

Chemistry Fsc Part 1 Chapter 5 Online Test

Sr	Questions	Answers Choice
1	How does heat of vaporization affect matter.	Answers Choice A. Cause decrease in kinetic energy B. Increases particle motion and separes tham into gas phase C. Decresease freedom of particles D. Strengthens intermolecular forces
2	Which of the following is a charcteristic of cyrstalline solids.	A. They do not have a definite geometrical shape B. They melt over a wide temerature range C. They have asharp melting point D. Their properties do not depend on the directionof measurement
3	Which of the following compounds exhibites permanent dipole permanent dipole intersctions.	A. HCI B. CI2 C. CH4 D. O2
4	Which of the following is NOT a property of liqud crystals.	A. Fludity B. Rigidity like crystals C. Complete isotropy D. Viscosity
5	According to kinetic molecualr theor liquid molecules are.	A. At rest B. Vibrating in fixed positions C. In constant random motion D. Only moving when heated
6	Why can gases be compressed easily.	A. Because they have strong intermolecular forces B. Becasue their particles are lightly packed C. Because most of their volume is empty space D. Beacuse they are denser than solids
7	What causes the lower compressibility of liquids compared to gaes.	A. Absence of intermolecular forces B. Larger mollecular size C. Stonger intermolecualr fores and less empty space D. High kinetic energy of molecules
8	What cause the polarity in amolecule like CHCl2	A. Equal sharing of electrons B. Non Polar Bond C. Uneven distribution of electrons creating positive and negative ceners D. Absence of hydronge atoms
9	Which of the following is a characteristic property of crystalling solids.	A. they have a range of melting points B. They are isotrople C. they have adefinite and sharp melting point D. They lack a regular arrangeness of particles
10	Why is water's surface tension so high.	A. Due to its ioic nature B. Because of hydrogen boding pulling surface molecules downward C. Because of hydrogen boding pulling surface molecules downward C. Beacue it has low boiling point

		D. Beacuse water molecules are very large
11	What is the SI unit of viscosity.	A. Pascal B. kg m-1 s-1 C. Joule D. Nm-2
12	At equilibrium in a clsoed system, which two processes occur at the same rate.	A. Melting and freezing B. Evaporation and boiling C. Evporatin and condensation D. Sublimation and condensation
13	The expasnsion of liqudis increae with temperature is	A. What happens to the kinetic energy of liquid molecules when cooled. B. Increases C. Remain unchanged D. Decreases, and the liquid may turn into solid E. Becomes Zero
14	Which of the following is a polar moleuce taht shows permanent dipole interactions.	A. CHCI B. CH4 C. CI2 D. H2
15	The expansion of liquids increase with temperatur eis.	A. Equa to that of gases B. Greater than that of gases C. Negligible copared to gasses D. Same as solids
16	Liquid crystals exhibit properties of	A. Only solid B. Only liquid C. Both solids and liquids D. Gases and solids
17	Earthenware pots keep water cool because.	A. They are made of clay B. They reflect sunlight C. They absorb water D. They allow water to waporate through pores
18	Which of the following has greater viscosity than water	A. Glycerine B. Hexane C. Acetone D. Methanol
19	When the vapour pressure of a liquid equals the external pressure, what phenomenon occurs.	A. Sublimation B. Bioling C. Freezing D. condensation
20	What happens to the voue of water when it freezes.	A. Because F is less electnegative B. Beacue it is non polar C. Because hydrogen bonding in HF traps the H+ ion D. Beacuse HF is a gas