Reaction Kinetics

Sr	Questions	Answers Choice
1	k has unit s-1 for	A. First order B. Second order C. Third order D. zero order
2	In pressure of catalyst Delta H	A. Increases B. Decreases C. Double D. Unchanged
3	Overall order is sum of	A. Coefficients B. Exponents in rate law C. Moles D. Products
4	Reaction rate by conductometry depends on	A. Ionic conductivity B. Pressure C. colour D. Temperature
5	Unit of k for second order reaction is	A. mol dm-3 B. s-1 C. mol-1 dm3 s-1 D. mol -2 dm⁶ s-1
6	Mechanism supports	A. Hess's law B. Kinetic data C. Le chatelier's principle D. Boyle's law
7	On a Boltzmann distribution curve, the activation energy is represented by	A. The height of the peak B. The area under the entire curve C. A vertical line drawn at a specific kinetic energy value D. The difference between the peak and the X axis
8	Rate constant can be determined by	A. Titration B. spectroscopy C. conductometry D. all of these
9	Which one reduces activation energy.	A. Catlyst B. Inhibitor C. Product D. Reactant
10	In a chemical reaction, a catalyst.	A. Alters Delta H B. Lowers activation energy C. Is consumed D. Forms product
11	Rate constant dependent of	A. Time B. Temperature C. Concentration D. Catalyst
12	In color change reactions, which method is best	A. Colorimetry B. Titration C. Conductometry D. Manometry
13	Which of the following methods measures reaction rate.	A. Conductivity B. Titration C. Volume of gas D. All of these
14	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed
15	The units of the rate constant (k) for a reaction depend on the	A. Activation energy of the reaction B. Temperature of the reaction C. Overall order of the reaction

		D. Stoichiometry of the balanced chemical equation
16	If we double the conentration of a reactant, the rate increases by four times, the reactionis.	A. First Order B. Second Order C. Third Order D. Zero order
17	Which factor affects rate of reaction.	A. Concentration B. Temperature C. Surface Area D. All of these
18	Unit of rae constant for zero order reaction is	A. mol dm-3 s-1 B. s-1 C. mol -1 dm³ s-1 D. mol⁻² dm⁶ s-1
19	The rate determinig step in a multistep reaction is	A. The slowest step B. Always the first step C. Always the last step D. The fastest step
20	Half life formula for 1st order reaction is	A. 0.693 /k B. k x t C. 1 /k D. 2 k