

MDCAT Chemistry Chapter 5 Chemical Energenitcs Online Test

Sr	Questions	Answers Choice
1	The crystal of diamond is	A. ionic B. molecular C. covalent D. metallic
2	The Cl ⁻ ion present at the corner of the unit cell is NaCl crystal, contributes	A. 1/8 th B. 1/4 th C. 1/2 th D. 1
3	In diamond, the carbon atoms are arranged in a	A. tetrahedral manner B. hexagonal manner C. square planar manner D. octahedral manner
4	The number of Na ⁺ , ions which surround each Cl ⁻ ion in the NaCl crystal lattice is	A. 8 B. 12 C. 6 D. 10
5	The existence of an element in more than one form is called	A. allotropy B. isomorphism C. polymorphism D. isotropy
6	One of the following is a ionic solid. Which is that?	A. Fe B. KBr C. Diamond D. Cr
7	The arrangement ABC, ABC is referred as	A. cubic close packing B. octahedral close packing C. hexagonal close packing D. tetrahedral close packing
8	Which attractive forces cause molecular solids to be formed?	A. Ionic B. Metallic C. Covalent D. van der Waals
9	Which of the following solids does not have a covalent bond?	A. Silica B. Copper C. Diamond D. Graphite
10	The hexagonal closed packing is associated with	A. Ag, Cu, Au B. Zn, Cd, Hg C. Li, Na, K D. NaCl, KBr
11	In crystal of sodium chloride, a Cl ⁻ ion present at the corner of cube is shared between how many cubes?	A. 8 B. 4 C. 6 D. 10
12	The examples of a hexagonal system is	A. sugar B. graphite(a=b not equal to c) {Alpha =beta. not equal to gamma} C. sulphur D. diamond
13	All the metal shine when they are freshly cut The reason is	A. the conductivity of the metal is increased B. the process of cutting gives energy to the metal atoms C. the electrons become less delocalized according to valance bond theory D. the electrons are excited at higher energy levels and emit the photons when they fall back
14	Polvmorphism is shown by AgNO3. Which one of the followinga options is true for AaNO3?	A. Orthorhombic and rhombohedral B. Cubic and orthorhombic

		C. Cubic and tetragonal D. Monoclinic and hexagonal
15	Ionic solids are characterized by which one of the following properties	A. moderately low pressure B. high vapour pressure C. good conductivity in solid state D. solubility in polar solvents
16	How many allotropic forms are present in carbon?	A. Two B. Four C. Three D. Five
17	in diamond a unit cell is tetrahedral and overall crystal structure is	A. face centred cubic B. body centred cubic C. tetrahedral D. hexagonal
18	Hardness of diamond is attributed to the	A. strength of the ionic bonds in the structure B. three-dimensional network of covalent bonds C. three-dimensional network of covalent bonds D. absence of valence electrons in carbon atoms
19	Plastics are amorphous solids and	A. have sharp melting points B. undergo clean cleavage when cut with knife C. do not undergo clean cleavage D. possess orderly arrangement over long distances
20	Which one is an isomorphous pair?	A. NaNO_3 , CaCO_3 , B. NaF , MgO C. K_2SO_4 , $\text{K}_2\text{Cr}_2\text{O}_7$ D. Zn , Cd