

MDCAT Chemistry Chapter 4 Chemical Bonding Online Test

Sr	Questions	Answers Choice
1	Liquid hydrocarbon is	A. methane B. propane C. ethane D. hexane
2	The boiling point of higher alkanes are greater than those of lower alkanes due to reason that	A. higher alkanes have greater number of atoms B. the polarizabilities of higher alkanes are greater C. higher alkanes have greater hydrogen bonding D. higher alkanes have zig-zag structures
3	The forces which are present between the ions and the water molecules are known as	A. dipole-induced dipole forces B. dipole-dipole forces C. ion-dipole forces D. London dispersion forces
4	Which of following factor affect vapour pressure of a liquid?	A. temperature B. inter molecules forces C. size of the molecules D. all of these
5	Strong dipole-dipole forces among the liquid molecules are responsible for	A. very high heat of vaporization B. very low heat of vaporization C. cannot be predicted D. negligible forces are these
6	Liquids evaporate at every temperature. When the temperature becomes constant for a liquid, then:	A. rate of evaporation is greater than the rate of condensation B. the rate of condensation is greater than the rate of evaporation C. The rate of condensation and evaporation become equal D. it depends upon the nature of the liquid
7	The B.P of glycerine at 760 torr pressure is	A. 200°C B. 290°C C. 250°C D. 262°C
8	Which of the following liquid has highest boiling point	A. HCl B. HBr C. H ₂ O D. Br ₂
9	The boiling point of glycerin at 1 atmospheric pressure is:	A. 290°C B. 390°C C. 190°C D. 210°C
10	The boiling points of the halogens	A. increases down the group B. decreases down the group C. remains constant D. can not be predicted
11	The weakest intermolecular forces present in a liquid may be	A. Dipole-induced dipole forces B. dipole-dipole forces C. instantaneous forces D. electrostatic forces between ions in a ionic solid
12	H-bonding is maximum in:	A. ethanol B. benzene C. diethyl ether D. water
13	The boiling point of H ₂ O is 100°C while that of C ₂ H ₅ -OH is 78.5°C. The reason is that:	A. H ₂ O molecules are small-sized B. the bond angles at oxygen atom are different C. C ₂ H ₅ -group is electron donating D. the number of H-bonds are greater in H ₂ O, than C ₂ H ₅ -OH

14	Hydrogen bonding is not present in which of following compound?	A. Ammonia B. Ethanol C. Ether D. Water
15	Dipole-induced dipole forces are also called	A. dipole-dipole forces B. ion-dipole forces C. Debye forces D. London-dispersion forces
16	The polarizabilities of elements mostly increase down the group due to the reason that	A. the atomic numbers increase B. number of protons increase C. number of shells increase along with increase of shielding effect D. the behaviour of the elements remain the same
17	Vapour pressure of a substance does not depend upon:	A. physical state of matter B. temperature C. intermolecular forces D. surface area
18	The nature of crystals formed due to London forces of interaction are	A. molecular B. metallic C. ionic D. covalent
19	The vapour pressure of a liquid depends upon	A. amount of the liquid B. surface area C. temperature D. size of container
20	Halogens form halogen acids. HF is the weakest among all of them This is due to the reason that	A. fluorine is a very small-sized atom B. fluorine is highly electronegative atom C. there is strong hydrogen bonding in HF D. the polarity of HF bond is less