined as a local corrosion damaged characterized by surface cavities.	A. Cracking<o:p></o:p> B. Pitting<o:p></o:p> C. Cavitation<o:p></o:p> D. Corrosion<o:p></o:p>
1038	Four elements A, B,C,D have atomic numbers Z , 1 ,Z, Z +1 and Z +2 Respectively If Z is 9, then bond between which pair of elements will be ionic.	A. A and C B. D and C C. D and B D. B and C
1039	Which is the purest form of iron.	A. Pig iron B. Cast iron C. Wrought iron D. Steel
1040	Pick out the incorrect statement regarding ozone.	A. O3 is an unstbale dark blue diamagnetic gas B. The central oxygen in O3 is sp3 hybridized C. It cause the tailing of mercury D. It does not react with KOH
1041	The element with atomic numebr greater than 100 are known as	A. Trans uranium elements B. Trans fermium elements C. Actinides D. Lanthanides
1042	The mole of photon is knonw as.	A. Quantum B. Einstein C. Energy packet D. None of the above
1043	Which of the following has the maximum tendency to form complexes.	A. K B. Na C. Rb D. Li
1044	Which of the following is most soluble in water.	A. BaSO4 B. Sr SO4 C. CaSO4 D. MgSO4
1045	Commercial or the phosphoric acid is pure.	A. 37.0% B. 82.98% C. 88.25% D. 90.12%

1046	The suffix "ate" at the end of the name of the compiled signifies that it is.	A. Cation B. Anion C. Neutral D. None of above
1047	The branch of chemistry which is concerned with the interrelation of electrical and chemical energy is called.	A. Reaction dynamics B. Electrochemisty C. Surface chemistry D. Kinetics
1048	A general trend in the properties of elements of carbon family shows that with increase in atomic number.	A. The tendency towards concatenation increases B. The tendency to show +2 oxidation state increase C. Metallic character decreases D. The tendency to form complexes with convalency higher than four decreases.
1049	Is a peroxy acid	A. H2SO5 B. H2S2O6 C. H2SO4 D. H2S2O7
1050	What ASTM test for tension is designated for plastics.	A. A 370 B. D 638 C. E 292 D. None of these
1051	Hydrogen at the moment of its generation is generally called.	A. Protium B. Nascent hydrogen C. Atomic hydorgen D. Heavy hydrogen
1052	The element having electronic configuration 1s2, 2s2, 3s2, 3p3 is.	A. Trivalent only B. Tetravalent only C. Trivalent and pentavalent D. Pentavalent only
1053	Which of the following statement is not correct with respect to electrometric effect.	A. It is permanent effect B. It is brought into play instantaneously at the demand of attacking reagent C. It proceeds a polar addition reaction D. The original electronic condition is restored after the removal of tacking regent.
1054	During the preparation of soap the liquid separated by distillation is	A. Sodium hydroxide B. Oil C. Fats D. Glycerol
1055	Separation of isotopes of uranium is carried out by	A. CaF2 B. SF6 C. HF D. All above
1056	The number of phases of mixtures of four gases enclosed in a container is	A. 1 B. 4 C. 4-1 D. zero
1057	Which of the following elements has the highest value of second ionization energy.	A. Lithum B. Beryllium C. Boron D. Magnesium
1058	Diamond and carbon are the _____ forms of carbon	A. Isotropic B. amorphous C. Allotropic D. Isomeric
1059	The specific gravity of H2SO4 is	A. 1.37 B. 1.84 C. 1.17 D. 1.57
1060	What is the ratio of stress to strain in a material loaded within its elastic ranger.	A. Poisson’s ratio<o:p></o:p> B. Refractive index<o:p></o:p> C. Modulus of elasticity<o:p></o:p>

		D. None of above<o:p></o:p>
1061	If the activation energy in the forwarded direction of an elementary step is 52 kj and the activation energy in the reverse direction is 74 kj. What is the energy of reaction Delta E for this step	A. 22 Kj B. -22 kj C. 52 kj D. -52 kj
1062	The increasing order of energies of various sub shells is	A. 1s < 2s<3s<2p<3p<4s<3d B. 1s <2s<2p<3s<3p<4s<3d C. 1s>2s>2p>3s>3p>4s>3d D. 1s>2s>2p>3p>3d>4s
1063	The speed of a chemical reaction	A. Is constant no matter what the temperature is. B. Is independent of the amount of contact surface of a solid involved C. Between gases should is all cases be extremely repid because the average kinetic energy of the molecules is great D. Between ions in aqueous solution is extremely rapid because there are no bonds that need to be broken
1064	Of all the noble gaes, easily available gases are	A. He & Ar B. He & Ne C. Ne & Ar D. Xe & Kr
1065	White Phosphorus is kept under	A. Cold water B. Ammonia liquor C. Ethanol D. Kerosene
1066	In German Silver copper is alloyed with which metal.	A. Zn B. Ni C. Ai D. Zn and Ni
1067	Which of the following has the greatest metallic character.	A. Na B. Mg C. Al D. Si
1068	The equivalent conductance of a 1 N solution of an electrolyte is nearly	A. The same as its specific conductance B. 10³ times more than its specific conductance C. 10-3 times its specific conductance D. 100 times its specific conductance.
1069	In reverse phase chromatography which of the analyte will be retained more on the stationary phase.	A. Semi polar B. Non polar C. Polar D. None of the above
1070	Hydrogen gas will not redue	A. Heated cupric oxide B. Headed ferric oxide C. Heated stannic oxide D. Heated aluminium oxide
1071	A molecule returns from the excited singlet state to the ground singlet state with emission of light This process is known as	A. Fluorescene B. Scattering C. Phosphorescence D. Chemiluminescence
1072	A 10% solution of sucrose contains 10 g of sucrose in how much volume of the solution.	A. 10 mL B. 100 mL C. 1000 mL D. 1 mL
1073	Xe reacts directly with	A. O2 B. Cl2 C. F2 D. Br2
1074	Who proved that all the six hydrogen atoms in benzen are equivalent.	A. Kekule B. Ladenburg C. Faraday D. Wohler
1075	The common temperature detecting device in DTA are.	A. Thermocouples B. Thermopiles C. Thermistore D. All
		A. Cu B. Hg

1076	Which metal burns in air at high temperature with the evolution of much heat.	C. Pb D. Al
1077	Which of the following linear polymer.	A. Polypeptide B. Protein C. Starch D. Phenol formaldehyde resin
1078	Inter halogens are of types.	A. 3 B. 4 C. 5 D. 6
1079	Nano technology in other words is.	A. Carbon engineering B. Atomic engineering C. Small technology D. Microphysics
1080	In Serpekr's process the ore is treated with which of the following.	A. Carbon B. Nitrogen gas C. Both A and B D. None of these
1081	Allotropic form of tin	A. White tin B. Grey tin C. Rhomic tin D. All above
1082	Sodium reacts with excess of oxygen to form	A. Na2O B. NaO2 C. Na2O2 D. NaO
1083	Blue color of glass of due to the presence of .	A. Cobalt (II) B. Chromium (III) C. Iron (III) D. copper (II)
1084	Which of the following property is not related to aluminum.	A. it is silvery white metal with brilliant lusture B. It is a very light metal with specific gravity as 2.7 C. It is good conductor of heat D. It is the least reactive element of III Group.
1085	Which of the following does not have an a,b, unsaturated carbonyl group.	A. Androsterone B. Oestrone C. Testosterone D. Progesterone
1086	Which of the following should have the largest dipole moment.	A. Carbon tetrachloride B. Cis-stibeue C. Trans-atibeue D. Cis-dichlorocthylene
1087	Glycine reacts with nitrous acid to form	A. Methyl amino B. Acetic acid C. Zwitter ion D. Glycollic acid
1088	A unit cell having dimension , a = b c, alpha, beta, gama = 90^{o} is known as.	A. Cubic B. Hexagonal C. Orthorhombic D. None of them
1089	Which of the following is component of the ecosystem.	A. Inorganic substances B. ORGANIC Substances C. Animal and plants only D. All above
1090	Which of the following is a mode of controlling pollution in big cities.	A. Cleanliness and less use of insecticides B. Proper disposal of organic wastes, sewage and industrial effluents C. Broader roads and shifting of factories out of the residential areas D. All of above
1091	Pick out incorrect statement about K2Cr2O7	A. It oxidizes acidified solution H2SO4 to S B. It oxidizes KI to I2 C. It oxidizes HCl to Cl2 D. It gives oxygen, when treated with cold conc. H2SO4
		A. 10-100 m long and 2 to 6 cm to diameter B. 1-10 m long and 0.2 to 0.6 cm in

1092	Which of the following information is correct about a typical packed column in GC.	diameter C. 0.1-1 m long and 0.02 to 0.00 cm in diameter D. None of the above
1093	According to Usanovich concept a base is defined as any species.	A. Capable of giving up anions B. Combining with cations C. Neutralizing an acid to give a salt D. All of above
1094	Ozone depletion in stratosphere will result in	A. Forest fires B. Increased incidence of skin cancer C. Global warming D. None of the above
1095	The titration involving oxidation reduction reactions is called.	A. Complex titration B. Simplex titration C. Redox titration D. Acid base titration
1096	Which of the following statement is not correct with respect to hardness of water.	A. It is due to soluble salts of Na B. it is due to soluble salts of Ca C. It is due to soluble salts of Mg D. It is due to soluble salts of Fe
1097	What prefix in steel identification means composition varies from normal limits.	A. E B. B C. X D. F
1098	Which of the following is not a true characteristics of a catalytic reaction.	A. The amount and chemical composition of the catalyst remains unchanged after the reaction B. The catalyst does not intiate a chemical reaction C. The reaction in which product also act as catalysis are called autocratically reactions. D. The catalyst shifts the equilibrium position of a reaction in a favorable direction
1099	Which of the following is the most abundant alkaline earth metal.	A. Be B. Mg C. Ca D. Sr
1100	VBT is unable to explain the nature of some of the complexes of.	A. Cobalt B. Copper C. Nickle D. Manganese
1101	When SI is dipped with As, it becomes	A. Superconductor B. p-type conductor C. N-type conductor D. None of these
1102	Which of the following does not form stable diatomic molecule.	A. Nitrogen B. Phosphorus C. Hydrogen D. Oxygen
1103	The particle would be stationary in a lattice only at.	A. 273 K B. 0 K C. 298 K D. 373 K
1104	Who was the first scientist to describe that substance having Nano dimensions possess altogether different and unique properties.	A. Richard Feynamann B. Erick Drexler C. Archimedes D. Michael Faraday
1105	Ozone layer of stratosphere requires protection from indiscriminate use of.	A. Fungicides, insecticides, bactericides and medicines B. Aerosols and high flying jets C. Atomic explosions and industrial wastes D. Weather ballons
1106	Which of the following factor is involved in band boarding that occur in column chromatography.	A. Number of theoretical plates B. Eddy diffusion C. In phase mass transfer D. All above
		A. Stray current corrosion<o:p></o:p> B. <p class="MsoNormal" style="margin-

1107	What corrosion occurs under organic coating on metals as fine wavy hairlines?	bottom:0in;margin-bottom:.0001pt;line-height: normal">Microbiological corrosion<o:p></o:p> C. Filiform corrosion<o:p></o:p> D. Simple corrosion<o:p></o:p>
1108	Steel is an alloy of iron and carbon with limits on the amount of carbon to less than _____ percent.	A. 2 B. 3 C. 1 D. 4
1109	Which is not true about polymers.	A. Polymers do not carry any charge B. Polymers have high viscosity C. Polymers scatter light D. Polymers have low molecular weight
1110	The group H steels can be used n what temperature range.	A. 600^oC to 1100 ^oC B. 1000 ^oC to 1500 ^oC C. 1100 ^oC to 2000 ^oC D. 200 ^oC to 800 ^oC
1111	After assimilation urea leaves behind in the soil	A. NH3 B. CO2 C. Both A and B D. None of above
1112	The secondary valency of Conc. CoCl3. 6NH3.	A. 2 B. 4 C. 6 D. 8
1113	C - O bond lengths in carboxylate anion are equal due to.	A. Resonance effect B. Inductive effect C. Resonance of identical contributing structures. D. Hyperconjugation
1114	The overall energy change during the Cannot cycle to.	A. Equal to zero B. Equal to Q C. Equal to W D. Maximum
1115	Which of the following group reagent is used for III group of basic radical.	A. Dilute HCl B. H2S + HCl C. NH4OH + NH4Cl D. NH4OH + H2S
1116	Which of the following compounds has fishing ordure	A. ammonia B. Organic sulphides C. Amines D. H2S
1117	Which of the following is not known.	A. KrF6 B. XeF6 C. XeO3 D. KrF2
1118	The reagent which can react with 1- chlorobutane to give substitution product is	A. Al Cl3 B. KOH -CH3OH C. NaCN D. Mg/ether
1119	Which of the following range is correct for macro analysis.	A. Minimum 100 mg B. Minimum 10 mg C. Minimum 1 mg D. Minimum 1000 mg
1120	Which of the following has cubic structure.	A. Sodium chloride B. Potassium Chloride C. Diamond D. All of above
1121	In the forth floatation process for the purification of ores, the ore particles float because.	A. They are light B. Their surface is not easily vetted by water C. They bear electrostatic charge D. They are insoluble
		A. Low temperature high pressure and

1122	Pick out the ideal conditions needed for the manufacture of H_2SO_4 by contact process.	A. Low temperature high pressure and high concentration of reactants B. Low temperature , low concentration of reactants and low pressure C. High temperature high pressure and high concentration of reactants D. Low temperature, low pressure and high concentration of reactants.
1123	H_2SO_4 acts as gent	A. Reducing B. Oxidizing C. Both A and B D. None of above
1124	A mixture containing S^{2-} and SO_4 ions on trating with dil HCl will produce	A. H_2S gas B. SO_2 gas C. H_2S and SO_2 gas D. CO
1125	Which of the following statement is not correct regarding galvanic cells.	A. Oxidation occurs at the anode B. Ions carry current inside the cell C. Electrons flow around the external circuit. form cathode to anode D. When the e.m.f. of the cell is positive cell reactionis spontaneous
1126	According to Fajns rules, which one of following results in increased ionic nature of the covalent bond.	A. Larger cation and smaller charges on anion B. Larger cation and larger charge on anion C. Smaller cation and smaller charge on anion D. Smaller cation and larger charge on anion
1127	Not a major contributor of engineering ceramics	A. SiC B. SiO_2 C. Si_3N_4 D. BH_3
1128	The bromine produced on commercial scale may contain impurities of.	A. Water B. Chloride C. iodine D. All above
1129	In an isochoric process	A. Energy remains constant B. Volume remains constant C. Pressure remains constant D. Temperature remains constant
1130	Silicon bronze contains how many percent of silicon.	A. 96% B. 3% C. 1 % D. 69 %
1131	A system which can exchange energy as well as natter with its surrounding is said to be a/an	A. Closed system B. Inert system C. Open system D. All of above
1132	Which of the following is not chemical characteristics of water.	A. pH B. COD C. BOD D. Colour
1133	Which of the following is not a component of AAS.	A. Hollow cathode lamp B. Burner C. Detector D. Tungsten lamp
1134	Which one of the following has the biggest electron affinity.	A. F_2 B. Cl_2 C. Br_2 D. I_2
1135	Visible light is just a portion of radiation emitted by atoms.Which of the following statements is not related with visible light.	A. visible light is electromagnetic in nature. B. It travels with the speed of light C. It is a mass D. The wave number of light is directly proportional to its wave length.
1136	Which of the following makes the motion of perpetual motion machine a physical impossibility.	A. First law of thermodynamics B. Second law of thermodynamics C. Third law of thermodynamics D. The Boltzmann law
1137	Which of the following is a natural polymer	A. Nylon B. Leucite C. Cellulose

		C. Cellulose D. Polystyrene
1138	Glucose and fructose react with which of the following reagent to give same product.	A. Tollen's reagent B. Phenyl hydrazine C. Hydroxyl amine D. All of these
1139	Which of the following is not an alkali metal	A. Rb B. Sb C. Cs D. Fr
1140	Smoke is a dispersion of	A. Gas in gas B. Gas in solid C. Solid in gas D. Liquid in gas
1141	Which substances is not used as an additive in paper industry.	A. Glucose B. Starch C. Alum D. None of these
1142	Which of the following statement is not related to the characteristics of gaseous state.	A. The inter molecular forces of attraction are not strong in gaseous state B. The gases do not have definite shape and volume C. The gases are characterized by low density. D. The gases have low comprehensibility
1143	The denationalization involving C - H sigma bond electrons is known as .	A. Conjugation B. Hyperconjugation C. Mesomerism D. Resonance
1144	The lightest alkali metal is.	A. Lithium B. Sodium C. Rubidium D. Caesium
1145	Conductometry is based on	A. Electric current B. Electrical potential C. Absorbance D. Electrical conductance
1146	Dry distillation of amino acids with barium hydroxide yields.	A. Acids B. Amines C. Alcohols D. Hydroxy acids
1147	Fish die in water bodies polluted by sewage due to.	A. Pathogens B. Clogging of gills by silt C. Reduction in dissolved oxygen D. Foul smell
1148	In a standard Weaton cell the cathode is	A. Cadmium amalgam B. Mercury C. Platinum D. Carbon
1149	Which of the following method is used for the coventrating of ores.	A. Gravity separation B. Magnetic concentration C. Fourth floatation D. Electrostatic concentration E. All
1150	Which of the following statement is not true with respect to electrode potential.	A. Feasibility of a chemicals reaction B. Rate of chemical reaction C. Nature of a chemical reaction D. Free energy of a chemical reaction
1151	Which of the following species have undistributed octahedral structure.	A. SF6 B. PF6 C. Si F₆²⁻ D. XeF6
1152	Which of the following statements is not related to the decomposition phenomenon occurring in nature.	A. Decomposition is due to autotrophic organisms B. Decomposition involves bacteria and fungi C. During decomposition organisms carry out specific reactions D. Many species of decomposer are present in the biosphere
1153	Which of the following pairs shows diagonal relationship	A. Li and Mg B. Na and K

1153	Which of the following pairs shows diagonal relationship	C. Zn and Cd D. Li and BE
1154	Carbon tetra chloried has no net dipole moment because of.	A. Its planar strcture. B. Its regular tetrahedral structures. C. Similar sizes of carbon and chlorine atoms D. Similar electron affinities of carbon and chlorine.
1155	The type of bonding in HCl is	A. Pure covalent B. Polar covalent C. Highly polar D. Hydrogen bonding
1156	The action of all the relations of all the organism to their environment is called	A. Biology B. Botany C. Ecology D. Archiology
1157	What refers to a shape achieved by allowing a liquid to solidify in a mold.	A. Casting<o:p></o:p> B. Molding<o:p></o:p> C. Forming<o:p></o:p> D. All of the choices<o:p></o:p>
1158	Which of the following pollutants results from chemicals petroleum and paper industries.	A. SO2 B. CO C. Hydrocarbons D. All above
1159	What is the possible number of optical isomers for a compound contained 2 dissimilar asymmetric carbon atoms.	A. 2 B. 4 C. 6 D. 8
1160	O2 molecule is.	A. Fermagnetic B. Forromagnetic C. Paramagnetic D. Diamagnetic
1161	Which of the following play significant role in depletion of ozone layer.	A. Oxides of nitrogen B. Oxides of carbon C. Oxides of sulphur D. None of above
1162	Which of the following reacts with excess oxygen to form a normal oxide.	A. Li B. Na C. K D. Rb
1163	Formation of nano particles involves process lime	A. Foramtion of metal nuclei on different sizes. B. Interaction among the formed particles C. Both A and B D. No interaction among the nano particles synthesized
1164	Green houses are responsible for keeping our plant warm and sustaining life on the earth.	A. CO2 & water vapours B. CO2 & CFC C. CO2 & H2O D. CO2 & CH4
1165	Ground state electronic configuration of valence shell in N2 m lecule is written as (a2s)2, (o*2s)2 , (pi sp)4, (o2p)2, Hence, the bond order of N2 molecule is.	A. 1 B. 2 C. 3 D. 0
1166	What is the most common alloying ingredient in copper?	A. Brass B. Zinc C. Cobalt D. Nickle
1167	The stationary and mobile phases in paper chromatography are.	A. ^{Liquid/Liquid} B. Solid /Liquid C. Liquid/Solid

		D. Gas/solid
1168	Metal are	A. Hard B. Ductile C. Malleable D. All
1169	The compounds whose formation require a host compound and a guest compound are called.	A. Exclusion compounds B. Inclusion compounds C. Crystal compounds D. None of the above
1170	When a strong beam of light is passed through a colloidal solution, the light will	A. Be reflected B. Be scattered C. Pass unchanged D. Be dispersed
1171	Natural fertilizers are materials derived from	A. Plants<o:p></o:p> B. Animal C. Algae D. All of above
1172	Low quality steels with an M. Suffix on the designation intended for non structural application is classified as.	A. Merchant quality B. Commercial quality C. Drawing quality D. Low quality
1173	The intensity of magnetization produced per unit strength of the applied magnetic field is called magnetic susceptibility., which of the following statements is not related with this phenomenon.	A. Confirmation of structure of given compound B. Complex storeochemistry C. Diamagnetic nature of molecules D. Paramagnetic nature of molecules.
1174	The degree of dissociation of weak acid increases with.	A. Decreasing pressure B. Increasing pressure C. Increasing concentration D. Decreasing concentration
1175	In whihc period, the element with least ionization enthalpy belong to	A. Group I B. Group 2 C. Group 17 D. Group 18
1176	Bioconversion of biomass can be used for.	A. Heating purposes B. Power production C. Methane production D. All of the above
1177	Which of the following statements is not correct about noble gases.	A. Their ionization energies are very high B. Their electron affinities are nearly zero C. They do not form any chemical compounds D. They are not easily liquefied
1178	The main active contaminants of nuclear reactors are.	A. Co- 60 B. Mn -54 C. Sr-60 D. All above
1179	Strength of H bond in inter mediate between	A. Van der Waals forces and covalent bond B. Ionic and covalent bond C. Ionic and metallic bond D. Metallic and covalent
1180	Which of the following water require zero hardness.	A. Boiler feed water B. Laundry water C. Paper will water D. Dyeing water
1181	Which of the following is not a naturally occurring dye.	A. Indogo B. Indigotin C. Alizarin D. Malachite green
1182	In the Lewis formula of which of the following species, the number of single double and dative bonds are equal.	A. N2O3 B. HNO3 C. SO2 D. SOCl2
1183	Molten iron withdrawn from the blast furnace is called.	A. Wrought iron B. Pig iron C. Bessemer iron D. Stainless steel

