Chemistry Fsc Part 1 Chapter 9 Online Test

Sr	Questions	Answers Choice
1	Two solution of NaCl and KCl are prepared separately by dissolving same moles of them in the fixed amount of solvent. Which of the following statements is true for these solution	A. KCl solution will have higher boiling point than NaCl solution B. Both the solutions have different boiling point C. KCl and NaCl solution possess same vapour pressure D. KCl solution possesses lower freezing point than NaCl solution
2	Which of the following solution s has the highest boiling point.	A. 5.85% solution of sodium chloride B. 18.0 % solution of glucose C. 6.0% solution of urea D. All have the same boiling point
3	The molal boiling point constant is the ration of the elevation in boiling point to	A. Molarity B. Molality C. Mole fraction of solvent D. Mole fraction of solute
4	The molar boiling point constant is the ratio of the elevation of boiling point to .	A. Molarity B. Molality C. Mole fraction of solvent D. Mole fraction of solute
5	In case of non volatile solute, lowering of vapour pressure is proportional to.	A. Mass fraction of solute B. Mole fraction of solvent C. Mole fraction of solute D. None of the above
6	Azeotropic mixture can be separated into pure components by	A. Distillation B. Fractional distillation C. Vacuum distillation D. None
7	Water and Phenol are partially miscible to each other at room temperature when both liquids are mixed together which is upper layer.	A. Water in Phenol B. Phenol and water C. Pure phenol D. Pure water
8	When an ionic compound is dissolved in water, it dissociate into positive and negative ions, which are surrounded by H2O molecule, This process is known as.	A. Hydrolysis B. Hydration C. Saturation D. solvolysis
9	The temperature which partially immiscible pair of liquid leads to the formation of a single phase in called.	A. Transition temperature B. Absolute temperature C. Consulate temperature D. Room temperature
10	A solution of sucrose is 34.2% The volume of solution containing one mole of solute.	A. 500 cm3 B. 1000 cm3 C. 342 cm3 D. 3420 cm3
11	Which of the following mixtures of liquids show negative deviation	A. Methyl alcohol water B. Hydrochloric acid water C. Carbon di sulphide chloroform D. Chlorobenzene bromobenzene
12	An aqueous solution of ethanol in water has vapour pressure.	A. Equal to that of water B. Equal to that of ethanol C. More than that of water D. Less than that of water
13	The molal boiling point constant is the ratio of elevation of boiling point to	A. Molarity B. Mole fraction of solvent C. Molality D. Mole fraction of solute
14	The molarity of 2% w/v NaOH solution is	A. 2 B. 0.25 C. 0.05 D. 0.5

15	18 g glucose is dissolved in 90 g o water, The relative lowering of vapour pressure is equal to.	A. 1/5 B. 5.1 C. 1/51 D. 6
16	In azeotropic mixture showing positive deviation from Raoult's law, the volume of the mixture is.	A. slightly more than the total volume of the components B. Slightly less than the total volume of the component C. Equal to the total volume of the components D. None of these
17	Which one of the following salts dissolved in water to form a solution with a pH greater than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
18	Which pair of mixture is called ideal solution.	A. Chlorobenzene and bromobenzene B. Water alcohol C. Water ether D. HCl and water
19	In a solution 7.8 g of benzene and 46 g of toluene is present The mole fraction of benzene is.	A. 1/2 B. 1/3 C. 1/5 D. 1/6
20	The liquid pair which is not completely miscible is	A. CH₃OH and water B. Alcohol and water C. Phenol and water D. Benzene and toluene
21	10 g of NaOH have been dissolved per kg of solvent The molality of solution.	A. 0.25 m B. 0.5 m C. 1.0 m D. 2.0 m
22	18 g of glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal to	A. 1/5 B. 5.1 C. 1/51 D. 6
23	The unit of mole fraction is	A. Moles dm⁻³ B. Moles kg⁻¹ C. Gram dm⁻³ D. None
24	Solubility curve of CaCl26H2O shows	A. Decrease in solubility with increase of temperature B. Increase in solubility with increase of temperature C. Discontinuous increase in solubility with temperature D. No effect of temperature on solubility
25	Unit of mole fraction is	A. mol dm-3 B. mol kg-1 C. g dm-3 D. No unit
26	Colligative properties are the properties of.	A. Dilute solution which behave as nearly ideal solutions B. Concentrated solutions which behave as nearly non ideal solutions C. Both a and b D. Neither a nor b
27	In azeotropic mixture showing positive deviation from Raoult's law the volume of the mixture is	A. Slightly more than the total volume of the components B. Slightly less than the total volume of the components C. Equal to the total volume of the components D. None of these
28	Molal boiling point elevation depends upon	A. Nature of solvent B. Natrue of solute C. Vapour pressure of solution D. None of these
29	Melting of ice can be forwarded by the use of.	A. LiCl B. BeCl2 C. NaCl D. Ag Cl
		A. 5.85% solution of sodium chloride B. 18.0% solution of glucose

