

MDCAT Chemistry Chapter 8 Reaction Kinetics Online Test

Sr	Questions	Answers Choice
1	Which one of the following metals can replace the Copper from aqueous solution of its salt more easily?	A. Cd B. Fe C. Zn D. Na
2	The value of oxidation number of chlorine in HClO ₄ is	A. +7 B. +5 C. -1 D. +3
3	Only those metals can replace Hydrogen from dilute acids, which have	A. High negative reduction potential B. Low negative reduction potential C. High positive reduction potential D. low positive reduction potential
4	Stronger the oxidizing agent, higher is	A. Redox potential B. Standard reduction potential C. Reduction potential D. Oxidation potential
5	In SO ₄ ²⁻ the oxidation number of sulphur is	A. -8 B. -6 C. +8 D. +6
6	When a metal rod is dipped in its one molar ionic solution	A. Electricity is produced B. Electricity is consumed C. Redox reaction occurs D. Potential difference is set up
7	Which of the following best describes the shape and polarity of the carbon disulphide molecule?	A. Bent and polar B. Linear and non-polar C. Pyramidal and polar D. Bent and non-polar
8	Which of the following statements is not correct about galvanic cell?	A. Anode is negatively charged B. Cathode is positively charged C. Reduction occurs at anode D. Reduction occurs at cathode
9	The reaction which is responsible for the production of electricity in the voltaic cell is	A. Hydrolysis B. Oxidation C. Reduction D. Redox
10	The standard electrode potential of hydrogen is arbitrarily taken at 298K is	A. 1.00 volt B. 0.10 volt C. 0.00 volt D. 10.0 volt
11	Electrolytic products of dilute aqueous solution of sodium sulphate is	A. Na, SO ₂ B. H ₂ , SO ₂ C. Na, O ₂ D. H ₂ , O ₂
12	The potential difference set up at 25°C and 1 atm when electrode is dipped in its one molar ionic solution is called	A. Single electrode potential B. electrode potential C. Standard electrode potential D. Standard hydrogen electrode
13	Total number of valence electrons in CH ₄	A. 8 B. 9 C. 10 D. 12
14	Rusting of iron metal Fe occurs when Fe gets converted into Fe ₂ O ₃ . What happens with Fe?	A. Fe is neutralized B. Fe is sublimed C. Fe is reduced D. Fe is oxidized
15	Which of the following bonds is not present in NH ₄ Cl?	A. Ionic bond B. Covalent bond C. Coordinate covalent bond D. De-localized covalent bond

16	Which of the following is an application of electrochemical series	A. Prediction of the feasibility of chemical reaction B. Calculation of the cell voltage C. Prediction of reaction of metal with dilute acid D. All of the above
17	In which molecule, all atoms are coplanar?	A. CH ₄ B. BF ₃ C. NH ₃ D. PH ₃
18	Which molecule is least ionic"	A. NaCl B. HCL C. HF D. CsF
19	The standard reduction potential of Zinc is	A. 0.76V B. 0.34 C. -0.34V D. -0.76V
20	The potential of SHE is taken as zero which is a value	A. Reference B. Arbitrary C. Exact D. Experimental