1184	Lithum silicide reacts with concentrated hydrochlric acid to give lithium chlride along with.	A. H2 and SI B. Si H 4 gas C. Disilane gas D. Si3H8
1185	Has maximum property of catenation.	A. C B. Si C. Sn D. Pb
1186	The number 7.43 is rounded to	A. 7,44 B. 7.4 C. 7.45 D. 7.3
1187	Process of separating the racemic mixture into optically active isomers is known as.	A. Resolution B. Racemisation C. Walden inversion D. Epimerization
1188	The pH Value 4.2 is of	A. Vinegar B. Lemons C. Oranges D. Tomatoes
1189	What is graphene.	A. A new material made from carbon nanotubes B. A one atom thick sheet of carbon C. This film made from fullerene D. A software tool to measure and graphically respresent nanoparticles.
1190	When a solute is dissolved in two immiscible solvents it will distributes itself between two phases and the ratio of the concentration of the solute in two phases will be constant, This is known as.	A. Starke law B. Distribution law C. Equilibrium law D. Snell's law
1191	The cooling of molten urea by air in the tower is called.	A. Prilling<o:p></o:p> B. Evaporation<o:p></o:p> C. Condensation<o:p></o:p> D. Distillation<o:p></o:p>
1192	The dye which is a constituent of Skiffs reagent used for detection formaldehyde group is.	A. Gentain violet B. Megneta C. Phenolphthalein D. Rosolic acid
1193	Solution with components which obeys Raoult's over the entire composition range are said to be.	A. Real solution B. Regular solutions C. Dilute solutions D. Ideal Solution
1194	Arrhenius concept explained	A. Constant heat of neutralization B. Quantitative determination of acid base strength C. Catalytic property of acid D. All above
1195	For a chemical reaction A _____ produce, the rat of the reaction doubles when the concentration of A is increased by 4 times the order of the reaction is.	A. 0 B. 1 C. 1/2 D. 4
1196	Which of the following is not a general property of amino acids.	A. They have high m.p. and b.p B. They are soluble in water C. Their dipole moments are high D. They are amorphous solids
1197	Photochemical smog consist of excessive amount of X in addition to aldehydes ketones, PAN etc. X is.	A. Methane B. Carbon monoxide C. Ozone D. Carbondioxide
1198	H-Bonding also ox in ling system like	A. Protein B. DNA C. Botha A and B D. None of above

		D. None of above
1199	The valence shell electronic configuration of group III A is.	A. ns1 p2 B. ns2 p1 C. ns3p2 D. ns2p2
1200	Give violet colour to flame	A. Gallium B. indium C. Thallium D. Aluminium
1201	Yellow green flame is observed with	A. Calcium salt B. Barium salt C. Strontium salt D. Sodium salt
1202	Which of the following is strongest reducing agent.	A. Be B. Mg C. Ca D. Sr
1203	The digits which are necessary to express the result of a measurement to the precision with which the measurement is made are called.	A. Non significant figures B. Mathematical figures C. Significant figures D. Reagent errors
1204	Ca H2 on reaction with water liberates	A. H2 B. O2 C. Botha of these D. None of these
1205	The colour of Ni2+ ion is.	A. Blue B. Green C. deep green D. Orange
1206	The formula of hexa borane is.	A. B4H10 B. B6 H10 C. B5 H9 D. B8 H12
1207	The layer containing petroleum oil and gas is.	A. Above that of water<o:p></o:p> B. Below water<o:p></o:p> C. Between water and sand<o:p></o:p> D. All of above<o:p></o:p>
1208	The multiplicity of the electronic state is equal to.	A. S + 1 B. 2S + 1 C. 2S - 2 D. 2S + 2
1209	Which one of the following elements shows the most stable oxidation state of +1	A. Al B. Ga C. In D. Tl
1210	Which of the following technique in current voltage technique	A. Amperometry B. Voltammetry C. Poteatiomertry D. Polarography
1211	The Schrodinger equation when solved for any system gives.	A. The mean force path B. The Polarizability C. The energy function D. The wave function
1212	The main constituent of glass is.	A. Silica B. Silicon C. Magnesia D. Alumina
1213	The correct increasing order of bond dissociation energy for N2, O2, F2 and Cl2 is	A. N2 < O2 < F2 < Cl2 B. F2 < Cl2 < O2 < N2 C. F2 < Cl2 < N2 < O2 D. N2 < Cl2 < F2 < O2
		A. Austenite

1214	What type of steel has 0.8 % carbon and 100% pearlite.	A. Austenite B. Eutectoid C. Hyper eutectoid D. Silver steel
1215	For each value of l. the number of m velocity are.	A. n² B. 2l C. (2l+1) D. (n+1)
1216	Which of the following is raw material not present on the cement.	A. Lime stone B. Gypsum C. Red lead D. Blast furnace slag
1217	All bond length in benzene are identical due to.	A. Resonance effect B. Inductomeric effect C. Electromeric effect D. Mesomeric effect
1218	Rotary spinning process is used to produce	A. Glass wool B. Optical fibre C. Glass marble D. None of above
1219	The atomic number of Potassium is 19 and that of manganese is 25. Although the coloured of MnO4 is dark violet yet the K+ is colourles.. This is due to the fact that	A. Mn is a transition element while K+ is not B. [MnO4]is negatively charged while K+ has a positive charge C. The effective atomicnumebr of Mn is [MnO4] is 26 while for K+ the atomic number is 18 D. The Mn in a high positive oxidation state allows charge transfer transitions
1220	Which of the following analytical method is used for the separation of dissolved components from solutions.	A. Chromatography B. Dialysis C. Solvent extraction D. Distillation
1221	Which of the following type of polymerization is used for the preparation of synthetic rubber.	A. Free radical B. Zingler natta C. Cationic D. Anionic
1222	Dry ice is	A. Solid CO B. Solid CO2 C. Solid NH3 D. Solid SO2
1223	What group of steels are molybdenum high sped steels.	A. Group A B. Group D C. Group M D. Group H
1224	Natural gas can be transported through	A. Cylinders B. Pipes C. Barriers D. All of above
1225	The pH of the 1.3 x 10^{-4} NH_4Cl is	A. 1.3 B. 4.0 C. 2.886 D. 3.886
1226	Which among the following elements has the highest value of IE.	A. Mg B. Na C. Ca D. Sr
1227	CNG is stored under	A. Power genertion B. Electric Generators C. Solvent D. All of above
1228	The size of E coli bacteria is._____ nm	A. 75000 B. 2000 C. 200 D. 5
1229	An organic liquid (X) containing C, H and H has a pleasent odour with a boiling point of 78 oC. On boiling X with conc. H2SO4 a colourless gas is produced which decolourless bromine water and alkaline KMnO4 One mole of this gas also takes one mole of H2. The organic liquid (X) is.	A. n-C3H7OH B. iso-C3H7OH C. C2H5CHO D. CH3CH2OH
1230	Urea is fertilizer	A. Nitrogen fertilizer B. Potash fertilizer C. Phosphorous fertilizer D. Complete fertilizer

1231	The colour imparted by lithium to the flame is.	A. Golden yellow B. Grasay green C. Violet D. Red
1232	Albumin is classified as	A. Simple protein B. Conjugated protein C. Lipoprotein D. Derived protein
1233	The ion that is isoelectornic with CO is	A. CN- B. O2+ C. CO2- D. N2+
1234	When calcium is heated in the flame of a Bunsen burner, the colour imparted to the flame is.	A. Golden yellow B. Brick red C. Crimson red D. Grassy green
1235	Carbon and Hydrogen are estimated by	A. Liebig's method B. Kjeldhal's method C. Carries method D. None of the above
1236	In which of the following characteristics does hydrogen resemble halogens.	A. Hydrogen is the lightest gas B. H atoms contains one electron each C. Hydrogen forms ionic hydrides with alkali metals D. Hydrogen has three isotopes.
1237	Which of the following statements is correct.	A. A sigma bond is weaker than a pi pond B. There are four coordinate bonds in the Lewis structure of NH4+ ion. C. The 1 covalent bond is directional in nature D. A single bond between the two atoms cannot be re bond.
1238	An optically active compound	A. _{Must contain at least favour carbons} B. When in solution rotate the plane of polarized light C. Most always contain an asymmetric carbon atom D. In solution always give negative reading in polarimetre
1239	Copper is mainly extracted from witch of the following ore	A. Sulphide ores B. Carbonate ores C. Oxides ores D. Non sulphide ores
1240	Which element amongst the following has the highest boiling point.	A. Na B. Mg C. Ca D. K
1241	In which of the following techniques the solvated molecules are separated according to their size by their ability to penetrate a sieve like structure.	A. Adsorption chromatography B. Partition chromatography C. Ion exchange chromatography D. Gel permeation chromatography
1242	The glow of the yellow phosphorous as a result of slow oxidation in air is called.	A. Chemiluminescene B. Luminescence C. Biolumineacence D. Photolysis
1243	The equation which relates the reaction rates and equilibrium constants of many reactions is known as.	A. Taft equation B. Hammett equation C. Differential equation D. Linear equation
1244	The hybridization of sulphur in sulphur dioxide is.	A. sp B. sp2 C. sp3 D. dsp2
1245	In the purification of bauxite , the ore is fused with sodium carbonate in the process	A. Baeyer's process B. Hall's process C. Serpeck's process D. Any of above
1246	What is prefix in steel identification means it is made in an electric furnaced.	A. E B. H C. B D. Z

1247	In compressive strength of a nanotube _____ its tensile strength.	A. I less than B. Is greater than C. Is equal to D. Less than or equal to.
1248	Which of the following haloacids is stronger acids.	A. FCH2COOH B. ClCH2COOH C. Br CH2COOH D. ICH2COOH
1249	Which of the following elements has the highest density.	A. Mg B. Na C. K D. Rb
1250	For dilute solutions colligative properties depend on.	A. The number of the particles of the solute and nature of solvent. B. The number of the solute particles and on their nature C. The number of the solute particles and nature of solute and solvent D. The number of the solute particles and irrespective of the nature of the solute and solvent.
1251	Which is the second most abundant element occurring the earth crust.	A. Iron B. Cu C. Cr D. Ni
1252	Which of the following chemical strong oxidizing agent is used in COD test.	A. KMnO4 B. H2SO4 C. CH3COOH D. K2Cr2O7
1253	Drained sewage has B.O.D.	A. More than that of water B. Less than that of water C. Equal to that of water D. None of the above
1254	Hypo is used in photography to.	A. Reduce AgBr to metallic silver B. Remove silver a silver salt C. Remove undecomposed silver bromide as soluble complex D. Remove reduced silver
1255	Greenish yellow gas with pungent irritating odour	A. Chlorine B. Fluorine C. Iodine D. Bromine
1256	Commercial detergents contain mainly _____	A. RCOON B. RONa C. RSNa D. All above
1257	Photochemical among is generally formed	A. In early hours of winters B. Around mid day in summer months C. When intensity of solar radiation sis very low D. When concentration of particulate matter is very low.
1258	Each of the following compound is an aromatic except.	A. Benzene B. Naphthalene C. Cyclopentadienyl cation D. Cyclopentadienyl anion
1259	VBT does not edplain	A. Absorption spectra B. Color of transition metal ion C. Heat of formation D. All above
1260	What term is used to denote a family of thermosetting polymers that are reaction products of alcohols and acids.	A. Alkaline B. Alkydes C. Alcocide D. Ketones
1261	Bond angle is minimum in	A. H2O B. CO2 C. NH3 D. CH4
1262	The number used in cancer therapy is.	A. Fe B. Co C. Ni D. Rn

1263	The molarity of a 500 mL solution containing 4 g NaOH	A. 0.1 B. 0.2 C. 0.3 D. 0.4
1264	The high oxidizing power of halogens is favored by.	A. Low heat od dissociation of X2 B. A high electron affinity of the atom C. A higher hydration energy of the ion D. All of above
1265	Which of the following pentahalides is not formed.	A. NF5 B. PF5 C. AsF5 D. BiF5
1266	A theoretical link between quantum mechanics and thermodynamic is.	A. Electrochemistry B. Kinetic theory of gases C. Spectroscopic analysis D. Statistical thermodynamics
1267	Reacts violently with water	A. AlH3 B. AlCl3 C. Lia lH4 D. Al2Cl6
1268	What is the most undesirable of all the elements commonly found in steels.	A. Sulphur B. Phosphorus C. Silocn D. Magnesium
1269	Keeping in view the periodic law and periodic table, suggest which of the following elements should have maximum electronegative character.	A. Oxygen B. Nitrogen C. Fluorine D. Astatine
1270	The variation of enthalpy of reaction with temperature is given by.	A. Hesse's law B. Clasisus Clapayron equation C. Kirchoffs equation D. Arrhenious equation.
1271	Magnesium burns in air to give.	A. MgO B. MgCO3 C. Mg3N2 D. Both A and C
1272	A minor constant is one whose amount in the sample is	A. 0.1 to 1 % B. 0.01 to 1% C. 1 to 10% D. None of the above
1273	Helium contents in the atmosphere by volume.	A. 0.0005% B. 0.0015% C. 0.0001% D. 0.00001%
1274	The point group of XeOF4 is.	A. C6v B. C4h C. D4h D. D2h
1275	Which of the following hydroxides has the maximum solubility in water.	A. Mg (OH)2 B. Ca (OH)2 C. Sr (OH)2 D. Ba (OH)2
1276	a-terpioneol is obtained on hydration of which of the following with dilute H2SO4.	A. Citral B. Myrcene C. Linalool D. Limonene
1277	Ziegler -Natta catalysta is	A. (C2H3)3 Al B. TiCl4 C. (C2H5)3 Al/TiCl4 D. (C2H3)3 B/TiCl2
1278	Which of the following is soluble in water.	A. AgF B. AgCl C. AgBr D. Ag I
1279	Hydrogen bond is not electrostatic in nature is stated by	A. Electrostatic approach B. Valence bond approach C. Molecular orbital approach D. None of the abvoe
1280	Which ionization Potential in the following equations involves the greatest amount of energy.	A. Na = Na+ + e B. K = K+ +e C. C2+ = C3+ +e D. Ca+ = Ca2+ + e

		D. Ca+ = Ca2+ + e
1281	Which element out of the following can exhibit a maximum co valency of seven.	A. Chlorine B. Fluorine C. Sulphur D. Both Cl and F
1282	The enrichment of chemical substance at the surface of a solid is called	A. Adsorption B. Absorption C. Sorption D. Isotherm
1283	The azimuth or angular quantum number (i) determines the number of sob shells in a given shall. the allowed values of l for a given value ffor n are.	A. 1.2.3..... B. 1,2,3 .....(n-1) C. 0.1.2.3.....(n-1) D. 2,4,6,.....(n-2)
1284	Which of the following acid radical is not interfering.	A. Phosphate B. Borate C. Flouride D. Sulphate
1285	Which of the following pairs does not represent Lowery acid base pair.	A. H2O+NH3 B. H2O +H2O C. HCL + H2O D. CH3NH2 + BF3
1286	The agricultural field that produces maximum methane gas into atmosphere is	A. Wheat field B. Paddy field C. Cotton field D. Groundnut field
1287	In a bucky ball each carbon atom in bound in _____ adjacent carbon atoms.	A. 1 B. 2 C. 3 D. 4
1288	Oxytocin, a pituitary hormone to	A. Amino acid B. Polypeptide C. Protein D. Conjugated protein
1289	They hydrolysis of methly acetate is a reaction of.	A. First order B. Second order C. Third oirder D. Fourth order
1290	The relative error is usually expressed as	A. Parts per ten B. Parts per one C. Parts per hundred D. Botha C and D
1291	Titanium dioxide shows the lattice strcuture.	A. Flluorite B. Rutile C. Wurtzite D. Zeolite
1292	Helium oxygen mixture is used by deep sea divers in preference to nitrogen oxygen mixture, because.	A. Helium is much less soluble in blood than nitrogen B. Nitrogen is much less soluble in blood than helium C. Due to high pressure deep under the sea, nitrogen and oxygen react to give poisonous nitric oxide. D. Nitrogen is highly soluble in wateer
1293	An ionic compound X + Y - is most likely to be formed if	A. Ionization enthalpy of X is high electron gain ethalpy of Y is low B. Ionization enthalpy of X is high electron gain enthalpy of Y is high C. Ionization enthalpy of X is low,electron gain enthalpy of Y is low D. Ionization enthalpy of X is low electron gain enthalpy of Y is high
1294	Which of the following compounds combines with hemoglobin.	A. CO2 B. CO C. NO D. N2
1295	What is abuckyball	A. A carbon molecule B. Nickname for Mercedes -Benz's futuristic concept car (CIII) C. Plastic explosives nanoparticle (C4) D. Concrete nanoparticle with a compressive strength of 20 nanonewtons(C20)
		A. Average of first and second ionization

1296	Electronegativity is given by	A. Average of first and second ionization energies. B. Average of first and second electron affinites C. Average of ionization energy and electron affinity D. None of the above
1297	Which of the following statement is not correct.	A. The element with highest IE belongs to group 18 B. In each period the element with lowest IE belongs to group I C. In each period the element with highest IE is a noble gas D. In the second period as we move from left to right, ionization energy increases regularly.
1298	Select a basic amino acid.	A. Glycine B. Cystine C. Alanine D. Lysine
1299	Which of the following species is very poor oxidizing agent	A. H+ B. Zn²⁺ C. Fe³⁺ D. MnO₄
1300	Environmental pollution effects.	A. Biotic component B. Plants only C. Humans only D. Both biotic and abiotic components of environment
1301	Regarding the internal energy of the molecules, which one of the following statements in not correct.	A. It is the sum of vibration rotational and electronic energy B. It is a path function C. It is a state function D. It is an exact differential
1302	Calcium cyanamide on treatment with steam under pressure gives NH3 and	A. Calcium carbonate B. Calcium hydroxide C. Calcium oxide D. Calcium bicarbonate
1303	Which trihalide is not hydrolysed by water	A. NF3 B. NCl3 C. PCl3 D. AsCl3
1304	Which of the following solution would have the largest depression in freezing point.	A. 1% glucose B. 1 % KCl C. 1 % AlCl3 D. 1 % BaCl2
1305	A drop of a liquid acquires spherical shape because of.	A. Its viscous nature B. Capillary action C. The tendency to acquire minimum surface are D. Its shape
1306	Red colour of glass of due to the presence of	A. Cu2O B. CoO C. MnO2 D. CdS
1307	Which of the following equations is the most general equation of state.	A. Vander Waal's equation B. Dielectric equation C. Clasuaiua equation D. Kamberling Onnes equation
1308	Which of the following orbitals has maximum penetration effect.	A. s B. p C. d D. f
1309	The units of coefficient of viscosity are.	A. kg m-1 n-1 B. gm -1, s-1 C. kgm-1, min -1 D. None of the above
1310	Biomass refers to all the organic material derived from	A. Photolysis B. Photosynthesis C. Electrolysis D. Oxidation
1311	Which substance is used as filler or additive in paper making.	A. Starch B. Glucose C. Cellulose

		D. Maltose
1312	Which of the following halogen exist in solid state.	A. F2 B. I2 C. Cl2 D. Br2
1313	Which of the following statement is not related to applications and limitations of first law of thermodynamics.	A. This law explains why chemical reactions proceed to completion B. It is silent about the source of heat C. It is silent about the direction of heat D. It does not tell us about the reversible process.
1314	The percentage of s-character in the hybrid orbitals sp, sp2 and sp3 follows the pattern.	A. sp3 > sp2 > sp B. sp > sp2 > sp3 C. sp = sp2 >sp3 D. sp = sp2 =sp3
1315	Which among the following is a false statement.	A. SiO2 has a structure similar to that of CO2 B. Natural Si exists only in the combined state C. Si can be prepared by reducing SiO2 with Mg D. Si does not exist in graphite like structure, but exists only ni diamond like structure.
1316	Which of the following is not an alum.	A. KAl (SO4)3 12 H2O B. NaAl (SO4)2 12 H2O C. NH4Fe (SO4)2 12H2O D. FeAl (SO4)2 . 12 H2O
1317	Identify a dye which was ot originally obtained from plant source.	A. Alizarin B. Tyrian purple C. Indigotin D. Quercitrin
1318	Which sequence of steps is correct in paper making machine	A. Pressing Dyring, Flow spreader Calender stock B. Flow spreader, Pressing, Prying Calender sock C. Drying ,Pressing , Flow spreader, Calender stock D. None of above
1319	Which of the following is not a proper use of Ni.	A. It is used as catalyst B. It is used in alloy formation C. It is used in the preparation of Monel metal D. It is attached by alkalis
1320	Disease caused by eating fish found in water contaminated with industrial waste having mercury is.	A. Minamata disease B. Bright's disease C. Hashimoto's disease D. Osteoaclerosis
1321	The extinction co efficient has the units.	A. cm2 mol -1 B. cm3 mol-1 C. mol cm-3 D. mol cm-1
1322	For an average exposure of 8 hours per day, the maximum permissible concentration limit of CO in the atmosphere is.	A. 50 ppm B. 500 ppm C. 10³ ppm D. 20 ppm
1323	Compounds consisting of two or more interlocked rings are called.	A. Inclusion compounds B. Cage compounds C. Catenanes D. Crown other
1324	The aluminium alloy used to make parts of aircrafts is.	A. Magnalium B. Aluminium bronze C. Duralumin D. All of the these
1325	Volta metric technique using a dropping mercury electrode is called.	A. Amperometry B. Coulometry C. Polarography D. Potentiometry
1326	Granulated sugar containing._____	A. Glucose<o:p></o:p> B. <p class="MsoNormal" style="margin-bottom:0in;margin-bottom: .0001pt;line-

		bottom:0in;margin-bottom:.0001pt;line-height: normal">Fructose<o:p></o:p></p> C. Maltose D. Sucrose
1327	Ingold's isoprene rule states that in tepenoids isoprene units are joined.	A. Head to tail B. Head to Head C. Tail to Tail D. In a random order
1328	Which of the following statements in not correct with respect to the important characteristics of aromatic compounds.	A. They are usually cyclic compounds B. They are resistant to usual addition reactions C. They usually undergo substitution reactions D. They are less stable
1329	Which of the following phenomenon are driven by solar energy.	A. Winds B. Water cycle C. Production of biomass D. All above
1330	How many varieties of commercial iron are known.	A. 1 B. 2 C. 4 D. 3
1331	The principal quantum number determines the overall size of the orbital and energy of the electron when it is associated with the orbital.It may have the values.	A. n = 1,3,5.....infinity B. n= 2,4,6.....infinity C. n = 1,2,3,4.....infinity D. None of the abvoe
1332	Identify the incorrect statement regarding crystallization from the following.	A. It is an important procedure for purifying solids B. The impurities are removed by filtering the solution C. Crystals are separated by filtration D. In crystallization method, the solid is dissolved in a solvvent in which it is soluble at all temperature.
1333	The equivalent conductance (^) and molar conductance (^ m) of BaSO4 are related as.	A. ^ = ^m/2 B. ^/2 = ^m C. ^ = ^m D. ^ = ^m/4
1334	Which of the following is not a biodegradable polymer.	A. Protein B. PVC C. Cellulose D. Nucleic acid
1335	Which of the following statements is not correct with respect to errors in flame photometry.	A. Errors rising form the phenomena developed in the Hollow cathode lamp B. Background effect C. Errors arising from test element itself D. Spectral interference
1336	Which of the following property has a higher value for trans isomer as compared to cis isomer.	A. Density B. Dipole moment C. Melting point D. Boiling point
1337	The IUPAC name of C2H3, CO , OC OC2H5 in	A. Prepanoic anhydride B. Ethanoic anhydride C. Diketoethoxy ether D. None of the above
1338	Among group IA elements, melting point	A. Increases down the group B. Decreases down the group C. Do not show any regular trend D. Remains constant
1339	The ions Sc3+ , Ca2+ and K+ have same electronic configuration as that of.	A. Neon B. Argon C. Krypton D. Xenon
1340	Nano particles may interact with the support to be.	A. Partially oxidized B. Partically reduced C. Both a and b D. None
1341	Which of the following statements is not true about potash alum.	A. Its empirical formula is KAI (SO4)2 12H2O B. Its aqueous solution is basic in nature C. It is used in dyeing industry D. On heating it melts in its water of crystallization

