

## Chemistry Fsc Part 2 Chapter 1 Online Test

Sr	Questions	Answers Choice
1	In which group of periodic table is the element which has atomic number 14.	A. II B. IV C. III D. VI
2	What are the total numbers of periods in the modern periodic table	A. 3 B. 5 C. 7 D. 8
3	Which property of hydrogen not resemble to alkali metals.	A. Electronic configuration B. Oxidation state C. Reaction with halogen D. Metallic nature
4	Ionic Hydrides react with water to form	A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions
5	What criteria did Mendeleev use in arranging his periodic table.	A. Atomic mass B. Atomic number C. Mass number D. Density
6	The basis of modern periodic table is	A. Electron affinity B. Atomic mass C. Ionization Potential D. Atomic number
7	Which of the following the highest hydration energy	A. $\text{Li}^{+}$ B. $\text{Na}^{+}$ C. $\text{K}^{+}$ D. $\text{Mg}^{++}$
8	Keeping in view size of atoms ,which order is the correct one.	A. $\text{Mg} > \text{Sr}$ B. $\text{Ba} > \text{Mg}$ C. $\text{Lu} > \text{Ce}$ D. $\text{Cl} > \text{I}$
9	In which group, melting point and boiling point increase downward in a group	A. IA B. II A C. VII A D. Both a and b
10	The first ionization energy of Na, Mg, Al and Si are in the order of.	A. $\text{Na} < \text{Mg} < \text{Al} < \text{Si}$ B. $\text{Na} > \text{Mg} > \text{Al} > \text{Si}$ C. $\text{Na} > \text{Mg} < \text{Al} < \text{Si}$ D. $\text{Na} < \text{Mg} > \text{Al} < \text{Si}$
11	Fluorine is in group VII A of periodic table. Its chemistry will most closely resembles that of.	A. Argon B. Boron C. Iodine D. Sulphur
12	Which class of elements shows low value of first ionization potential.	A. Alkali metals B. Alkaline earth metals C. Halogens D. Noble gases
13	Variable valency is generally exhibited by	A. Transition elements B. Alkali metals C. s-block elements D. Gaseous elements
14	Which of the following are alkaline earth metals?	A. Be, Mg, Ca B. Li, Na, K C. Fe, CO, Ni D. B, Al, Ga
15	Hydrogen resembles in properties with	A. IA, IV A and VII A elements B. III A, IV A and V A elements C. II A, IV A and VI A elements D. II A, III A and VII A elements

16	The element of 2nd period, which has highest ionization energy from the following is	A. Be B. C C. N D. O
17	Which one of the following is a metalloid.	A. sulphur B. Antimony C. Mercury D. Zinc
18	Which element has lowest melting point	A. Beryllium B. Magnesium C. Calcium D. Barium
19	Which of the following statement is correct	A. Na atom is smaller than $\text{Na}^{+}$ B. Na atom is larger than K atom C. F atom is smaller than $\text{F}^{-}$ D. F atom is larger than $\text{F}^{-}$
20	As you proceed across a period in the periodic table the first ionization energy	A. Decrease B. Increase C. Remains constant D. First increase up the middle of period and then decreases