

## PPSC Chemistry Part III Inorganic Chemistry Online Test

Sr	Questions	Answers Choice
1	Noble gases are sparingly soluble in water owing to.	A. Dipole -dipole interactions B. Dipole -induced dipole interactions C. Hydrogen bonding D. Induced dipole -instantaneous dipole interactions
2	What is considered as the general purpose oldest type and widely used case iron.	A. Grey iron B. Alloy iron C. Black iron D. Ductile iron
3	Which of the following unit cells has least symmetry.	A. Monocline B. Cubic C. Triclinic D. Tetragona
4	In Dumas method, the volume of the gas collected is equivalent to which of the following gases set free from the compound.	A. Ammonia B. O <sub>2</sub> C. N <sub>2</sub> D. NO
5	Inert pair effect is that	A. When an element shows inertness in chemical combination B. When higher oxidation state is more stable than lower oxidation state C. When an electron pair is present on the atom of an element D. When two s -electrons or outermost shell remain paired and do not participate in bonding.
6	In a one -component system the maximum number of phase that can consist in equilibrium is.	A. 1 B. 2 C. 3 D. 4
7	Which is true for DDT it is.	A. <p style="margin: 0;">Not a pollutant</p> B. <p style="margin: 0;">An antibiotic</p> C. <p style="margin: 0;">A non degradable pollutant</p> D. <p style="margin: 0;">A pesticide</p>
8	Glycine reacts with nitrous acid to form	A. Methyl amino B. Acetic acid C. Zwitter ion D. Glycollic acid
9	An example of acyclic monoterpenoid is	A. a -pinene B. Camphor C. Geranial D. Citral
10	By applying an external force the ionic solid can be easily broken to powder form so the ionic solid are highly	A. Hard B. Brittle C. Tough D. Soft

A.  $\text{CH}_4\text{N}_2$  and  $\text{CO}_2$

11	Which one of the following set of raw material is most suitable for manufacture of urea.	<p>A. <math>\text{H}_2\text{CO}_2</math> and <math>\text{H}_2\text{O}</math></p> <p>B. <math>\text{H}_2\text{O}</math>, <math>\text{N}_2</math> and <math>\text{H}_2</math></p> <p>C. <math>\text{H}_2\text{O}</math>, <math>\text{N}_2</math> and <math>\text{KCl}</math></p> <p>D. <math>\text{H}_2\text{CO}_2</math> and <math>\text{H}_2\text{O}</math></p>
12	The internal resistance to flow possessed by a liquid is called its.	<p>A. Fluidity</p> <p>B. Viscosity</p> <p>C. Surface tension</p> <p>D. Turbidity</p>
13	The relative lowering of vapour pressure of a solution on the addition of non-volatile solute.	<p>A. Is equal to the mole fraction of solute</p> <p>B. Is equal to the sum of the mole fraction of the solute and solvent</p> <p>C. Depends upon the nature of the solute</p> <p>D. Depends upon the mole fraction of the solvent</p>
14	Chlorine is used in	<p>A. Sterilization of water</p> <p>B. Extraction of gold</p> <p>C. Bleaching of cotton</p> <p>D. All above</p>
15	1 meter = _____ nm	<p>A. <math>10^9</math></p> <p>B. <math>10^{-9}</math></p> <p>C. <math>10^{10}</math></p> <p>D. <math>10^{-10}</math></p>
16	Anything that influences the valence electrons will affect the chemistry of the element. Which of the following factors does not affect the valency shell.	<p>A. Valence principle quantum number</p> <p>B. Nuclear charge (Z)</p> <p>C. Nuclear mass</p> <p>D. Number of core electrons</p>
17	An equal volume mixture explodes with violence	<p>A. <math>\text{H}_2</math> &amp; <math>\text{N}_2\text{O}</math></p> <p>B. <math>\text{H}_2</math> &amp; <math>\text{NO}</math></p> <p>C. <math>\text{H}_2</math> &amp; <math>\text{N}_2\text{O}_4</math></p> <p>D. <math>\text{H}_2</math> &amp; <math>\text{N}_2\text{O}_3</math></p>
18	Which of the following metals is the most abundant in the earth's crust.	<p>A. Mg</p> <p>B. Ca</p> <p>C. K</p> <p>D. Na</p>
19	Which of the following bonds will be non-polar.	<p>A. N - H</p> <p>B. O - H</p> <p>C. C - H</p> <p>D. Cl - Cl</p>
20	Which substance has the greatest lattice energy.	<p>A. CuBr</p> <p>B. MgO</p> <p>C. KI</p> <p>D. NaF</p>