1342	Which librates H2 with NaOH	A. B B. Al C. Zn D. All
1343	Setting of cement is improved by	A. Lime stone B. Clay C. Gypsum D. Water
1344	In which polymer the strength of inter molecular forces is maximum	A. Elastomers B. Thermoplastic C. Fibre D. Cross linked polymer
1345	CFSE for d^7 ion is.	A. 0.8 B. -0.8 C. -1.8 D. 1.8
1346	To obtain cement dry powder, lime stones and shales or their slurry, is burnt in a rotary kiln at a temeperature between	A. 1100 ^o and 1200 ^oC B. 1200 ^o and 1300 ^oC C. 1400 ^o and 1500 ^oC D. 1900^o and 2000 ^oC
1347	The law of triads is applicable to	A. Lithium, beryllium, boron B. Fluorine, chlorine, bromine C. Chlorine, bromine, iodine D. Sodium, potassium, Rubidium
1348	When borax is strongly hented, it gives	A. B2O4 B. Na2B4O7 C. NaBO2 D. NaBO2 + B2O3
1349	Which of the following reagent cannot be used to detect the phenolic group.	A. Neutral FeCl3 B. I2/NaOH C. NaOH solution D. Br2/H2O
1350	Pick out the incorrect statement about K2Cr2O7	A. It is thermally stable B. It dissolves in alkali to form chromate C. It oxidizes acidified FeSO4 solution to Fe2(SO4)3 D. It is used as cleansing agent for glassware, etc. when mixed wiht cold conc. H2SO4
1351	The number of coordinates required to specify the position of all the atoms in a molecule is called number of degree of freedom. The vibration degrees of freedom of a linear molecule containing N atoms are	A. 2N-5 B. 2N -6 C. 3N -5 D. N-6
1352	Which of the following statements is false about enantiomers.	A. Rotate plane of polarized light B. Are superimposable mirror images C. Nbonsuperimposable mirror images D. All of the above
1353	Pick out the incorrect statement for Xe F6	A. XeF6 is hydrolyzed practically to form XeOF4 B. It reacts with SiO2 to form Xe F4 C. On complete hydrolysis, it forms XeO3 D. It acts as F acceptor when treated with alkali metal fluoride, but cannot act as F donor to form complexes.
1354	Sodium silicate is used	A. In the paint industry B. For fixing labels to glass C. In a soap industry D. All above
1355	Chief source of water and soil pollution in	A. Mining of ores B. Thermal power plant C. Agro industry D. All the above
1356	Which of following is used as make up chemical in Kraft process.	A. Na2CO3 B. KCl C. Na2SO4 D. NaOH
1357	An element with high electronegativity has	A. High IE and high EA B. High IE and low EA

		C. Low IE and High EA D. Low IE and low EA
1358	Result of ozone hole is.	A. Acid rain B. Global warming C. Increased amount of CO2 D. Greater exposure of earth to U.V. rays.
1359	An impure sample of camphor contaminated with sand, can be purified by	A. Distillation B. Sublimation C. Steam distillation D. None of the above
1360	The nitrogen present in some fertilizers helps plants.	A. To fight against diseases<o:p></o:p> B. To fight against diseases<o:p></o:p> C. To undergo photosynthesis<o:p></o:p> D. To produce protein<o:p></o:p>
1361	Which of the following to non -auditory effect of noise on human body.	A. Changes in the Vascular tone B. _{Increase in the blood pressure} C. Wakening of the coloured vision D. All above
1362	Vitamin D1 is chemically known as	A. Ergocalciferol B. Tocopherol C. Aserphthol D. Phylloquinone
1363	Chlorine gas acts as a bleaching agent only in presence of.	A. dry air B. Moisture C. Sunlight D. Pure oxygen
1364	The atomic number of potassium is 19 and that of mangness is 25 Although the colour of MnO4 is dark violet yet the K+ is colourless this is due to the fact that.	A. Mn is a transition element while K+ is not B. [MnO4]- is negatively charged while K+ has positive charge C. The effective atomic number of Mn is [MnO4]- is 26: while for K+ the atomic number is 18 D. The Mn is a high positive oxidation state allows charge transfer transitions.
1365	A thionic acid	A. H2S2O3 B. H2S2O6 C. H2S2O8 D. H2S2O7
1366	Finely divided iorn combines with CO to give	A. Fe(CO)3 B. Fe2 (CO)9 C. F33(CO)12 D. Fe(CO)6
1367	Not a Characteristic property of ceramic material	A. High temperature stability B. High mechanical strength C. Low elongation D. Low hardness
1368	Permanent hard water is softened by addition of.	A. Na2CO3 B. CaCO3 C. MgCO3 D. ZnCO3
1369	In the extraction of iron, the furnace change consists of iron ore, coke and limestone. The function of limestone is to act as.	A. An oxidizing agent B. A reducing agent C. Flux D. Slag
1370	The device that convert the chemical energy of fuel directly into electrical energy is called.	A. Galvanic cell B. Electrolytic cell C. Fuel cell D. Concentration cell
1371	Which of the following device is used to measure the surface tension.	A. Polarimeter B. Viscometer C. Refractometer

		D. Stalagnameter
1372	Which of the following elements has the highest melting point.	A. Magnesium B. Calcium C. Strontium D. Berylliium
1373	Ingold's isoprence rule states that in terpenoids isopren units are joined.	A. Head to tail B. Head to head C. Tail to tail D. In a random order
1374	The number of significant figures in the number 0.216 is	A. 1 B. 2 C. 3 D. <strike>4</strike>
1375	Which of the following is not a property of aluminium.	A. An efficient electrical conductor B. A low density compared to other metals C. Is amphoteric D. Toxic to humans
1376	Which of the following technique describes titrations in which a standard iodine solution is need.	A. Iodometry B. Iodimetry C. potentiometry D. Argentometry
1377	Which of the following is not an extensive property.	A. Work B. Entropy C. Free energy D. Volume
1378	The addition of Br2 to cis 2-butene produces.	A. (+) 2,3 - dibromobutane only B. (-) 2,3 -dibromobutane only C. (+) 2,3, dibromobutane D. meso-2,3, -dibromobutane
1379	Which of the following combination is used to make buffer.	A. NaOH and HCl B. KOH and H2SO4 C. CH3COOH and CH3COONa D. CH3COOH and NH4OH
1380	Polyethylene Glycols are used in the preparation of which tye of detergetns.	A. Cationic detergents B. Anionic detergents C. Non ionic detergent soaps D. None of above
1381	Which of the following substance is not weak electrolyte.	A. CH3COOH B. NH4OH C. Oxalic Acid D. NaCl
1382	BCl3 is an example of hybridization	A. sp B. sp2 C. sp3 D. None of above
1383	Which of the following sets of quantum number is possible.	A. n = 4, l = 3, m = -3 ,s = 0 B. n = 4, l = 0, m = 0, s =+1/2 C. n = 4, l = 4, m = -4, s = -1/2 D. None of these
1384	When a lead a storage battery is discharged .	A. SO2 is evolved B. PbS is consumed C. Pb is formed D. H2SO4 is consumed
1385	According to Arrhenius theory an acid is defined as substance which	A. Accepts and electron pair B. Donatees H+ ion in ammonia C. Contains Cl- ions D. Furnishes H3O + ion in water
1386	Which of the following term is not used in pulping.	A. Kappa number B. Copper number C. Bromine Number D. Octane Number
1387	Hydrocarbon X (C6H12) on oxidation with hot alkaline (KMnO4) gives a mixture of prop ionic acid and dimethyl ketone. The structure of compound X is	A. CH3CH = CHCH2CH2CH3 B. (CH3)2 C = CH CH2 CH3 C. CH3CH2CH = CHCH2CH3 D. (CH3)2 C = C (CH3)2
1388	Cobalt salt imparts which colour to the borax bead	A. Blue B. Green C. Red D. Yellow
		A. CO2

1389	Which of the following molecules has the lowest average speed at 273 K.	B. CO C. CH4 D. O2
1390	During reaction of copper with aqueous solution of silver nitrate	A. Silver atoms are reduced B. Cu2+ ions are reduced C. Silver ions are reduced D. No³ ions are reduced
1391	Which of the following devoice is used to measure potential difference between celctrodes.	A. Polarimetre B. Conductometer C. Voltmeter D. Photometer
1392	Which of the following radical is not a member of IV group.	A. Mg²⁺ B. Co²⁺ C. Ni²⁺ D. Mn²⁺
1393	What field of study encompasses procurement and production of metals.	A. Metallurgy<o:p></o:p> B. Geology<o:p></o:p> C. Material science<o:p></o:p> D. Metalgraphy<o:p></o:p>
1394	The elements with highest electron affinity belongs to.	A. Period 2 , group 17 B. Period 3, group 17 C. Period 2, group 18 D. Period, 2 , group 1
1395	Boron does not form B3+ ion because.	A. It has small size and high ionization energy B. It has high electromagnetically C. It has high charge density D. None of the above
1396	The equilibrium constant value for a chemical reaction is 5 x 10^{20}which of the following statement is true with respect to this value.	A. Reaction will be reversible B. Reaction will proceed in backward direction C. Reaction is at equilibrium D. Reaction will proceed in the forward direction
1397	Pick out the incorrect statements for transition metals.	A. They have low melting and boiling points B. 5d-element have higher energies than 3d or 4 d elements C. Zr and Hf have almost identical atomic and ionic radii D. They form interstitial compounds.
1398	Which of the following effects best explains that o-nitro phenol is insoluble in water.	A. Inductive effect B. Resonance effect C. Intramolecular H -bonding D. Isomeric effect
1399	Which of the following are anionic detergents.	A. Sodium salts of sulfonated long chain alcohol B. Ester of stearic acid and polythlene glycol C. Quaternary ammonium salt of amine with acetate ion D. Sodium salts of sulfonated long chain hydrocarbons
1400	Which of the following statement is incorrect.	A. An alloy is a mixture of two or more metals B. An alloy is a mixture of two or more metal and non metal elements that have metallic properties C. An alloy has a fixed composition D. An amalgam is an alloy containing Hg
1401	Equal volumes of all gases, under similar conditions of temperature and pressure, contain equal number of molecules. This is a statement of.	A. Graham's law B. Dalton's law C. Avogadro' law D. Boyle's law

1402	A device which is used to measure the interracial angle is known as	A. Voltmeter B. Potentiometer C. pH Meter D. Goniometer
1403	The noble gases are used due to having property	A. Chomical inertness B. Low boiling point C. Any of a or b D. Both a and b
1404	Metal crystallize is system having co ordination number	A. 8 B. 12 C. 14 D. any one of above
1405	The process of heating to redness and then slow cooling in known as	A. Tempering B. Annealing C. <blockquote style="margin: 0 0 0 40px; border: none; padding: 0px;">Quenching</blockquote> D. Hardening
1406	Example of intra molecular hydrogen bonding.	A. O-nitrophenol B. O-hydroxy benzaldehyde C. O- hydroxy benzoic acid D. All of the above
1407	Which of the following statements is NOT true .	A. About 10% of the earth's crust is iron B. Pure iron does not have significant industrial use because it is too weak and soft. C. Steel in an alloy of carbon and iron with limits on the amount of carbon D. None of above
1408	Acid rain is caused due to increase in the concentration of _____ in the atmosphere	A. Ozone and dust B. CO2 and CO C. SO3 and CO D. SO2 and NO2
1409	Sodium Tetra borate is used	A. As alkaline buffer in dying & bleaching process B. In manufacture of opticl glass C. in enameling and making glaze D. All above
1410	Which of the following is not biological characteristics of water.	A. COD B. Animals C. Plants D. Viruses
1411	All the member of group III A are metals except.	A. B B. Ai C. Ga D. In
1412	The most important conditions for the formation of ionic bond are.	A. High ionization energy of the metallic atom and high electron affinity of the non metallic atom. B. Low ionization of the metallic atom and low electron affinity of the non metallic atom. C. Low ionization energy of metallic atom and high electron affinity of the non metallic atom D. High ionization energy of the metallic atom and high electron affinity of non metallic atom.
1413	Which of the following factors effect the strengths of acids and bases.	A. Inductive effect B. Romance effect C. Hydrogen effect D. All above
1414	Major achievement of CFT is	A. Interpreting the color B. Adsorption spectra C. Both A and B D. None of above
1415	Different arrangement of groups in space which can be converted into one another by rotation around a single bond are caled.	A. Conformations B. Metameres C. Enatiomers D. All of the above
1416	The exchange equilibrium in gas chromatography depends on.	A. Solubility or absorbability of he sample B. The polarity of he stationary phase and analyte

		C. The degree of H bonding D. All above factors
1417	The forces which holds the atoms together in a molecule is called	A. Ionic bond B. Covalent bond C. Co ordinate bond D. Chemical bond
1418	Which of the following statement is related with CO.	A. It is a colorless and tasteless gas B. It has less affinity to words hemoglobin C. It has a boiling point of -192 ^oC D. It is a dangerous asphyxiant
1419	Glycerol on dehyeration gives	A. Allyl alcohol B. Aerolein C. CHOH = C= CHOH D. -CHO -CHOH -CH2OH
1420	Branch of chemistry that deals with the basic principles governing energy changes during various processes is called.	A. Wave mechanics B. Chemical kinetics C. Chemical thermodynamics D. Electro chemistry
1421	Which one of the following set of raw material is most suitable for manufacture of urea.	A. CH4N2 and CO2<o:p></o:p> B. H2CO2 and H2O<o:p></o:p> C. H2O N2 and H2<o:p></o:p> D. H2O N2 AND KCl<o:p></o:p>
1422	Toluene is o/p -orienting with respect to an electrophilic substitution reaction due to.	A. +1 effect of the methyl group. B. +1 as wellas +H effect of the methyl group C. Hyper conjugatin between the methyl group and phenyl ring. D. + R effect of the methyl group
1423	Which of the following pollutant result from combustion of fossil fuels.	A. SO2 B. NOx C. CO D. All above
1424	H2SO4 is manufactured by	A. The lead chamber process B. The contact process C. Both A and B D. The Ostwald's process
1425	Potassium sulphate with 48% to 52% potash, is made from.	A. Potassium phosphate<o:p></o:p> B. Potassium Chloride C. Potassium Nitrate D. None of these
1426	Which is true for DDT it is.	A. Not a pollutant<o:p></o:p> B. An antibiotic<o:p></o:p> C. A non degradable pollutant<o:p></o:p> D. A pesticide<o:p></o:p>
1427	The stabilization of the dispersed phase in a lyophobic sol is due to	A. Liking for the dispersion medium B. The surface tension of the medium C. The formation of an electrical layer between the two phases D. The viscosity of the medium

1428	Fluorine is.	A. Powerful oxidizing agent B. Most reactive element C. Used as refrigerants D. All of above
1429	Which of the following statements not correct with the concept of Bronsted concept of acids and bases.	A. An acid can donate a proton B. A base can accept a proton C. This concept has many bases that have OH- ions D. This concept is more general
1430	Chlorination of benzene with excess chlorine in the presence of FeCl3 as Lewis acid gives.	A. Chlorobenzene as a major product B. o-dichlorobenzene as major product C. p-dichloro benzene as an only product D. A mixture of 0- and p- dichloro benzene
1431	Refining is	A. Extracting petroleum gas<o:p></o:p> B. Separation of various fraction<o:p></o:p> C. Heating of coal<o:p></o:p> D. All of above<o:p></o:p>
1432	Which is not an ore of aluminium.	A. Baxuite B. Cryolite C. Monazite D. Corundum
1433	What is the equilibrium temperature of transformation of austenite to pearlite	A. 1000 F B. 1333 F C. 166 6 F D. 1222 F
1434	Out of seven crystal system, how many can have body centered unit cell.	A. 3 B. 4 C. 2 D. 7
1435	Which of the following is an example of super octet molecules.	A. C1F3 B. IF7 C. PCl5 D. All the three
1436	The process of removing disolved impurities from a colloidal system, by means of diffusion through a suitable membrane under the influence of an electric field , is called.	A. Electrosmosis B. Electrodialysis C. Electrophoresis D. Peptization
1437	Among the solvents given below, with dielectric constant (E) given in parentheses which has highest solubility of KCl?	A. Benzene (E=O) B. Carbon disulphide (E = O) C. Methanol (E =32) D. Acetone (E = 2)
1438	Xenon reacts best with	A. The most electropositive elements B. The most electronegative elements C. The hydrogen halides D. Non metals
1439	A terpenoid which has as alcoholic group in the molecule is.	A. Citral B. Camphor C. Menthol D. Carvone
1440	Which of the following pollutant is not primary pollutant.	A. Ash B. Smoke C. SO3 D. SO2
1441	Which of the following statement is not related with industrial ecology.	A. Study of interactions between human activities and its environment B. Industrial ecology seeks to optimize the total industrial materials cycle from virgin material to finished product C. Industrial impacts on the environment D. Economic system are viewed in isolating from their surrounding

1442	The electronegativity of phosphorus is.	A. 3.0 B. 2.1 C. 2.0 D. 1.9
1443	Which of the following process is involved in getting back nitrogen into atomosphere.	A. Nitrification B. Denitricication C. Ammonification D. All above
1444	Fats and oil are_____	A. Acids B. Alcohols C. Salts D. Base
1445	Which of the following statements is false about transition metals.	A. They form complexes B. They show variable valency C. All transition metal compounds are paramagnetic D. They form coloured ions
1446	Which metal can produce dihydrogen gas by reaction with dil H2SO4	A. Ag B. Fe C. Cu D. Pt
1447	The fluoride tooth paste contains	A. SnF2 and Sn2P2O7 B. NaF C. CaF2 D. None of these
1448	When fullerencs were discovered they were thought to be	A. First example of spherical aromatic molecule B. First example of spherical non aromatic molecule C. First example of diamond lime molecule D. None of the above
1449	The maximum noise level at which a man can work for 8 hours is.	A. 80 dB B. 70 dB C. 90 dB D. 60 dB
1450	The fertilizers which provide single nutrient from NPK are called _____ fertilizer	A. Straight<o:p></o:p> B. compound C. Both A and b D. None of above
1451	Which of the following is not a characteristics of terponoids.	A. They are pleasant smelling liquids B. They are steam volatile C. They are nitrogenous bases D. They are insoluble in water
1452	Which of the following solids is a better conductor of electricity.	A. Pore NaCl cyrstal B. Diamond C. Graphite D. Marble pieces
1453	The electronic configuration of chromium (Z =24) in the ground state is.	A. [Ar] 4s² , 3d⁴ B. [Ar] 3d⁶ C. [Ar]4s¹, 3d⁵ D. [Ar]
1454	Which of the following is the most suitable catalyst for ammonia synthesis.	A. Pt<o:p></o:p> B. ZnO+ Cr2O3<o:p></o:p> C. Fe in fused mixture of Al2O3 + SiO2+ MgO<o:p></o:p> D. All of above
1455	HS- is a conjugate base of.	A. S2- B. H2S C. H2SO3 D. H2SO4

