## MDCAT Chemistry Chapter 6 Electrochemistry Online Test

| Sr | Questions | Answers Choice |
|---|---|---|
| 1 | Higher the surface area available for reaction | A. slower the reaction<br>B. faster the reaction<br>C. constant the reaction<br>D. lower the Ea |
| 2 | The rate of reaction for a reaction is 30 mol dm-3sec-1 if the product of concentration of 10.reactant is unity the specific rate constant is | A. 25<br>B. 2.5<br>C. 30<br>D. 15 |
| 3 | In the reaction A+B→ Products, if B is taken in excess, then it is an example of | A. Second order reaction<br>B. zero order reaction<br>C. Pseudo first order reaction<br>D. first order reaction |
| 4 | The rate of reaction between A and B increases by a factor of 100, when the concentration of A is increased 10 folds, the order of reaction with respect to A is | A. 10<br>B. 1<br>C. 4<br>D. 2 |
| 5 | Unit of the rate constant depends upon the | A. Molecularity of reaction<br>B. Order of reaction<br>C. Concentration terms<br>D. Number of reactants |
| 6 | When the concentration of reactants is taken as unity the rate of reaction is equal to | A. average rate<br>B. concentratian of reactant<br>C. instantaneous rate<br>D. specific rate constant |
| 7 | The reaction which is zero order | A. Decomposition ofN2O5<br>B. Formation of Glucose in plant<br>C. Formation of FeI2<br>D. Chorination of methane in sunlight |
| 8 | The number of atoms or molecules whose concentrations determines the rate of a chemical reaction is called the | A. Molecularity of the reaction<br>B. specific activity of the reaction<br>C. Order of the reaction<br>D. rate constant of the reaction |
| 9 | The number of reacting molecules whose concentration change during reaction is called | A. Activated molecule<br>B. Rate of reaction<br>C. Order of reaction<br>D. half-life |
| 10 | Spectrometry method is applicable if a reactant or a product absorbs radiation | A. Ultraviolet<br>B. Visible<br>C. Infrared<br>D. Any of these |
| 11 | The order of reaction provides valuable information about of reaction | A. Condition<br>B. Concentration<br>C. Mechanism<br>D. Parameters |
| 12 | Doubling the pressure in a liquid phase reaction | A. Will double the rex<br>B. Will increase the rex<br>C. Will decrease the rex<br>D. Will not alter the concentration of reactant |
| 13 | Amount of product formed increases with time, this statement is true for reactions-------with kinetics | A. 1s order<br>B. 3rd order<br>C. zero order<br>D. Any order |
| 14 | Half-lives required to convert 100% reactant to produet for a first order reaction are | A. 10<br>B. 1000<br>C. 100<br>D. Infinity |
| 15 | The collision which results in chemical reaction | A. Effective collision<br>B. Ineffective collision<br>C. Useless collision |

| 16 | Which property of liquid is measured by polarimeter | A. Conductance<br>B. Optical activity<br>C. Refractiye Indéx<br>D. Change in volume |
|----|-----------------------------------------------------|-----------------------------------------------------------------------------------|
| 17 | Substance which is formed as well as consumed during a chemical reaction and have temporary existence. | A. Reactant<br>B. product<br>C. Catalyst<br>D. Intermediate |
| 18 | If the energy of the activated complex lies close to energy of reactants, it means that reaction is | A. Slow<br>B. Exothermic<br>C. Endothermic<br>D. Exothermic and fast |
| 19 | The conversion of molecules of A to B follows a second order kineties. Doubling the concentration of A will inerease the rate of formation of B by a factor of | A. 2<br>B. 4<br>C. 1/2<br>D. 1/4 |
| 20 | The radioactive disintegration of $238U92$ is | A. First order<br>B. Second order<br>C. Third order<br>D. Zero order |