Chemical Equilibrium

Sr	Questions	Answers Choice
1	Position of equilibrium is affected by	A. Temperature B. Catalyst C. Inert gas D. Surface area
2	If Kc> 1, the reaction.	A. Favors products B. Favors reactants C. Is at equilirbium D. Does not occur
3	Consider the gas phase equilibrium sysem represented by teh equation 2 H2O..... 2H2+O2 Givent that the forward reactio is endothermic, which of the following chagnes will decrease the equilibrium amount of H2O.	A. Adding more oxygen B. Adding a solid phase catalyst C. Decreasing the volume of the container D. Increaseing the temperature at constant pressure
4	Catalyst used in Haber process is	A. Fe B. Pt C. Cu D. Zn
5	Kc is expressed in terms of.	A. Pressure B. Mole fraction C. Concentration D. Volume
6	A saturted solution represents a dynamic equilibrium Macroscopically, the concentration of dissolved solute is constant, Microscopically this occurs because.	A. No more solute particles are dissolving B. The rate of dissolution of solute is zero C. Solute particles are dissolving and precipitating at the same rate D. All solute particles have dissolved
7	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reation left C. Shifts reaction right D. No effect
8	In a reversible reaction.	A. Products do not reform reactants B. Rate of forward reaction is always higher C. Both forward and reverse reactions occur D. Products are in excess
9	Increase in concentratio of reactants	A. Increase K B. Decrease K C. Shifs equilibrium forward D. Stops reverse reaction
10	Which one of the following is not an example of reversible reaction.	A. Formation of ammonia B. Foramtion ow water C. Decomposition of PCl5 D. Decomposition of NO2
11	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversibel systems D. Combustion reactions
12	Which is NOT a feature of dynamic equilibrium.	A. Closed system B. Constant Temperature C. Uequal reaction rates D. No net change
13	The reversibel reation cannot be achieved in	A. Open system B. Closed system C. Both a and b D. None of these
14	Dynamic equilirbrium is attained when	A. Forward reaction stops B. Reverse reaction stops C. Forward and reverse reactions

	[illegible]	continue at equal rates D. Concentration of products become zero
15	For a specific reaction the valu eof the equilibrium constant, Kc?	A. Always remains the same at differrent reaction conditions B. Increases if the concentration of one ofthe product is increased C. Changes with changes in the temperature D. Increasees if the concentration of one of the ratants is increased
16	Dynamic equilibrium occurs.	A. In open systems B. Only in gases C. In closed systems D. Only at low temperature
17	Catalyst affects	A. Value of K B. Equilibrium position C. Activation energy D. Final concentrations
18	A reversibel reaction is one which	A. Proceeds to completion B. Occurs only in one direction C. Proceeds in both directions D. Has no products
19	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperture
20	Increasing temperatur efavors	A. Exothermic reaction B. Endothermic reaction C. Formation of solid D. Reverse in all cases