Stoichiometry

Sr	Questions	Answers Choice
1	A gas has a density of 1.43 g dm3 at STP. What is the molar mass of the gas?	A. 14.3 g mol-1 B. 32.0 g mol-1 C. 22.4 g mol-1 D. 64.0 g mol-1
2	The largest numebr of molecules are present in	A. 3.6 g of H2O B. 4.6 g of C2H5OH C. 2.8 g of CO D. 5.4 g of N2O5
3	A solution contains 4.0 g of sodium hydroxide in 250 cm3 of solution. What is the molar concentration of this solution	A. 0.10 mol dm-3 B. 0.20 mol dm-3 C. 0.40 mol dm-3 D. 0.80 mol dm-3
4	What is the ratio of volumes of 2 g of 2 to the volum of 16 g CH4 both volume areat STP	A. 1:1 B. 1:8 C. 1:2 D. 2:1
5	Which one of the followng gases will have maximum volume at STP.	A. 88 g of N2O B. 22 g of CO2 C. 28 g of CO D. 28 g of N2
6	What the mass of oxygen obtaingng from 72 g of pure water.	A. 16 g B. 32 g C. 64 g D. 72 g
7	One mole of SO2 conatains.	A. 6.02 x 10 ²³ atoms of oxygen B. 6.022 x 10 ²³ atoms of sulphur C. 4 g atoms of SO2 D. 18.1 X 10²³ moles of SO2
8	Which of the following has maximum mass	A. 2 moles of P B. 5 moles of H2O C. 2 mole of Na2CO3 D. 1 mole of glucose
9	The reatant which consumes earlier and gives lest quantity of product is called.	A. Reactant B. Limiting reacant C. Stiochiometry D. Stoichiometric amount
10	A container holds 0,5 moles of an ideal gas at STP, What is the volume of the gas in dm3	A. 11.2 dm3 B. 22.4 dm3 C. 44.8 dm3 D. 12.2 dm3
11	The volume of occupied by 1.4 g of N1 at	A. 1.12 dm3 B. 2.24 dm3 C. 22.4 dm3 D. 112 dm3
12	Which one of the following Statments is incorrect	A. One mole of nitrogen gas contains Avogadro's numebr of moleclues B. One mole of ozone gas contains avogadr's number of molecules C. One mole of ozone contains avogadro's number of O atom D. One moleof hydrogen gas contains Avogadro's numebrof molecules
		A. 5.3 g of Na2CO 3

13	4.0 g of NaOH (molar mass 40 g mol-1) contains same numebr of sodium ions as are present in	B. 58.5 g of NaCl C. 76 g Na2 SO4 D. 85 g of NaNO3
14	The mass of one mole chlorine gas is.	A. 71 g B. 32 g C. 35.5 g D. 23 g
15	The mass of 0.5 mole of Al is	A. 13.5g B. 12g C. 14 g D. 2.7 g
16	One mole of water contains.	A. 81 g water B. 6.02 x 10²³ atom C. 6.02 x 10²³ ions D. 6.02 x 10²³ molecule E. 6.02 x 10²³ atom
17	1 Mole of OH-1 ion is equal to	A. 18 g B. 17 g C. 16 g D. 10 g
18	The ratio of numerbof olecues of 2 g H2 gas to numebr of molecues of 64 g gaseous oxygen is	A. 1:2 B. 1:1 C. 1:32 D. 1:4
19	Mass of moles of CO2 is.	A. 44 g B. 88 g C. 40 g D. 50 g
20	A sample of nitrogen gas has a mass of 14.0 g,How many nitrogen atoms are present in the this sample	A. 3.01 x 10 ²⁵ atoms B. 6.02 x 10 ²³ atoms C. 12.01 x 10 ²⁵ atoms D. 2.40 x 10 ²⁵ atoms