

## Chemistry Fsc Part 2 Chapter 1 Online Test

Sr	Questions	Answers Choice
1	The element of 2nd period, which has highest ionization energy from the following is	A. Be B. C C. N D. O
2	Which class of elements shows low value of first ionization potential.	A. Alkali metals B. Alkaline earth metals C. Halogens D. Noble gases
3	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
4	Number of elements in the first period of the periodic table is	A. 2 B. 8 C. 14 D. 18
5	Which of the following has highest M.P	A. Aluminium B. Silicon C. Phosphorus D. Sulphur
6	Which statement is incorrect.	A. All the metals are good conductors of electricity B. All the metals are good conductor of heat C. All the metal form positive ions D. All the metal form acidic oxides
7	What are the total numbers of periods in the modern periodic table	A. 3 B. 5 C. 7 D. 8
8	Variable valency is generally exhibited by	A. Transition elements B. Alkali metals C. s-block elements D. Gaseous elements
9	Ionic Hydrides react with water to form	A. Proton B. Hydride ions C. Hydroxide ions D. Hydronium ions
10	Majority of the elements of the periodic table are.	A. Semi metals B. Non metals C. Metals D. Noble metals
11	Which one of the following is a metalloid.	A. sulphur B. Antimony C. Mercury D. Zinc
12	The oxides of metal are generally	A. Acidic B. Basic C. Neutral D. Amphoteric
13	In which group, melting point and boiling point increase downward in a group	A. IA B. II A C. VII A D. Both a and b
14	The basis of modern periodic table is	A. Electron affinity B. Atomic mass C. Ionization Potential D. Atomic number
		A. $\text{Cl}^{\text{sup}} > \text{Cl}^{\text{atom}}$ B. $\text{Cl}^{\text{sup}} > \text{Cl}^{\text{(ion)}}$ and $\text{Cl}^{\text{atom}}$

15	Which the correct statement	(atom) are equal in size C. $\text{Na}^{+}$ is smaller than Na atom D. $\text{Na}^{+}$ is larger than Na atom
16	Which one of the following sets has coinage metals is it.	A. Cu, Hg, Au B. Cu, Ag, Au C. Ag, Au, Hg D. Cu, Fe, Au
17	Who gave the law of Triads in 1829?	A. Dobereiner B. Mosely C. Newland D. Mendeleev
18	Mark the correct statement.	A. $\text{Na}^{+}$ is smaller than Na atom B. $\text{Na}^{+}$ is larger than Na atom C. $\text{Cl}^{-}$ is smaller than Cl atom D. $\text{Cl}^{-}$ and Cl are equal in size
19	Which property of hydrogen not resemble to alkali metals.	A. Electronic configuration B. Oxidation state C. Reaction with halogen D. Metallic nature
20	Elements of Groups IIA are called	A. Alkali metals B. Alkaline earth metals C. Coinage metals D. Halogens
21	Which of the following the highest hydration energy	A. $\text{Li}^{+}$ B. $\text{Na}^{+}$ C. $\text{K}^{+}$ D. $\text{Mg}^{++}$
22	Across a period from left to right in the periodic table, the melting and boiling point.	A. Decrease B. Increase C. Remain constant D. First increase upto the middle of period and then decrease
23	Which one of the following elements has the largest second ionization energy.	A. O B. Na C. F D. Ne
24	As you proceed across a period in the periodic table the first ionization energy	A. Decrease B. Increase C. Remains constant D. First increase up the middle of period and then decreases
25	In which group of periodic table is the element which has atomic number 14.	A. II B. IV C. III D. VI
26	Which one of the following elements burns in air to form an oxide which, when shaken with water, give a solution with a pH greater than 7.	A. Carbon B. Magnesium C. sulphur D. Hydrogen
27	Which of the following element has lowest ionization energy	A. Beryllium B. Boron C. Carbon D. Oxygen
28	What are the total numbers of periods in the modern periodic table.	A. 3 B. 5 C. 7 D. 8
29	Hydrogen resembles in properties with	A. IA, IV A and VII A elements B. III A, IV A and V A elements C. II A, IV A and VI A elements D. II A, III A and VII A elements
30	Correct order according to atomic size in the following is	A. $\text{Na} > \text{K}$ B. $\text{Be} > \text{Mg}$ C. $\text{O} > \text{N}$ D. $\text{Cl} > \text{F}$
31	Which of the following statement is correct	A. Na atom is smaller than $\text{Na}^{+}$ B. Na atom is larger than K atom C. F atom is smaller than $\text{F}^{-}$ D. F atom is larger than $\text{F}^{-}$

