

Group IIA & IVA Elements

Sr	Questions	Answers Choice
1	The structure of white phosphorus is	A. Square planar B. Pyramidal C. Tetrahedral D. Trigonal planer
2	Solvay process is used in the manufacture of	A. Na_2CO_3 B. NaHCO_3 C. CaCl_2 D. All
3	Borax is a common mineral of alkali metal sodium. Its formula is	A. $\text{Na}_2\text{B}_4\text{O}_7$ B. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ C. $\text{Na}_2\text{B}_3\text{O}_6 \cdot 10\text{H}_2\text{O}$ D. $\text{Na}_2\text{B}_4\text{O}_7 \cdot 5\text{H}_2\text{O}$
4	Red P can be obtained form white P by	A. Heating it with a catalyst in an inert atmosphere B. Distilling it in an inert atmosphere C. Dissolving it in CS_2 and cryslizing D. Melting it ad pouring the liquid into water
5	Which shows maximum catenation property?	A. S B. Se C. Te D. O
6	Which salt is used for preserving food	A. BaCl_2 B. CaCl_2 C. NaCl D. Na_2SO_4
7	Lime, calcium oxide, is used in agriculture for	A. Adding ca metal in soil B. Making soil acidic C. Neutralizing acidic soil D. Adding oxygen to soil
8	White phosphorus is	A. A mono atomic gas B. P_4 , a tetrahedral solid C. P_8 , a crown D. A linear diatomic molecule
9	Chemical composition of colemnite is:	A. $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$ B. $\text{CaB}_4\text{O}_7 \cdot 4\text{H}_2\text{O}$ C. $\text{Na}_2\text{B}_7\text{O}_{17} \cdot 4\text{H}_2\text{O}$ D. $\text{CaNaB}_5\text{O}_9 \cdot 8\text{H}_2\text{O}$
10	The oxide of beryllium is	A. Acidic B. Amphoteric C. Superoxide D. Basic
11	Francium is an element at the bottom of Group I in the Periodic Table. Which one of the following predication is likely to be correct?	A. It will react with water to liberate oxygen B. Its hydroxide will be a strong alkali in water C. Its carbonate will decompose on heating to give carbon dioxide D. Its nitrate on heating will give nitrogen dioxide and oxygen
12	Which one of the following is not an alkali metal	A. Francium B. Caesium C. Rubidium D. Radium
13	Sodium forms largely	A. Normal oxides B. Per-oxides C. Superoxides D. None of these
14	Which of the following sulphates is not soluble in water	A. Sodium sulphate B. Potassium sulphate C. Zinc sulphate D. Barium sulphate

15	The ore $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ has the general name	A. Gypsum B. Dolomite C. Calcite D. Epsom salt
16	A compound used as eye wash:	A. Borax B. Boric acid C. Metabolic acid D. Pyroboric acid
17	The chemical formula of Trona is	A. $\text{KCl} \cdot \text{MgCl}_2 \cdot 6\text{H}_2\text{O}$ B. $\text{Na}_2\text{CO}_3 \cdot 2\text{NaHCO}_3 \cdot 2\text{H}_2\text{O}$ C. $\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$ D. KCl
18	Which of the following elements is not present abundantly in earth's crust?	A. Silicon B. Aluminium C. Sodium D. Oxygen
19	Down's cell is used to prepare	A. Sodium carbonate B. Sodium bicarbonate C. Sodium metal D. Sodium hydroxide
20	When sulphur is boiled with Na_2SO_3 solution, the compound formed is	A. Sodium sulphides B. Sodiums sulphates C. Sodium persulphate D. Sodium thiosulphate
21	Which property is not present in Al?	A. Reacts with acid B. Reacts with bases C. Changes litmus paper D. Changes methyle orange colour
22	The element which has a simple cubic lattice in solid state is	A. Se B. Te C. Po D. None of these
23	Chile saltpetre has the chemical formula	A. NaNO_3 B. KNO_3 C. $\text{Na}_2\text{B}_4\text{O}_7$ D. $\text{Na}_2\text{CO}_3 \cdot \text{H}_2\text{O}$
24	Which of the following phosphorus is most reactive?	A. Red phosphorus B. White phosphorus C. Scrllet phosphorus D. Violet phosphorous
25	_____ gives peroxide	A. Li B. Ba C. Sr D. Be
26	Which of the following salt is used as purgative	A. CaSO_4 B. MgSO_4 C. BeSO_4 D. NaCl
27	What is the formula of asbestos?	A. $\text{CaMg}_3(\text{SiO}_3)_4$ B. CaSiO_3 C. Na_2SiO_3 D. $\text{Mg}_3\text{H}_2(\text{SiO}_3)_4$
28	Al is the most element in earth crust:	A. O B. Si C. Al D. Pb
29	Sodium is never found free in nature because of its	A. Chemical reactivity B. Small ionic size C. Small atomic volume D. None of these
30	Alkaline earth metals are usually	A. Reducing agent B. Oxidizing agent C. Amphoteric D. Acidic
31	Which on of the following compounds does not exist?	A. NCl_5 B. AsF_5 C. SbCl_5 D. PF_5

