Acid-Base Chemistry

Sr	Questions	Answers Choice
1	Which indicator is typically used for titration involving strong acids and strong bases.	A. Methyl red B. Phenolphthalein C. Bromothymol blud D. Litmus solution
2	Which is amphoteric	A. H2O B. HCl C. NaOH D. NH3
3	The pKa value for HCOOH is.	A. 4.74 B. 3.78 C. 4.78 D. 4.24
4	Which has the highest pH	A. 0.01 M NaOH B. 0.1 M NaOH C. 0.1 M HCl D. 0.01 M HCl
5	Which species is both a Bronsted acid and base.	A. H2O B. Na+ C. OH- D. Cl-
6	Which of the following is a weak base.	A. KOH B. NaOH C. NH3 D. Ca(OH)2
7	A salt that hydrolyzes in water in from	A. Strong acid and base B. Weak acid or base C. None D. Both strong
8	pH of 10^{-9} M HCl is	A. 5 B. 9 C. Slightly below 7 D. Exactly -9
9	Litmus turns which color in a basic solution.	A. Red B. Orange C. Blue D. Colorless
10	NaOH is not considered an Bronsted Lowry acid because.	A. It's neutral B. It's not soluble C. Doesn't donate H+ D. Doesn't produce H+
11	pH of 1 x 10 -4 M HCl is	A. 4 B. 1 C. 3.5 D. 7
12	Strong bases have	A. Low Kb B. High Kb C. Low pH D. High Ka
13	Which of the following is a tribasic acid.	A. H3PO4 B. HNO3 C. HCl D. H2SO4
14	Acid rain is mainly due to	A. SO3 and NO3 B. CO2 C. H2 D. CH4
15	CH3COONA is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid

16	The conjugate acid of NH_3 is	A. NH_4 B. NH_2 C. NO_3 D. N_2H_4
17	Strong acids completely dissociate in water because.	A. High solubility B. High Ka C. Forms complex D. Low Ka
18	A solution with equal $[H^+]$ and $[OH^-]$ is	A. Basic B. Acidic C. Neutral D. Salt
19	If the concentration of Cl^- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero
20	The sgtrength of base is measured by its.	A. Ka B. Kb C. Pkw D. Kw