

## MDCAT Chemistry Chapter 6 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Half-lives required to convert 100% reactant to product for a first order reaction are	A. 10 B. 1000 C. 100 D. Infinity
2	When the concentration of reactants is taken as unity the rate of reaction is equal to	A. average rate B. concentration of reactant C. instantaneous rate D. specific rate constant
3	For a chemical reaction to occur	A. The vessel shall be open B. Reacting molecules should have less energy than $E_a$ at time of collision C. Reacting molecules must be properly oriented and energy more than or equal to $E_a$ D. The reacting molecules must not collide with each other
4	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Exothermic C. Endothermic D. Exothermic and fast
5	Unit of the rate constant depends upon the	A. Molecularity of reaction B. Order of reaction C. Concentration terms D. Number of reactants
6	In dilatometric method is directly proportional to extent of reaction	A. Change in concentration B. Change in pressure C. Change in volume D. Change in temperature
7	Doubling the pressure in a liquid phase reaction	A. Will double the $r_{ex}$ B. Will increase the $r_{ex}$ C. Will decrease the $r_{ex}$ D. Will not alter the concentration of reactant
8	The radioactive disintegration of $^{238}\text{U}$ is	A. First order B. Second order C. Third order D. Zero order
9	If the rate of the reaction is equal to the rate constant, the order of the reaction is	A. 3 B. 1 C. 0 D. 2
10	Substance which is formed as well as consumed during a chemical reaction and have temporary existence.	A. Reactant B. product C. Catalyst D. Intermediate
11	The order of reaction provides valuable information about of reaction	A. Condition B. Concentration C. Mechanism D. Parameters
12	By increasing the concentration of reactants, the rate of reaction	A. Decreases B. Increases C. Remains constant D. Not predicted
13	When the concentration of product is increased the instantaneous rate of reaction with reference to reactants will be	A. Positive B. Negative C. the same D. falling curve
14	The collision which results in chemical reaction	A. Effective collision B. Ineffective collision C. Useless collision D. All of these

		D. All of the above
15	The reaction kinetics concerned with the	A. Rate of reaction B. Direction of reaction C. Factor effecting rate of reaction D. both a & b
16	The number of atoms or molecules whose concentrations determines the rate of a chemical reaction is called the	A. Molecularity of the reaction B. specific activity of the reaction C. Order of the reaction D. rate constant of the reaction
17	In the reaction $A+B \rightarrow \text{Products}$ , if B is taken in excess, then it is an example of	A. Second order reaction B. zero order reaction C. Pseudo first order reaction D. first order reaction
18	The slope of the graph is steepest at the beginning of reaction showing	A. Rapid decrease in concentration of reactants B. Rapid increase in concentration of reactants C. Fast rate of reaction D. All of the above
19	The increase in reaction rate as a result of increase in temperature from 10K to 90K is	A. 512 B. 256 C. 400 D. 112
20	The reaction which is zero order	A. Decomposition of $N_2O_5$ B. Formation of Glucose in plant C. Formation of $FeI_2$ D. Chlorination of methane in sunlight