A. NH_4Cl

32	Which of the following has highest dipole moment?	<p>A. PH_3</p> <p>B. PH_3</p> <p>C. AsH_3</p> <p>D. SbH_3</p>
33	_____ is used in breathing equipments for mountaineers and in space craft	<p>A. Li_2O</p> <p>B. BeO</p> <p>C. N_2</p> <p>D. KO_2</p>
34	Crystalline form of sulphur stable at room temperature is	<p>A. Rhombic sulphur</p> <p>B. Monoclinic sulphur</p> <p>C. Plastic sulphur</p> <p>D. Prismatic sulphur</p>
35	The number of electron that are paired in oxygen molecule are	<p>A. 16</p> <p>B. 12</p> <p>C. 14</p> <p>D. 7</p>
36	Which is not the form of silica?	<p>A. Amethyst quartz</p> <p>B. Rose quartz</p> <p>C. Smoky</p> <p>D. None of these</p>
37	LiOH _____ soluble than NaOH	<p>A. More soluble</p> <p>B. Less soluble</p> <p>C. Equally soluble</p> <p>D. None</p>
38	NaHCO_3 is commonly called	<p>A. Soda ash</p> <p>B. Baking soda</p> <p>C. Washing soda</p> <p>D. None of these</p>
39	Dipole moment of CO molecule is:	<p>A. 0.0</p> <p>B. 1.112 D</p> <p>C. 0.112 D</p> <p>D. 2.112 D</p>
40	Calcium carbide reacts with water to produce	<p>A. Acetylene</p> <p>B. Methane</p> <p>C. Ethylene</p> <p>D. Ethane</p>
41	Ozone is not	<p>A. An allotrope</p> <p>B. A powerful oxidizing agent</p> <p>C. Paramagnetic</p> <p>D. A bent molecule</p>
42	The acid which has a peroxy linkage is	<p>A. Sulphurous acid</p> <p>B. Pyrosulphuric acid</p> <p>C. Dithionic acid</p> <p>D. Caro's acid</p>
43	Which of the following salt is soluble in water	<p>A. CaCO_3</p> <p>B. CaSO_4</p> <p>C. MgSO_4</p> <p>D. BaSO_4</p>
44	Plaster of paris is obtained by heating	<p>A. Gypsum</p> <p>B. Epsom</p> <p>C. Lime stone</p> <p>D. Dolomite</p>
45	Which of the following fluorides does not exists?	<p>A. NF_5</p> <p>B. PF_5</p> <p>C. AsF_5</p> <p>D. SbF_5</p>
46	Hybridized in oxygen is:	<p>A. sp</p> <p>B. sp^2</p> <p>C. sp^3</p> <p>D. dsp^3</p>
47	Lithium reacts with air to form	<p>A. Peroxide</p> <p>B. Normal oxide</p> <p>C. Superoxide</p> <p>D. None of these</p>
48	Density of aluminium is (g cm^{-3}):	<p>A. B</p> <p>B. Al</p> <p>C. Si</p> <p>D. Ge</p>
49	Which element among the following belongs to Group IV-A of the periodic table?	<p>A. Barium</p> <p>B. Iodine</p> <p>C. Lead</p> <p>D. Oxygen</p>

