MDCAT Chemistry Chapter 6 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	The concentration of product is increasing from 30 mole/dm3 to 40mol/dm3 in 0.5 sec then rate of reaction will be------moledm-3sec-1	A. 0 B. 20 C. 15 D. 25
2	Which of the following reactions are usually slow?	A. Neutralization of acids and bases B. Displacement Reactions C. Organic substitution reaction D. Free radical reactions
3	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of the above
4	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
5	Doubling the pressure in a liquid phase reaction	A. Will double the rex B. Will increase the rex C. Will decrease the rex D. Will not alter the concentration of reactant
6	Which of the following statement about the order of reaction is true?	A. The order of reaction can only be determined by experiment B. a second order reaction is also bimolecular C. The order of reaction is always non-zero D. The order of reaction increases with increasing temperature
7	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same D. falling curve
8	The reaction takes place among the molecules when they have:	A. Activation energy B. Properly oriented C. Concentrated D. Activation energy and proper orientation
9	By increasing the concentration of reactants, the rate of reaction	A. Decreases B. Increases C. Remains constant D. Not predicted
10	In the reaction A+B→ Products, if B is taken in excess, then it is an example of	A. Second order reaction B. zero order reaction C. Pseudo first order reaction D. first order reaction
11	In which of the following techniques rate of reaction is directly related with number ofions	A. Spectrometry B. Dilatometric method C. Conductometric method D. Refractometric method
12	Consider gas is measure in bars then the units of rate of reaction is	A. Mole dm-3 sec B. Bars sec C. Mole dm-3 sec-1 D. Bars sec-1
13	The reaction which is zero order	A. Decomposition ofN2O5 B. Formation of Glucose in plant C. Formation of FeI2 D. Chorination of methane in sunlight
		A. P is catalyst in this reaction B. P molecules do not possess sufficient energy to react

14	If the reaction "P+Q→R+S" is described as being of zero order with respect to P, it means that	C. The concentration of P does not change during the reaction D. The rate of reaction is independent of the concentration of P
15	If the rate of the reaction is equal to the rate constant, the order of the reaction is	A. 3 B. 1 C. 0 D. 2
16	Reaction kinetics is important to discover the---under which reaction will proceed most economically:	A. rate constant B. Conditions C. volume D. equilibrium point
17	All the Hydrolytic reactions are	A. First order B. Second order C. Third order D. pseudo-first order
18	For reaction of methane and chlorine light is not available then	A. Reaction will take place rapidly B. No Reaction take place C. Reaction occurs at double the rate D. May all cases occur
19	Unit of the rate constant depends upon the	A. Molecularity of reaction B. Order of reaction C. Concentration terms D. Number of reactants
20	Higher the surface area available for reaction	A. slower the reaction B. faster the reaction C. constant the reaction D. lower the Ea