30	Which of the following solutions has the highest boiling point	D. 10.0% solution of glucose C. 6.0% solution of urea D. All have the same boiling points
31	Depression in the F.P is directly proportional to	A. Molarity of solution B. Molarity of solvent C. Molality of solvent D. Molality of solution
32	Solubility of which substance decreases by increasing temperature.	A. NaNO3 B. KNO2 C. NaCl D. Ce2(SO4)3
33	The vapour pressure of an aqueous solution of glucose is.	A. Equal to vapour pressure of water B. Independent of temperature C. More than vapour pressure of pure water D. Less than vapour pressure of pure water
34	Upper consulate temperature for water phenol system is.	A. 150 ^oC B. 65.9 ^oC C. 120 ^oC D. 130 ^oC
35	18 g glucose is dissolved in 90 g of water. The relative lowering of vapour pressure is equal to.	A. 1/5 B. 5.1 C. 1/51 D. 6
36	What amount of NaOH is required to prepare 500 g of 0.5 molal solution.	A. 10 g B. 20 g C. 30 g D. 40 g
37	Which one of the following salts dissolved in water to form a solution with a pH lesser than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
38	Molal boiling constant for water is 0.52 ^{o}C. If 6 g of urea is dissolved in 100 g of water, what will be its boiling point.	A. 100.52 ^oC B. -100.52^oC C. 100^oC D. 99^oC
39	The relative lowering of vapour pressure is equal to the mole fraction of the solute. This law is known as	A. Ostwald dilution law B. Raoult's law C. Vant hoff's law D. Henry's law
40	Salt of weak acid with strong base when dissolved in water gives.	A. Acidic solution B. Basic solution C. Neutral solution D. None of above
41	Colligative properties are used to determine the	A. Freezing points B. Boiling point C. Atomic mass of an element D. Molar mass of solute
42	10 g of NaOH has been dissolved per dm3 of solution. The molarity of solution is.	A. 0.5 M B. 0.25 M C. 1 M D. 2 M
43	Heat of solution of an ionic compound is equal to.	A. Hydration energy B. Lattice energy C. Slum of both 'a' and 'b' D. Difference of both a and b
44	Colligative properites are the properties of	A. Dilute solutions which behave as nearly ideal solution B. Concentrated solution which behave as nearly non-ideal solution C. Both a and b D. Neither a nor b
45	Depression of freezing point method is used for determination of molar masses of	A. Electrolytes B. Non-volatile solids C. Volatile solids D. Volatile liquids
46	Which of the following solutions will have the highest boiling point	A. 0.1 M NaCl B. 0.1 M CaCl₂ C. 0.1 M FeCl₃ D. 0.1 M glucose
		A. Obey Raoult's law

47	Azeotropic mixture	B. Do not obey Raoult's law C. Boils at low temperature only D. Boils at high temperature only
48	Molarity of pure water is	A. 1 B. 18 C. 55.5 D. 6
49	Azeotrpic mixture of two liquids boils at a lower temperature than either of them, when	A. It is saturated B. Is shows positive deviation from Raoult's law C. It shows negative deviation from Raoult's law D. Is is metastable
50	Solubility of which substance decreases by incressing temperature.	A. NaNO3 B. KNO2 C. NaCl D. Ce2(SO4)3
51	A solution of glucose is 10% The volume in which 1 g mole of it dissolved will be.	A. 1 dm3 B. 1.8 dm3 C. 900 cm3 D. 200 cm3
52	An azeotropic mixture of two liquids boils at lower temperature than either of them when.	A. It is saturated B. It shows positive deviation from Raoult's law C. It shows negative deviation from Raoult's law D. It is metastable
53	Which solution is an example of solid in gas	A. Fog B. Steel C. smoke D. Air
54	Relative lowering of vapour pressure is equal to.	A. Mole fraction of solute B. Mole fraction of solvent C. Mole fraction of solute and solvent D. Molality of solution
55	When an ionic compound is dissolved in water, it dissociate into positive and negative ions, which are surrounded byH2O molecule, This process is known as.	A. Hydrolysis B. Hydration C. Saturation D. solvolysis
56	Two solutions of NaCl and KCl are prepared separately by dissolving same amount of the solute in water. Which of the following statements is true for these solution.	A. KCl solution will have higher boiling pint than NaCl solution B. Both the solutions have different boiling point C. KCl and NaCl solutions possess same vapour pressure D. KCl solution possesses lower freezing point than NaCl solution
57	Butter is solution of	A. Liquid in liquid B. Solid and liquid C. Liquid and solid D. Liquid and gas
58	Which one of the following salts dissolved in water to form a solution with a pH greater than 7	A. NaCl B. CuSO4 C. Na2CO3 D. NH4Cl
59	A negative deviation from Raoult's law in solution means, the solution has	A. High boiling point and high vapour pressure B. High boiling point and low vapour pressure C. Low boiling point and low vapour pressure D. Low boiling point and high vapour pressure
60	The molal boiling point constant is the the ratio of the elevation in boiling point to.	A. Molarity B. Molality C. Mole fraction of solvent D. Mole fraction of solute
61	Relative lowering of vapour pressure is equl to.	A. Mole fraction of solute B. Mole fraction of solvent C. Molarity D. Molality
62	An aqueous solution of ethanol in water has vapour pressure	A. Equal to that water B. Equal to that of ethanol C. More than that of

		H₂O D. less than that of water
63	An aqueous solution of ethanol is water has vaporu pressure.	A. Equal to the of water B. Equal to that of ethanol C. More than that of H2O D. Less than that of water
64	Which concentration unit is independent of temperature.	A. Molarity B. Molality C. ppm D. both a and b
65	A solution of glucose is 10% to volume in which 1 g mole of it is dissolved will be	A. 1 dm³ B. 1.8 dm³ C. 200 cm³ D. 900 cm³
66	The sum of mole percent of all the components of solution is always equal to.	A. Unity B. 100 C. Less than one D. Less than 100
67	Which one of the following salts do not hydrolyses	A. CuSO4 B. Na2CO3 C. NaCl D. AlCl3