

MDCAT Chemistry Chapter 8 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	The products of electrolysis of which of the followings are known	A. Fused electrolyte B. Aqueous solution of electrolyte C. Solid electrolyte D. Solid metal
2	In an electrochemical series, elements are arranged on the basis of	A. pH scale B. pKa scale C. pOH scale D. Hydrogen scale
3	The standard reduction potential of Zinc is	A. 0.76V B. 0.34 C. -0.34V D. -0.76V
4	Which of the following salts would give the same products irrespective of whether its molten form or concentrated aqueous solution is electrolysed?	A. Magnesium bromide B. Magnesium sulphate C. Copper sulphate D. Copper chloride
5	Rusting of iron metal Fe occurs when Fe gets converted into Fe ₂ O ₃ What happen with Fe?	A. Fe is neutralized B. Fe is sublimed C. Fe is reduced D. Fe is oxidized
6	The reduction potentials of non-metals are A =+0.54V, B=+1.08V, C=+1.36V. D= +2.87V Which non -metal can displace all other from aqueous solution of their salts	A. A B. C C. B D. D
7	The reaction which is responsible for the production of electricity in the voltaic cell is	A. Hydrolysis B. Oxidation C. Reduction D. Redox
8	Which of the following statements is not correct about galvanic cell?	A. Anode is negatively charged B. Cathode is positively charged C. Reduction occurs at anode D. Reduction occurs at cathode
9	The standard electrode potential of hydrogen is arbitrarily taken at 298k is	A. 1.00volt B. 0.10 volt C. 0.00 volt D. 10.0 volt
10	If a strip of Cu metal is placed in a solution of FeSO ₄	A. Cu will be deposited B. Cu and Fe both dissolve C. Fe is precipitated out D. No reaction take place
11	In all oxidation reactions, atoms of an element in a chemical species lose electrons and increases their	A. Oxidation states B. Reduction states C. Electrode D. Negative charges
12	Stronger the oxidizing agent, higher is	A. Redox potential B. Standard reduction potential C. Reduction potential D. $\text{Oxidation potential}$
13	Which of the following molecules has angle of 120°	A. BeCl ₂ B. BF ₃ C. CH ₄ D. NH ₃
14	Most reactive among the following	A. Li B. Mg C. Ca D. Na
15	The working condition/s for SHE	A. 1atm pressure B. 1M H ⁺ -solution C. 298K temperature D. All of these

16	If Cl ₂ is passed through hot NaOH. NaClO ₃ is formed and the oxidation number of Cl changes from	A. -1 to 0 B. 0 to +5 C. 0 to -1 D. 0 to +1
17	In MgCl ₂ , the oxidation state of Cl is	A. Zero B. -2 C. +2 D. -1
18	Only those metals can replace Hydrogen from dilute acids, which have	A. High negative reduction potential B. Low negative reduction potential C. High positive reduction potential D. low positive reduction potential
19	The value of oxidation number of chlorine in HClO is	A. +7 B. +5 C. -1 D. +3
20	The electrolyte used in fuel cell is	A. KOH B. NaCl(aq) C. NaNO ₃ D. Molten NaCl