Chemical Equilibrium

Sr	Questions	Answers Choice
1	Na2CO3 in wate gives.	A. Acidic solution B. Basic Solution C. Neutral solution D. Buffer
2	Removing a reactant shifts equilibrium to	A. Left B. Right C. No shift D. Depends on temperture
3	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversibel systems D. Combustion reactions
4	Lowering temperatur ein an exothermic reaction.	A. Favors reverse B. Favors forward C. No effect D. Stop the reaction
5	If Q=K , then	A. Reaction is irrversible B. System is at equilibium C. Forward reaction dominates D. Reverse reaction dominates
6	Dynamic equilibrium occurs.	A. In open systems B. Only in gases C. In closed systems D. Only at low temperature
7	Pressure used in Haber Process.	A. 2 atm B. 300 atm C. 200 atm D. 5 atm
8	Active mass means	A. Moles B. Volume C. Mass D. Molar concentration
9	Increase in concentratio of reactants	A. Increase K B. Decrease K C. Shifs equilibrium forward D. Stops reverse reaction
10	Contact process is used for	A. Sulfuric acid production B. Ammonia Synthesis C. Nitric Acid production D. Hydrogenation
11	Catalyst used in contact process	A. Fe B. V2O5 C. Ni D. Al2O3
12	The reversibel reation cannot be achieved in	A. Open system B. Closed system C. Both a and b D. None of these
13	Untis of Kc depend on	A. Catalyst B. Reaction stoichimometty C. Activation energy D. Delta H
14	Which of the following is an irreversible reaction.	A. Haber process B. Precipitation of Ag Cl C. Synthesis of ammonia D. Esterification
15	Which is NOT a feature of dynamic equilibrium.	A. Closed system B. Constant Temperature C. Uequal reaction rates D. No net change

16	Which one of the following is not an example of reversible reaction.	A. Formation of ammonia B. Foramtion ow water C. Decomposition of PCl5 D. Decomposition of NO2
17	Which statement is true at dynamic equilibrium.	A. No reactio is occurring B. Concentrations are changing C. Rates of forward and reverse reactions are equal D. Rate forward reaction < reverse
18	Le Chatelier's Principle applies to.	A. Irreversibel reactions B. Static equilibrium C. Dynamic equilibrium D. Precipitation reactions
19	A large K value indicates.	A. Prodcuts are favored B. Reactants are favored C. No reaction D. Slow reaction
20	Removing product from equilibrium	A. Shifts equilibrium left B. Stops reaction C. Shifts rquilibrium right D. Has no effect