

## PPSC Chemistry Full Book Test

Sr	Questions	Answers Choice
1	Strength of H bond in inter mediate between	A. Van der Waals forces and covalent bond B. lonic and covalent bond C. lonic and metallic bond D. Metallic and covalent
2	According to Arrhenius theory an acid is defined as substance which	A. Accepts and electron pair B. Donatees H+ ion in ammonia C. Contains Cl- ions D. Furnishes H3O + ion in water
3	In 1952 who popularized the use of CFT for inorganic chemist	A. Bethe B. Orge C. Van Vleck D. Werner
4	The electronic configuration of chromium (Z =24) in the ground state is.	A. [Ar] 4s <sup>2</sup> , 3d <sup>4</sup> B. [Ar] 3d <sup>6</sup> C. [Ar]4s <sup>1</sup> , 3d <sup>5</sup> D. [Ar]
5	When Phosphate rock Ca3(PO4) 2 is converted to phosphorus.	A. One of the products of the reaction is water     B. Sulphuric acid is added to generate insoluble calcium sulphate     C. Hydrogen is used to reduce the phosphate to phosphorus     D. Silica is added to form a calcium silicate slag
6	The light absorbed in UV and visible region causes.	<ul><li>A. Vibrational energy changes</li><li>B. Rotational energy changes</li><li>C. Electronic excitation</li><li>D. All of these</li></ul>
7	Which of the following statement is incorrect.	A. An alloy is a mixture of two or more metals B. An alloy is a mixture of two or more metal and non metal elements that have metallic properties C. An alloy has a fixed composition D. An amalgam is an alloy containing Hg
8	When a concentrated solute of an electrolyte is diluted.	A. Its specific conductance increases     B. Its equivalent conductance     decreases     C. The specific conductance     decreases and equivalent     conductance increases     D. Both specific and equivalent     conductance increase
9	BCl3 is an example of hybridization	A. sp B. sp2 C. sp3 D. None of above
10	Types of carides	<ul><li>A. lonic carides</li><li>B. Covalent carbides</li><li>C. Interstitial carbides</li><li>D. All above</li></ul>
11	A colloidal system in which both the dispersion phase and dispersed phase are liquid is.	A. Smoke B. Emulsion C. Whipped cream D. Mist
12	The number of mole of the solute dissolved per dm3 of the solution is called.	A. Molality B. Formality C. Normality D. Molarity

13	Which one of the following statements if false with respect to CFT.	A. In an octahederal crystal field, the d electron on a metal ion occupy the ex of orbitals before they occupy the t98 not of orbitals.  B. Diamagnetic metal ions cannot have an odd number of electrons C. Low spin complexes can be paramagnetic D. Low spin complexes contain strong field ligands.
14	Which of the following hydroxides is most stable.	A. Mg (OH)2 B. Ca(OH)2 C. Sr (OH)2 D. Ba (OH)2
15	The one in which the acceptor atom is of low positive charge, Large size and has several outer electrons which can be easily excited is a.	A. Soft base B. Hard Base C. Soft acid D. Hard acid
16	Which of the following is most soluble in water	A. CaSO4 B. Sr SO4 C. MgSO4 D. BaSO4
17	Which of the following process is involved in nitrogen flotation	A. Non symmetric fixation of nitrogen     B. Fixation by soil bacteria     C. Fixation by yeast     D. Fixation by blue green algae     E. All above
18	Which of the following contains both covalent and ionic bond.	A. CCI4 B. NH4CI C. CaCI2 D. H2O
19	The compounds whose formation require a host compound and a guest compound are called.	A. Exclusion compounds     B. Inclusion compounds     C. Crystal compounds     D. None of the above
20	Which of the following is the weakest base.	A. KOH B. NaOH C. LIOH D. RbOH