50	As a fixing agent in photography, sodium thiosulphate is used for	A. Dissolving out unreacted silver bromide B. Converting silver C. Reducing solubility of AgBr D. Preventing overdeveloping and fogging
51	P ₂ O ₅ is heated with water to give	A. Hypophosphorus acid B. Phosphorus acid C. Hypophosphorus acid D. Orthophosphorus acid
52	The oxyacids of phosphorus in which phosphorus has the lowest oxidation state is	A. Hypophosphorus acid B. Orthophosphorus acid C. Pyrophosphorus acid D. Metaphosphorus acid
53	Which colour is imparted by sodium	A. Yellow B. Violet C. Red D. Crimson
54	Which of the following compounds is explosive in nature?	A. Phosphorus trichloride B. Nitrogen trichloride C. Hyponitrous acid D. Nitrosyl chloride
55	Bond angle is minimum for	A. H ₂ O B. H ₂ S C. H ₂ Se D. H ₂ Te
56	Basicity of orthophosphoric acid is	A. 2 B. 3 C. 4 D. 5
57	Which of the element is not an alkaline earth metal	A. Beryllium B. Strontium C. Barium D. Caesium
58	When a colourless gas is passed through bromine water only decolourisation takes place the gas is	A. SO ₂ B. HBr C. HCl D. H ₂ S
59	Phosphide ion has the electronic structure similar to that of	A. Nitride ion B. Fluoride ion C. Sodium ion D. Chloride ion
60	All the following decompose easily on heating to give oxygen except	A. Lead nitrate B. Potassium chlorate C. Mercuric oxide D. Manganese dioxide
61	CaC ₂ on hydrolysis form	A. CH ₄ B. C ₂ H ₄ C. C ₂ H ₂ D. C ₆ H ₆
62	Which salt is used for the treatment of hyperacidity in stomach	A. NaCl B. KCl C. NaHCO ₃ D. Na ₂ CO ₃
63	Chief ore of aluminium is:	A. Na ₃ AlF ₆ B. Al ₂ O ₃ · 2H ₂ O C. Al ₂ O ₃ D. Al ₂ O ₃ · H ₂ O
64	Phosphorus pentoxide finds use as	A. An oxidizing agent B. A reducing agent C. A bleaching agent D. A dehydrating agent
65	Corundum is ore of:	A. Li B. Be C. B D. Al
66	HNO ₂ acts as an/a	A. Acid B. Oxidizing agent C. Reducing agent D. All the three
67	Which one of the following properties is not correct for ozone?	A. It oxidizes lead sulphides B. It oxidizes potassium iodide C. It oxidizes mercury

