

## MDCAT Chemistry Chapter 7 Chemical Equilibrium Online Test

Sr	Questions	Answers Choice
1	By convention, the standard heat of formation of all elements is assumed to be	A. Zero B. positive C. Negative D. Infinity
2	How much heat is absorbed by 100 g of water when its temperature decreases from 25°C to 5°C? (heat capacity is 4.2 $J/gK$ )	A. 84,000J B. 2000/4.2J C2000/4.2j D8400J
3	The net heat change in a chemical reaction is the same whether it is brought about in two or more different ways in one or several steps.it is known as	A. Henry's law B. Hess's law C. joule's law D. Law of conservation of energy
4	Most of the reactions which give stable products are	A. Endothermic B. Exothermic C. Isothermal D. Non of these
5	What is not correct about ΔHF	A. It is always negative B. Its value gives an idea about the relative stability of reactants and the products. C. Its value can be greater or less than zero D. Value depends upon nature of bonds
6	NaOH+HCl- NaCl+ H2O. Enthalpy change in the above reaction is called	<ul><li>A. Enthalpy of reaction</li><li>B. Enthalpy of Neutralisation</li><li>C. Enthalpy of formation</li><li>D. Enthalpy of combustion</li></ul>
7	ΔH° represent the enthalpy change at	A. 0°C and 1 atm pressure B. 25°Cand 1atm C. 0K and I atm pressure D. 25°C and 2 atm pressure
8	Choose from the followings the correct statement about Born Haber cycle	A. Born Haber cycle is different from Hess's law B. The energy changes in a cyclic process is not zero C. The lattice energy of crystalline substances can be calculated easily D. None
9	The value of $\Delta V$ being very small. The term P $\Delta V$ can be neglected for process involving	A. Liquid and gas B. Solids and gases C. Liquid and solid D. None of these
10	For an endothermic reaction, enthalpy of reactants	A. Is smaller than that of the products B. Is greater than that of the products C. Must be greater or smaller than that of the products D. Is equal to that of the products
11	The enthalpy of formation of a compound is	A. Positive B. Either positive or negative C. Negative D. None
12	One of the best applications of Hess's law to calculate the lattice energy of ionic compound is	A. Measurement of enthalpy change in a calorimeter B. Studying of first law of thermodynamics C. Measurement of a heat of formation of a compound D. Born-Haber cycle
13	Total heat content of a system is called	A. Internal energy B. Entropy C. Enthalpy D. All of these

14	What is correct about heat of combustion	A. It is applicable to gaseous substances only B. It is always negative C. It is always positive D. It is positive in some cases while negative in other
15	A state function which describes together the internal energy and product of pressure and volume is called	A. Enthalpy B. internal energy C. Work D. Kinetic energy
16	Enthalpy of a reaction can be measured by	A. Glass calorimeter B. Barometer C. Manometer D. Thermometer
17	Hess's law is analogous to	A. Law of heat summation B. law of increasing entropy C. Law of heat exchange D. lst law of thermodynamics
18	The exothermic process is	A. Evaporation B. Sublimation C. Respiration D. Boiling
19	Enthalpy of a system can be calculated by which of following relationship	A. q=ΔE B. q=m×S×ΔT C. q=pv D. q=m×v×ΔT
20	Calorie is equivalent to	A. 0.4184J B. 4.184J C. 418.4J D. 40.18J