States and Phases of Matter

Sr	Questions	Answers Choice
1	What typ of molecular shape in communly found in liquid crystals.	A. Spherical B. Rod like and elongated C. Spiral D. Cubic
2	Which of the following is a characteristic property of crystalling solids.	A. they have a range of melting points B. They are isotrople C. they have adefinite and sharp melting point D. They lack a regular arrangeness of particles
3	What happens to the volue of water when it freeze	A. It contracts by 9% B. It remains the same C. It expends by 9% D. It evaporates completely
4	Which of the following is a polar moleuce taht shows permanent dipole interactions.	A. CHCl B. CH4 C. Cl2 D. H2
5	Which of the folloowing halogens has the highest boiling point.	A. I2 B. Cl2 C. F2 D. Br2
6	What is vapour pressure.	A. Atomospheric pressure. B. Pressure exerted by aliqudi at 0 oC C. pressure exercted by vaporus in equilibrium with their liqudi at a given temperature D. Pressrue inside the liquid only
7	Earthenware pots keep water cool because.	A. They are made of clay B. They reflect sunlight C. They absorb water D. They allow water to wvaporate through pores
8	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The partiles are charged D. Solids are made of gases
9	Which of the following materials is an example of an amorphous solid	A. Glass B. Ice C. Diamon D. Sodium Chaloride
10	Surface tensin of liquid is due to.	A. Inword pull of surface molecule B. Upard pull from the surface C. Collision of molecules D. Repulsive forces
11	Which of the following statemetns about ideal gases is true.	A. they have strong intermolecular forces B. Their particles have significant volume C. Their volume is mainly due to particle size D. They have negligible intermolecular forces
12	What is the role of a pressur ecooker in boiling water.	A. It lowers atmospheric pressure B. It allows water to evaporate quickly

		C. It increases external pressure raising boiling point D. It cools the steam
13	Liquid crystals exhibit properties.	A. Only like solids B. Only like liquid C. Between solids and liquids D. Unliek solids or liquids
14	Liquid are approximately how many times less compressible than gases.	A. 10 B. 105 C. 100 D. 1000
15	Which of the following is essential for hydrogen bond formation.	A. Hydrogen bonded to a metal B. Hydrogen bonded to high electongtive atoms F,O, or N C. Hydogen boned to non pola ratom D. Present of pi bond
16	amorphous solids are typically	A. Rigid and hard B. Found in lump or fine powder form C. Have a shrp melting point D. Always crystalline in structure
17	What is the value of water's enthalpy of vaporization.	A. 21 kJ/mol B. 41 kJ /mol C. 30 kJ /mol D. 60 kJ /mol
18	When does a liqudi start boiling.	A. When temperatur reahes 100 oC B. When its vaporu pressur eequals atmospheric pressure C. When all molecules turn into vapour D. When surface tension becomes zero
19	Why does ice float on water.	A. It has lower density than water B. It is heavier than water C. It forms ionic bonds D. It denser than water
20	The expasnsion of liqudis increae with temperature is	A. What happens to the kinetic energy of liquid molecules when cooled. B. Increases C. Remain unchanged D. Decreases, and the liquid may turn into solid E. Becomes Zero