Acid-Base Chemistry

Sr	Questions	Answers Choice
1	pH of 10^{-9} M HCl is	A. 5 B. 9 C. Slightly below 7 D. Exactly -9
2	Which of these is NOT a property of acids.	A. Conduct electricity B. Turn litmus red C. Sour taste D. Feel slippery
3	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Elecron pair donors
4	Which dialyzing solution has highes pH	A. Vinegar B. Lemon juice C. Ammonia solution D. Coffee
5	pH of blood is around.	A. 6.0 B. 7.4 C. 5.4 D. 8.0
6	A salt from stron acid+weak base gives.	A. Neutral solution B. Acidic solution C. Precipitate D. Basic solution
7	The conjugate base of HClO4 is	A. ClO4 B. ClO4 ²⁻ C. hClO3 D. H2ClO4+
8	Which of the following pair form a buffer solution .	A. HCl and NaCl B. CH3COONa and CH3COOH C. NaOH and HCl D. NH3 and Na2SO4
9	A buffer solution is.	A. Stron acid+ strong base B. Weak Acid + Its salt C. Weak base + Its salt D. Weak base+ salt of strong acid
10	Which one is a monoprotic acid	A. H3PO4 B. H2SO4 C. HCl D. H2CO3
11	HCl +NH3---------- NH4Cl shows	A. Redox reaction B. Acid base reaction C. Bronsted acid base D. Precipitation
12	An acid and base react to give a.....acid and base	A. Amphoteric B. Salt C. Conjugate D. Neutral
13	The sgtrength of base is measured by its.	A. Ka B. Kb C. Pkw D. Kw
14	Acid rain is mainly due to	A. SO3 and NO3 B. CO2 C. H2 D. CH4
15	Which of the following is the conjugate base of water.	A. OH- B. OH+ C. H2O D. H2O+

16	In a titration, the equivalence point is when	A. No acid left B. pH = 7 C. Moles of acid= moles of base D. Buffer forms
17	The logarithmic nature of pH means.	A. Each pH unit = 2 x H+ B. 10 X H+ Per unit C. H+ drop linearly D. No change in H+
18	Which is amphoteric	A. H2O B. HCl C. NaOH D. NH3
19	CH3COONA is a salt of.	A. Strong acid+ Strong base B. Weak acid+Strong base C. None D. Weak base+Weak acid
20	A solution with equal [H+] and [OH-] is	A. Basic B. Acidic C. Neutral D. Salt