Acid-Base Chemistry

Sr	Questions	Answers Choice
1	A solution with equal [H+] and [OH-] is	A. Basic B. Acidic C. Neutral D. Salt
2	Which of the following is a diprotic acid	A. HNO3 B. H2SO4 C. HCl D. CH3COOH
3	The definition of amphoteric is	A. Only acid B. Only base C. Botha acid and base D. Neither
4	Which dialyzing solution has highes pH	A. Vinegar B. Lemon juice C. Ammonia solution D. Coffee
5	Which of the following is a weak base.	A. KOH B. NaOH C. NH3 D. Ca(OH)2
6	Which has the highest pH	A. 0.01 M NaOH B. 0.1 M NaOH C. 0.1 M HCl D. 0.01 M HCl
7	Which has the lowest pH?	A. 0.1 M CH3COOH B. 1 M HCl C. 1 M NH3 D. 0.1 M HCl
8	The pH of a buffer doesn't chagne much when	A. Diluted with water B. Acid or base is added C. Heated D. Temp changes
9	Strong bases have	A. Low Kb B. High Kb C. Low pH D. High Ka
10	If Ka of acetic acid is 1.8 X 10^{-5}, IT IS	A. Strong acid B. Base C. Weak acid D. Neutral
11	A 0.01 M solution of a strong acid has a pH of	A. 3 B. 2 C. 5 D. 4
12	If the concentration of Cl- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero
13	The pKa value for HCOOH is.	A. 4.74 B. 3.78 C. 4.78 D. 4.24
14	Which acid is present in vinegar	A. CH3COOH B. HNO3 C. H2SO4 D. HCl
15	The solubility product of AgCl is 2.0 x 10^{-10} mol2 dm-6. The maximum concentration of Ag+	A. 2.0 x 10⁻¹⁰ mol dm⁻³ B. 1.41 x 10⁻⁵ mol dm⁻³

	ions in the solution is.	C. 1.0 x 10⁻¹⁰ mol dm⁻³ D. 4.0 x 10⁻²⁰ mol dm⁻³
16	The pH of distilled water at 25 oC is	A. 6 B. 7 C. 5 D. 4
17	Which of the following is the conjugate base of water.	A. OH- B. OH+ C. H2O D. H2O+
18	Which is a weak acid	A. HNO3 B. H2SO4 C. HF D. HCl
19	KOH is an example of	A. Strong Base B. Weak acid C. Neutral salt D. Weak base
20	A salt from stron acid+weak base gives.	A. Neutral solution B. Acidic solution C. Precipitate D. Basic solution