		D. It cannot act as a bleaching agent
68	NaHCO_3 is prepared by	A. Down's process B. Solvay's process C. Nelson's process D. None of these
69	BiCl_3 on hydrolysis forms a white precipitate of	A. Bismuthic acid B. Bismuth oxychloride C. Bismuth pentachloride D. Bismuth hydroxide
70	Which of the following statement is not related to Solvay's process of Na_2CO_3	A. Cheap materials B. Pure product C. Continuous process D. Harmful by-products
71	Heavy water is obtained by	A. Prolonged electrolysis of water B. Dissolving heavy salt in water C. Simple distillation of water D. Removing impurities of calcium and magnesium from water
72	Which will have the maximum value of heat of hydration	A. Na^{+} B. Cs^{+} C. Ba^{2+} D. Mg^{2+}
73	Pb has inert pair of electrons:	A. One B. Two C. Three D. Four
74	The alkaline earth elements have in their s-orbital	A. One electron B. Two electron C. No electron D. Three electron
75	S-block elements consist of	A. Alkali metals B. Alkaline earth metals C. Alkali and alkaline earth metals D. None of these
76	What is going to replace the petroleum?	A. Silica B. Silicates C. Silicones D. Silicon
77	All the elements of oxygen family are	A. Non metals B. Metalloids C. Radioactive D. Polymorphic
78	Borax is hydrated:	A. Penta B. Deca C. Hepta D. Octa
79	Which metal is used in the thermit process because of its activity?	A. Iron B. Copper C. Aluminium D. Zinc
80	Which is litharge or massicot?	A. PbO B. Pb_2O C. Pb_3O_4 D. PbO_2
81	Which of the following is a tetrabasic acid?	A. Orthophosphoric acid B. Hypophosphorous acid C. Metaphosphoric D. Pyrophosphoric acid
82	Which element from group 15 gives most basic compound with hydrogen?	A. Nitrogen B. Bismuth C. Arsenic D. Phosphorus
83	The chemical formula of magnesite is	A. MgCl_2 B. $\text{Mg}(\text{HCO}_3)_2$ C. MgCO_3 D. None of these
84	Some of the elements of a period show similar behavior with the elements of next group in next period this is called	A. Vertical relationship B. Oblique relationship C. Diagonal relationship D. None

A. P

85	Which of the following elements is most metallic?	B. As C. Sb D. Bi
86	Which electronic configuration corresponds to an element of Group IZ-A of the periodic table?	A. $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2$ B. $1s^2, 2s^2, 2p^6, 3s^2, 3p^1$ C. $1s^2, 2s^2, 2p^6$ D. $1s^2, 2s^2, 2p^6, 3s^2, 3p^3$
87	When SO_2 is passed through acidified $K_2Cr_2O_7$ solution	A. The solution turns blue B. The solution is decolourised C. SO_2 is reduced D. Green $Cr^{3+}(SO_4)_3$ is formed
88	Phosphine is not obtained by the reaction when	A. White P is heated with NaOH B. Red P is heated with NaOH C. Ca_2P_2 reacts with water D. PH_4I is boiled with water
89	Tincal is a mineral of:	A. Al B. B C. Si D. C
90	Potassium superoxide has a use in breathing equipment in space crafts. The balanced equation for the reaction is	
91	The formula of lime stone is	A. $CaCl_2$ B. $MgCO_3$ C. Na_2CO_3 D. $CaCO_3$
92	Which is used to remove air bubbles form metals?	A. B B. Be C. Mg D. Al
93	Which of the following equations represents the action of heat on $NaHCO_3$	
94	Sodium is manufacture by the electrolysis of fused sodium chloride and not from an aqueous solution of sodium chloride because	A. Sodium chloride does not ionize in the water solution B. Sodium chloride is not soluble in water C. Sodium deposited at the cathode may react with water to form sodium hydroxide D. Electricity does not pass through aqueous NaCl
95	Which of the following is formed by the action of water on sodium peroxide?	A. H_2 B. N_2 C. O_2 D. CO_2
96	Identify the incorrect statement with respect to ozone	A. Ozone is formed in the upper atmosphere by a photochemical reaction involving dioxygen B. Ozone is more reactive than dioxygen C. Ozone is diamagnetic whereas dioxygen is paramagnetic D. Ozone protects the earth's inhabitants by absorbing gamma-radiations
97	Alkaline earth metals posses two electrons in their outermost	A. F-orbital B. D-orbital C. S-orbital D. P-orbital
98	Which is an essential constituent of chlorophyll	A. Be B. Fe C. Mg D. Ca
99	Which is used in navigational equipments?	A. B B. Be C. Mg D. Al
100	Copper oxide is dedected by borax bead test with colour:	A. Blue B. Red C. Yellow D. Black
101	Polyanion formation is maximum in	A. Nitrogen B. Oxygen C. Sulphur D. Boron
102	Orthophosphoric acid is	A. Monobasic B. Dibasic C. Tribasic

