Chemistry Fsc Part 1 Chapter 7 Online Test

Sr	Questions	Answers Choice
1	Standard enthalpy of combustion of carbon is -394 kJ mol-1 than, which is the standard enthalpy of formation of CO2	A. +394 kJ B. -394 kJ C. 0 kJ D. +197 kJ
2	The enthalpy of combustion is	A. Positive B. Negative C. Either positive or negative D. No correlation
3	Standard enthalpy change when one mole of compound is formed from their elements at standard state is.	A. Heat of formation B. Standard heat of formation C. Heat of combustion D. Standard Heat of neutralization
4	Question Image	A. Heat of reaction B. Heat of sublimation C. Heat of neutralization D. Heat of combustion
5	In endothermic reactions, the heat content of the	A. Products is more then that of reactants B. Reactants is more than that of products C. Both a and b D. Reactant and products is equal
6	A chemical change always involve	A. Absorption of heat B. Evolution of heat C. Either absorption or evolution of heat D. The liberation of heat and light energy
7	Which of the following value of heat of formation indicates that the product is least stable.	A. -94 KJ B. -231.6 KJ C. +21.4 KJ D. +70 KJ
8	The optimum temperature for the synthesis of NH3 by Haber's process is.	A. 200 ^oC B. 300 ^oC C. 400 ^oC D. 500 ^oC
9	Question Image	
10	For a given process, the heat change at constant pressure (q_p) and at constant volume (q_v) are related to each other as	A. q_p= q_v B. q_p< q_v C. q_p> q_v D. q_p= q_v/2
11	Calorie is equivalent to	A. 0.4184 J B. 41.84 J C. 4.184 J D. 418.4 J
12	Born Haber cycle is used to determine the	A. Lattice energy B. Enthalpy of formation C. Enthalpy of ionization D. Enthalpy of dissociation
13	Which substance have delta E = Delta H and no pressure - volume work.	A. Liquids only B. Solids only C. Gases only D. Liquids and solids
14	For a given process, the heat chagnes of constant pressure and at constant volume are related to each other as.	A. qp= qv B. qp<qv C. qp>qv D. None of these
15	The change in heat energy of a chemical reaction at constant temperature and pressure is called	A. Enthalpy change B. Bond energy C. Heat of sublimation

		D. Internal energy change
16	Which one of the following enthalpies is always an exothermic process.	A. Enthalpy of atomization B. Enthalpy of neutralization C. Enthalpy of ionization D. Enthalpy of dissociation
17	The standard heat of formation is measured at 1 atmosphere and	A. 0 ^oC B. 100 ^oC C. 293 ^oC D. 25 ^oC
18	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is knnown as	A. Henry's law B. Hess's law C. Joule's principle D. Law of combustion
19	Which is not is state function	A. Enthalpy B. Entropy C. Pressure D. work
20	The S.I. units for the molar heat capacity are	A. Joule Cm⁻³degree⁻¹ B. Joule deg⁻¹atm⁻¹ C. Joule deg⁻¹mol⁻¹ D. Joule deg⁻¹kg⁻¹
21	The pressure of oxygen inside the bomb calorimeter is.	A. 100 atm B. 50 atm C. 25 atm D. 20 atm
22	A material or a collection of materials which is under study is called	A. State function B. Degree and joule C. Degree and ergs D. Calorie and joule
23	Which process is endothermic and spontaneous	A. Neutralization of NaOH with HCl B. Formation of NH3 from H2 and N2 C. Formation of H2O from H2 and O2 D. Evaporation of sea water
24	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air	A. Remains constant B. Increase C. Decrease D. Increase a bit
25	The heat contents of the system of known as.	A. Entropy B. Enthalpy C. Work D. Free energy
26	Most of the reaction which give stable products are	A. Exothermic B. Endothermic C. Isothermal D. None of these
27	the value of delta H and delta E for liquids and solids is almost same because.	A. No change in temperature B. Heat absorbed C. No change in volume D. Heat evolved
28	One of the following statements about Born-Haber cycle is correct. Which is that statement	A. Born-Haber cycle is different from Hess's law B. The energy change in a cyclic process is not zero C. The lattice energy of the crystalline substances can be calculated easily D. Heat of formation of the product and the lattice energy of the substance can be calculated simultaneously
29	The born Haber cycle is the best application of law.	A. Boyle's B. Dalton's C. Hess's D. Graham's
30	The net change in a chemical reaction is same whether it takes place directly or indirectly is	A. Henry's law B. Charles's C. Hess's lass D. Graham's law
		A. condensation of steam

31	Which one of the following process is endothermic	A. condensation of steam B. Freezing of water C. electrolysis of water D. None of these
32	In endothermic reactions, the heat content of the.	A. Products is more than that of reactants. B. Reactants is more than that of products C. Both a and b D. Reactants and products are equal
33	The sum of all kinds of energies of atoms, ions or molecules of a system is known as.	A. Kinetic energy B. Potential energy C. Internal energy D. Solar energy
34	The net heat change in a chemical reaction is same whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Hess's law C. Joule's principle D. Law of conservation of energy
35	Spontaneous processes are mostly	A. Reversible B. Irreversible C. Not irreversible D. None of these
36	Whenever a reaction is endothermic, then it means that	A. Heat is transferred from surrounding to the system B. Heat is transferred system to the surrounding C. Heat content of the product is greater than that of reactants D. Heat content of the reactants is greater than the products
37	At constant volume of a system remains constant and the heat is absorbed by the system, them amount of heat absorbed is called	A. Enthaply change of the system B. Internal energy change of the system C. Total enthalphy of the system D. Total internal energy of the system
38	The change in heat energy of a chemical reaction at constant temperature and pressure is called.	A. Enthalpy change B. Bond energy C. Heat of sublimation D. Internal energy change
39	What happens to the enthalpy change when the coeffficients of a chemical equation are douubled	A. It doubles B. It becomes half C. It does not change D. It cannot be predicted
40	If an endothermic reaction is allowed to take place very rapidly i the air, the temperature fo the surrounding air.	A. Remains constant B. Increase C. Decrease D. Remain unchanged
41	Which of the following statement is contrary to the first law of thermodynamics	A. An equivalent amount of heart energy can neither be created nor destroyed B. One form of energy can be transferred into an equivalent amount of other kinds of energy C. In an adiabatic process, the work done is independent of its path D. Continuous production of mechanical work without supplying an equivalent is amount of heat possible