Stoichiometry

Sr	Questions	Answers Choice
1	A limiting reactatn is the one which	A. Is taken on lesser quantity in gram to compared to other reactionts B. Is taken in lesser quantity in volume as compared to the other reactants C. Gives the maximum amount o fthe produce which is required D. Gives teh minimum amountof the proudcut under considerationn
2	1 Mole of OH-1 ion is equal to	A. 18 g B. 17 g C. 16 g D. 10 g
3	The largest numebr of molecules are present in	A. 3.6 g of H2O B. 4.6 g of C2H5OH C. 2.8 g of CO D. 5.4 g of N2O5
4	Mass of moles of CO2 is.	A. 44 g B. 88 g C. 40 g D. 50 g
5	Which of the following has minimum mass	A. 3 x 10 ²³ atom of C B. 1 mole of S C. 79 gram of Ag D. 2 gram atomsof N
6	A ring has 6 g of diamond in it. What is the number of atoms in it.	A. 3.01 x 10²³ B. 6.02 x 10²³ C. 6.02.01 x 10⁹ D. 12.01 x 10²³ E. 3.01 x 10²³
7	What the mass of oxygen obtaingng from 72 g of pure water.	A. 16 g B. 32 g C. 64 g D. 72 g
8	How many moles of water are proeuced by buringing 4 moles of H2 with excess of oxygen.	A. 1 mole B. 2moles C. 3 moles D. 4 moles
9	The mass of one mole chlorine gas is.	A. 71 g B. 32 g C. 35.5 g D. 23 g
10	What is the ratio of volumes of 2 g of 2 to the volum of 16 g CH4 both volume areat STP	A. 1:1 B. 1:8 C. 1:2 D. 2:1
11	32 g of oxygen gas contains.	A. 6.02 x 10²³ moles B. 6.02 x 10²³ atoms C. 12.4 x 10²³ molecules D. 24.4 x 10²³ atoms
12	The mass of one mole of electron is	A. 1.088 mg B. 0.55 mg C. 0.184 mg D. 1.67 mg

13	Number of molesin 100 g of KClO3	A. 0.76 B. 0.56 C. 0.816 D. 0.014
14	One mole of SO2 conatains.	A. 6.02 x 10 ²³ atoms of oxygen B. 6.022 x 10 ²³ atoms of sulphur C. 4 g atoms of SO2 D. 18.1 X 10²³ moles of SO2
15	The mass of 1.55 x 10 ^{23} atoms of Si	A. 0.5 g B. 8 g C. 0.7 g D. 0.6 g
16	If 16 g of O2 react with excess C2H6,how many gamof CO2 will be formed. 2C2H6+7O2----------4CO2 +5H2O	A. 22 g B. 13 g C. 9 g D. 7g
17	4.0 g of NaOH (molar mass 40 g mol-1) contains same numebr of sodium ions as are present in	A. 5.3 g of Na2CO 3 B. 58.5 g of NaCl C. 76 g Na2 SO4 D. 85 g of NaNO3
18	Which of the following contains same number of particles as present in 12 g of carbon	A. 28 f od ieon (Atomic mass of Fe= 56) B. 48 g of magnesium (Atomic mass of Mg =24) C. 32 g of S Molecules (Atomic mass S= 32) D. 44 g of carbon dioxide (molar mass of CO2= 44)
19	The ratio of numerbof olecues of 2 g H2 gas to numebr of molecues of 64 g gaseous oxygen is	A. 1:2 B. 1:1 C. 1:32 D. 1:4
20	What is the number of moless of oxygen in 11 g of CO2	A. 0.25 B. 0.50 C. 0.75 D. 1.0