		D. Tetrabasic
103	Each of the following is true of white and red phosphours except that they	<p>A. Are both soluble in CS_2</p> <p>B. Can be oxidized by heating in air</p> <p>C. Consist of the same kind of atoms</p> <p>D. Can be converted into one another</p>
104	The chemical formula of Epson salt is	<p>A. MgSO_4</p> <p>B. MgCl_2</p> <p>C. $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$</p> <p>D. $\text{MgCl}_2 \cdot 7\text{H}_2\text{O}$</p>
105	In a group IIA from top to bottom as the atomic number increases, there is regular decreases in	<p>A. Ionic size</p> <p>B. Atomic size</p> <p>C. Ionization potential</p> <p>D. None of these</p>
106	The alkali metals form	<p>A. Ionic bond</p> <p>B. Covalent bond</p> <p>C. Co-ordinate bond</p> <p>D. H-bond</p>
107	Which oxide of nitrogen is obtained on heating ammonium nitrate at 250°C ?	<p>A. Nitric oxide</p> <p>B. Nitrous oxide</p> <p>C. Nitrogen dioxide</p> <p>D. Dinitrogen tetraoxide</p>
108	Which metal of Group-II A of the periodic Table, will form the least ionic chloride	<p>A. Be</p> <p>B. Mg</p> <p>C. Ca</p> <p>D. Sr</p>
109	Alkali metals react violently with halogens to form	<p>A. Hydrides</p> <p>B. Halides</p> <p>C. Anyhydrides</p> <p>D. None of these</p>
110	Which of the following element is most reactive	<p>A. Li</p> <p>B. Na</p> <p>C. K</p> <p>D. Cs</p>
111	Silicon atom is hybridized:	<p>A. sp</p> <p>B. sp^2</p> <p>C. sp^3</p> <p>D. dsp^2</p>
112	Sodium should be stored in	<p>A. Air free from moisture</p> <p>B. Air free form carbon dioxide</p> <p>C. Under water</p> <p>D. Under kerosene oil</p>
113	Sulphuric acid has great affinity for water because	<p>A. It hydrolyses the acid</p> <p>B. It decomposes the acid</p> <p>C. Acid forms hydrates with water</p> <p>D. Acid decomposes water</p>
114	Permonosulphuric acid is known as	<p>A. Marshall's acid</p> <p>B. Carlo's acid</p> <p>C. Sulphuric acid</p> <p>D. None of these</p>
115	Sodium hexametaphosphate is known as	<p>A. Calgon</p> <p>B. Permutite</p> <p>C. Natalite</p> <p>D. Nitrolim</p>
116	Oleum is	<p>A. Castor oil</p> <p>B. Oil of vitriol</p> <p>C. Fuming of H_2SO_4</p> <p>D. None of them</p>
117	Which of the following are electropositive in nature	<p>A. Alkali metals</p> <p>B. Alkaline earth metals</p> <p>C. Halogens</p> <p>D. Alkali and alkaline earth metals</p>
118	Sulphuric acid reacts with PCl_5 to give	<p>A. Thionyl chloride</p> <p>B. Sulphur monochloride</p> <p>C. Sulphuryl chloride</p> <p>D. Sulphur tetrachloride</p>
119	Which element has highest oxidation potential	<p>A. Li</p> <p>B. Be</p> <p>C. Ba</p> <p>D. Ra</p>
120		<p>A. PF_5</p> <p>B. AsF_5</p>

