Chemical Equilibrium

Sr	Questions	Answers Choice
1	Kc = 0.040 at 450 oC for the teh given reaction, evaluate Kp for the reaction. Pcl5 ..... PCl2 + Cl2	A. 0.40 B. 2.4 C. 0.64 D. 0.052
2	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversibel systems D. Combustion reactions
3	According to law of mass action rate of reaction is proportional to.	A. Temperature B. Pressure C. Product of active masses D. Atomic mass
4	Which of the following statemetns correctly describes the effect of temperature on the equilirium constant.	A. Kc is directly proportional to temperature B. Kc is inversely proportional to temperature. C. Kc depends on the enthalpy change of the reaction D. Temperature has no effect on the value of Kc
5	If Delta n is positive, then Kp is	A. Greater than Kc B. Less than Kc C. equal Kc D. Zero
6	Catalyst affects	A. Value of K B. Equilibrium position C. Activation energy D. Final concentrations
7	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
8	Increasing temperatur efavors	A. Exothermic reaction B. Endothermic reaction C. Formation of solid D. Reverse in all cases
9	Reaction in Haber process is	A. Endothermic B. Exothermic C. Irrversible D. Neutral
10	Optimum temperatur ein Haber process is	A. 50 oC B. 450 oC C. 200 oC D. 1000 oC
11	Active mass means	A. Moles B. Volume C. Mass D. Molar concentration
12	CH3COONa in water forms	A. Acidic solution B. Basic Solution C. Neutral Solution D. Salt bridge
13	Removing product from equilibrium	A. Shifts equilibrium left B. Stops reaction C. Shifts rquilibrium right D. Has no effect
14	Contact process is used for	A. Sulfuric acid production B. Ammonia Synthesis C. Nitric Acid production D. Hydrogenation

15	Dynamic equilibrium occurs.	A. In open systems B. Only in gases C. In closed systems D. Only at low temperature
16	NH_4Cl in water makes solution	A. Neutral B. Acidic C. Basic D. amphoteric
17	Salts of weak acid + strong base are.	A. Basic B. Acidic C. Amphoteric D. Neutral
18	The reversibel reation cannot be achieved in	A. Open system B. Closed system C. Both a and b D. None of these
19	Lowering temperatur ein an exothermic reaction.	A. Favors reverse B. Favors forward C. No effect D. Stop the reaction
20	If $Q > K$, the reaction	A. Shift left B. Moves forward C. At equilirbium D. Stops