

PPSC Chemistry Part II Organic Chemistry Online Test

Sr	Questions	Answers Choice
1	RNA is involved in the synthesis of	A. Protein B. Nucleic acid C. Carbohydrates D. Fats
2	One arm of each t-RNA terminates in the base sequence.	A. UGU B. GGC C. ACT D. CCA
3	In which period, the element with least ionization enthalpy belongs to	A. Group I B. Group 2 C. Group 17 D. Group 18
4	Which of the following mixture is used as most popular flame in AAS.	A. Acetylene air B. Acetylene O ₂ C. Hydrogen air D. Hydrogen O ₂
5	Which of the following statement is not related to the characteristics of gaseous state.	A. The inter molecular forces of attraction are not strong in gaseous state B. The gases do not have definite shape and volume C. The gases are characterized by low density. D. The gases have low comprehensibility
6	The reduction in ozone layer would lead to	A. Temperature changes B. Rainfall failure C. Increase uv radiation on earth D. All above
7	Argillaceous material does not include.	A. <p class="MsoNormal" style="margin-bottom:0in; margin-bottom:.0001pt; line-height: normal">Vlay</o:p></p> B. <p class="MsoNormal" style="margin-bottom:0in; margin-bottom:.0001pt; line-height: normal">Marine shells<o:p></o:p></p> C. <p class="MsoNormal" style="margin-bottom:0in; margin-bottom:.0001pt; line-height: normal">Slate<o:p></o:p></p> D. <p class="MsoNormal" style="margin-bottom:0in; margin-bottom:.0001pt; line-height: normal">Blast furnace slag<o:p></o:p></p>
8	Major achievement of CFT is	A. Interpreting the color B. Adsorption spectra C. Both A and B D. None of above
9	The oxidation Number of I in HIO ₄ is.	A. +6 B. +7 C. + 3 D. + 14
10	A chemical reaction resulting in a change in the electric charge on the reacting particles may be called as.	A. Add ion reaction B. Redox reaction C. Elimination reaction D. Chain reaction
11	The possible sub levels in the n = 4 energy level are.	A. s,p,d B. s,p,d,f C. s D. s,p
		A. C4 H10 B. C4 H10

12	Which of the following hydrocarbon cannot be obtained on reacting chloromethane with sodium metal in the presence of dry ether.	B. C ₂ H ₆ C. C ₂ H ₄ D. C ₃ H ₈
13	The Lewis formula of SOCl ₂ the total number of bond pairs and lone pairs of electrons around sulphur are.	A. 2,1 B. 2,2 C. 3,1 D. 3,0
14	The decreasing order of the second ionization energies of K, Ca and Ba is	A. K > Ca > Ba B. Ca > Ba > K C. Ba > K > Ca D. K > Ba > Ca
15	Which of the following is not alloy of aluminium.	A. Aluminium bronze B. Magnalum C. Duralumin D. Stellite
16	The sugar present in DNA is	A. D- Ribose B. D-Glucose C. 2- Deoxy D-Ribose D. 3-Deoxy D-ribose
17	The molarity of a 500 mL solution containing 4 g NaOH	A. 0.1 B. 0.2 C. 0.3 D. 0.4
18	The reason why phenylamine is a much weaker base than ammonia when each is in aqueous solution to that.	A. The ion pair of electron on two nitrogen atom of phenylamine is delocalised over the benzene ring. B. The phenylamine molecule is too large to capture hydrogen ion easily C. Phenylamine is much less soluble in water than is ammonia D. The benzene ring has a tendency to increase the acidity of its substituents.
19	Atomicity of which of the following pair of elements is not same as hydrogen.	A. Phosphorus, Nitrogen B. Nitrogen, Argon C. Nitrogen, iodine D. Iodine, sulphur
20	Which can be purified by sublimation	A. F ₂ B. Cl ₂ C. Be ₂ D. I ₂