120	Which one of the following pentafluorides cannot be formed?	<p>C. SbF_5</p> <p>D. BiF_5</p>
121	Which of the following is the hardest metal among following	<p>A. Li</p> <p>B. Na</p> <p>C. Rb</p> <p>D. K</p>
122	Hypo is used in photography for	<p>A. Developing picture</p> <p>B. Picture printed</p> <p>C. The colour of picture</p> <p>D. The fixation of picture</p>
123	The chemical formula of chile salt peter is	<p>A. Na_2CO_3</p> <p>B. KNO_3</p> <p>C. NaNO_3</p> <p>D. NaNO_2</p>
124	Boric acid cannot be used:	<p>A. An antiseptic in medicine</p> <p>B. For washing eyes</p> <p>C. In soda bottles</p> <p>D. For enamels and glazes</p>
125	The acid used in lead storage cells is	<p>A. Phosphoric acid</p> <p>B. Nitric acid</p> <p>C. Sulphuric acid</p> <p>D. Hydrochloric acid</p>
126	Sometimes a yellow turbidity appears while passing H_2S gas even in the absence of II group radicals. This is because	<p>A. Sulphur is present in the mixture as impurity</p> <p>B. IV group radicals are precipitated as sulphides</p> <p>C. Of the oxidation of H_2S gas by some acid radicals</p> <p>D. III group radicals are precipitated as hydroxides</p>
127	Carbonates of alkali metals dissolves freely in water to form	<p>A. Acidic solutions</p> <p>B. Neutral solution</p> <p>C. Alkaline solution</p> <p>D. None of these</p>
128	Which carbonate of alkali metals is insoluble in water	<p>A. Na_2CO_3</p> <p>B. K_2CO_3</p> <p>C. Li_2CO_3</p> <p>D. Cs_2CO_3</p>
129	Which of the element is not alkali metal	<p>A. Lithium</p> <p>B. Rubidium</p> <p>C. Francium</p> <p>D. Magnesium</p>
130	Which of the following oxides is peroxide?	<p>A. Na_2O_2</p> <p>B. MnO_2</p> <p>C. BaO</p> <p>D. SO_2</p>
131	Group VA of the periodic table consists of the elements N, P, As, Sb and Bi. On passing from N to Bi, the oxides of the elements of general formula M_2O_3 become	<p>A. Stronger reducing agents</p> <p>B. More ionic</p> <p>C. More basic</p> <p>D. More volatile</p>
132	Hybridized in carbon is:	<p>A. sp</p> <p>B. sp^2</p> <p>C. sp^3</p> <p>D. d^2sp^3</p>
133	Which of the following is acidic?	<p>A. SO_3</p> <p>B. N_2O</p> <p>C. BeO</p> <p>D. HgO</p>
134	Which one of the following does not belong to alkaline earth metals	<p>A. Be</p> <p>B. Ra</p> <p>C. Ba</p> <p>D. Rn</p>
135	Which of all following compound is not possible	
136	_____ is called milk of magnesia	<p>A. NaOH</p> <p>B. KOH</p> <p>C. LiOH</p> <p>D. None</p>
137	Baking powder has which one of the following formula	<p>A. Na_2CO_3</p> <p>B. Na_2SO_4</p> <p>C. NaHCO_3</p> <p>D. K_2CO_3</p>
138	Which of the following is lighter	<p>A. Li</p> <p>B. k</p> <p>C. Na</p> <p>D. Ca</p>