1456	Detergents are known to pollute revers and water ways. However, detergents can be made biodegradable and pollution free by taking.	A. cyclic hydrocarbon chain B. Shorter hydrocarbon chain C. Unbranched hydrocarbon chain D. Benzenoid hydrocarbons
1457	The branch of physics that mathematically describes the wave properties of electron in atomic is called.	A. Statistical Mechanics B. Quantum Mechanics C. Chemical statistics D. Thermodynamics
1458	The possible sub levels in the n = 4 energy level are.	A. s,p,d B. s,p,d,f C. s D. s,p
1459	Usually the rate of the reactions is expressed as.	A. mol dm-1 B. mol dm-3 s-1 C. mol dm-2 s-1 D. mol 2 dm-3 s-1
1460	Which of the following statement is not correct regarding Lewis acids and bases.	A. NH_3 and H_2O both behaves as Lewis bases B. Substances which donate a pair of electrons are called Lewis bases C. All Lewis bases are also Bronsted bases D. Lewis base must contain an atom having less than an octet of electron.
1461	A trace constituent is one whose amount in the sample is.	A. < 10% B. < 010% C. < 1.0% D. < 0.01 %
1462	Co ordinate compounds are	A. Polar B. Non polar C. Dem polar D. None of above
1463	Which of the following is renewable energy source.	A. Moon B. Wind C. Sun D. Ocean
1464	The three dimensional silicate anion $(Si_2O_5^{2-})n$ is present in	A. Beryl B. Silica C. Asbestos D. Clays
1465	Which of the following groups exert -1 effect.	A. - NO_2 B. - CN C. -COOH D. <C = 0
1466	The pair of molecules or ions having identical geometry is.	A. BCl_3, PCl_3 B. BF_3, NH_3 C. $CHCl_3$, CCl_4 D. $SiCl_4$, CCl_4
1467	Which of the following statements is not true for both B and Al	A. They burn in oxygen to give oxides at high temperature B. Their halides are Lewis acids C. They combine with nitrogen to form nitrides D. They react with HCl to form chlroides.
1468	Concentration polarization arises because of the	A. Different concentrations of solutions in the two half cells B. Changes in the concentration of electrolyte around the electrode from bulk concentration C. Reversible nature of the cell D. Variation in temperature during measurements
1469	Which of the following ions does not have the electronic configuration same as that of neon.	A. F- B. O^{2-} C. Na+ D. Ca^{2+}
1470	Which of the elements of group II A has the highest value of IE.	A. Mg B. Be C. Ca D. Sr
1471	The formula of Bauxite is.	A. Al_2O_3 B. $Al_2O_3.2H_2O$ C. AL_2O_3, H_2O D. Na_3AlF_6

		D. Na_3AlF_6
1472	Pick out the incorrect statement for XeF_4	A. XeF_4 disproportionate violently with water B. It is used as fluorinating agent C. It has octahedral structure for geometry D. It oxidizes I to I₂
1473	Phosphoric acid is the most important of the phosphorus oxy acids. Industrially phosphoric acid is prepared by.	A. The Ostwald process B. The Haber's process C. The reaction of phosphate rock with sulphuric acid D. The reaction P_4O_{10} with water.
1474	Which of the following is an azo dye.	A. Congo red B. Rhodamine B C. Ertythroein D. Paraosaniline
1475	In propagation step the reaction intermediate of radical polymerization is	A. Carbocation B. Carbonion C. Free radical D. Carbene
1476	The IUPAC suffix used for _______ NC group is	A. Cyanide B. Isocyanides C. Carbylamines D. Nitrite
1477	Beer's law is followed in	A. Flame photometry B. Atomic absorption spectrophotometry C. Mass spectrometry D. Potentiometry
1478	Which of the following expressions represent the equivalent conductance.	A. A = I s x 1000/V B. A = Ls x 1000/C C. A = Ls I/A D. A = Ls/V
1479	Which of the following process is not involved in the purification of bauxite.	A. Bayer's process B. Serpek's process C. Hall's process D. Goldsmith's process
1480	Which of the following is not correct criteria for an idea solution.	A. Enthalpyh of mixing = 0 B. Volume of mixing = 0 C. Free energy of mixing = 0 D. Obeys Reoult's law
1481	Which of the following material is a constituent of crop residue.	A. Cull B. Fruit C. vines D. Bagasse E. All above
1482	Pig iron is also called.	A. Cast iron B. Steel C. Wrought iron D. Stainless steel
1483	The hardness of water i s due to the presence of dissolved soluble salts of.	A. Calcium B. Megnesium C. Iron D. All above
1484	When two atoms of hydrogen combine to form a molecule of hydrogen gas the energy of the molecule.	A. higher than that of the separate atoms B. Equal to that of the separate atoms C. Lower than that of the separate atoms D. Sometimes lower and sometime higher.
1485	Helium is used for	A. The preservation of food B. Filling electrical transformer C. Pressuring agent in rockets D. All of above
1486	The light source in AAS used is	A. UV light B. Visible light C. Radio waved D. Hollow cathode lamp
1487	Iodination of benzene takes place in the presence of iodine and	A. HNO_3 B. HIO_3 C. HgO D. All of these
		A. Dennis method

1488	On industrial scale chlorine is prepared by	B. Deacon's process C. Plantner's process D. Aludels process
1489	Which of the following has the highest lattice energy	A. LiCl B. NaCl C. KCl D. CaCl
1490	On the basic of CFT the bonding between the metal and ligand is totally	A. Ionic B. Covalent C. Coordinate D. Metallic
1491	Particulate from soil and mineral primarily contains	A. Sodium compounds B. Calcium compounds C. Silicon compounds D. Calcium, aluminum and silicon compounds
1492	The oxidation state shown by phosphorus is.	A. - 3 B. + 3 C. + 3 and +5 D. -3 ,+ 3 and +5
1493	According to the Langumir isotherm , when the pressure of the gas is very large, the adsorption.	A. Is directly proportional to pressure B. Is inversely proportional to pressure C. Is directly proportional to the square of the pressure. D. Is independent of pressure
1494	The steroid which plays an important role in carbohydrate metabolism is.	A. Oestrone B. Progestrrone. C. Androsterone D. Cortisone
1495	SO2 acts as	A. Lewis base B. Lewis acid C. Botha A and B D. None of above
1496	Lithium shows diagonal relationship with	A. Beryllium B. Sodium C. Magnesium D. Calcium
1497	When to a solution of weak electrolyte a strong electrolyte with a common ion is added, the dissociation of weak electrolytes is suppressed . This is known as.	A. Stark effect B. Salt effect C. Common ion effect D. Zeman effect
1498	Which one of the following statements if false with respect to CFT.	A. In an octahederal crystal field. the d electron on a metal ion occupy the ex of orbitals before they occupy the t98 not of orbitals. B. Diamagnetic metal ions cannot have an odd number of electrons C. Low spin complexes can be paramagnetic D. Low spin complexes contain strong field ligands.
1499	Chlorine when attached to benzene has	A. +1 and + R effect B. -1 and - R effect C. -1 and +R effect D. None of the above
1500	Which of the following statements is not correct. with respect to resonance.	A. The position of atomic nuclei mus be same B. The limiting structures must have same number of paired and unpaired electrons. C. The energy of the various limiting structures must contribute equally D. All above
1501	In sodium chloride type lattice, the ratio of coordination number of cation to anion is.	A. 6:6 B. 7:7 C. 4:8 D. 4:4
1502	Which of the following cast irons is a high carbon, iron carbon silicon alloy.	A. Deorizers B. Deoxidizers C. Deoxifiers D. Deterrent
1503	Compounds formed when noble gasses get entrapped in the cavities of crystal lattices of certain organic and inorganic compounds are called.	A. Interstitial compounds B. Hydrates C. Clathrates

		D. Picrates
1504	Hydrolysis of nucleoprotein result in the formation of.	A. Proteins B. Nucleic acids C. Both A and B D. They do not hydrolyse
1505	The glow of yellow phosphorous as is result of slow oxidation in air is called.	A. Luminescence B. Chemiluminescene C. Bioluminescence D. Photolysis
1506	Cement is a mixture of	A. Clay and clinker B. Clay limestone and gypsum C. Limestone and gypsum D. Binder
1507	Oxidation state of the chromium [Cr(NH3)6]3+ complex ion is	A. +2 B. +3 C. +4 D. +5
1508	The magnitude of electron affinity depends on.	A. Atomic size B. Nuclear charge C. Electronic configuration D. All of the above
1509	Zinc oxide is.	A. A basic oxide B. An amphoteric oxide C. An acidic oxide D. A neutral oxide
1510	Which of the following process is not sorbent separation technolgy.	A. Penex B. Parex C. Molex D. Olex
1511	Which of the following is the best indicator for titration of NH4OH with HCl.	A. Methyl red B. Methyl orange C. Eosin D. Phenolphthalein
1512	Which of the following analytical technique is used for separating similar substance by preferential adsorption or partition between two phases.	A. Distillation B. Dialysis C. Chromatography D. Solvent extraction
1513	Which of the following techniques is used for the separation of macromolecules polymers.	A. Size exclusion chromatogrphy B. TLC C. GLC D. HPLC
1514	If reaction A has an activation energy of 250 kj and reaction activation energy of 100 kj, which of the following statements must be correct.	A. If reaction A is exothermic and reaction B is endothermic then reaction A is favored kinetically B. At the same temperature the rate of reaction B is greater than the rate of reaction A C. The energy of reaction A must be greater than the energy of reaction B. D. The energy of reaction B must be greater than the energy of reaction A
1515	Which of the following molecules belongs to C_{av} point group.	A. H2O B. H2S C. NH3 D. BF3
1516	Which of the following gas is not used as carrier gas in GC.	A. Argon B. Nitrogen C. Helium D. CO2
1517	A molecule MX4 has a square planar shape, The number of non bonding pairs of electrons around M is .	A. 2 B. 1 C. 0 D. 3
1518	The number 8.47 is rounded to	A. 8.5 B. 8.4 C. 8.7 D. 8.6
1519	The following ceramic product is mostly used as pigment in paints.	A. TiO2 B. SiO2 C. uo2 D. ZrO2
		A. White crystalline solid

1520	Aluminium halides is.	A. White crystalline solid B. Hygroscopic C. Sublimes at 180^oC D. All above
1521	The most electronegative and the most electromotive elements of the first period is	A. H and He B. Na and Cl C. Li and F D. H and He
1522	The technique which involves measurement of the changes in conductance of the solution by employing high frequency alternating current in known as.	A. Potentiomerty B. Polarography C. Oscillometry D. Conductomerty
1523	Classical smog occurs in place of.	A. Excess concentration of SO2 B. Low temperature C. High temperature D. Excess concentration of ammonia
1524	Proper proportioning of concrete, ensures	A. Resistance to water B. Desired durability C. Water tightens of the structure. D. All
1525	Electron gas theory fails to explain	A. Specific heat of metals B. Electrical and thermal conductivity C. Paramagnetic behavior of metals D. All of the above
1526	In the Mendeleev's periodic table elements are arranged in the increasing order of their .	A. Numbers of neutrone. B. Atomic number C. Atomic mass D. Atomic volume
1527	Which of the following is an acceptable value fo the molecularity.	A. 0 B. 2 C. 6 D. 3/2
1528	Environmental pollution refers to.	A. Peeling of top soil B. Dissipation of energy C. Release of toxic materials in environment D. None of the above
1529	Which of the following is not related to crystallography .	A. Law of rational indices B. Law of anymmetry C. Law of constancy of interfacial angel D. Henry's law
1530	The law which relates the solubility of a gas to its pressure is called.	A. Raoult's law B. Nernst law C. Ostwald's law D. Hennery's law
1531	Which of the following is the most stable towards heat.	A. CaCO3 B. BaCO3 C. Na2CO3 D. MgCO3
1532	The half life period of any first order reaction.	A. Is half the specific rate constant B. Is independent of the initial concentration C. Is always the same whatever the reaction D. Is directly proportional to the initial concentration of the reactant
1533	The coordination number of closely packed hexagonal is.	A. 4 B. 6 C. 8 D. 12
1534	The branch of chemistry which deals with the analysis of chemical products is known as.	A. Physical chemistry B. Organic chemistry C. Inorganic chemistry D. Analytical chemistry
1535	Boric Acid is used	A. In manufacture of pottery glaze B. In medicine as an antiseptic C. In tanning industry D. All above
1536	Which of the following is not evoked in quantum theory?	A. Schrodinger wave equation B. The rigid rotor approximation C. The particle in a box D. Boltzmann distribution
		A. Carbon

1537	Which of the following elements forms maximum number of compounds.	A. Carbon B. Silicon C. Hydrogen D. Fluorine
1538	Which of the following pollutant is generated from combustion of fuel.	A. Smoke B. SO2 C. CO2 D. Metallic oxides E. All above
1539	The nature of bonds in compounds of carbon and silicon is mostly	A. Covalent B. Electrovalent C. Metallic D. Both A and B
1540	Which of the following is a non degradable pollutant.	A. Long chain phenoiics B. DDT C. Mercuric salts D. All above
1541	The Hall process involves the reduction of Al2O3 to aluminium by	A. Carbon B. Carbon monoxide C. Molecular hydrogen D. Electrolysis
1542	Ferrochrum contains Cr up to	A. 60-70% B. 70-80% C. 80-90% D. 40-50%
1543	The process of identifying the component present in a sample is called.	A. Quantitative analysis B. Qualitative analysis C. Volumetric analysis D. Gravimetric analysis
1544	Aluminothermy used for on the spot welding of large iron structures is based upon the fact that.	A. As compared to iron, aluminium has greatest affinity for oxygen. B. As compared to aluminium, iron has greater affinity for oxygen C. Reaction between aluminim and oxygen endothermic D. Reaction between iron and oxygen is endothermic
1545	Glass was first made by about_____	A. 40 BC B. 400 BC C. 4000 BC D. 100 BC
1546	During the titration of weak acid against NaOH the conductance of the solution after the neutralization point.	A. Is constant B. Decreases C. Varies irregularly D. Increase
1547	Which of the following element has six electrons in the valence shell but cannot exhibit a maximum co valency of six.	A. Sulphur B. Oxygen C. Selenium D. Both A and B
1548	Metals are	A. Transparent B. Tranalucant C. Opaque D. None of above
1549	Which of the following does not represent Lewis acid.	A. ZnCl2 B. FeCl2 C. BF3 D. BuI4
1550	Who prepared and explained nano tubes for the first time.	A. Sumio Tijima B. Richard Smaley C. Erick Drexler D. Richard Feynamann
1551	Arrange the hydrides of group 15 in the correct order of reducing nature.	A. NH3 < PH3 < AsH3 < Sb H3 < BiH3 B. NH3 > PH3 >AsH3 >Sb H3 > BiH3 C. PH3 < AsH3 < SbH3 < BiH3 < NH3 D. PH3 > AsH3 > SbH3 > BiH3 > NH3
1552	The soap and detergent are source of organic pollutants like.	A. Glycerol B. Polyphosphates C. Sulphonated hydrocarbons D. All of these

1553	Which of the following is not a component of flame photometer.	A. Pressure regulator and flow meter B. The atomizer C. The burner D. Hallow cathode lamp
1554	Which of the following oxide formed in appreciable quantity in the atmosphere.	A. NO B. NO2 C. N2O D. All above
1555	Example of inter molecular H-bonding is	A. NH3 and H2O B. HF C. CH3COOH D. All of abvoe
1556	Which of the following is not a characteristic of dye.	A. It must have suitable colour B. It must be able to fix to fibre C. It must be fast to wash and lights D. It must be highly soluble in water
1557	Which of the following statement is not related with environmental pollution.	A. Direct or indirect change in any component of the biosphere B. Undesirable change in the physical characteristics of the air C. Undesirables change in the biological characteristics of the soil D. not affecting adversely the industrial progress
1558	Which one of the following is natural polymer.	A. Starch B. Nylon-6 C. Neoprene D. Buna-S, SBR
1559	Xenon diflouride is obtained by irradiating a mixture of xenon and fluorine with light from a high pressure.	A. Mercury arc B. Tungeston arc C. Xenon arc D. None of above
1560	Organic substance responsible for the smell of the Flowers etc are grouped together in chemistry as.	A. Perfumes B. Terphenoids C. Flavonoids D. Alkaloids
1561	Glass obtained by placing a layer of butyral plastic with a suitable adhesive between two layers of glass and cementing them by heat and pressure is called.	A. Glass wool B. Safety glass C. Optical glass D. Jena glass
1562	Which of the following is not a chemical pollutant.	A. Solid waste B. Noiso C. Insecticides D. Liquid waste
1563	Which of the following alloys contains Cu and Zn	A. Bronze B. Brass C. Gun metal D. Type metal
1564	Which of the following is not a polysaccharide	A. Cellobiose B. Cellulose C. Insulin D. Amylase
1565	Dibornae is used	A. Fro high energy fuel B. For welding torches C. as reducing agent D. All above
1566	Used for sterilizationof drinking water	A. F B. Br C. Cl D. I
1567	Which type of the solids are generally good conductors of electricty.	A. Covalent B. Ionic C. Metallic D. Molecular
1568	Which of the following state is not correct regarding Langmunir adsorption theory.	A. Adsorbent has specific equivalent sites B. One site can adsorbs only one molecule C. Adsorbed molecules cannot interact with each other D. Adsorption is a static process
		A. Molecules must collide with each other

1569	Which of the following statement is not related to collision theory.	to do a chemical reaction B. Molecules must posses a minimum amount of energy C. Molecules must have proper orientation D. Collision theory is applicable to liquid only.
1570	In emulsions , the dispersed phase and the dispersion medium are.	A. Both solids B. Both liquids C. Both gases D. Phase is liquid and medium is solid.
1571	Which of the following gas is lightest.	A. Dihydrogen B. Helium C. Dinitrogen D. Dioxygen
1572	For one mole a gas, the total kinetic energy is equal to.	A. 2/3 Rt B. 3/2 RT C. 2/3 kT D. 3/2 kT
1573	Ethylene belongs to.	A. C_{2v} group B. D_{2h} group C. C₂ group D. D_{ah} group
1574	Transition elements, in general, exhibit the following properties, except one, Name that property.	A. Variable oxidation state B. Natural radioactivity C. Tendency to form complexes D. Formation of alloys
1575	The concept of telluric helisx was developed by	A. Lother meyer B. A.E. de Chancourtois C. New lands D. Doberieiner
1576	Which of the following is not a redox indicator.	A. Ferroin B. Diphaylamine C. Phenolphthalein D. Methyl blue
1577	Which of the following techniques does not belong to column chromatographyy	A. TLC B. HPLC C. Electrophoresis D. Ion exchange
1578	Which of the ionic possesses highest bond energy.	A. C-C B. Si -Si C. Ge - Ge D. Sn -Sn
1579	Which of the following analytical technique is based on the refraction of radiation.	A. Conductometry B. Refractometry C. Coulometry D. Potentiaometry
1580	A high frequency sound has frequency	A. 100 Hz B. 200 HZ C. 300 Hz D. 500 Hz
1581	The chief ore of aluminium is.	A. Cryolite B. Bauxite C. Kaolin D. Carnalite
1582	Of the following an amphoteric hydroxide is.	A. Ca(OH)2 B. NaOH C. Be (OH)2 D. Li OH
1583	Which configuration has lowest potential energy.	A. Eclipsed B. Staggered C. Skew D. All have same energy
1584	1-Chlorobutane on reaction with alcohols potash gives.	A. 1- butane B. 1-butanol C. 2- butane D. 2- butanol
1585	The formula of bleaching powder is.	A. Ca OCl2 B. CaClO3 C. Ca(ClO)3)2 D. CaOCl
		A. The colloidal particle has a charge

1586	Which of the following statement is not correct regarding the stern theory of charge on colloidal particles.	A. The colloidal particle has a charge distribution at its surface. B. In the immediate vicinity of the colloidal particles there is an excess of counter ions C. The greater the concentration and charge of ions in the diffused electrical double layer. the larger is the thickness of the layer D. At large distance from the colloidal particles, the concentration of co-ions and counter ions are almost equal
1587	Which of the following glass transmits the maximum light.	A. Serrated glass B. Clear glass C. Milk glass D. Opalescent glass
1588	Group IV A consist elements.	A. 3 B. 4 C. 5 D. 6
1589	Which parameter of a chemical reaction will change with the use of a catalyst.	A. Delta F, change in free energy B. Delta S, change in entropy C. Delta E , change in internal energy D. K, the rate constant
1590	What is called black gold.	A. Petroleum B. Coal C. Coal tar D. Natural gas
1591	As it passes into food chain, the concentration of DDT	A. Remains same B. Decreases C. Increases D. Unpredictable
1592	Steel that are used for axles, gears, and similar parts requiring medium to high and strength are known as.	A. Medium carbon steel B. Low carbon steel C. Very high carbon D. High carbon steel
1593	Recrystallization is the most common technique of purification of solid organic substances Which of the following statements is not related with characteristics of a suitable solvent.	A. It dissolves the substance on heating B. It readily allows it to separate out in the form of crystal on cooling C. It does not react chemically with substance D. It does dissolve the impurities.
1594	Which of the following can act both as a Bronsted acid and a Bronsted base.	A. Na2CO3 B. OH- C. HCO3- D. NH3
1595	The correct order of acidic strength is.	A. HF < HCl< HI < HBr B. HI <HBr < HCl < HF C. HI < HBr < HF < HCl D. HF < HCl < HBr < HI
1596	Which of the following is not related to the limitations of Bohr's model.	A. It does not applicable to more than one electron system. B. It does not explain the extra lines obtained in the H-spectrum C. It considers the electron as particle D. It considers the electron as a wave.
1597	SO3 exists in form	A. a -so3 B. b-SO3 C. gama SO3 D. All above
1598	Explosive trioxide XeO3 is produced when	A. XeOF4 reacts with water B. XeOF4 reacts with silica C. XeF4 reacts with water D. Any of above statements
1599	Which of the following statements is not true with respect to atomic spectroscopy.	A. Atoms are simplest form of matter B. Atoms cannot rotate or vibrate as molecules do C. Only electronic transitions within atoms take place D. Band spectra are observed
1600	In the Friedel-Craft acylation, the amount of AlCl3 tha tmust be taken is	A. In catalytic amount B. One equivalent C. More then one equivalent D. Amount does not matter

