

MDCAT Chemistry Chapter 10 Chemical bonding Online Test

Sr	Questions	Answers Choice
1	The reaction between Cu and conc. H ₂ SO ₄ produces	A. Cu+2 B. SO ₂ C. SO ₃ D. H ₂
2	How magnesium reacts with water?	A. In frozen ice water B. With cold water C. In with steam D. In hot state
3	Which of the following compounds is not known?	A. SbCl B. NCl ₃ C. NI ₃ D. NCl ₅
4	Which element differs from rest of elements of its group?	A. Ba B. Mg C. Ca D. Be
5	What is the formula of magnesite?	A. PbS B. MgSO ₄ · 7H ₂ O C. MgCO ₃ D. CaCO ₃
6	Which compound gives carbon when heated with conc. H ₂ SO ₄ .	A. Starch B. Ethyl alcohol C. Oxalic acid D. Formic acid
7	Elements of group II-A are called	A. f-block elements B. s-block elements C. p-block elements D. d-block elements
8	Asbestos is commonly used in making	A. wall board B. black board C. soft board D. hard board
9	Which of the following does not give flame test?	A. Li B. Ba C. Mg D. Sr
10	What is the formula of silica?	A. Si ₂ O ₃ B. SiO ₂ C. Si ₃ O ₄ D. SiO-
11	Lithium differs from rest of members of its group due to which of following reasons	A. High E.N of Li+1 B. Small radius C. High charge density D. All above are correct
12	Which of the following are not known to form compounds in more than one oxidation state?	A. Transition metals B. Halogens C. Alkali metals D. Noble gases
13	One of the following does not give the flame test. Which is that	A. Sr B. Ba C. Be D. Na
14	Which alkaline earth metal makes peroxide?	A. Ba B. Be C. Mg D. Ca
15	The general electronic configuration of group IV-A elements is	A. ns ² , np ⁶ B. ns ² , np ⁴ C. ns ² , np ³ D. ns ² , np ²

16	Which pair of following pair is metalloid?	A. Antimony and bismuth B. Phosphorous and arsenic C. Nitrogen and phosphorous D. Arsenic and antimony
17	Which of the following elements is most metallic	A. Bi B. sb C. As D. P
18	Plaster of Paris is obtained from	A. marble B. bauxite C. gypsum D. limestone
19	One of the following metals is the most reactive and form super oxide. Indicate that	A. Mg B. K C. Be D. Li
20	Which of the following belongs to alkaline earth metals	A. Cu B. Zn C. Sn D. Mg
21	Compared with alkaline earth metals, the alkali metals exhibit.	A. lower ionization energies B. greater hardness C. high boiling point D. smaller ionic radii
22	Formula of sodium beryllite is	A. $\text{Na}_2\text{B}_4\text{O}_7$ B. Na_2BeO_2 C. BeONa D. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$
23	Which one of the following elements is most electropositive out of group I -A and II-A group?	A. K B. Mg C. Na D. Ca
24	Which is the least reactive of all the alkali metals	A. Li B. Na C. K D. Cs
25	What is the formula of bauxite?	A. $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ B. Al_2O_3 C. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$ D. $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$
26	Elements of II-A group are called alkaline earth metals due to the reason that	A. they occur in earth only B. they form divalent cations only C. they have ns^2 electronic configuration D. their oxides and hydroxides are alkaline in nature and metals are present in earth crust
27	Which one of the following doesn't exhibit allotropy?	A. Bi B. As C. N D. P
28	What is the formula of cryolite?	A. $\text{Al}_2\text{O}_3 \cdot 2\text{H}_2\text{O}$ B. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10 \text{H}_2\text{O}$ C. Na_3AlF_6 D. $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$
29	Corrundam is ore of which element?	A. Al B. Th C. In D. Mg
30	Which set of elements is good loser of electrons	A. F ₂ , Cl ₂ , Br ₂ B. N, P, As C. O, S, Se D. Li, Na, K
31	Beryllium differs from other elements of group II-A due to	A. high charge density B. comparatively high nuclear charge C. small radius D. all above
32	About 25% of earth crust mass is made up of element	A. Oxygen B. Silicon C. Aluminium D. Aluminates

33	What is the formula of dolomite?	A. $\text{CaMg}_3(\text{SiO}_3)_4$ B. MgCO_3 C. $\text{MgCO}_3 \cdot \text{CaCO}_3$, D. MgSO_4
34	Elements of group IV-A are	A. neither strongly electropositive nor strongly electronegative B. strongly electropositive C. strongly electronegative D. none of these
35	Which is not mineral of Al?	A. Diaspore B. Corundum C. Bauxite D. Galena
36	What is name of hydrated variety of quartz?	A. Rose quartz B. Smokey quartz C. Silica D. Opal
37	Which of the following element has high m.p and b.p, it acts as a reducing agent, and can react with bases?	A. Sr B. Ca C. Be D. Mg
38	Which of the following configurations corresponds to alkaline earth metals?	A. $[\text{Ar}] 3d^{10}, 4s^2$ B. $[\text{Ne}] 3d^2, 3p^2$ C. $[\text{Ar}] 4s^2$ D. $[\text{Ar}], 3d^{10}, 4s^1$
39	Soda lime is often employed to remove both	A. H_2O and NO_2 B. CO_2 , and NO_2 , C. H_2O and CO_2 D. H_2S and CO_2
40	Which noble gas is used in mixture used for breathing by divers?	A. Ge B. Ar C. Kr D. He
41	Which one of following property is not true about alkali metals?	A. Strongest bases due to their hydrides B. Low ionization energy C. Oxidation number more than +1 D. Form acidic oxides
42	Which one of the following elements is not an alkali metal?	A. Na B. Sr C. Cs D. Rb
43	The element which exhibits maximum catenation property is	A. C B. Pb C. Ge D. Sn
44	Fluorine is largely used in	A. rocket fuels B. making Teflon C. making freon D. All
45	What is the formula of clay?	A. Asbestos B. Talc C. $\text{H}_2\text{Al}_2(\text{SiO}_4)_2 \cdot \text{H}_2\text{O}$ D. Na_2SiO_3
46	What is the formula of talc or soapstone?	A. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ B. $\text{H}_2\text{Mg}_3(\text{SiO}_3)_4$ C. Cu_2S D. NaNO_3
47	Which element of group V-A and VII-A does not use d-orbital?	A. Nitrogen B. Sulphur C. Arsenic D. Chlorine
48	The elements of group I-A react violently with water and make the solution	A. neutral B. amphoteric C. acidic D. alkaline
49	The alkaline earth metals are called so because they	A. form alkaline solution and are present in earth crust as minerals B. form alkaline solution and are found in nature states C. are present in earth crust D. are present in earth crust as their minerals

50

Li resembles with Mg, because

- A. the ratio of their charge to size is nearly the same
- B. both have nearly same size
- C. both are metallic in nature
- D. both are found together in nature