MDCAT Chemistry Chapter 6 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Half-lives required to convert 100% reactant to produet for a first order reaction are	A. 10 B. 1000 C. 100 D. Infinity
2	In dilatometric method is directly proportional to extent of reaction	A. Change in concentration B. Change in pressure C. Chang in volume D. Change in temperature
3	The reaction takes place among the molecules when they have:	A. Activation energy B. Properly oriented C. Concentrated D. Activation energy and proper orientation
4	The rate of reaction between A and B increases by a factor of 100, when the concentration of A is increased 10 folds, the order of reaction with respect to A is	A. 10 B. 1 C. 4 D. 2
5	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of the above
6	For a chemical reaction in which one of the reactant also act as solvent, the order will be	A. First order B. Third order C. Second order D. pseudo-first order
7	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same D. falling curve
8	For a chemical reaction which can never be a fractional no	A. order B. molecularity C. half-life D. rate constant
9	In which of the following techniques rate of reaction is directly related with number ofions	A. Spectrometry B. Dilatometric method C. Conductometric method D. Refractometric method
10	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
11	Substance which is formed as well as consumed during a chemical reaction and have temporary existence.	A. Reactant B. product C. Catalyst D. Intermediate
12	When the concentration of reactants is taken as unity the rate of reaction is equal to	A. average rate B. concentratian of reactant C. instantaneous rate D. specific rate constant
13	The concentration of product is increasing from 30 mole/dm3 to 40mol/dm3 in 0.5 sec then rate of reaction will be------moledm-3sec-1	A. 0 B. 20 C. 15 D. 25
14	The rate of reaction for a reaction is 30 mol dm-3sec-1 if the product of concentration of 10.reactant is unity the specific rate constant is	A. 25 B. 2.5 C. 30 D. 15
15	When does average rate become equal to instantaneous rate of reaction	A. At the start of reaction B. time interval is zero C. at the end of reactior

		D. time interval approaches zero
16	For reaction of methane and chlorine light is not available then	A. Reaction will take place rapidly B. No Reaction take place C. Reaction occurs at double the rate D. May all cases occur
17	The reaction kinetics concerned with the	A. Rate of reaction B. Direction of reaction C. Factor effecting rate of reaction D. both a &b
18	The radioactive disintegration of 238U92 is	A. First order B. Second order C. Third order D. Zero order
19	The order of reaction provides valuable information about of reaction	A. Condition B. Concentration C. Mechanism D. Parameters
20	The conversion of molecules of A to B follows a second order kineties. Doubling the concentration of A will inerease the rate of formation of B by a factor of	A. 2 B. 4 C. 1/2 D. 1/4