1601	Since the acid gives both acidic and normal salts so the acid is.	A. di acid B. di basic C. double salt D. Any of above
1602	Which of the following has the maximum ionic character.	A. HF B. HCl C. HI D. HBr
1603	The role of the mineral cryolite Na2AlF6 in the Hall process for aluminum production is.	A. It is the source of aluminum B. it is a chemical reducing agent C. It forms a slag to remove impuriteis D. In the molten state, it is a solvent for alumina Al2O3
1604	Group III A of the periodic table consist of elemetns.	A. 3 B. 4 C. 5 D. 6
1605	Which of the following does not represent Lewis base.	A. Pyridine B. NaNH2 C. PCl3 D. NaOH
1606	Which of the following is a thermometric method.	A. TGA B. DTA C. DTG D. All
1607	Solid sodium chloride does not conduct electricity be cause.	A. In solid NaCl, no ions are present B. Solid NaCl is covalent in nature C. In solid NaCl, there is no mobility of ions D. In solid NaCl, there are no electrons.
1608	The Lewis structure of which of the following does not have coordinate bond.	A. SO2 B. HNO3 C. H2SO4 D. HNO2
1609	Apoenzyme is	A. Hydrolytic enzyme B. Oxidative enzyme C. Coenzyme D. Protein part of enzyme after removal of coenzyme
1610	The SI unit of pressure is Pascal it is define da sa force per unit are of 1N/m2 one atmosphere of pressure is equal to.	A. 760 mm of Hg B. 1 bar C. 101 k Pa D. 760 torr E. All are correct
1611	Acute toxicity is expressed by the term	A. LD50 B. IC50 C. I 1/2 D. Mean life
1612	Which of the following has the highest melting poing.	A. NaCl B. Li Cl C. KCl D. Rb Cl
1613	Bromine is used as	A. Fungicides B. Herbicides C. Germicides D. Insecticides
1614	Deviation in a particular measurement is the difference between the measured value and the average value The arithmetic mean of the different deviations observed in several measurements of the same quantity is known as.	A. The standard deviation B. The average deviation C. Relative mean deviation D. variance
1615	Type of hybrid orbitals used by the chlroine atom in ClO2 is.	A. sp2 B. sp3 C. sp D. None of these
1616	Essential oils are purified by which of the following methods.	A. Steam distillation B. Sublimation C. Crystallization D. Fractional crystallization
1617	What is the following is incorrect.	A. Water is more polar than H2S B. H2O2 is a planar molecule C. Heavy water is produced by the exhauative electrolysis of water made acidic D. H2O2 act both as oxidising as well as

		D. H2O2 act both as oxidising as well as reducing agent in acidic medium
1618	Which of the following substance is not present in acid rain.	A. Sulphuric acid B. Nitric acid C. Acetic acid D. Sulphurous acid
1619	_____ are the extensions of bucky balls.	A. Goodesic domes B. Hexagons C. Carbon nanotubes D. AFM and STM
1620	Eosin dye belongs to the group of dyes known as.	A. Nitroso syes B. Triphenylmethane dyes C. Diphenylmethane dyes D. Phthalein dyes
1621	Finely divided iron combines with CO to give.	A. Fe(CO)5 B. Fe2(CO)9 C. Fe(CO)12 D. Fe(CO)6
1622	The number 7.65 is rounded to.	A. 7.6 B. 7.7 C. 7.5 D. 7.8
1623	The formula of borax glass is.	A. Na2B4O7 10H2O B. Na2B4O7 5H2O C. Na B4 O7 D. None of above
1624	Graphite is a good conductor of electricity because is.	A. Has sp2 hybridized carbon stoms B. Has free electrons C. Is crystalline D. Has free atoms
1625	The dimensions for first order rate constant are.	A. s-1 B. s mol-1 C. mol-1 s-1 D. s
1626	The concentration of OH- ions in a certain household ammonia solution is 0.0025. This ammonia solution is.	A. Basic B. Acidic C. Neutral D. None of above
1627	How many stereoisomers are possible for CH3CH = CHCHCH(Br) CH3	A. 2- geometrical isomers B. 2- optical isomers C. 2- geometricla nad 2- optical isomers D. 2- geometrical and 1 optical isomers
1628	The liquor is screened to exclude _____ material	A. Fibrous<o:p></o:p> B. Polymers<o:p></o:p> C. Maltose<o:p></o:p> D. Sucrose<o:p></o:p>
1629	Alkyl cyanide and alkyl isocyanides are	A. Tautomers B. Metamers C. Functional isomers D. None of the above
1630	HClO2 gives the structure of a.	A. Linear B. Angular C. trigonal pyramidal D. Tetra hddral
1631	Arrangement of peptide chains of protein in spec to form helix stucture is referred to as.	A. Primary structure B. Secondary structure C. Tertiary structure D. Quaternary structure
1632	The common oxidation state of lanthanides is.	A. +3 B. +2 C. +1 D. +4
		A. Epoxy methane

1633	The IUPAC name of ethylene oxide is.	A. Epoxy methane B. Oxcethene C. Methoxymethane D. All of the above
1634	The process of extracting a metal in pure form its ores is known as.	A. Crushing B. Grinding C. Dressing D. Metallurgy
1635	Eutrophication is process which involves	A. Depletion of ozone layer B. Increase in the concentration of ozone in water C. Decrease in the concentration of dissolved oxygen in water by algae D. Decrease in the level of SO2 in air
1636	The reduction in ozone layer would lead to	A. Temperature chages B. Rainfall failure C. Increase uv radiation on earth D. All above
1637	The only oxidation state of alkali metals in their compounds is.	A. +1 B. +2 C. -1 D. 0
1638	According to SHAB, Lewis acid are divided into.	A. Two classes B. Three classes C. Four classes D. None of above
1639	Nitric acid is used in the manufacturing of.	A. Dyes B. Drugs C. Artificial silk D. All above
1640	According to recent view which is the correct representation of hydrated proton in aqueous solutions.	A. H+ B. H9O3+ C. H9O4+ D. H2O+
1641	Which of the following substance act as photochemical oxidant	A. Ozone B. NOx C. peroxyacetyl nitrate D. All above
1642	The ionization energy of N is more than that of oxygen becasue.	A. Nitrogen has half filled p orbitals B. Nitrogen atom is smaller in size than oxygen atom C. Nitrogen contains less number of electrons D. Nitrogen is less electronegative
1643	In bi sulphate ion, the formal charge on sulphru atom is.	A. +1 B. +2 C. +4 D. +6
1644	Beside the common silica based SPE particles, polymer supports are also available They have advantages over silica based SPE particles, Which of the following reason is possible.	A. These are stable over a wide pH range. B. These do not possesses residual silica groups C. These are designed to be wettable and have high capacity than silica base particles. D. All above
1645	The efficiency of a reversible heat engine depends only on the	A. Temperature of the heat sink B. Temperature of the heat source C. Temperature of the heat source and sink D. Pressure of the fluid
1646	Dull red flame is observed with	A. Calcium salt B. Barium salt C. Strontium salt D. Sodium salt
1647	Which of the following give higher fibre strength.	A. Eucalyptus B. Pine C. Bagnasse D. Sugar cane
1648	Which among the following is insoluble in water.	A. LiOH B. KOH C. NaOH D. RbOH
		A. Davy

1649	"Acids are substance whose aqueous solutions turned blue litmus red and tasted sour" stated by	A. Davy B. Liebig C. Boyle D. Rouelle
1650	Which of the following is the statement of third law of thermodynamics.	A. Entroy of perfectly crystalline substance is zero at T = 0 B. Entropy of a perfectly crystaline substance is zero at standard state conditions C. Entropy and enthalpy of a substance become equal at T = 0 D. Free energy of a crystalline substance is zero at T = 0
1651	Consider to violet colored compund.[Cr(OH)2)6 Cl3 and the yellow compound. [Cr(NH3)6]C;2 which of the following statements is false.	A. Both chromium metal ions are paramagnetic with 3 unpaired electrons. B. [Cr(NH3)6)3+ is calculated directly form the energy of yellow light C. For [Cr(OH)2)6]3+ is less than for [Cr(NH3)6]3+ D. The two complexes absorb their complementary colors.
1652	The proper number of significant figures in the number 0.0780 is.	A. 3 B. 1 C. 4 D. 2
1653	A mordants is substance which in	A. Coloured B. Leuco -base of a dye C. Fixes dye on the fabric D. All of these
1654	Which of the following technique is useful to remove metal ions from an interfering matrix.	A. Solvent extraction B. Electrophorests C. Cataphorests D. Gel permeation
1655	Which of the following statement is not related with direct use of solar energy.	A. It is used for space heating of buildings B. It can be used to produce electrical power using photovoltaic cells C. It can be used to produce hydrogen gas D. It can be used start motor vehicle
1656	Which of the following is not true as compared with alkaline earth metals.	A. Alkali metals are more reactive B. Alkali metals have lower density C. Alkali metals are more electro positiveness D. Alkali metals have stronger metallic bonds
1657	In a one -component system the maximum number of phase that can consist in equilibrium is.	A. 1 B. 2 C. 3 D. 4
1658	The angle of rotation in a polarimeter depends on.	A. Nature of the compound B. Nature of the solvent C. Wavelength of the light used D. All above factors.
1659	The volume of given mass of gas at constant pressure is directly proportional to the absolute temperature.This is a statement of.	A. Charles's law B. Boyle's law C. Avogadro's law D. Dalton's law
1660	_____ is used as stablizer.	A. CaO B. SiO2 C. NaCl D. None of these
1661	Which of the following is planar?	A. CH2Cl2 B. CHCl3 C. CCl4 D. C2H2
1662	The value of comprehensibility factor (z) = pV/nRTfor an ideal gas is equal to.	A. R B. 1 C. 2 D. 3
1663	Which one of following is paramagnetic and has the bond order equal to 0.57	A. N2 B. H2+ C. O2 D. F2

1664	Aluminium reacts with boiling water to liberatedi hydrogen gas along with the formation of.	A. Aluminium oxide B. Aluminium hydroxide C. Aluminium suboxide D. Aluminium superoxide
1665	The common oxidation state of elements of group V A is.	A. -3 B. +3 C. +5 D. Any above
1666	Which of the following is domain of industrial ecology.	A. The materials extractor B. The materials processor C. The consumer D. All of above
1667	How many unpaired electron are there in a strong field iron (II) octahederal compled.	A. 0 B. 1 C. 2 D. 4
1668	Arrange the following in order of increasing boiling point.	A. CH3OH < CH3Cl < RbCl < CH4 B. CHOH < CH4 < CH3Cl < RbCl C. RbCl < CH3Cl < CH3OH < CH4 D. CH4 < CH3Cl < CH3OH < RbCl
1669	Indigotin is a dye obtained from indigo plant which belongs to the group of.	A. Substantive dyes B. Mordant dyes C. Vat dyes D. Disperse dyes
1670	Group VII A of periodic table consist of elements.	A. 4 B. 5 C. 6 D. 7
1671	Iodine is a grey black solid and its vapours are in color	A. Grey B. Black C. Yellow D. Violet
1672	Metallic bond is treated essentially as in character	A. Ionic B. Covalent C. Polar D. Non polar
1673	Which of the following process is not physical in nature.	A. Mixing B. Flocculation C. Sedimentation D. Activated sludge process
1674	Aluminum is an active metal but does not corrode as iron does because.	A. Ai does not react with O2 B. A protective layer of Al2 O3 forms on the metal surface C. Al is harder to Oxidize than is Fe D. Aluminium has a high tensile strength
1675	Which of the following statements is not correct.	A. The conductance of one cm3 of a material is called specific conductance B. Specific conductance increase while equivalent conductance decreases on progressive dilution C. The limiting equivalent conductance of weak electrolytes cannot be determined by extrapolation of the plot of A against concentration D. The conductivity of metals is due to the movement of elctrons.
1676	At extremely low pressures, the van der Waals equations for one mole may be written as.	A. PV = RT + Pb B. PV = RT C. PV = RT - a/V D. (P +a) (V-b) = RT
1677	Which is major component of Bordeaux mixture.	A. Copper sulphate<o:p></o:p> B. Sodium chloride<o:p></o:p> C. Calcium chloride<o:p></o:p> D. Magnesium sulphate<o:p></o:p>

		sulphate 4s.p³ 4s.p³ 4p³
1678	For a single -component system, the maximum degree of freedom in	A. 1 B. 2 C. 3 D. Between 3 and 6
1679	Which of the following exists as polymeric chains in solid state.	A. Sr Cl2 B. Ba Cl2 C. MgCl2 D. BeCl2
1680	Chlorofluorocarbon are widely used as coolants in.	A. Air conditioners B. Clearing solvents C. Aerosol propellant's D. All above
1681	Which of the following is a buffer solution.	A. CH3COOH + NH4OH B. CH3 COOH + HCl C. CH3COOH + NaOH D. CH3COOH + CH3COONa
1682	Which of the following species is very good oxidizing agent.	A. MnO₄ B. H+ C. Zn²⁺ D. Fe³⁺
1683	The common host compound for the formation of inclusion compound is.	A. Urea B. Thiourea C. Cholic acid D. All above
1684	Dolmite is a mineral whose formula is.	A. CaCO3 B. Mg CO3 C. CaCO3, MgCO3 D. CaSO4
1685	Which of the following process is used for the conversion of matte is to nickel.	A. Orford process B. Mond's process C. Electrolytio process D. All
1686	Which of the following iso electronic spices has the highest IE.	A. Ne B. Na+ C. F D. O2-
1687	Which of the following is not true for metalloids.	A. They are borderline elements B. They usually act as electron during with non metals. C. B, Si, and Ge D. They are all solids at room temperature.
1688	Law of octaves was proposed by	A. Lother meyer B. D.I.Mendeleev C. J.A.R. Newlands D. J.W. Dobereiner
1689	Which of the following parameter is not involved in calculations based on Born Haber Cycle.	A. Ionization enthalpy B. Electron gain enthalpy C. Electronegativity D. Bond dissociation energy
1690	Atomicity of which of the following pair of elements is not same as hydrogen.	A. Phosphorus, Nitrogen B. Nitrogen, Argon C. Nitrogen, iodine D. Iodine, sulphur
1691	Sea water is converted into fresh water bases upon the phenomenon of.	A. Plasmolysis B. Sedimentation C. Diffusion D. Osmosis E. Reverse osmosis
1692	When a large block of silicon wafer is reduced to smaller component and hence non material is formed this approach is called.	A. Bottom up B. Top down C. Left to right D. Right to left
1693	The purification of Bauxite can be carried out.	A. Baeyer's process B. Hall's process C. Serpek's process D. Any of above
1694	A stable molecule is a group of atoms held together by	A. Chemical forces B. Physical forces C. Valence force D. None of above

1695	The most common beta brase with a composition of 60 % copper and 40% zinc is called.	A. Yellow brase B. Red brase C. Muntz metal D. None of above
1696	The central metal atom or ion and the ligands that are directly attached to it are enclosed in a square bracket called.	A. Coordiantion complex B. Coordination sphere C. Coordination number D. Coordination compounds
1697	Which one of the following has a linear structure.	A. H2O B. CO2 C. NO2 D. SO2
1698	A colloidal system in which a liquid is dispersed in a solid is called a/an	A. Emulsion B. Sol C. Gel D. Precipitate
1699	Which of the following is major sink for carbo monoxide.	A. Water B. Soil C. Animal respiration D. Salts dissolved in ocean water
1700	Which of the following allows charge transfer through the solution but prevents mixing of the solution.	A. Anode B. Cathode C. Electrode cell D. Salt bridge
1701	Considering the element B,C, N, G and Si, the correct order of their non- metallic character is.	A. B > C > Si > N > F B. Si > C > B > N > F C. F > N > C > B > Si D. F > N c > > SI > B
1702	The reaction of toluene with chlorine in the presence of light gives.	A. Benzoyl chloride B. Benzyl chloride C. m-chlorotoluene D. Mixture of 0 and p -chlorotoluene
1703	Reaction in which molecules absorbing light do not themselves react but induce other molecules to react are called.	A. Chain reactions B. Photosenaitized reactions C. Reversible reactions D. Free radical reactions
1704	Which of the following substances act as pollutant.	A. Oils B. Greases C. Toxins D. All above
1705	Which of the following is a component of soap.	A. Sodium sulphate B. Sodium stearate C. Sodium chloride D. Sodium bromide
1706	The minimum amount of energy that the reacting molecules must posses at the time of collations in under to produce effective collisions is called.	A. Free energy B. Threshold energy C. Activation energy D. External energy
1707	Which of the following compounds liberates CO2 on heating.	A. Li2CO3 B. Na2CO3 C. K2CO3 D. All liberate CO2 on heating.
1708	The following are primary alloying ingredients of Group H steel except.	A. Malybdenum B. Cobalt C. Chromium D. Tungsten
1709	In a period, the element with biggest electron affinity belong to.	A. Group 1 B. Group 2 C. Group 17 D. Group 18
1710	Who coined the word nanotechnology.	A. Eric Drexler B. Richard Feynamann C. Sumio tijma D. Richard smalley
1711	The isotonic nucleotide X and Y have mass numbers 35 and 37 respectively if the atomic number of X is 17 the atomic number of Y will be.	A. 15 B. 17 C. 19 D. 18
		A. Cyanogen B. Thiocyanogen

1712	Example of peseudohalonge group.	B. Thiocyanogen C. Selenocyanogen D. All above
1713	Which among the following hydride is ionic in nature.	A. Ammonia B. Protium oxide C. Calcium hydride D. Sulphane
1714	Nitric acid is used in manufacturing of.	A. Explosive B. H2SO4 C. Fertilizer D. All above
1715	Solid substances consist of an ordered any of ions and solid as a whole is electrically.	A. Conductor B. Neurtal C. Acidic D. Basic
1716	Which of the following is not a property of Cr.	A. it is brilliant silvery metal B. it is malleable C. It can take very high polish D. Its surface is tarnished easily
1717	Which of the following statement is not relat4ed wiht SO2	A. It is a colourless gas B. It has sharp and pungent odour C. It is moderately soluble in water D. It is reduced slowly in clear air to H2S
1718	Granulated sugar contains _____% sucrose	A. 80 B. 99.30 C. 60 D. 90
1719	The name hydrogen was proposed by.	A. Lavoisier B. Rutherford C. Henry Cavandish D. Scheele
1720	The halide which is inert to water is	A. PCl5 B. SiCl4 C. BCl3 D. SF3
1721	Which one of the following does not exhibit paramegnetion.	A. NO B. NO2 C. CiO2 D. ClO-2
1722	Alkaline hydrolysis of chloroform produces.	A. HCCO B. HCOO - + CO C. H3COH D. CHCL2 OH
1723	The commonly used catalyst in the manufacture of H2SO4	A. Fe2O3 with a little CuO B. v2O5 C. Platinized asbestos and MgSO4 D. All above
1724	The compound (CH3)3COH according to IUPAC is known as.	A. Tert Butanol B. 2,2 -Dumethyl -Propanol C. 2- Methyl -2-propanol D. Tert Alcohol
1725	Which of the following system has low as well as upper consolute temperature.	A. Nicotine - water B. Aniline -water C. Triethlylamine -water D. Phenol -water
1726	Which of the following elements with excess oxygen to form proxides.	A. Ca B. Mg C. Li D. Ba
1727	Estimation of nitrogen in proteins is generally arrived out by the method.	A. Duma's method B. Van Slyke method C. Kjeldahl's method D. Carius method
1728	Which can be purified by sublimation	A. F2 B. Cl2 C. Be2 D. I2
1729	The point at which the reaction is observed to be complete is called.	A. The equivalence point B. The end point C. The triplet point D. The equilibrium point
		A. Increase

1730	A considerable number of atoms pertaining to the surface_____ with the decreasing the particle size.	A. Increase B. Decrease C. No effect D. Both a and b
1731	Which one of the following sets of elements has the strongest tendency to form negative ions in gaseous state.	A. NA, Mg, Al B. Ca, V, Cr C. N , O, F D. Ga, In, tl
1732	A covalent bond which is formed between two atoms by the overlap of atomic orbitals along their axis is called.	A. Pi bond B. Sigma bond C. Polar bond D. Non polar bond
1733	Bases and reducing agents are electron giving agents and also called as.	A. Electrodotic B. Electrophile C. Nucleophile D. None of above
1734	Perdisulphuric acid is.	A. Marshal acid B. Caro acid C. None of above D. Any of above
1735	The total number of bond pairs around sulphur and total number of lone pairs around oxygen atoms in the Lewis structure of sulphate ion are respectively.	A. 4, 12 B. 8,12 C. 12,4 D. 6 ,12
1736	Which of the following bonds between carbon -carbon is teh strongest.	A. Sigma bond B. Pi bond C. Double bond D. Triple bond
1737	Which of the following configuration of an ionic species represents psedue noble gas configuration.	A. ns2 B. ns2 np6 C. ns2 np6 nd 10 D. ns2 np3
1738	Which of the following has maximum number of unpaired electrons.	A. Fe3+ B. Fe2+ C. Co2+ D. CO3+
1739	The Lewis formula of SOCl2, the total number of bond pairs and lone pairs of electron around sulphur are.	A. 2 ,1 B. 2,2 C. 3,1 D. 3,0
1740	The strongest acid is.	A. HNO2 B. HNO3 C. H2N2O2 D. HNOS
1741	The Ostwald process is the main method for the manufacture of nitric Acid in the fist step in this process is.	A. Nitrogen and hydrogen react to form NH3 B. Ammonia is burned in O2 to generate N2 and H2O C. Nitrogen and oxygen react to form NO2 D. Ammonia is burned with O2 to generate NO and H2O
1742	All naturally occurring processes spontaneously in a direction leads to.	A. Decrease of entropy B. Increase of entropy C. Decrease in free energy D. Increase in free energy
1743	Which of the following species is determined by complex metric titrations.	A. K+ B. Na+ C. Ca+ D. Cl-
1744	It is known that AgCl is insoluble in HNO3 but dissolves readily in NH4OH solution .Which of the following statement is not correct.	A. Ag ion reacts to form complex with NH4OH solution B. The concentration of Ag ion decreases C. Ionic product is less than the solubility product D. Ionic product is greater than solubility product
1745	Which of the following disposal method is used or municipal wastes.	A. Compaction B. Composting C. Recycling D. Chemical processing E. All above

