

MDCAT Chemistry Chapter 8 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	Zinc reacts with dilute acids to liberate hydrogen. This is because:	A. Zn ²⁺ ion is a powerful oxidising agent than H ⁺ ion B. H ⁺ ion is a powerful oxidising agent than Zn ion C. Zn ²⁺ ion is a powerful reducing agent than H ⁺ ion D. H ⁺ ion is a powerful reducing agent than Zn- ion
2	If Cl ₂ is passed through hot NaOH. NaClO ₃ is formed and the oxidation number of Cl changes from	A. -1 to 0 B. 0 to +5 C. 0 to -1 D. 0 to +1
3	By the electrolysis of CuCl ₂ using inert electrodes of platinum which species is deposited at cathode	A. H ₂ B. O ₂ C. Cu D. Cl
4	The products of electrolysis of which of the followings are known	A. Fused electrolyte B. Aqueous solution of electrolyte C. Solid electrolyte D. Solid metal
5	The standard electrode potential of hydrogen is arbitrarily taken at 298k is	A. 1.00volt B. 0.10 volt C. 0.00 volt D. 10.0 volt
6	Rusting of iron metal Fe occurs when Fe gets converted into Fe ₂ O ₃ What happen with Fe?	A. Fe is neutralized B. Fe is sublimed C. Fe is reduced D. Fe is oxidized
7	The oxidation state of carbon in C ₂ O ⁻² is	A. +4 B. -4 C. +3 D. +2
8	Stronger the oxidizing agent, higher is	A. Redox potential B. Standard reduction potential C. Reduction potential D. Oxidation potential
9	Total number of valence electrons in CH ₄	A. 8 B. 9 C. 10 D. 12
10	In which molecule. all atoms are coplanar?	A. CH ₄ B. BF ₃ C. NH ₃ D. PH ₃
11	During oxidation process, oxidation number of an element	A. Decreases B. Increases C. Remains constant D. Both a and b
12	The reaction which is responsible for the production of electricity in the voltaic cell is	A. Hydrolysis B. Oxidation C. Reduction D. Redox
13	The electrochemical series is based on	A. pH scale B. Redox scale C. Hydrogen scale D. Arrhenius scale
14	In all oxidation reactions, atoms of an element in a chemical species lose electrons and increases their	A. Oxidation states B. Reduction states C. Electrode D. Negative charges

15	Which of following is oxidation state of oxygen in peroxides?	A. -2 B. 1/2 C. -1 D. +2
16	The reduction potentials of non-metals are A =+0.54V, B=+1.08V, C=+1.36V. D= +2.87V Which non -metal can displace all other from aqueous solution of their salts	A. A B. C C. B D. D
17	The cell which converts electrical energy to chemical energy is called	A. Electrochemical cell B. Voltaic cell C. Galvanic cell D. Down's cell
18	Electrolytic products of dilute aqueous solution of sodium sulphate is	A. Na. SO ₂ B. H ₂ . SO ₂ C. Na. O ₂ D. H ₂ ,O ₂
19	If a salt bridge is removed from two half cells the emf is	A. Increased B. Decreased C. Dropped to zero D. Electrodes will be reversed
20	The standard reduction potential of Zinc is	A. 0.76V B. 0.34 C. -0.34V D. -0.76V