Physics Fsc Part 1 Chapter 6 Online Test

Sr	Questions	Answers Choice
1	No entrpy chagne is associated with	A. isothermal process B. Adiabatic process C. Isobaric process D. Isochoric process
2	KE of molecules of an ideal gas at absolutely zero will be	A. 0 B. Infinite C. Very High D. Below zero
3	Effeciency of a carnot engine is.	A. Infinite B. Ten C. Greater than 1 D. Less Than
4	SI unit of entropy is	A. J/Kg B. J/K C. K gms-1 D. JK
5	Internal energy of a substance, is directly proportional to	A. T B. V C. W D. P
6	No entropy change is associated with	A. Isothermal process B. Adiabatic process C. Isobaric process D. Isochoric process
7	Which is not exmple of ddiablatic process.	A. Rapid escape of air from burst tyre B. Rapid expansion of air C. Conversion of water into ice in refrigerator D. Cloud formation in the atmosphere
8	What one is not an isothermal change.	A. Melting of solid B. Bioling of liquid C. Bursting of bicycle tyre D. Slow expansion of gas
9	In carnot engine, each process is.	A. Reversible B. Preferable Reversible C. Irreversible D. Perfectly irreversible
10	What can be calculated from the curve under PV graph.	A. Heat B. Work done C. Temperatures D. Internal energy
11	Pressure of a gas is directly proportional to average.	A. Potential energy B. Rotational energy C. Translation K.E D. Compressed P.E.
12	In reversible cyclic process the change in entropy of system.	A. Remains constant B. Increase C. Decrease D. Becomes zero
13	The chagne in inernal energy is defined as.	A. Q- W B. Q- T C. Q+ P D. Q - P
14	Carnot Cycle is	A. Reersible B. Irreversible C. Both

		D. Cp - Cv = R
15	Entropy of a system in reversibel process	A. Fluctuates B. Increases C. Is infinite D. Decrease
16	Adiabatic change occurs hen the gas expands or compressed.	A. Rapidly B. Slowly C. Gradually D. Moderately
17	The efficiency of carnot eninge is always.	A. Greater than real engine B. Less than real engine C. Equal to the real engine D. Both a and b
18	The sum of all forms of molecular energies of substance is termed as	A. Kinetic energy B. Potential energy C. Internal energy D. Heat energy
19	Thermodynamics mostly deals with.	A. Measuremetn of quantity B. Transfer of quantity of heat C. Change of state D. Conversion of heat to other forms of energy
20	According to kinetic theory of gases, a a finite volume of a gas consts of very	A. Large numebr of molecules B. Small number ofmolecules C. Both a and b D. Large no of ions L
21	Work done by the system is takn as	A. Positive B. Negative C. Undefined D. None of these
22	What happens to internal energy of an object when its temperature.	A. Decreses B. Increases I C. Fluctuates D. Remains Constant
23	Boyle's law states that " The volume of a given mass of a gas is....."	A. Directly proportional to absolute temperature B. Invesely proportional to absolute temperature C. Directly proportional to density D. Inversely proportional topressure
24	In an isothermal chagne, internal energy.	A. Decrease B. Increase C. Remain same D. Becomes zero
25	'R' is called	A. Universal constant B. Universal per molecule constant C. Universal gas constant D. All of the above
26	The process in which no heat enters or leaves the system is known as.	A. Isothermal process B. Adiabatic process C. Isochoric process D. Isobaric process
27	Chagne in entropy is maximum when temperature of source is that.....of sink	A. Greater than B. Less than C. Equal to D. Zero
28	The efficiency of carnot engine depends upon	A. Sink temperature B. Source temperature C. Both a and b D. The working substance
		A. Postive and on the system

29	When th systme is expanded by additing heat energy, then the work done will be	B. Negative and on the system C. Postive and by the system D. Negative and by the system
30	A system does 600 J of work and at the same time has its internal energy increased by 320 J. How much heat has been supplied.	A. 920 J B. 280 J C. 600 J D. 200 J
31	Collision between gas molecules are perfectly C	A. Elastic E B. Inelastic C. Neither elastic nor inelastic D. All of these
32	In all natural prcoessees where heat flow from one system to another there is always a net increase in	A. Pressure B. Entropy C. Work D. Volume
33	When hot and cold water are mixed the entropy	A. Decrease B. Increase C. Remains constant D. Zero
34	In an adiabatic process, there is no.	A. Change in temperature B. Exchagne of heat C. Chagne is internal energy D. Work done
35	Internal energy is similar to the	A. Vibrational K.E. B. Gravitational P.E. C. K.E. D. All of these
36	The process which is carried out at constant temperature is called.	A. Adiabatic process B. Isothermal process C. Isochoic process D. Isobaric process
37	K = R/NA , Where k is called.	A. Rydsberg constant B. Boltzmann constant C. Stefan constant D. Planck's constant
38	According to kinetic theory of gases, the size of the molecule is.	A. Much smaller than the separation between molecule M B. Much larger than the separation btween molecules C. Botha a and b D. Much larger ahtn the separation between atom
39	When two obnects are made in thermal contact heaving same temperature then they are at.	A. Thermal Equilibrium B. Chemical equilibrium C. Mechanicla Equilibrium D. Physical Equilibrium
40	A device which covnerts thermal eergy into mechanical energy is called.	A. Turbine B. Heat engine C. Carnot engine D. Rerigerator
41	An addition of 400 J of heat casues the increase in internal energy of system is equal to 300 J, then work done will be	A. 100 J B. 200J C. 300 J D. 400 J
42	First law of thermodynamics is based upon law of conservatio of.	A. Mass B. Momentum C. Energy D. Charge
		A. Linear motion

43	The gas molecule sare in	B. Random Motion C. Brownian motion D. Circulatory motion