1746	Yellow colour of the flame is observed with	A. Calcium salt B. Barium salt C. Sodium salt D. Potassium salt
1747	Urea an enzyme used to estimate urea is a	A. Hydrolytic enzyme B. Oxidative enzyme C. Reductive enzyme D. Iso me rising enzyme
1748	The tyndall effect was used by Zsigmondy to device.	A. The ultramicroscope B. The ultracentrifuge C. The osmometer D. Electrodialysis
1749	What is the effect of manganese in cast iron.	A. ^{To affect the machinability ductility and shrinkage depending on form} B. To reduce hardness by combining with sulfur below 0.5% and increase hardness above 0.55 C. To dioxide molten cast iron D. Have no effect
1750	The variable valency is generally observed in case of.	A. Transition elements B. Inert gases C. Normal elements D. Non- metallic elements
1751	Earth is protected from U.V. radiations by	A. Carbon dioxide layer B. Oxygen layer C. Ozone layer D. Troposphere
1752	The emission of light characteristics of metal and correlation of intensity of the light emitted with concentration of that metal forms the basis of.	A. Roman spectroscopy B. IR spectroscopy C. Flame photometry D. Rotational spectroscopy
1753	Which of the following orbitals does not make sense.	A. 6f B. 4f C. 7s D. 2d
1754	Which of the following statement is not related to BVT	A. individual orbitals lose their indention B. VBT uses the concept of resonance C. VBT does not explain the paramanhetic nature of molecule D. it uses only valence electron
1755	Which of the following functional groups is not involved in ion exchange chromatography.	A. Weak acids B. Strong acids C. Strong bases D. Carbohydrates
1756	Which of the following process is used for the removal of gases.	A. Precipitation B. Chemical reaction in the atmosphere C. Absorption D. All above
1757	Which of the following statement is not correct with respect to limitations of Hammett equation.	A. It is only applicable to aromatic systems B. Only applicable to aliphatic systems C. It is not valid for m-substituent
1758	Argon is used in filling of.	A. Discharge tubes B. Luminous tube C. Fluorescent tubes D. None of above
1759	At higher altitudes, the boiling point of water is lowered because.	A. Atmospheric pressure is low B. Temperature is low at high altitude C. Atmospherics pressure increase D. None of the above
1760	Two solids A and B have appreciable different solubility in water but their m.p. are very close. The mixture A and B can be separated by.	A. Sublimation B. Distillation C. Fractional crystallization D. Specific rotation
1761	Which of the following statements is not correct with respect to applications of H-bonding.	A. It explains the usual b.p. and m.p of certain class of compound. B. It explains the solubility of certain organic compounds in hydroxylic solvents C. It explains the lack of ideal behavior in gases and solutions D. It has stonrg influence on the configuration of certain molecules

		configuration of certain molecules.
1762	Which of the following can be used as drying agent of ammonia.	A. CaO B. Anhydrous CaCl2 C. P2O5 D. Conc. H2SO4
1763	When some quantity of electricity is passed though two electrolytic cells. The ratio of the mass of the products obtained at the cathode is the same as the ratio of their	A. Densities B. Atomic masses C. Equivalent masses D. Atomic numbers
1764	Valence bond theory is also called as	A. Electron pair theory B. Band theory C. Electron gas theory D. Electron pool theory
1765	The vibration degrees of freedom for a linear and non liner poly atomic molecule of seven atoms each an respectively	A. 30 and 29 B. 30 and 32 C. 28 and 29 D. None of above
1766	Retarded reaction are those	A. <span style="font-family:"Times New Roman","serif"; mso-ascii-theme-font:major-bidi;mso-hansi-theme-font:major-bidi;mso-bidi-theme-font: major-bidi">In which the rate of the reaction is independent of pressure<o:p></o:p> B. In which products are strongly adsorbed on the surface of the solid catalyst C. Which are reversible under all conditions<o:p></o:p> D. For which G is positive<o:p></o:p>
1767	What is use of the addition of brine solution in the production of soap from castor oil and sodium hydroxide.	A. To speed up the reaction B. To lower the solubility of soap C. To remove unreacted castor oil and sodium hydroxide D. To increase the purity of the soap obtained.
1768	Which of the following adsorption indicator is used for any of the halides at pH.	A. Fluorescein B. Eosin C. Thorin D. Rhodamine 6 G
1769	The large increase in the rate of a reaction on rise in temperature is due to.	A. The lowering of activation energy B. The decreases in mean free path C. The increase in collision frequency D. The increase in the number of molecules having more than the threshold energy
1770	Which of the following pose threat to historical monument Taj.	A. Floods in Yamuna river B. Temperature mediated spoilage of marble C. Air pollutants from Mathura refinery D. Weathering of marble
1771	Number of unpaired electrons in Cu2+ ions are.	A. 1 B. 2 C. 3 D. 4

1772	Naphthalene balls are obtained from	A. Carbon B. Coke C. Coal Tar D. All of above
1773	The electrical conductivity of a nano tube is _____ times that of copper.	A. 10 B. 100 C. 1000 D. 1/100
1774	All the halogen form oxyacide, except	A. Florine B. Chlorine C. Bromine D. Iodine
1775	The decrease in electron density at one position accompanied by a corresponding increase at other position is called.	A. Inductive effect B. Asymmetric effect C. Electromeric effect D. Resonance effect
1776	Which of the following detector is used in HPLC system.	A. Differential refractometer detector B. UV detector C. Diode array detector D. All above
1777	Which of the following class of compounds follow the criteria of aromatically.	A. The compounds must have high degree of unsaturation B. they must have the property to undergo addition reactions C. They must have the property to undergo substitution reactions D. They must have the ability to sustain an induced current in NMR
1778	In average composition of a good sample of cement the percentage of silica is.	A. 18.5% B. 20.5% C. 22.5% D. 24.5%
1779	In Dumas method, the volume of the gas collected is equivalent to which of the following gases set free from the compound.	A. Ammonia B. O2 C. N2 D. NO
1780	Maximum desirable concentration of fluorides according to international standard is.	A. 10-100 ppm B. 1 ppm C. 100-200 ppm D. 10-20 ppm
1781	pKa value of hyponitrous acid is.	A. -7.0 B. 8.9 C. 4.1 D. 6.6
1782	Oil of turpentine contains	A. a-pinene B. b- pinene C. Both A and B D. Name of these
1783	The alternate feasible fuel for existence of mankind to	A. Uranium B. Wood C. Bontonite D. Cloth residues
1784	The matrix is usually in the form of.	A. Sand B. Limestone C. Rocks D. All
1785	Potassium reacts with excess of oxygen to form	A. K2O B. K2O2 C. KO2 D. K2O3
1786	The decomposition of dimethyl ether at 504 oC is first order with a half -life of 1570 second. What fraction of an initial amount of dimethyl ether remain after 47-10 seconds.	A. 1/3 B. 1/6 C. 1/8 D. 1/16
1787	Ten elements Sc (Z = 21) to Zn (Z = 30) fill their 4s orbitals first and then 3d orbitals are called elements. of.	A. 3 d series. B. 4d Series C. 5d Series D. None of above
1788	Long diseases are about four times more in urban areas as compared to rural areas. This is due to the presence of which of the following. in atmosphere.	A. CO2 B. NO2 C. O2

		D. N2
1789	In each period, the most electropositive element belongs to group.	A. 18 B. 17 C. 1 D. 2
1790	A system is said to be in the colloidal state if the particle size of the dispersed phase ranges from	A. _o l to 10 A B. 10 to 10000 A C. 10 to 100 A D. 1000 to 10000 A
1791	Iso-osmotic solutions are those which have the same.	A. Vapour pressure lowering B. Osmotic pressure C. Molality D. Boiling point elevation
1792	The noble gases which does not I do not form any clathrates is.	A. He B. Ne C. Argpm D. Both He and Ne
1793	If there are only two components in a solution with mole fraction X_A and X_B then which of the following relation is correct.	A. X_A + X _B = 0 B. X _A + X _B > 1 C. XA = XB < 1 D. XA = 1 - XB
1794	1-Butyne on oxymereuration -demercuration would give.	A. Butanone B. Butanal C. Propanol and methanol D. Propanoic acid and formic acid
1795	An acid base titration involves a neutralization reaction in which an acid is reacted with an equivalent amount of base The titrant is always a strong acid or base The analyte may be	A. Strong acid B. Strong base C. Weak acid and Weak base D. All above
1796	The gases that are responsible for green house effect are.	A. CO2 & CH4 B. CFC C. N2O D. All above
1797	The major role of Flurospar which is added in small quantities in the electrolytic reduction alumina dissolved in fused cryolite is.	A. As a catalyst B. To make the fused mixture very conducting C. To lower the temperature of the melt D. To decreases the rate of oxidation of carbon at the anode
1798	Soap is soluble in grease because it	A. Is non polar B. Has a hydrophobic head C. Has a hydrophobic tail D. Has an ionic head and a hydrocarbon tail
1799	Which of the following molecules have centre of symmetry.	A. H2O B. HCl C. CO2 D. H2SO4
1800	Which of the following state is not true with respect to copper.	A. it is malleable and ductile B. It is a best conductor of heat and electricity C. It forms alloys easily D. Moltan copper absorbed carbon dioxide
1801	Inert pair effect is best shown by	A. Si B. Z C. Sn D. Pb
1802	The decreasing order of the second ionization energies of K, Ca and Ba is	A. K > Ca > Ba B. Ca > Ba > K C. Ba > K > Ca D. K > Ba > Ca
1803	Which of the following detector is used in GC analysis	A. Thermal conductivity detector B. Flame ionization detector C. Mass spectrometer D. All above
1804	Which of the following physical property forms the basis of radio chemicals methods of analysis.	A. Absorption of light B. Emission of light C. Radioactivity D. Thermal conductivity
		A. CCl4

		A. CCl_4
1805	Which of the following contains both covalent and ionic bond.	B. NH_4Cl C. $CaCl_2$ D. H_2O
1806	For an elementary reaction 2A + B ---------- C + D The molecularity of the reaction is.	A. 1 B. 2+ C. 3 D. 4
1807	a- pinene hydrochloride on warming rdarranges to form bornyl chloride.The rearrangement is known as.	A. Pinacol pinacolone B. Hofinann C. Wager Mecrwein D. Wolff
1808	The maximum number of electrons in s,p,d and f sub shells are.	A. 2 in each B. 2, 6, 10, 18 C. 2,6 ,10,14 D. 5 in each
1809	The dye obtained from madder root	A. Indogotin B. Indanthrene C. Alizarin D. Acriflavin
1810	Which type of elements form ionic hydrides.	A. Transitionelements B. Metalloids C. Elements with high electronegativity D. Elements with high electropositivity.
1811	A silver iodide and was prepared by mixing KI and $AgNO_2$ solution with the $AgNO_2$ in slight excess. Which of the following descriptions is correct regarding is not particles.	A. Negatively charged because of the excess of NO-3 ions B. Positively charged because of the excess of Ag+ ions in the AgI lattice C. Negatively charged because I ions are adsorbed from the KI solution D. Neutral
1812	The function of boiling the sodium extract with conc. HNO_3 before testing the halogens is	A. To make solution clear B. To make the solution acidic C. To bring common ion effect D. To destroy CN- and S2- ion
1813	The oxidation number Xe in $XeOF_2$ is	A. 0 B. +2 C. +4 D. +3
1814	In the formation of H_2O molecule, the oxygen atom makes use of.	A. 2p orbitals B. sp hybrid orbitals C. Sp2 hybrid orbitals D. Sp3 hybrid orbitals
1815	Among the following a good solvent for a Grignard reagent formation would be.	A. t- butanol B. dimethyl ether C. difiuoro ethane D. tetrahyudroform
1816	The designation of an orbital with n =4 and l = 1 in	A. 4 s B. 4 p C. 4 d D. 4 f
1817	Select the major product obtained from the addition of HBr to I -Methyl cyclohexene	A. 1-bromo -2- methyl cyclohexane B. 6- bromo-i- methyl cyclohex -i- ene C. 3- bromo -1- methyl cyclohex - 1- ene D. I-bromo -I- methyl cyclohexane
1818	Which of the following statement is no true with respect to nitrogen dioxide.	A. It is produced by the oxidation of NO B. Its small concentration has been detected to lower stratosphere C. It is major pollutant D. It does not absorb sunlight.
1819	A gas obeying the van Waals equation will closely resemble and ideal gas if	A. The parameters 'a' and 'b' are small B. 'a' is small but 'b' is large C. 'a' is large but 'b' is mall D. None of the above
1820	H_2SO_4 is used	A. In the preparation of aqua regia B. In the purification of gold and silver C. In the dental filling D. None of above
1821	Co ordinate covalent bond found is formed by the	A. Transference of electrons B. Sharing of electrons C. Donation of electrons D. None of these

1822	Air pollution is not caused by	A. Pollen grains B. Hydroelectric power C. Industries D. automobiles
1823	Ca2+ is isolelectronic with.	A. Mg2+ B. Kr C. Ar D. Na+
1824	The branch of chemistry dealing with the study of reactions in the Uv visible region of the spectrum in known as.	A. Kinetics B. Photo chemistry C. Surface chemistry D. Catalysis
1825	At constant temperature and pressure, the decrease in Gibbs free energy (F) in equal to	A. Increase in entropy B. Decrease in entropy C. Reversible work done by the system D. All types of work except the work of expansion
1826	Which of the following is not a characteristics of solids.	A. Definite shape B. Definite mass C. Definite volume D. Fluidity
1827	Ionization potential of carbon is.	A. 11.2 B. 7.8 C. 8.1 D. 7.3
1828	Which of the following statement is not correct with respect to hydrolytic cycle.	A. It is the major constituent of the lithosphere B. Water covers about 83% of the earth's surface C. it is essential requirement of all the organisms D. Water covers about 73% of the ear5th's surface.
1829	The diameter of hydrogen atom is._____nm	A. 10 B. 1 C. 0.1 D. 0.01
1830	Is a chain silicate	A. Olivine B. Tremolite C. Beryl D. Zeolite
1831	Galvanized steel are steel products coated with	A. Carbon B. Sulphur C. Zinc D. Iron
1832	The contact process is mainly used when acid is required for the manufactures of.	A. Explosives B. Fine chemicals C. Lead accumulators D. All above
1833	In which of the following compound valency of carbon in 4 but its oxidation number is zero	A. Methane B. Carbon dioxide C. Carbon monoxide D. Formaldeyde
1834	Which one of the following statements regarding BF3 is not correct.	A. It is an ionic compound B. It is an electron deficient compound C. It is a Lewis acid D. It forms adducts
1835	The penultimate shells have pscudo inert gas type configuration.	A. Ga B. In C. TI D. All above
1836	In the metallurgy of iron, when limestone is added to the blast furnaced, the calcium ion ends up in	A. Slag B. Gangue C. Metallic calcium D. Calcium carbonate
1837	When steam is passed over red bot coke The product formed is	A. Hydrogen and carbon dioxide B. Mixture of hydrogen and carbon monoxide C. Mixture of hydrogen and oxygen D. Heavy hydrogen
1838	Which of the following is capable of forming zwitter ion.	A. Amino acids B. Halo acids C. Hydroxy acids

		D. All of these
1839	Among the elements A, B,C and D having atomic numbers 7,8,9 AND 12 Respectively, the element with smallest size and highest IE is.	A. A B. B C. C D. D
1840	The emission of light in a biological reaction in known as.	A. Fluorencence B. Phosphoreacence C. Biolumineacence D. Phtolysis
1841	Which law of thermodynamics helps in calculating the absolute entropies of varies substances.	A. Zeroth law B. 1st law C. Second law D. Third Law
1842	Which of the following process always involve the decrease in oxidation number.	A. Hydrolysis B. Elecomposition C. Oxidation D. Reduction
1843	The gases H2, N2, O2 and NH3 . H2 =2 , N2 = 28, O2 = 32 and NH3 =17 will effuse in the order.	A. H2> N2> O2 > NH3 B. NH3 >O2 > N2 > H2 C. H2 > N2 > NH3 > O2 D. H2 >NH3 > N2 > O2
1844	iodine is used as	A. Tincture of iodine B. Iodex and antiseptic C. Treatment of goiter D. All above
1845	2- Butanol is optically active because a contains	A. An asymmetric carbon atom B. A plane of symmetry C. Centre of symmetry D. A hydroxyl group
1846	Soap and detergent remove the direct form clothes due to.	A. Osmosis B. Gravity C. Lowering of interfacial tension D. Diffusion
1847	The width of a carbon nano tube is._____ nm	A. 1 B. 1.3 C. 2.5 D. 10
1848	The compound which does not act as Lewis acid is.	A. BF3 B. AlCl3 C. BeCl2 D. SnCl4
1849	Which of the following health effect is caused by mercury.	A. Nerve damage B. Brain damage C. Kidney damage D. All above
1850	A molecule returns from the excited singlet state to the ground singled state with emission of light , This process is known as.	A. Fluorescene B. Scattering C. Phosphorescence D. Chemiluminescence
1851	Which of the following oxides is amphoteric..	A. CaO B. BaO C. BeO D. MgO
1852	Neon is used in neon signs for advertising purpose because.	A. Neon lights are visible from long distance B. Neon light are visible though fog & mist C. Both A and B D. None of the above
1853	Which of the following does not apply to metallic bond.	A. Overlapping valence orbitals B. Mobile valency electron C. Delocalized electrons D. Highly directed bonds
1854	Bromination of n-butane produces.	A. I-bromobutane as the major product B. 2- bromobutane as the major product C. Both I - bromo and 2- bromobutane with equal percentage D. Both i-bromo and 2-bromo products whose percentage depends upon temperature.
1855	Xenan hexaflouride at 47.7 ^{o}C is	A. Colorless solid B. yellow solid

	Xenon hexafluoride at 47.7° C is	C. Yellow liquid D. Colorless liquid
1856	Phosphorus is detected by fusing the organic compound with -------followed by extraction with H2O	A. HNO3 B. H2SO4 C. Sodium per oxide D. Ozone
1857	Which of the following is not an intensive property.	A. Melting point B. Refractive index C. Entropy D. Density
1858	D(+) glyceraldebydes has the absolute configurtion.	A. E- B. S- C. E- D. Z-
1859	Which of the following health effect is caused by lead.	A. Cancer B. Neurotoxin C. Hypertension D. Kidney damage
1860	Conjugation of chromophore	A. Deepens the colour B. Lightene the colour C. Shifts absorption to shorter wavelength D. All of these
1861	Fertilizers are classified into	A. Two major categories<o:p></o:p> B. Three major categories<o:p></o:p> C. Four major categories<o:p></o:p> D. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri","sans-serif";mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:Arial;mso-bidi-theme-font:minor-bidi;mso-ansi-language: EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">None of above
1862	Which one of following is non polar	A. CH2Cl2 B. CCl4 C. CHCl3 D. CH3Cl
1863	Alpha hematite nano tubes show dimensional magnetic ordering at temperature laser than 300 K.	A. 0 B. 1 C. 2 D. 3
1864	Which of the following unit cells has least symmetry.	A. Monocline B. Cubic C. Triclinic D. Tetragona
1865	Which of the following properties of a system does not change in a state of equilibrium.	A. Density B. Pressure C. Colour D. All above properties
1866	Which of the following statement is not related with high quantum yield reasons.	A. Formation of reactive intermediates which may act as catalyst B. The active molecules may collide with other molecules and activates these molecules. C. The reaction may be exothermic and heat evolve may activate other molecule D. The primary photo chemical process may be reversed
1867	In statistical mechanics, there exists a function which contains all the information about a macroscopic system. This function is known as.	A. Eigen function B. Wave function C. Partition function D. Distribution function
		A. Adsorptoin on solids is reversible in

1868	Which of the following characteristics of adsorption is wrong.	A. Adsorption on solids is reversible in nature B. Adsorption, in general increase with increase in temperature. C. Adsorption is generally selective in nature. D. Both enthalpy and entropy of adsorption are negative
1869	Which of the following substance is a volatile metals.	A. Lead B. Zinc C. Mercury D. Sodium
1870	The phenomenon of x-ray diffraction was studied by	A. Huygen B. Bragg C. Max Planck D. None of above
1871	Formula of orthophosphoric acid	A. H_2PO_4 B. H_3PO_2 C. H_3PO_2 D. $H_2P_2O_5$
1872	Concentrated aqueous sodium hydroxide can separate a mixture of.	A. Al^{3+} and Sn^{2+} B. Al^{3+} and Fe^{3+} C. Al^{3+} and Zn^{2+} D. Zn^{2+} and Pb^{2+}
1873	Copper occurs in nature as.	A. Native B. Combined C. Both native and combined D. None of the above
1874	The pH of 0.001 N HCl is	A. 1 B. 2 C. 3 D. 4
1875	RNA is involved int eh synthesis of	A. Protein B. Nucleic acid C. Carbohydrates D. Fats
1876	In group 17, the element with highest first ionization enthalpy belongs to.	A. Period 1 B. Period 2 C. Period 7 D. Period 6
1877	Which show maximum number of oxidation states in 3d series.	A. Mn B. Ni C. Co D. Zn
1878	Which of the following statements is not related with chemical equilirbium.	A. The properties of the system become constant B. The equilibrium can be approached from either direction C. The chemical equilibrium is static is nature D. A catalyst can hasten the approach towards equilibrium
1879	Which of the following element has six electrons in the valance shell but cannot exhibit a maximum co valency of six.	A. Sulpher B. Oxygen C. Salenium D. Both A and B
1880	Zero group elements are called as	A. Inert gases B. Rare gases C. Noble gases D. All of above
1881	TLC belongs to which of the following chromatographic techniques.	A. Ion exchange B. Partities chromatography C. Adsorption chromatography D. Gel permeation
1882	The study of coiled long peptide chains of protein to give a 3 dimensional structure is the study of.	A. Primary structure B. Secondary structure. C. Tertiary structure D. Quaternary structure.
1883	The pH of 0.01 N NaOH is.	A. 12 B. 13 C. 14 D. 11
		A. 75% lime stone and 25% clay B. 25% lime stone and 75% clay

