

## MDCAT Chemistry Chapter 4 Chemical Bonding Online Test

Sr	Questions	Answers Choice
1	The nature of crystals formed due to London forces of interaction are	A. molecular B. metallic C. ionic D. covalent
2	Polarizability is responsible for intermolecular forces and it	A. increases down the group B. decreases down the group C. almost remains the same D. increased along a period
3	Strong dipole-dipole forces among the liquid molecules are responsible for	A. very high heat of vaporization B. very low heat of vaporization C. cannot be predicted D. negligible forces are these
4	The boiling points of the halogens	A. increases down the group B. decreases down the group C. remains constant D. can not be predicted
5	At freezing point of water, the density decreases due to	A. change of bond angles B. change of bond lengths C. cubic structure of ice D. empty spaces present in the structure of ice
6	The B.P of glycerine at 760 torr pressure is	A. 200°C B. 290°C C. 250°C D. 262°C
7	H-bonding is maximum in:	A. ethanol B. benzene C. diethyl ether D. water
8	Saturated hydrocarbons having carbon atoms more than 20 in a molecule are solids due to	A. higher densities B. higher molar masses C. the chain, are more zig-zag D. all are correct
9	Vapour pressure of a substance does not depend upon:	A. physical state of matter B. temperature C. intermolecular forces D. surface area
10	Ice floats on water because	A. the hydrogen bonding in ice is stronger than that of in water B. empty spaces are left in ice C. ice has two-dimensional structure D. the bond length of the oxygen and hydrogen bond is different in water and ice
11	The weakest intermolecular forces present in a liquid may be	A. Dipole-induced dipole forces B. dipole-dipole forces C. instantaneous forces D. electrostatic forces between ions in a ionic solid
12	The B.P. of compound is mostly raised by	A. dipole-induced dipole interactions B. london dispersion forces C. intramolecular H-bonding D. intermolecular H-bonding
13	The long chains of amino acids are coiled around one another into a spiral by	A. ionic bond B. Van der Waal's forces C. hydrogen bonding D. overlapping of orbitals
14	The forces which are present between the ions and the water molecules are known as	A. dipole-induced dipole forces B. dipole-dipole forces C. ion-dipole forces D. London dispersion forces

15	Oxygen and sulphur are present in VI-A group of the periodic table The hydride of oxygen i.e., H <sub>2</sub> O is liquid at room temperature but the hydride of sulphur (H <sub>2</sub> S) is a gas. This is due to	<p>A. greater bond angle of water than H<sub>2</sub>S</p> <p>B. greater bond lengths in HS than H<sub>2</sub>O</p> <p>C. hydrogen bonding in water</p> <p>D. acidic character of HS</p>
16	Liquid hydrocarbon is	<p>A. methane</p> <p>B. propane</p> <p>C. ethane</p> <p>D. hexane</p>
17	Dipole-induced dipole forces are also called	<p>A. dipole-dipole forces</p> <p>B. ion-dipole forces</p> <p>C. Debye forces</p> <p>D. London-dispersion forces</p>
18	The polarizabilities of elements mostly increase down the group due to the reason that	<p>A. the atomic numbers increase</p> <p>B. number of protons increase</p> <p>C. number of shells increase along with increase of shielding effect</p> <p>D. the behaviour of the elements remain the same</p>
19	The boiling point of higher alkanes are greater than those of lower alkanes due to reason that	<p>A. higher alkanes have greater number of atoms</p> <p>B. the polarizabilities of higher alkanes are greater</p> <p>C. higher alkanes have greater hydrogen bonding</p> <p>D. higher alkanes have zig-zag structures</p>
20	Ice occupies more space than liquid water	<p>A. 9%</p> <p>B. 10%</p> <p>C. 11%</p> <p>D. 12%</p>