

MDCAT Chemistry Chapter 9 Period Online Test

Sr	Questions	Answers Choice
1	The shielding effect of inner electron is responsible for	A. Having no effect on ionization energy B. Decreasing ionization energy C. Increasing ionization energy D. Increasing electronegativity
2	Low IE is a symbol of	A. high electronegativity B. small size C. High electron affinity D. Metallic character
3	Covalent bonds are	A. directional B. Bidirectional C. Multidirectional D. Non directional
4	The ionization energy of hydrogen atom is	A. Zero B. 131.3kJmole C. 13.13kJ/mole D. 1313kJ/mole
5	In a period the atomic radii	A. increase B. decrease C. remain same D. first increase, then decreased
6	Geometry of simple molecule with sp2 hybridization	A. Triangular planar B. Trigonal C. Square planner D. Pyramidal
7	Mostly ionic compounds are produced between elements of group	A. IA and IIA B. IB and VIB C. IA. IIA and VII-A D. IA and IB
8	lonic bond is produced after complete transfer of	A. nucleus B. neutrons C. electrons D. protons
9	Which one of the following has zero dipole moment	A. NH3 B. CHCl3 C. H2O D. BF3
10	Carbon-Carbon double bond length in C3H6	A. 154 pm B. 134 pm C. 120 pm D. 105 pm
11	For formation of ionic bond, electronegativity difference should be	A. Equal to zero B. Equal to 0.5 C. More then 1.7 D. Less than 1.7
12	Polarity of a molecule is expressed in terms of	A. Bond strength B. Dipole moment C. Bond length D. Shape
13	pi-bond can be formed by sideways overlap of	A. s-orbital B. d-orbital C. p-orbital D. sp orbital
14	Elements having high ionization potential values are	A. metals B. non- metal C. liquids D. solid
15	what is the exact value of angle in BF3	A. 90 B. 104.51 C. 119.5

	D. 120°
A covalent bond may be	A. 100% covalent B. Partial ionic C. 100% ionic D. Both a and b
Among the following molecules, which one has coordinate covalent (dative) bond?	A. CCI4 B. CO2 C. CO D. CH4
What will be the shape of a molecule which contains two sigma bond pairs and one lone pair?	A. Linear B. V shape C. Tetragonal D. Triangular
The no. of lp's on oxygen in CO are	A. 1 B. 3 C. 4 D. 2
Which one of the followings has polar covalent bonds hut is overall nom-polar molecule:	A. HF B. CO ₂ C. CH ₄ D. N ₂
	Among the following molecules, which one has coordinate covalent (dative) bond? What will be the shape of a molecule which contains two sigma bond pairs and one lone pair? The no. of lp's on oxygen in CO are