1884	The composition of mixture of clay and lime stone in the raw for cement material is.	B. 25% lime stone and 75% clay C. 15% lime stone and 55% clay D. 30% limes stone and 79% clay
1885	Which statement is true.	A. Resonance hybride are inherently unstable. B. Resonance hybride are more static than any individual resonance form C. Resonance hybride are average of all resoance forms resembling the more stabel forms D. None of the above
1886	The greater stability of bonzyl carbonium ion as compared to t-butyl carbonium ion is due to.	A. Inductive effect B. Resonance effect C. Electrometric effect D. All above
1887	Which of the following is biodegradable pollutant.	A. Domestic waste B. DDT C. Mercury salta D. Aluminum foil
1888	The rise of a liquid in capillary tube is due to.	A. Osmosis B. Diffusion C. Surface tension D. Viscosity
1889	Which of the following ions is smallest in size.	A. F- B. Cl- C. I- D. Br-
1890	For which of the following compounds is the rate of hydrolysis by aqueous alkali most likely to be independent of the hydroxide ion concentration.	A. 1-Chlorobutane B. 2- Bromobutane C. 1- Iodobutane D. 2- Bromo -2- methyl butane
1891	Temporary hardness of water is due to.	A. Bicarbonates of K B. Bicarbonates of Na C. Carbonates of Ca D. Bicarbonates of Ca
1892	The photoelectric effect is the ejection pf electrons from the surface of metal when light falls on it. Which of the following statements is not correct about the phenomenon.	A. The kinetic energy of photo electron depends upon the frequency of the incident radiation B. Electrons are ejected only when the frequency of light exceeds a certain threshold value C. The higher the energy of the photon greater the kinetic energy of the ejected electron. D. The threshold frequency of all the metals is same.
1893	Used in Geiger counter to detect radioactivity	A. He B. Ne C. Ar D. Kr
1894	Ozone is stratosphere is depleted by	A. CF2Cl2 B. C7F16 C. C6H6Cl6 D. C6F6
1895	A chromophore is an isolated fractional group which has	A. Coloured appearance B. Absorption in UV visible region C. Only sigma bonds D. Absorption in the region
1896	What refer to the casehardening process by which the carbon content of the steel ear the surface of a part is increased?	A. Carburizing<o:p></o:p> B. Annealing<o:p></o:p> C. Normalizing<o:p></o:p> D. None of these
1897	Which of the following statements is wrong.	A. Covenant compounds are generally soluble is polar solvents. B. Covalent compounds have low melting and boiling points C. Lower than that of separate H atoms D. Sometimes lower and sometimes

		higher than that of separate H
1898	Which of the following substance has been advocated as fuel of future.	A. O2 B. N2 C. H2 D. H2O
1899	Enfleurage process is used to extract the essential oils from	A. Back of plant B. Seeds of plant C. Leaves of plant D. Flowers of plant
1900	Correct order of increasing _____ I effect of groups is	A. ----- NO2 >----------CN > ----------COOH > ----------F B. --------F > ----------COOH>----------CN>--------NO2 C. -----F> ----------CN>--------NO2 > COOH D. -------CN > ----------COOH > ----------NO2 >---------F
1901	The special chrome steels of the stainless variety contain how many percent of chromium.	A. 4 to 8<o:p></o:p> B. 11 to 17<o:p></o:p> C. 9 tp 10 D. 12 to 15
1902	Nanoscience can be studied with the help of	A. Quantum mechanics B. Newtonain mechanics C. Macro dynamics D. Grophysics
1903	The IUPAC name of HCOOCH3 is.	A. Methoxy methanol B. Ethanoic acid C. Methyl methanoate D. Methoxy methane
1904	C is -2 butene on reaction with bromino give 2,3 -dibromobutane which is	A. Recemic mixture B. Meso isomer C. Dextoroisomer D. Levoisomer
1905	Which of the following molecule contains two dative bonds according to Lewis structure.	A. NH3 B. SO3 C. PCl5 D. BF3
1906	Which of the following basic process is involved in the separation of the complex mixture by chromatographic techique.	A. Partition B. Adsorptions C. Ion exchange D. All of the above processes
1907	Sugar and common salt in a mixture can be separated through then process of.	A. Sublimation B. Distillation C. Ion exchange D. Crystallization from solution in ethanol
1908	Which of the following agrochemical acts as pollutant.	A. Fertilizers B. Weedicides C. Herbicides D. All above
1909	20 micron =_____ nm	A. 20 x10⁻⁹ B. 20000 C. 200 D. 20x10⁹
1910	a -amino acids when heated alone form	A. Cyclic lactum B. a-b-unsaturated acid C. Fatty acids D. Diketopiperzines
1911	An example of acyclic monotterpenoid is	A. a -pinene B. Camphor C. Geranial D. Citral
1912	'A line, a point or a plane about which a symmetry operation is performed, is known as.	A. Symmetry operation B. Symmetry element C. Reflection D. Inversion
		A. Native

1913	Aluminum occurs in nature as.	B. Combined form C. Both native and combined D. Free
1914	The most harmful components of incomplete combination are generally groped as particulate polyclic matter organic (PPOM) These materials are derivatives of .	A. Benzene B. Naphthalene C. Benz a pyrene D. None of the above
1915	Which of the following source is commonly used as excitation source in fluorimeter.	A. Tungsten lamp B. Mercury vapour lamp C. Nernst vapur lamp D. Radio souse
1916	Which of the following specie is stronger acid than formic acid, HCOOH, in aqueous solution.	A. CH3COOH B. NH+4 C. H2SO-3 D. H4P2O7
1917	Which of the following acid radical give organic layer test.	A. Cl- B. CO3 C. I D. S²⁻
1918	Noble gases are sparingly soluble in water owing to.	A. Dipole -dipole interactions B. Dipole -induced dipole interactions C. Hydrogen bonding D. Induced dipole -instantaneous dipole interactions
1919	The most promising technique for solar production of electricity is.	A. Dry cell B. Battery C. Solar cell D. None of above
1920	An indicator for an acid base titration is a	A. Weak acid B. Weak base C. Strong acid D. Strong base E. Both A and B
1921	On hybridization of one s and one p orbitals we get.	A. Two mutualy perpendicular orbitals B. Two orbitals at 180^o C. Four orbitals directed tetrahedrally D. Three orbitals in a plane
1922	The most abundant metal in earth's crust is.	A. Fe B. Al C. Ti D. Ca
1923	In DTA , theriac effect may be exothermic of endothermic These are cause by	A. Fusion B. Crystal structure inversion C. Destruction of crystal lattice D. All of above
1924	Which one of the following woul dmake an S_N2 mechanism more likely	A. Bulky substituents near the halogen B. A polar solvent C. A tertiary carbocation intermediate D. A reactive nucleophile
1925	Which element are non metals.	A. N & P B. Sb & Bi C. As & Sb D. N & Bi
1926	Which of the following techniques involves ion exchange phenomenon.	A. Size exclusion chromatography B. Ion exchange chromatography C. GLC D. HPLC
1927	Pick out incorrect statement.	A. NF3 molecule has trigonal pyramidal structure. B. It is practically incoluble in water and is only hydrolyzed, an electric spark is passed through a mixture with water vapour. C. Dipole moment of NF3 is more than that of NH3 D. Nitrogn (III) oxide (N2O3) is an acidic oxide.
1928	In manufacturing of cement crystallization of amorphous dehydration products of clay	A. 500 ^oC to 800 ^oC B. 900 ^oC to 1200^oC C. 1250 ^oC ti 1400 ^oC D. 1000 to 1100 ^oC

		D. 1000 to 1100 ⁰C
1929	The bonding of transition metal complex was not well understand until the pioneer work of.	A. Ps JAISWAL B. GS MANKU C. BR thukral D. Alfred Weriner
1930	NH4OH in the presence of H2S is used as a group reagent for which of the following group.	A. Group I B. Group II C. Group III D. Group IV
1931	Which of the following has the highest melting point.	A. NaCl B. KCl C. MgO D. BaO
1932	Which of the following mixture is used as most popular flame in AAS.	A. Acetylene air B. Acetylene O2 C. Hydrogen air D. Hydrogen O2
1933	Which of the following statement is not correct with respect to group theory.	A. Two elements of a group combine to form a third element of a group B. An element combines with itself to form another element of the group. C. Each element of the group obey associative law of combination D. Each group element has no reciprocal
1934	In TGA, the width loss curve depends on the which instrumental factors.	A. Furance heating rate B. Recording or chart speed C. Furnace atmosphere D. All
1935	A type of a chemical bond which is formed by the mutual sharing of electrons between combining atoms of the same or different elements is called.	A. Ionic bond B. Covalent bond C. Co ordinate Covalent bond D. Metallic bond
1936	SAN is a polymer of	A. Styrene B. Acrylonitrile C. Both A and B D. Vinyl chloride
1937	Which of the following is not a characteristic of phthalocyanine dyes.	A. They are metal complex B. the are insoluble in water C. They have porphin nucleus D. They are used in photographic plates
1938	Poise is a unit of.	A. Refractive index B. Optical activity C. Fluidity D. Viscosity
1939	Which of the following is not obtained when Br2 is added to ethylene in the presence of aqueous NaCl solution.	A. Br CH2CH2Br B. Br CH2CH2Cl C. ClCH2 CH2Cl D. ClCH2CH2Cl
1940	Which of the following process is involved in the purification of crude metals.	A. Liquation process B. Oxidation process C. Distillation process D. Electro refining
1941	The property associated in thermometric tittration is	A. Change in weight B. Rate of change in weight C. Heat evolved or absorbed D. Change in temperature
1942	Variable oxidation states is shown by	A. Normal eleemnts B. Metallic elements C. Non metallic elements D. Transition elements
1943	Which of the following analytical techniques can be used to extract metal ion chelates.	A. Solvent extractions B. Evaporation C. GC D. Distillation
1944	The number of moles of solute dissolved in 1000 gram of the solvent is called	A. Formality B. Molality C. Molarity D. Mole fraction
1945	Potassium crystallizes in a body centered lattice. Hence, the coordination number of potassium in potassium metal is.	A. 4 B. 6 C. 8 D. 12

1946	The concept is also known as proton donor acceptor system.	A. Bronsted Lowery B. Lewis C. Lux Flood D. Usanovich
1947	In a system of designating wrought aluminum alloys. what does the second digit represents.	A. The purity of aluminum B. The identity of the alloy C. The modification of the alloy group or impurity limits D. None of above
1948	Which of the following cells is used to produce electricity from chemical reaction	A. Electroytic cell B. Galvanic cell C. Voltaic cell D. Fuel cell E. Both C and D
1949	The process of transfer of genetic message from DNA to m-RNA is known as	A. Refplication B. Translation C. Transcription D. Transference
1950	In terms of the amount of the substance adsorbed per gram of the adsorbent (x/m), and pressure p of the gas, the Freudlich adsorption isotherm is represented as.	A. x/m = k/pⁿ B. x/m = kpⁿ C. p = k (x/m)ⁿ D. x/m = (k/p)ⁿ
1951	If for a solution of an electrolyte. It is the transport number of the cation, then the transport number of the anion I, is equal to	A. t/2 B. 1 - t+ C. 1 + t+ D. (i -t) /2
1952	H-Bond has a preferred bonding direction like	A. Ionic bond B. Covalent bond C. Co ordinate bond D. None of these
1953	Which of the following statements is not correct with respect to limitations of flame photometry.	A. Low energy of the exciting source B. Liquid samples are generally used C. Can be employed for direct detection halides or inert gases D. It does not provide informatin about the molecular forms of metals.
1954	Which is not a pollutant from the exhaust of motor.	A. Hydrocarbons B. Carbon monoxide C. NOx D. Fly ash
1955	The electrophile in the sulphonation of benzene is.	A. SO3 B. SO3H C. HSO4 D. SO2
1956	Which of the following hydroxide is amphoteric.	A. B(OH)3 B. Al(OH)3 C. Ga (OH)3 D. In (OH)3
1957	High density polyethylene has which type of structure.	A. Linear B. Branch chain C. Cross linked D. Any one of these
1958	Which of the following solution has pH= 11?	A. 1 X 10⁻¹¹ m NaOH B. 1 x 10⁻¹¹ m HCl C. 1 x 10⁻³ M NaOH D. 1 X 10 ³ M NaOH
1959	Which two atoms of hydrogen combine to form a molecule of hydrogen gas. the energy of the hydrogen molecule is.	A. Higher than that of separate H atoms B. Equal to that of separate H atoms C. Lower than that of separate H atoms D. Sometimes lower and sometimes higher than that of separate H.
1960	Helium is used in weather balloons and airships instead of H2 becuse it is.	A. Lighter than hydrogen B. Incombustible C. More abundant than hydrogen D. Radiactive
1961	The property measured in TGA is	A. Change in weight B. Rate of change in weight C. Heat envolved and absorbed D. Change of temperature.
1962	In 1952 who popularized the use of CFT for inorganic chemist	A. Bethe B. Orge C. Van Vleck

		D. Werner
1963	Elements of group 14 have the electronic configuration of their outer shell as	A. ns2 np3 B. ns2 np2 C. ns2 np6 D. ns2
1964	The first step of formation of sugar is	A. Extraction B. Washing C. Cutting D. Clarifying
1965	What ASTM test for shear strength is designated for plastics.	A. D 732 B. D 790 C. D 695 D. D 638
1966	The different layers in graphite are held together by	A. Ionic bonding B. Metallic bonding C. Covalent bonding D. Van der Waals forces
1967	Aromatic amine (X) was treated with alcoholic potash and another compound (Y) when foul smelling gas was formed with formula C2H3N (Y) was formed by reacting a compound (Z) with Cl2 in the presence of slaked lime . The compound (Z) is	A. C6H5NC B. CHCl3 C. CH3CH2OH D. C6H5NH2
1968	In proper rotation (Cn) an object is rotated through an angle of.	A. a/n radians B. 2n/n radians C. 3n/n radians D. 4n/n radians
1969	The substance that can form the glassy non-crystalline structure is called.	A. Stabilizers B. Fluxes or modifiers C. Formers D. None of these
1970	The diameter of a bucky ball is about _____	A. o 1 A B. 1 nm C. o 100 A D. 10 nm
1971	Given A + 3B _____ 2C + D This reaction is first under with respect to reactant A and second order with respect to reactant B . If the concentration of A is doubled and the concentration of B is halved, the rate of the reaction would _____ by factor of _____	A. Increase ,2 B. Decrease ,2 C. Increase ,4 D. Decrease ,4
1972	Has the highest value of electronegativity	A. F B. Cl C. Br D. I
1973	The correct order of second ionization potential of carbon , nitrogen, oxygen and fluorine is.	A. C > N > O > F B. O > N > F > C C. O > F > N > C D. F > O > N > C
1974	The noble gas which was discovered first in the sun and then on the earth is.	A. Helium B. Neon C. Argon D. Xenon
1975	In each period the most electro negative element belongs to.	A. ^{Group -1} B. Group -17 C. Group -2 D. Group -18
1976	In reverse phase chromatography which of the analyte will be eluted more readily.	A. Polar B. Non polar C. Semi polar D. All above
1977	The nutrients which are required in very small amount for the normal growth of plants are called.	A. Nitrogenous fertilizers<o:p></o:p> B. Micronutrients<o:p></o:p> C. Phosphorus fertilizer<o:p></o:p> D. All of the above<o:p>

		</o:p></p>
1978	Which of the following represent the fuming sulphuric acid	A. H2S2O4 B. H2S2O3 C. H2S2O6 D. H2S2O7
1979	Dyes used in photographic plates to make them panchromatic is.	A. Cyanine dyes B. Azine dyes C. Phthalocyanine dyes D. Acridine dyes
1980	The diameter of fly ash particles is._____ micro meter	A. 5-10 B. 10-20 C. 20-30 D. 100
1981	The sum of pH and pOH is aqueous solution is equal to.	A. 14 B. 7 C. zero D. pKw
1982	The electrode Pt/Fe2+ (C1) Fe 3+ (C2) belong to the type.	A. Gas electrodes B. Inert metal electrodes C. Magam electrodes D. Metal metal insoluble salt electrode
1983	In which paper some additive is not added.	A. Carbon paper B. Filter paper C. Glazed paper D. Art paper
1984	Which of the following elements display maximum tnedency to form P Pi - p PI multiple bonds with itself and with carbon and oxygen.	A. N B. p C. Bi D. As
1985	HClO4, HNO3 and HCL are all strong acids in aqueous solution inglacial acetic acid medium, their acid strength is such that.	A. HClO4 > HCL > HNO3 B. HNO3 > HClO4 > HCl C. HCl > HClO4 > HNO3 D. HCl > HClO4 > HNO3
1986	Which of the following alkaline earth metals occurs in radioactive form in nature.	A. Ca B. Mg C. Ra D. Ba
1987	BCl3 is a planar molecule because B atom is.	A. sp2 hybridized B. Sp3 hybridized C. sp hybridzed D. sp3 d hybridized
1988	The theoretical plate in chromatography is represented by how many equilibrium step	A. One B. Two C. Three D. Four
1989	The electronic configuration of chromium is 4s1, 3d5, The elements tungsten (W) belongs to the same group and has atomic number 74.The configuration of its valence shell is.	A. 5s1, 4d5 B. 6s1, 5d5 C. 6s1, 5d6 D. 6s1, 5d4
1990	Equilibrium constant Kp and Kc are related as	A. Kc = Kp (RT)delta n B. Kp = Kc (RT)an C. Kp = (Kc/RT)delta n D. Kp -Kc = (RT)delta n
1991	Most electronegative element is.	A. C B. Si C. Pb D. Sn
1992	Stable metal ions strictures are.	A. Noble gas structure B. Is electron group structure C. Transition metal in structure D. All of the above
1993	Anything that influence the valence electrons will affect the chemistry of the element Which of the following factors does not affect the valency shell.	A. Valence principle quantum number in B. Nuclear charge (Z) C. Nuclear mass D. Number of core electrons
1994	Lux -Flood concept is a dono-acceptor system of.	A. Proton B. Electron pair C. Neurtron D. Oxide ion
		A. It has for arrangement of Na+ B. Na+ and Cl- ions have coordination number of 6:6

1995	Which of the following statement is incorrect about rock salt type	number of 6:6 C. A unit cell of NaCl metals have rock salt type structure. D. None of them
1996	Which of the following statement is not relevant with nitrous oxide.	A. It is a colorless and odourless gas B. It is non toxic gas C. It is present in the atmosphere in higher concentration D. It has high reactivity in the lower atmosphere
1997	Which of of the following statement is not correct in respect of Arrhenius concept.	A. This concept is applicable only for aqueous systems. B. Neutralization takes place in aqueous medium only C. H+ ion concept remain as such in water D. This concept is applicable for non aqueous system only.
1998	Which of the following statement is true.	A. Ferromagnetic separation is used to remove rion impurities from bauxite. B. Aluminium is an amphoteric element which means that it can act as an oxidizing agent and as a reducing agent C. Aluminium has a strong attinity for oxygen D. Aluminothermic reactions are endothermic
1999	Transition metal possess	A. Definite color B. Catalytic power C. Both A and B D. None of above
2000	Which of the following will be most effective int he coagulntion of Fe (OH)3 sol.	A. NaCl B. MgSO4 C. AlCl3 D. Mg3 (PO4)2
2001	The expected specific waste of paper and allied products industry is.	A. Chemicals B. Paper and fibre residues C. Links D. All above
2002	Which of the following elements has the highest third ionization energy.	A. Sodium B. Magnesium C. Aluminum D. Silicon
2003	Lother Meyer plotted a graph showing variation of.	A. Atomic volume with increase in atomic number B. Atomic volume with increase in atomic weight C. Atomic redii with increase in atomic weight. D. Atomic weight which increase in atomic number
2004	Arrange the hydrides group 15 in the order of increasing boiling point.	A. PH3 < AsH3 <SbH3 < BiH3 <NH3 B. PH3 <AsH3< SbH3< NH3 < BiH3 C. PH3 < AsH3<NH3 <SbH3<BiH3 D. NH3 < PH3 < AsH3 < Sb H3 < BiH3
2005	AlCl3 fumes in air because of.	A. Hydrolysis B. Dehydration C. Hydration D. Oxidation
2006	Which of the following halide has lowest melting point.	A. NaCl B. NaF C. NaBr D. NaI
2007	Which of the following interaction is involved in solid phase extraction technique.	A. Van der Waals forces B. Dipolar attraction C. H bonding D. All of above
2008	Point out the incorrect statement.	A. Rate law is an experimental fact whereas law of mass action is a theoretical in nature. B. Rate law is always different from the expression of law of mass action C. Rate law is more informativeness than law of mass action

		law of mass action D. Order of the reaction is equal to the sum of the exponents of concentration terms in the case law.
2009	Which of the following energy is trapped by the autotrophic organisms.	A. Mechanical energy B. Electrical energy C. Radiant energy D. Electronic energy
2010	The thermal conductivity of an SWNT along length is _____ watt/(m.k)	A. 35 B. 330 C. 386 D. 3500
2011	The noble gas used or treatment of cancer is	A. Helium B. Argon C. Radon D. Krypton
2012	Glass electrode cannot be used to measure the pH of pure.	A. Acetic acid B. Ethyl alcohol C. Gelatin D. All above
2013	Which of the following relation corresponds to Faraday law of electrolysis.	A. m = ZI t B. E = mc² C. E = ho D. None of the above
2014	In Dannis's method the end of the copper caps into which graphite electrode are fixed with cement.	A. Portiant B. Bakelite C. Asbestos D. All of above
2015	1 nanometre =_____ cm	A. 10⁻⁹ B. 10⁻⁸ C. 10⁻⁷ D. 10⁻⁶
2016	Colour in transition metal compounds in attributed to	A. Small sized metal ions B. Absorption of light in UV region C. Complete ns sub shell D. incomplete (n-1) sub shell
2017	The constant temperature and pressure,the rates of effusion of various gases vary inversely as square root of their denature. This is a statement of.	A. Boyle's law B. Charles's law C. Graham's law D. Daltons's law
2018	Which of the following detector is used for compounds containing electronegative atoms.	A. Mass specdtrometer B. ECD C. TCD D. UV-detector
2019	Types of carides	A. Ionic carides B. Covalent carbides C. Interstitial carbides D. All above
2020	The geometry of Xe F2 is	A. Triangular planar B. Square planar C. Linear D. Trigonal bipyramidal
2021	Monomer of Nylon -6 is	A. Adipic acid B. Hexamethylenediamine C. Caprolactam D. All of these
2022	The entropy change accompanying any physical or chemical transformation approaches zero as T approaches zero. This statement refers to.	A. Helmholtz law B. Third law of thermodynamics C. Second law of thermodynamics D. Nernat heat theorem
2023	In smelting process the ore is mixed with	A. Silica B. Coke C. Limestone D. All
2024	The process in which ore is heated generally in the presence of air, at temperature below their melting points is called.	A. Calcination B. Roasting C. Fourth floatation D. besemerization
2025	What is clinker.	A. Roasted calcareous material B. Roasted argillaceous material C. Roasted calcareous and argillaceous material D. Roasted gypsum

