Acid-Base Chemistry

Sr	Questions	Answers Choice
1	Which of the following is a diprotic acid	A. HNO3 B. H2SO4 C. HCl D. CH3COOH
2	If Ka of acetic acid is 1.8 X 10^{-5}, IT IS	A. Strong acid B. Base C. Weak acid D. Neutral
3	Which species is both a Bronsted acid and base.	A. H2O B. Na+ C. OH- D. Cl-
4	An acid and base react to give a.....acid and base	A. Amphoteric B. Salt C. Conjugate D. Neutral
5	The pH of distilled water at 25 oC is	A. 6 B. 7 C. 5 D. 4
6	Which indicator is typically used for titration involving strong acids and strong bases.	A. Methyl red B. Phenolphthalein C. Bromothymol blud D. Litmus solution
7	The solubility product of AgCl is 2.0 x 10^{-10} mol2 dm-6. The maximum concentration of Ag+ ions in the solution is.	A. 2.0 x 10⁻¹⁰ mol dm⁻³ B. 1.41 x 10⁻⁵ mol dm⁻³ C. 1.0 x 10⁻¹⁰ mol dm⁻³ D. 4.0 x 10⁻²⁰ mol dm⁻³
8	Which one of the following compound is added in purficatio of NaCl in common ion efffect.	A. HCl B. H2SO4 C. HNO3 D. HF
9	In an acid base titration, the equivalence point is reached when.	A. pH of the solution is 7.0 B. The indicagor changes color C. Equal volumes of acid and base have been added D. The reaction stops
10	The solubility prodct of Pb SO4 is	A. 1.6 x 10⁻⁸ B. 1.4 x 10⁻⁴ C. 1.6 x 10⁻⁶ D. 1.6 x 10⁻²
11	pH of 1 x 10 -4 M HCl is	A. 4 B. 1 C. 3.5 D. 7
12	The pH of a buffer doesn't chagne much when	A. Diluted with water B. Acid or base is added C. Heated D. Temp changes
13	The definition of amphoteric is	A. Only acid B. Only base C. Botha acid and base D. Neither
14	If the concentration of Cl- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero

15	Which one is a monoprotic acid	A. H3PO4 B. H2SO4 C. HCl D. H2CO3
16	The conjugate base of HClO4 is	A. ClO4 B. ClO4 ²⁻ C. hClO3 D. H2ClO4+
17	Which has the highest pH	A. 0.01 M NaOH B. 0.1 M NaOH C. 0.1 M HCl D. 0.01 M HCl
18	CH3COONA is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid
19	Which of the following is the conjugate base of water.	A. OH- B. OH+ C. H2O D. H2O+
20	Which one of the following is not the applicatio of solubility products.	A. Determination of solubility from Ksp B. Common ion effect C. Predicting precipitation D. Dissociation of bonds