

## MDCAT Chemistry Chapter 8 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The reaction which is responsible for the production of electricity in the voltaic cell is	A. Hydrolysis B. Oxidation C. Reduction D. Redox
2	If $\text{Cl}_2$ is passed through hot $\text{NaOH}$ . $\text{NaClO}_3$ is formed and the oxidation number of Cl changes from	A. -1 to 0 B. 0 to +5 C. 0 to -1 D. 0 to +1
3	Stronger the oxidizing agent, higher is	A. Redox potential B. Standard reduction potential C. Reduction potential D. $\text{Oxidation potential}$
4	If a salt bridge is removed from two half cells the emf is	A. Increased B. Decreased C. Dropped to zero D. Electrodes will be reversed
5	The element which has greatest value of Reduction potential is used as	A. Strongest reducing agent B. Weak oxidizing and strong reducing agent C. Strongest oxidizing agent D. None of these
6	Which of the following salts would give the same products irrespective of whether its molten form or concentrated aqueous solution is electrolysed?	A. Magnesium bromide B. Magnesium sulphate C. Copper sulphate D. Copper chloride
7	The electrochemical reactions occurring at both the electrodes along with the electrolytic conduction constitute	A. Oxidation B. reduction C. Redox reaction D. electrolysis
8	On ascending the electrochemical series strength as reducing agent	A. Increases B. Decreases C. Remains same D. not determinable
9	Which molecule is least ionic"	A. $\text{NaCl}$ B. $\text{HCl}$ C. $\text{HF}$ D. $\text{CsF}$
10	Coinage metals Cu, Ag and Au are the least reactive because they have	A. Negative reduction potential B. Negative oxidation potential C. Positive reduction potential D. Positive oxidation potential
11	Which of the following statements is not correct about galvanic cell?	A. Anode is negatively charged B. Cathode is positively charged C. Reduction occurs at anode D. Reduction occurs at cathode
12	If a strip of Cu metal is placed in a solution of $\text{FeSO}_4$	A. Cu will be deposited B. Cu and Fe both dissolve C. Fe is precipitated out D. No reaction take place
13	Stronger is the oxidizing agent, stronger is the	A. emf of cell B. Oxidation potential C. Reduction potential D. Reduction potential
14	During oxidation process, oxidation number of an element	A. Decreases B. Increases C. Remains constant D. Both a and b
		A. $\text{Zn}^{2+}$ ion is a powerful oxidising agent than $\text{H}^+$ ion B. $\text{H}^+$ ion is a powerful oxidising agent

15	Zinc reacts with dilute acids to liberate hydrogen. This is because:	<p>than Zn ion</p> <p>C. <math>\text{Zn}^{2+}</math> ion is a powerful reducing agent than <math>\text{H}^+</math> ion</p> <p>D. <math>\text{H}^+</math> ion is a powerful reducing agent than Zn- ion</p>
16	The oxidation state of carbon in $\text{C}_2\text{O}_4^{2-}$ is	<p>A. +4</p> <p>B. -4</p> <p>C. +3</p> <p>D. +2</p>
17	SHE acts as anode when connected with Cu electrode but act as cathode with Zn electrode	<p>A. Zn has less reduction potential than hydrogen and Cu</p> <p>B. Zn has high reduction potential than hydrogen and Cu</p> <p>C. Zn is below electrochemical series than hydrogen and Cu</p> <p>D. Zn has least tendency to lose electron</p>
18	In voltaic cell a salt bridge is used in order to	<p>A. Pass the electric current</p> <p>B. Prevent the flow of ions</p> <p>C. Mix solutions of two half cells</p> <p>D. Allow movement of ions between two cells</p>
19	The emf produced by galvanic cell is called	<p>A. Cell potential</p> <p>B. Oxidation potential</p> <p>C. Redox potential</p> <p>D. Reduction potential</p>
20	Geometry of $\text{NH}_3$ is	<p>A. Tetrahedral</p> <p>B. Square planar</p> <p>C. Pyramidal</p> <p>D. Linear</p>