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32	Which one is an incomplete period	A. 4th B. 5th C. 6th D. 7th
33	The concept of atomic number was introduced by	A. Alrazi B. Mendeleev C. Moseley D. Dobereiner
34	Which of the following are alkaline earth metals?	A. Be, Mg, Ca B. Li, Na, K C. Fe, CO, Ni D. B, Al, Ga
35	The first ionization energy of Na, Mg, Al and Si are in the order of.	A. Na < Mg < Al < Si B. Na > Mg > Al > Si C. Na > Mg < Al < Si D. Na < Mg > Al < Si
36	Electron affinity is a measure of energy	A. Required to remove the electron B. Released by adding an electron C. Required to excite an electron D. Released by removing an electron
37	Which has greater hydration energy.	A. Li <sup>+</sup> B. Na <sup>+</sup> C. K <sup>+</sup> D. Mg <sup>2+</sup>
38	In which compound, oxidation state of sulphur is +6	A. H <sub>2</sub> S B. H <sub>2</sub> SO <sub>4</sub> C. H <sub>2</sub> SO <sub>3</sub> D. SO <sub>3</sub>
39	Which hydride is intermediate in nature.	A. NaH B. BeH <sub>2</sub> C. NH <sub>3</sub> D. HCl
40	Fluorine is in group VIIA of periodic table. Its chemistry will most closely resemble that of.	A. Argon B. Boron C. Iodine D. Sulphur
41	Ionization energy of calcium is.	A. Lower than that of barium B. Lower than that of magnesium C. Higher than that of beryllium D. Lower than that of strontium
42	Which element when reacts with chlorine forms a polymeric halide.	A. Na B. Be C. Ba D. P
43	Keeping in view size of atoms, which order is the correct one.	A. Mg > Sr B. Ba > Mg C. Lu > Ce D. Cl > I
44	In which one of the following sets do all three particles have the same number of total electrons	A. F <sup>-</sup> , Cl <sup>-</sup> , Br <sup>-</sup> B. Li <sup>+</sup> , Na <sup>+</sup> , K <sup>+</sup> C. N <sup>3-</sup> , O <sup>2-</sup> , F <sup>-</sup> D. Na <sup>+</sup> , Mg <sup>2+</sup> , K <sup>+</sup>
45	Which is the longest periodic table	A. 4 B. 5 C. 6 D. 7
46	Which oxide is more basic in nature.	A. BeO B. MgO C. CaO D. BaO
47	Transition elements are those	A. Which comes before uranium B. Which comes after uranium C. Which are prepared artificially D. Which are in between s-block and p-block elements
48	Which one of the following elements burns in air to form an oxide which, when shaken with water, gives a solution with a pH greater than 7.	A. Carbon B. Magnesium C. Sulphur D. Hydrogen
		A. MgO

49	Which one of the following oxides is amphoteric in nature.	B. $\text{Na}_2\text{O}$ C. $\text{SO}_2$ D. $\text{ZnO}$
50	In which block of periodic table non metals are present.	A. s B. p C. d D. f
51	In modern periodic table VI period contains elements	A. 8 B. 18 C. 10 D. 32
52	In a period, from left to right in the periodic table, the size of atom generally.	A. Increases B. decreases C. Remains constant D. First increase upto the middle of period and then decreases
53	What criteria did Mendeleev use in arranging his periodic table.	A. Atomic mass B. Atomic number C. Mass number D. Density
54	Which element has lowest melting point	A. Beryllium B. Magnesium C. Calcium D. Barium
55	Fluorine is in group VII A of periodic table. Its chemistry will most closely resembles that of.	A. Argon B. Boron C. Iodine D. Sulphur
56	Mark the incorrect statement	A. Metallic character increase down the group B. Metallic character increase from left to right along a period C. Metallic character decrease from left to right along a period D. Metallic character remains the same down the group
57	Element of which group reacts with hydrogen and form ionic hydrides.	A. II A B. IV A C. V A D. VI A