States and Phases of Matter

Sr	Questions	Answers Choice
1	What else is requrid on the electornegative atom for hydrogen bonding besides bonding to H.	A. Resonance B. Lone pari of electrons C. Doyble bond D. Positive charge
2	According to Avodgaro's Law, volume is directly propertional to.	A. Pressure B. Temperature C. Number of moles D. Density
3	What causes the lower compressibility of liquids compared to gaes.	A. Absence of intermolecular forces B. Larger mollecular size C. Stonger intermolecualr fores and less empty space D. High kinetic energy of molecules
4	The ideal gas equation is given as.	A. PV =nRT B. PV=RT C. PV = nRT/V D. None of these
5	Why is the compression of solids not possible	A. The particles are fixed in place and cannot move closer B. Their particles are widely spaced C. The partiles are charged D. Solids are made of gases
6	Why is water's surface tension so high.	A. Due to its ioic nature B. Because of hydrogen boding pulling surface molecules downward C. Beacue it has low boiling point D. Beacuse water molecules are very large
7	Why can gases be compressed easily.	A. Because they have strong intermolecular forces B. Becasue their particles are lightly packed C. Because most of their volume is empty space D. Beacuse they are denser than solids
8	What happens to the volue of water when it freeze	A. It contracts by 9% B. It remains the same C. It expends by 9% D. It evaporates completely
9	What types of force is hydrogen bonding.	A. Ionic bond B. A special type of dipole dipole force C. Metalic Bond D. London Dispersion force
10	Liquid crystals are used in thermometers becasue.	A. They arehighly reflective B. They glow at high temperatures C. Their color changes with temperature D. They conduct electricity
11	Liquid crystals exhibit properties of	A. Only solid B. Only liquid C. Both solids and liquids D. Gases and solids
12	Whv does water boil at 69 oC on Mount Everest	A. Due to hgih temperature B. Due to low vapour pressure

		C. Due to low external pressure D. Due to presence of snow
13	What happens to the voue of water when it freezes.	A. Because F is less electnegative B. Beacue it is non polar C. Because hydrogen bonding in HF traps the H+ ion D. Beacuse HF is a gas
14	What is the SI unit of viscosity.	A. Pascal B. kg m-1 s-1 C. Joule D. Nm-2
15	Which of the following is essential for hydrogen bond formation.	A. Hydrogen bonded to a metal B. Hydrogen bonded to high electongtive atoms F,O, or N C. Hydogen boned to non pola ratom D. Present of pi bond
16	What is the effect of decresing external pressur eon boiling point of a liquid.	A. It increases B. It remains constant C. It decrease D. It fluctuates
17	What is the value of water's enthalpy of vaporization.	A. 21 kJ/mol B. 41 kJ /mol C. 30 kJ /mol D. 60 kJ /mol
18	The expasnsion of liqudis increae with temperature is	A. What happens to the kinetic energy of liquid molecules when cooled. B. Increases C. Remain unchanged D. Decreases, and the liquid may turn into solid E. Becomes Zero
19	Why can a water molecule formtwo hydrogen bonds on avarage.	A. Beacuse it is polar B. Due to its temperature C. It has 2 hydrogen atm and 2lone pairs D. Becuse it has a covalent bond
20	Surface tension usually with interasig temperature.	A. Increase B. Decrease C. Remain constant D. Doubles