		D. Roasted gypsum
2026	Oil of turpentine contains.	A. a-pinene B. p- pinene C. Both A and B D. None of these
2027	The tyndall effect is not observed in	A. Suspensions B. Emulsions C. Colloidal solutions D. True solutions
2028	The correct order of increasing polar character is.	A. H2O < NHE < H2S < HF B. H2S < NH3 < H2O < HF C. NHE < H2O < HF < H2O D. HF < H2O < NH3 < H2S
2029	Which of the following is the best indicator for titration of CH3COH with NaOH	A. Methyl orange B. Methyl red C. Phenophthalein D. Eosin
2030	Which of the following is class of nanorods	A. metals B. alloys C. Metal oxide and Metal sulphite D. All of the above
2031	In urea the amount of nitrogen is	A. 82.0% B. 46.0% C. 33.0% D. 21.0%
2032	The process in which ore is heated generally in the absence of air, to expel water from a hydrated oxide at temperature below their melting points is called.	A. calcination B. Roasting C. Froth floatation D. Bessemerization
2033	For the respiration of sea divers mixture is used.	A. He & O2 B. Ar & O2 C. Ne & O2 D. Kr & O2
2034	Artificial nitrogen fixation may occur by the formation of.	A. Nitric acid B. Ammonia C. Nitrides D. Any of above
2035	Which of the following have identical bond order.	A. CN- and O2- B. CN - and NO+ C. O2- and CN+ D. NO+ and CN+
2036	The key element to be considered when evaluating a health hazard is.	A. The amount of material the employee is exposed. B. The total time of exposure C. The toxicity of the substance D. All above
2037	Which of the following metals form volatile carboyl with CO below 80 ^{o}C	A. Cu B. Fe C. CO D. Ni
2038	Which of the following substance is released into environment in the nuclear power plants.	A. Iodine -131 B. Argon - 41 C. Sr-90 D. Cs- 137 E. All above
2039	The separation efficiency of a column can be expressed in terms of number of.	A. Solvents used B. Theoretical plates C. Stationary phases D. Mobile phases
2040	Enzymatic action is heat at a fixed	A. Temparature B. pH C. Both of these D. None of these
2041	Covalent compound are soluble in	A. Polar solvents B. Non polar solvent C. Concentrated acids D. All solvent
2042	When the concentration of reactant molecules is increased the rate of reaction increases. The best explanation is As the reactant concentration increases.	A. The average kinetic energy of molecules increases. B. The frequency of molecular collisions increases C. The rate constant increase D. The activation energy increases

		D. The activation energy increases
2043	If 20 ml of 0.5 N salt solution is diluted in one litre. what is the new concentration.	A. 0.01 N B. 0.001 N C. 1 N D. 10 N
2044	Enantiomers have which of the following characteristics.	A. Rotate ordinary light B. Have the same melting point C. Are superimposable mirror images D. React with optically active molecule at the same rate
2045	Which of tetra chloride is resistant to hydrolysis.	A. CCl4 B. SiCl4 C. GeCl4 D. SnCl4
2046	Aluminium is used for.	A. Making ultensile & framea B. Making alloys C. Reducing agent D. All above
2047	Which of the following products is obtained when but 2-ene is treated with perchloric acid.	A. CH3CHO only B. CH3COOH only C. CH3CHO and CH3COOH D. CH3CH2COOH + HCOOH
2048	The angle between corresponding planes forming the external surfaces of the crystal remains constant for a given substances This is known as.	A. Steno's law B. Henry's law C. Bragg law D. Pascal law
2049	Which of the following is always true for the adiabatic expansion of gas.	A. Temperature rises B. Pressure rises C. W=0 D. Q = 0
2050	Among the elements of third period, the element with lowest boiling point belongs to group.	A. 1 B. 14 C. 16 D. 18
2051	Which one has a co ordinate bond.	A. Al2Cl6 B. BF3 C. NaCl D. O2
2052	The current voltage characteristics forms the basis of.	A. Thermal analysis B. Potentiometry C. Polarography D. Colorimetry
2053	The compound contains two types of X and Y its crystal structure is a cubic lattice with X-atoms at the corners of the unit cells and Y-atom t the body centre, The simplest formulae of this compound is.	A. X2Y B. XY C. XY2 D. X8Y
2054	Which of the following concentration term is used in respect of standard solutions.	A. Normality B. Formality C. Molarity D. All of above
2055	Which of the following statement is not true for carbon.	A. Its forms compounds with multiple bonds B. Its ionization energy is very high C. It undergoes catanation D. It shows inert pair effect
2056	Which of the following impurities are present with the bauxite.	A. Silica B. Ferric oxide C. Alumina D. Both silica and ferric oxide
2057	The IUPAC name of HOCCH2CH2CH2COOH is	A. 4- formylbutanoic acid B. 5- formylpentanoic acid C. 4- carboxybutanal D. 5- carboxypentanal
2058	The chemical method of separation in which the analytes to be separated are distributed between two phases, one of which is stationary phase while the other moves in a definite direction This technique is known as.	A. Electrophoreals B. Chromatography C. Solvent extraction D. Catachreals
2059	Which among the following is least soluble in water.	A. NaF B. LiF C. KF D. CsF
		A. Corundum

2060	Oxides ores of Aluminium	A. Corundum B. Bauxite C. Diaspore D. All above
2061	The temperate of a gas below which only the gas cools when allowed to expand is know as.	A. Inversion temperature B. Ideal temperature C. Critical temperature D. Joule Thomaon temperatu4re
2062	What is considered as the general purpose oldest type and widely used case iron.	A. Grey iron B. Alloy iron C. Black iron D. Ductile iron
2063	Which isotope of hydrogen is radioactive in nature.	A. Protium and deuterium B. Tritium only C. Tritium and deuterium D. Only deuterium
2064	Which of the following statement is not true in case of catalytic reforming.	A. High temperature results in loss of reformate yield B. High naphthenic stock require high space velocity C. Presence of water decrease the hydrocracking activity. D. None of above
2065	In the fourth flotation process for the purification of ores, the ore particles floats because.	A. They are light B. Their surface is not easily wetted by water C. They bear electrostatic charge D. They are insoluble
2066	The criteria for aromatically is presence of	A. Uneaturations B. Cyclic structure C. Presence of 4nx electrons D. Presence of 4n + 2n electrons
2067	Pick out the incorrect statement.	A. Red phosphorus consists of a complied chain structure and black phosphorus has a layer structure. B. Nitrogen shows a little tendency for catenation, because N-N a single bond is very strong. C. The xamimum number of covalent bonds formed by nitrogen is four, since it has no d-orbitals in its valence shell D. The group 15 elements do not form M5+ ions, but +5 oxidatin state is realized only through covalent bonding.
2068	Which of the following metal ion cannot be catimated by gravimetric analysis.	A. K+ B. Ca²⁺ C. Al3+ D. Zn2+
2069	The phase rule was deduced by	A. Gibbs B. Thomson C. Troution D. Henry
2070	The group of steel are water hardened tool steels.	A. Groups S B. Groups W C. Groups O D. Group F
2071	Which of the following alkyl halide undergoes nucleophilic substitution reaction via the formation of a carboncation.	A. 1-chloro -2 methyl propane B. 2- chlro-2-methyl propane C. 2- chloro butane D. 1-Chloro, 3,3- dimethyl pentane
2072	Which of the following analytical technique is not concerned with atomic spectroscopy.	A. Flame photometry B. Flame emission spectrometry C. Atomic absorption spectrometry D. I-R spectrophotometry
2073	Compounds HCN and HNC are.	A. Tautomers B. Metamers C. Functional isomers D. Conformers
2074	Which of the following is a source of energy but does not cause pollution.	A. Gaslone B. Nuclear power plant C. Fossil fuels D. Sun
2075	Which of the following statement represent advantages of sanitary Landfill	A. Economical method B. Low initial investment C. Flexible daily capacity

		C. Flexible daily capacity D. All above
2076	_____is best in its cleaning action.	A. Soap B. Detergents C. Surfactant D. None of these
2077	To increase the life of filament and to low the heat conductivity a mixture in filled in electric bulb.	A. Ar & N2 B. Ar & Kr C. Kr & N2 D. Xe & N2
2078	Zero group of the periodic table consists of.	A. Four elements B. Five elements C. Six elements D. Eight elements
2079	Which of the following statements in incorrect.	A. Sodium hydride is ionic B. Beryllium chloride is covalent C. CCl4 gives a white ppt with AgNO3 solutions. D. Bonds in NaCl are non directional
2080	Identify an oxygenated cyclic terpenoid	A. a- pinene B. Camphor C. Citral D. Geranial
2081	Which of the following trihalides of nitrogen behaves as the weakest base.	A. NF3 B. NCl3 C. NBr3 D. Nl3
2082	In which of the following species the bonds are non directional.	A. NCl3 B. RbCl C. BeCl2 D. BCl3
2083	Which of the following salt is soluble in water.	A. BaCO3 B. SrCO3 C. CaCO3 D. K2CO3
2084	Atomic volume of C, N, O and F are in the order	A. C > N > F > O B. C > N > O > F C. F > O > N >C D. N > C > O >F
2085	Which of the following health effect is caused by cadmium.	A. Hypertension B. Cardiovascular problem C. Kidney damage D. All above
2086	Monel metal is a alloy of Ni which constrains Ni uptown	A. 50% B. 60% C. 70% D. 80%
2087	According to the Grothus -Draper law	A. Only absorbed light is effective in producing photo chemical changes B. Only light between certain wavelengths C. Light is effective only for photo chemical reactions is solution D. The light absorbed in proportional to its intensity
2088	Lewis concept explain the formation of	A. Ionic bond B. Covalent bond C. Co-ordinate bond D. Chemical bond
2089	The number of hydrogen bond present in G -C pair is	A. 1 B. 2 C. 4 D. 3
2090	d2 sp3 is oriented in a manner	A. Trigonal B. Tetrahedral C. Octahedral D. Trigonal bipyramidal
2091	The number of hydrogen bonds boding A _____ T pair is	A. 1 B. 2 C. 3 D. 4
		A. Bauxite B. Corundum

2092	Aluminum is usually extracted from	B. Corundum C. Feldepar D. Alumite
2093	Aque regia is made by dissolving a mixture of HNO3 and HCl with ratio.	A. 1 : 1 B. 1 : 3 C. 1 : 2 D. 1 :10
2094	Which of the following statements correct regarding copper.	A. It is used in electroplating B. Its salts are used as insecticides C. Its salts are used as coloring materials D. All are correct
2095	Which of the following is the cause of Brownian movement of colloidal particles.	A. Convection currents int he fluid B. Bombardment by the molecules of the dispersion medium C. Setting of dispersed phase under gravity. D. Thermal gradient in the medium
2096	Flourine differs from the other members of its own group due to.	A. Its small size and low bond energy B. Its higher electornegativity C. None-availability of d-orbitals in its valence shell D. All the above
2097	The number of optically active compounds in the isomers of C3H5Br3 is.	A. 1 B. 2 C. 3 D. 4
2098	Pick out the incorrect statement for ClF3	A. It has trigonal planar geometry B. It is used to make gaseous UF6 which is useful in making enriched U-235 fuel C. It is used as powerful fluorinating agent for inorganic compounds D. ClF2 has been used as fuel in short range rockets reating with hydrazine.
2099	The pKa of acetic acid is 4.74 which implies that.	A. pH of 1N solution is 4.74 B. At pH 4.74 the dissociation of acetic acid is maximum C. At pH 4.74 half of the acetic acid molecules are dissociated in the solution. D. At pH 4.74 the dissociation of acetic acid is minimum.
2100	Which of the following gas protects us form harmful effect of uv radiation.	A. SO2 B. NO2 C. CO D. O3
2101	Borax exist in the form	A. Ordinary borax B. Octahdral borax C. Borax glass D. All above
2102	Both the elements shows allotropy	A. B & Ai B. B & Si C. Al & Si D. Any of above
2103	Iron is said to be abundant in nature. About how many percent of the earth's crust is iron.	A. 10% B. 5% C. 20% D. 8%
2104	For highly paraffine , crude oil, the characterization factor will be in range of	A. 11.5-12.5 B. 12.5-13.0 C. 13.5-14.0 D. 13.4-15.0
2105	Example of linear geometry	A. XeF2 B. F2 and HgCl2 C. CdI2 AND Ag Cl3 D. All of the above
2106	Which of the following ahs non zero dipole moment.	A. NH3 B. SF6 C. BF3 D. CO2
2107	pH of pure water at 25 ^{o}C. kw = 1 x 10^{-4}	A. 0 B. 7 C. 14 D. None of above

2108	Fluorine form Fluorides reacting with	A. Metals B. Non metals C. Metalloide D. Any of above
2109	Cytosine a pyrimidine base pairs with	A. Guanine B. Thymine C. Adenine D. Any of these
2110	Dyes which can be applied to cellulosic fibre from water solution are called.	A. Ingrain dyes B. Substantive dyes C. Mordant dyes D. Vat dyes
2111	The rusting of iron is catalysed by which of the following.	A. Fe B. O2 C. Zn D. H+
2112	The energy gap between two bands so large that it effectively prevents the promotion of electron from the lower to the higher band such energy gap all called.	A. Ionization zone B. Dissociation zone C. Distinction zone D. Forbidden zone
2113	According to the VSEPR theory, the shape of the SO3 molecule is.	A. Pyramidal B. Tetrahedral C. Trigonal planar D. Distorted totrahedron
2114	The addition of As to Ge makes the latter a	A. Metallic conductor B. Ionic conductor C. Intrinaic conductor D. Extrinsic semiconductor
2115	Which of the following compound does not following octet rule.	A. CS2 B. PBr3 C. IBr D. Br F3
2116	What refer by the ability of steel to be hardened through to its centre in large section?	A. Malleability<o:p></o:p> B. Hardenability<o:p></o:p> C. Ductility<o:p></o:p> D. Rigidity<o:p></o:p>
2117	Which of the following technique is based on deposition of the analyte at appropriate electrode by the passage of the electric current.	A. Chromatography B. Dialysis C. Electrodeposition D. Solvent extraction
2118	Which of the following proportion is associated with the covalent nature of the compound.	A. It conducts electricity in molten stater or aqueous state B. It is a non electrolyte C. It has high m.p. D. It is a compound of a metal and non metal.
2119	If Steel is heated to a temperature well below red heat and to then cooled slowly the process is called.	A. Annealing B. Quenching C. Tempering D. Nitriding
2120	Which one of the following oxides is basic.	A. MnO B. Mn2O3 C. MnO2 D. Mn2O7
2121	Which of the following process is used for the removal of particulates.	A. Wet removal by precipitation B. Sedimentation C. Diffusion and impaction D. All above
		A. Molar rotation

2122	The rotation of plane polarized light when it posses through 1 dm of a solution containing 1 gram of the substance per cm3 of the solution is called.	B. Molar refraction C. Specific refraction D. Specific rotation
2123	Duralumin is an alloy of.	A. Mg + Al B. Ai+ Mg+Mn C. Mg + Ai + Cu D. Mg + Ai + Cu + Mn
2124	Which of the following is the correct order of interactions.	A. Covalent < hydrogen bonding < Van Der Waal's < dipole -dipole B. Van der Waal's < hydrogen bonding < dipole -dipole < covalent C. Van der Waal's < dipole -dipole < hydrogen bonding < covalent D. Dipole-dipole < Van der Waal's < hydrogen bonding < covalent
2125	The percentage of nitrogen in ammonium sulphate is_____%	A. 27 B. 21 C. 23 D. 19
2126	The link between classical thermodynamics and quantum mechanics is prevented by	A. Statiatical mechanics B. Boltzmann law C. Wave mechanics D. Matrix mechanics
2127	SO2 is generated from which of the following industry.	A. Drying and packing B. Paper C. Pulp D. paper and pulp
2128	The instrument used for measuring fluorescence is known as.	A. Fluorimeter B. Potentiometer C. Flame photometer D. Mass spectrometer
2129	Which of the following methods is the most common method for separation of liquid components from a mixture.	A. Dialysis B. Solvent extraction C. Precipitation D. Distillation
2130	The pink colour of phenolphthalen in in basic medium is due to the	A. Cationic form B. Anionic form C. Natural form D. OH- ions of the base
2131	The molecule returns from the first excited triplet sate to the ground state singlet. The light emitted in known as.	A. Inter system crossing B. Phosphorescence C. Fiuoreacence D. Quenching
2132	In the Lewis formula of which of the following species, the number of single double and dative bonds are equal	A. N2O5 B. HNO3 C. SO2 D. SOCl2
2133	Which of the following statements is not a part of Bohr's theory of the hydrogen atom.	A. An electron in an atom revolves aroung the nucleus only in circular paths. B. An electron does not absorb energy in the stationary orbit C. An electron does not emit energy in the stationary orbit D. Energy is emitted or absorbed in a discrete amount from the stationary orbit
2134	Which of the following techniques is bulk technique.	A. Powder XRD B. Single Crystal XRD C. SEM D. TEM
2135	Which of the following statements is not true with reference to ionic conductors.	A. Ionic conductance is due to movement of the ions B. It involves the transfer of matter C. It involves oxidation reduction reactions D. It decreases with rise in temperature.
2136	Burning of fossil fuels is the main sources of which of the following pollutant.	A. Nitrogen oxide B. Nitric oxide C. Nitrous oxide D. Sulphur dioxide
2137	The migration of positively charged colloidal particles, under an electrical field , towards the cathode is called.	A. Cataphoresis B. Electroamosis C. Sedimentation D. Electrodialysis

2138	The most important problem regarding nano chemistry	A. Elucidation of relationship between also and chemical reactivity of particle B. Determination of size of particle C. Determination of reactivity of particle D. Determination of physical properties of nano particles.
2139	The dyes which are produced on the fibre in suit by reactions are known as.	A. Mordant dyes B. Fast dyme C. Ingrain dyes D. Disperse dyes
2140	All cycle engines working reversibly between same temperature of source and sink have the same efficiency This is the statement for the.	A. Carnot cycle B. Carnot theorem C. Narnst theorem D. Second law of thermodynamics
2141	Water pollution is due to	A. Agricultural discharges B. Swages and other wastes C. Industrial effects D. All the above
2142	Which of the following will exhibit variable electro Valency due to intert pair effect.	A. Fe B. Sn C. K D. Both Fe and Sn
2143	The alpha iron will become paramagnetic at temperature above	A. 770 C B. 550 C C. 660 C D. 440 C
2144	Among oxides of nitrogen all are gases except.	A. N_2O_5 B. N_2O C. NO D. N_2O_3
2145	Ionic bond are also forces called as.	A. Polar bond B. Electrovalent bond C. None polar bond D. Both A and B
2146	Beillstein test is used for.	A. Cl B. N_2 C. CO_2 D. Na
2147	The Lambert beer law states that	A. Transmission is directly proportional to path length B. Transmission is directly proportional to concentration C. Absorbance is inversely proportional to transmission D. Absorbance is directly proportional to concentration.
2148	Sodium metal cannot be stored under	A. Hexane B. Benzene C. Kerosene D. Ethanol
2149	An example of cyclic polyterpenoid is	A. Myrcene B. Alcoholic C. Synthetic rubber D. Natural rubber
2150	Which of the group 13 element does not form M (III) idodie.	A. Al B. Ga C. Ti D. In
2151	Which of the following is not a characteristic of covalent compound.	A. They have low melting and boiling points. B. They ionize on dissolution in polar solvents C. Their molecules have definite geometry D. They are generally insoluble in water
2152	Monomer of natural rubber is	A. 1,3-Butadiene B. 2-Methyl -1,3-butadiene C. 1,2 -Butadiene D. 1,3 - Pentadiene
2153	An stereospecific enzyme in one which catalyses	A. Formation of one stercolsomer B. Reaction of one stereoisomer only C. Both of these D. None of these

2154	The following alloys are the chief alloys that are die cast except.	A. Zinc alloys B. Magnesium alloys C. Manganese alloys D. Nickel alloys
2155	Which of the following pollutant is not secondary pollutnat.	A. SO3 B. NO2 C. SO2 D. Ozone
2156	Most effective pesticide is	A. Carbonates<o:p></o:p> B. Organophosphates<o:p></o:p> C. Organ chlorines<o:p></o:p> D. All of these<o:p></o:p>
2157	Carbon dioxide content in atmosphere is	A. 0.0034% B. 0.034 % C. 0.34 % D. 3.4 %
2158	Which of the following carbides reacts with H2O to form propane.	A. Al4C3 B. CaC2 C. SiC2 D. Sic
2159	An example of acyclic polytropenoid is	A. Myrcone B. Buna -S C. Synthetic rubber D. Natural rubber
2160	Fluorine does not show variable oxidation state because of.	A. its high electronegativity B. Its small sixe C. low dissociation energy of F-F bond D. Non availability of d-orbitals
2161	The bond formed by complete transfer of electrons from electropositive to more electronegative atom is called.	A. Ionic bond B. Covalent bond C. Mettalic bond D. Co ordinates bond
2162	CFT was originally applied to.	A. Ionic crystal B. Liquid crystal C. Solid crystal D. All above
2163	The co ordination number of atoms in a hexagonal closed packed structure is	A. 2 B. 6 C. 12 D. 4
2164	The entropy of the universe	A. Tends towards a maximum B. Tend towards a maximum C. Tends to be zero D. Remains constatn
2165	Coulometry is based on the measurement of	A. Electrical current B. Electrical potential C. Electrical conductance D. Dielectric constant
2166	The dipole moments of the given species are such that.	A. BF3 > NF3 > NH3 B. NF3 > BF3 > NH3 C. NHE > NF3 >BF3 D. NH3 > BF3 > NF3
2167	Solar energy mainly light originates from sun due to.	A. Addition relations B. Displacement reactions C. Thermonuclear reactions D. Substitution reactions
2168	Used in TV sets and sound movies to give ready response to electrical potential	A. He B. Ne C. Ar D. Kr
	Organic farming is the technique of raising crops through uses of	A. Manures

2169 Organic farming is the technique of raising crops through uses of:

B. Biofertilizers
C. Resistant varieties
D. All of these