

Chemistry Fsc Part 1 Chapter 7 Online Test

Sr	Questions	Answers Choice
1	The born Haber cycle is the best application of law.	A. Boyle's B. Dalton's C. Hess's D. Graham's
2	At constant volume of a system remains constant and the heat is absorbed by the system, then amount of heat absorbed is called	A. Enthalpy change of the system B. Internal energy change of the system C. Total enthalpy of the system D. Total internal energy of the system
3	A material or a collection of materials which is under study is called	A. State function B. Degree and joule C. Degree and ergs D. Calorie and joule
4	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Hess's law C. Joule's principle D. Law of combustion
5	What happens to the enthalpy change when the coefficients of a chemical equation are doubled	A. It doubles B. It becomes half C. It does not change D. It cannot be predicted
6	In endothermic reactions, the heat content of the	A. Products is more than that of reactants B. Reactants is more than that of products C. Both a and b D. Reactant and products is equal
7	Which one of the following process is endothermic	A. condensation of steam B. Freezing of water C. electrolysis of water D. None of these
8	Which is not a state function	A. Enthalpy B. Entropy C. Pressure D. work
9	Question Image	
10	Which one of the following enthalpies is always an exothermic process.	A. Enthalpy of atomization B. Enthalpy of neutralization C. Enthalpy of ionization D. Enthalpy of dissociation
11	For a given process, the heat change at constant pressure (q_p) and at constant volume (q_v) are related to each other as	A. $q_p = q_v$ B. $q_p < q_v$ C. $q_p > q_v$ D. $q_p = q_v/2$
12	the value of ΔH and ΔE for liquids and solids is almost same because.	A. No change in temperature B. Heat absorbed C. No change in volume D. Heat evolved
13	The enthalpy of combustion is	A. Positive B. Negative C. Either positive or negative D. No correlation
14	Spontaneous processes are mostly	A. Reversible B. Irreversible C. Not irreversible D. None of these
15	For a given process, the heat changes of constant pressure and at constant volume are related to each other as.	A. $q_p = q_v$ B. $q_p < q_v$ C. $q_p > q_v$ D. $q_p = q_v/2$

		D. None of these
16	Which of the following value of heat of formation indicates that the product is least stable.	A. -94 KJ B. -231.6 KJ C. +21.4 KJ D. +70 KJ
17	Which substance have $\Delta E = \Delta H$ and no pressure - volume work.	A. Liquids only B. Solids only C. Gases only D. Liquids and solids
18	Question Image	A. Heat of reaction B. Heat of sublimation C. Heat of neutralization D. Heat of combustion
19	Which process is endothermic and spontaneous	A. Neutralization of NaOH with HCl B. Formation of NH_3 from H_2 and N_2 C. Formation of H_2O from H_2 and O_2 D. Evaporation of sea water
20	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air	A. Remains constant B. Increase C. Decrease D. Increase a bit
21	The heat contents of the system is known as.	A. Entropy B. Enthalpy C. Work D. Free energy
22	Whenever a reaction is endothermic, then it means that	A. Heat is transferred from surrounding to the system B. Heat is transferred system to the surrounding C. Heat content of the product is greater than that of reactants D. Heat content of the reactants is greater than the products
23	Born Haber cycle is used to determine the	A. Lattice energy B. Enthalpy of formation C. Enthalpy of ionization D. Enthalpy of dissociation
24	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Hess's law C. Joule's principle D. Law of conservation of energy
25	Most of the reaction which give stable products are	A. Exothermic B. Endothermic C. Isothermal D. None of these
26	In endothermic reactions, the heat content of the.	A. Products is more than that of reactants. B. Reactants is more than that of products C. Both a and b D. Reactants and products are equal
27	Standard enthalpy change when one mole of compound is formed from their elements at standard state is.	A. Heat of formation B. Standard heat of formation C. Heat of combustion D. Standard Heat of neutralization
28	Standard enthalpy of combustion of carbon is -394 kJ mol^{-1} than, which is the standard enthalpy of formation of CO_2	A. +394 kJ B. -394 kJ C. 0 kJ D. +197 kJ
29	The change in heat energy of a chemical reaction at constant temperature and pressure is called	A. Enthalpy change B. Bond energy C. Heat of sublimation D. Internal energy change
30	Which of the following statement is contrary to the first law of thermodynamics	A. An equivalent amount of heart energy can neither be created nor destroyed B. One form of energy can be transferred into an equivalent amount of other kinds of energy C. In an adiabatic process, the work done is independent of its path D. Continuous production of mechanical work without supplying an equivalent is amount of heat possible

31	The pressure of oxygen inside the bomb calorimeter is.	A. 100 atm B. 50 atm C. 25 atm D. 20 atm
32	Calorie is equivalent to	A. 0.4184 J B. 41.84 J C. 4.184 J D. 418.4 J
33	If an endothermic reaction is allowed to take place very rapidly i the air, the temperature fo the surrounding air.	A. Remains constant B. Increase C. Decrease D. Remain unchanged
34	The change in heat energy of a chemical reaction at constant temperature and pressure is called.	A. Enthalpy change B. Bond energy C. Heat of sublimation D. Internal energy change
35	One of the following statements about Born-Haber cycle is correct. Which is that statement	A. Born-Haber cycle is different from Hess's law B. The energy change in a cyclic process is not zero C. The lattice energy of the crystalline substances can be calculated easily D. Heat of formation of the product and the lattice energy of the substance can be calculated simultaneously
36	The standard heat of formation is measured at 1 atmosphere and	A. 0 ^o C B. 100 ^o C C. 293 ^o C D. 25 ^o C
37	The S.I. units for the molar heat capacity are	A. Joule Cm ⁻³ deg ⁻¹ B. Joule deg ⁻¹ atm ⁻¹ C. Joule deg ⁻¹ mol ⁻¹ D. Joule deg ⁻¹ kg ⁻¹
38	The net change in a chemical reaction is same whether it takes place directly or indirectly is	A. Henry's law B. Charles's C. Hess's lass D. Graham's law
39	The optimum temperature for the synthesis of NH ₃ by Haber's process is.	A. 200 ^o C B. 300 ^o C C. 400 ^o C D. 500 ^o C
40	A chemical change always involve	A. Absorption of heat B. Evolution of heat C. Either absorption or evolution of heat D. The liberation of heat and light energy
41	The sum of all kinds of energies of atoms, ions or molecules of a system is known as.	A. Kinetic energy B. Potential energy C. Internal energy D. Solar energy