# Physics ICS Part 1 Chapter 6 Online Test

| Sr | Questions | Answers Choice |
|---|---|---|
| 1 | In carnot engine, each process is. | A. Reversible <br>B. Preferable Reversible <br>C. Irreversible <br>D. Perfectly irreversible |
| 2 | Entropy of a system in reversibel process | A. Fluctuates <br>B. Increases <br>C. Is infinite <br>D. Decrease |
| 3 | Thermodynamics mostly deals with. | A.    Measuremetn of quantity   <br>B.  Transfer of quantity of heat   <br>C.   Change of state <br>   <br>D.   Conversion of heat to other forms of energy <br>   |
| 4 | The process which is carried out at constant temperature is called. | A. Adiabatic process <br>B. Isothermal process <br>C. Isochoic process <br>D. Isobaric process |
| 5 | 'R' is called | A. Universal constant <br>B. Universal per molecule constant <br>C. Universal gas constant <br>D. All of the above |
| 6 | A system does 600 J of work and at the same time has its internal energy increased by 320 J. How much heat has been supplied. | A. 920 J <br>B. 280 J <br>C. 600 J <br>D. 200 J |
| 7 | SI unit of entropy is | A. J/Kg <br>B. J/K <br>C. K gms-1 <br>D. JK |
| 8 | No entrpy chagne is associated with | A. isothermal process <br>B. Adiabatic process <br>C. Isobaric process <br>D. Isochoric process |
| 9 | Adiabatic change occurs hen the gas expands or compressed. | A. Rapidly <br>B. Slowly <br>C. Gradually <br>D. Moderately |
| 10 | An addition of 400 J of heat casues the increase in internal energy of system is equal to 300 J, then work done will be | A. 100 J <br>B. 200J <br>C. 300 J <br>D. 400 J |
| 11 | The chagne in inernal energy is defined as. | A. Q- W <br>B. Q- T <br>C. Q+ P <br>D. Q - P |
| 12 | First law of thermodynamics is based upon law of conservatio of. | A. Mass <br>B. Momentum <br>C. Energy <br>D. Charge |
| 13 | What happens to internal energy of an object when its temperature. | A. Decreses <br>B.  Increases I <br>C.  Fluctuates <br>   <br>D. Remains Constant <br>   |
| | | A.  Potential energy |

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| 14 | Pressure of a gas is directly proportional to average. |   <br>B.  Rotational energy <br>   <br>C.   Translation K.E <br>   <br>D. Compressed P.E. |
| 15 | In all natural prcoessees where heat flow from one system to another there is always a net increase in | A. Pressure <br>B. Entropy <br>C. Work <br>D. Volume |
| 16 | Boyle's law states that " The volume of a given mass of a gas is....." | A. Directly proportional to absolute temperature <br>B. Invesely proportional to absolute temperature <br>C. Directly proportional to density <br>D. Inversely proportional topressure |
| 17 | In an adiabatic process, there is no. | A.   Change in temperature   <br>B.   Exchagne of heat <br>   <br>C.   Chagne is internal energy     <br>D. Work done   |
| 18 | Internal energy is similar to the | A. Vibrational K.E. <br>B. Gravitational P.E. <br>C. K.E. <br>D. All of these |
| 19 | The sum of all forms of molecular energies of substance is termed as | A. Kinetic energy <br>B. Potential energy <br>C. Internal energy <br>D. Heat energy |
| 20 | According to kinetic theory of gases, a a finite volume of a gas consts of very | A.   Large numebr of molecules   <br>B.    Small number ofmolecules   <br>C.  Both a and b <br>   <br>D.   Large no of ions <br> L |