139	_____ reacts with alkalis to give hydrogen	A. Be B. Mg C. Ca D. None
140	Electron affinity of sulphur is	A. More than O and Se B. More than O but less than Se C. Less than O but more than Se D. Equal to O and Se
141	Which of the following form normal oxide	A. K B. Li C. Na D. None
142	Which of the following compound is industrially prepared by the electrolysis of solution of NaCl	A. Na_2CO_3 B. NaHCO_3 C. NaOH D. NaOCl
143	Gypsum is applied to the soil to provide calcium and	A. Oxygen B. Nitrogen C. Phosphorous D. Sulphur
144	Aluminium oxide is:	A. Acidic oxide B. Basic oxide C. Amphoteric oxide D. None of these
145	The chemical formula of gypsum is	A. $\text{CaSO}_4 \cdot 5\text{H}_2\text{O}$ B. $\text{CaSO}_4 \cdot 4\text{H}_2\text{O}$ C. $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ D. None of these
146	The number of unpaired electrons in the p-subshell of oxygen atom	A. 1 B. 2 C. 3 D. 4
147	What causes nitrogen to be chemically inert?	A. Multiple bond formation in the molecule B. Absence of bond distance C. Short internuclear distance D. High bond energy
148	Where lime is not used	A. In refining of metals B. In paper industry C. In glass industry D. In the preparation of NaOH
149	Which element forms an ion with charge 3+?	A. Beryllium B. Aluminium C. Carbon D. Silicon
150	Which of the following sulphates has the highest solubility in water	A. BaSO_4 B. CaSO_4 C. MgSO_4 D. BeSO_4
151	What is chrome yellow?	A. PbO B. Pb_2O C. PbCrO_4 D. Pb_3O_4
152	Marble is chemically	A. CaCl_2 B. CaCO_3 C. Na_2CO_3 D. NaHCO_3
153	Commercial common salt becomes slightly damp on storing because	A. Common salt is hygroscopic B. Common salt contains some impurity, which is hygroscopic C. Salt is efflorescent D. Salt is crystalline
154	The electrolytic cell used for the production of metallic sodium is known as	A. Down's cell B. Solvay's cell C. Haber's cell D. None of these
155	Which one of the following equations represents the reaction that occurs when calcium nitrate is heated strongly	
156	Which one of the following elements occurs free in nature?	A. N B. P C. As D. Sb

157	The most acidic of the following compounds is	<p>A. P_2O_3</p> <p>B. Sb_2O_3</p> <p>C. B_2O_3</p> <p>D. As_2O_3</p>
158	Which of the following is not the reactions of lithium	
159	The metal with highest electrical resistance at room temperature is	<p>A. Pb</p> <p>B. Te</p> <p>C. Po</p> <p>D. Fe</p>
160	The maximum number of electrons in the outermost shell of s-block elements is	<p>A. One</p> <p>B. Two</p> <p>C. Three</p> <p>D. Four</p>
161	Which has soapy touch?	<p>A. $\text{Na}_2\text{B}_4\text{O}_7$</p> <p>B. H_3BO_3</p> <p>C. $\text{Ca}_2\text{B}_6\text{O}_{11}$</p> <p>D. HBO_2</p>
162	On alkali and alkaline earth metals down the group, there is decreasing trend in	<p>A. m.p.</p> <p>B. b.p.</p> <p>C. Ionization potential</p> <p>D. All of these</p>
163	The silver bromide in hypo ($\text{Na}_2\text{S}_2\text{O}_3$) solution is	<p>A. Soluble</p> <p>B. Not soluble</p> <p>C. Precipitated</p> <p>D. Not effect</p>
164	White P when boiled with strong solution of caustic soda produces	<p>A. Phosphine</p> <p>B. Phosphic acid</p> <p>C. Phosphorous acid</p> <p>D. None</p>
165	The oxidation number of each element of group II-A is	<p>A. 0</p> <p>B. +1</p> <p>C. +2</p> <p>D. -1</p>
166	When ammonia is heated with cupric oxide, a molecule of ammonia will	<p>A. Gain 3 electrons</p> <p>B. Lose 3 electrons</p> <p>C. Gain 2 electrons</p> <p>D. Lose 2 electrons</p>
167	The element caesium bears resemblance with	<p>A. Ca</p> <p>B. Cr</p> <p>C. Rubidium</p> <p>D. None of the above</p>
168	Gypsum is a common mineral of	<p>A. Magnesium</p> <p>B. Strontium</p> <p>C. Calcium</p> <p>D. Barium</p>
169	In the earth crust sodium is	<p>A. 2.50%</p> <p>B. 2.30%</p> <p>C. 2.40%</p> <p>D. 3.50%</p>
170	The oxidation number of each element of group I-A is	<p>A. 0</p> <p>B. +1</p> <p>C. +2</p> <p>D. -1</p>