

Reaction Kinetics

Sr	Questions	Answers Choice
1	The rate determinig step in a multistep reaction is	A. The slowest step B. Always the first step C. Always the last step D. The fastest step
2	Half life is time is which	A. Rate becomes zero B. Concentration reduces to 1/4 C. Concentratio reduces to half D. Product is half formed
3	Rate of reaction is measured as.	A. Increase Temperature B. Change is concentration per unit time C. Total eergy of reactants D. Change in volume per mole
4	Delta G tells about	A. Speed B. Mechanism C. Feasibility D. All of these
5	Which one reduces activation energy.	A. Catlyst B. Inhibitor C. Product D. Reactant
6	On Boltzmann distribution curve, the area under the curve represents.	A. Activation energy of the rreaction B. Total numebr f moleclues in the sample C. Average kinetic energy of the molecules D. Rate constant of the reaction
7	Which of the following methods measures reaction rate.	A. Conductivity B. Titration C. Volume of gas D. All of these
8	Unit of k for second order reaction is	A. mol dm-3 B. s-1 C. mol-1 dm3 s-1 D. mol -2 dm ⁶ s-1
9	Homogenous catalysts are in	A. Same phase as reactionts B. Different phase C. Solid only D. Gaseous only
10	Which increases with rise in temperature?	A. Activation energy B. Enthalpy C. Collision frequency D. Molecular weight
11	Half life of first order reaction	A. Become zero B. Increases with [A] C. Is independent of [A] D. Decreases with [A]
12	For gaseous reactions, rate is ofter monitored by	A. Mass change B. pH change C. Pressure change D. Temperature change
13	If we double the conentration of a reactant, the rate increases by four times, the reactionis.	A. First Order B. Second Order C. Third Order D. Zero order
14	Which order has constant half life	A. First B. Zero C. Second D. Third
		A. Overall rate

15	Rate determining step controls	B. I emperature C. Volume D. Final yield
16	Rate constant dependent of	A. Time B. Temperature C. Concentration D. Catalyst
17	On a Boltzmann distribution curve, the activation energy is represented by	A. The height of the peak B. The area under the entire curve C. A vertical line drawn at a specific kinetic energy value D. The difference between the peak and the X axis
18	Which unit is common for all orders	A. mol dm-3 B. No common unit C. J mol -1 D. s-1
19	Higher pressur eincreases rate for	A. Solids B. Gases C. Liquids D. All phases
20	First order reactions are best studies using.	A. log [A] vs time B. Pressure vs time C. [A] vs time D. Titriation