

Chemistry Fsc Part 1 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	When water freezes at 0 °C, its density decrease due to.	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
2	The boiling point of glycerin at one atm is.	A. 280 °C B. 290 °C C. 100 °C D. 110 °C
3	In order to mention the boiling point of water at 110 °C, the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure
4	The long chains of amino acids are coiled about one another onto a spiral by	A. Ionic bond B. Van der Waals forces C. Hydrogen bonding D. Overlapping of orbitals
5	Which of the following is pseudo solid	A. CaF ₂ B. Glass C. NaCl D. All
6	Long chains of amino acids are coiled about one another into a spiral by	A. Covalent bond B. Ionic bond C. Hydrogen bond D. Van Der Waals forces
7	The repulsion of electronic clouds of the molecules are responsible for the attractive forces among the molecules. These forces are	A. Dipole-induced dipole forces B. Ion-dipole forces C. Instantaneous dipole-induced dipole forces D. Dipole-dipole forces
8	When water freezes, its volume increase.	A. 12% B. 9% C. 15% D. 18%
9	Diamond is a bad conductor of electricity because	A. It has a tight structure B. It has a high density C. There are no free electrons present in the crystal of diamond to conduct electronics D. None of these
10	The boiling point of water at Murree Hills.	A. 90 °C B. 98 °C C. 100 °C D. 120 °C
11	Diamond is a bad conductor because.	A. It has tight structure. B. It has a high density C. There is no free electron present in the crystal of diamond to conduct electricity D. None of the above
12	The shape of diamond crystal is.	A. cubic B. Hexagonal C. Tetragonal D. Orthorhombic
13	Exceptionally low acidic strength of HF is due to.	A. Strong polar bond between H and F B. Smaller size of fluorine C. Strong hydrogen bonding D. electronegativity of fluorine
14	NaCl is face centered cubic structure. The Na ion at the face of the unit cell is shared by	A. 2-unit cells B. 4-unit cells C. Only one unit cell D. 8-unit cells

15	Boiling point of a liquid is high when	A. There is no hydrogen bonding B. Dipole moment is zero C. Inter molecular forces are weak D. Hydrogen bonding is present
16	Select the correct answer out of the following alternative suggestions London dispersion forces are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium is gaseous state at high temperature. C. Molecules of solid I ₂ D. Molecule of H-Cl gas
17	Which is not use of liquid crystals.	A. Temperature sensor B. Liquid crystal display C. Skin thermography D. Energy supply in electrical devices.
18	Which has greater enthalpy of vaporization	A. F ₂ B. Cl ₂ C. Br ₂ D. I ₂
19	In which system all the three axes are of equal length and all angles are at right angle.	A. Cubic B. Tetragonal C. Orthorhombic D. Hexagonal
20	Diamond is bad conductor because.	A. It has a tight structure B. It has a high density C. It is transparent to light D. There are o free electros present in the crystal of diamond to conduct electricity.
21	In order to mention the B.P of water at 110 °C the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure
22	Which liquid has low boiling point with.	A. Less intermolecular force and higher V.P B. Greater intermolecular forces and low V.P C. Bigger size and greater polarizability D. High hydrogen bonding in it
23	Vapour pressure of a substance does not depend upon.	A. Temperature B. Intermolecular forces C. Surface area D. Physical state of water
24	How many allotropic forms are present in carbon	A. Two B. Three C. Four D. Five
25	In order to mention the B.P. of water at 110°C, the external pressure should be	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 760 torr D. Any value of pressure
26	Allotropy is the property of.	A. Compound B. Element C. Atom D. Mixtuer
27	The number of Na ⁺ ions which surround each Cl ⁻ ion the NaCl crystal lattice is	A. 8 B. 12 C. 6 D. 4
28	Which pair of compound are isomorphic in nature.	A. NaCl and KNO ₃ B. KNO ₃ and MgO C. MgO and NaF D. CaF and CaCO ₃
29	Dipole-dipole forces are present among.	A. Molecules of Iodine B. Atoms of Neon i gaseous state C. Chloroforms' molecules D. CCl ₄ molecules
30	Acetone and chloroform are soluble in each other due to	A. Intermolecualr hydrogen bonding B. Dipole-dipole interaction C. Instantaneous dipoles D. All of the above

31	Which of the given has hydrogen bonding.	A. CH ₄ B. CCl ₄ C. NH ₃ D. NaCl
32	Amorphous solids.	A. Have sharp melting points. B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms.
33	Molecular crystals are generally	A. Hard B. Relatively soft C. Unstable D. do not exist
34	Acetone and chloroform are soluble in each other due to.	A. Intermolecular hydrogen bonding B. Dipole dipole interaction C. Instantaneous dipoles D. All of the above
35	Polymorphic substances have	A. Same physical and chemical properties B. Different physical and chemical properties. C. Same physical but different chemical properties D. Different physical and same chemical properties.
36	The polarizabilities of elements mostly increase down the group due to the reason that	A. The atomic numbers increase B. Number of protons increase C. Number of shells increase along, with increase of shielding effect D. The behavior of the elements remain the same
37	Ionic solid are characterized by.	A. Low melting point B. Good conductivity in solid state C. High vapours pressure D. solubility in polar solvent
38	Isomorphous crystals show	A. Same chemical properties B. Same physical properties C. Same crystalline form D. Same melting point
39	A pressure cooker reduces cooking time because.	A. Heat is uniformly distributed B. Boiling point of water rises C. A large flame is used D. Vapour pressure of liquid reduces
40	The molecules of CO ₂ in dry ice form the.	A. Ionic crystals B. Molecular crystals C. Amorphous D. Covalent crystals
41	NH ₃ shows a maximum boiling point among the hydrides of the group V elements due to.	A. Very small size of nitrogen B. Long pair of electrons present on nitrogen C. Enhanced electronegative character of nitrogen D. Pyramidal structure of NH ₃
42	The distillation of liquid under reduced pressure is called.	A. Destructive distillation B. Vacuum distillation C. Simple distillation D. Fractional distillation
43	Which of the following elements in its crystalline form will have the lowest enthalpy change of vaporization	A. Chlorine B. Argon C. Phosphorus D. Silicon
44	The boiling point of pure water at 1 atm pressure is.	A. 98 ^o C B. 100 ^o C C. 69 ^o C D. 120 ^o C
45	Which substances has diffused melting point.	A. Crystalline solids B. Amorphous solids C. Metallic solids D. Covalent solids
46	NH ₃ shows a maximum boiling point among the hydrides of V-A group elements due to.	A. Very small size of nitrogen B. Lone pair of electron present on nitrogen C. enhanced electronegative D. Pyramidal structure of NH ₃

		character of nitrogen D. Pyramidal structure of NH_3
47	The lightest value of lattice energy is for which one of these ionic compounds.	A. NaI B. NaF C. NaBr D. NaCl
48	London dispersion force are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecule of solid iodine D. Molecules of hydrogen chloride gas
49	Crystal to diamond is.	A. Ionic B. Molecular C. Covalent D. Metallic
50	Ionic solids are characterized by.	A. Low melting points B. High vapour pressures C. Good conductivity in solid state D. Solubility in polar solvents
51	When water freezes at 0°C , its density decreases due to	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
52	Ice float over water because.	A. Its structure is diamond like B. Its density is maximum at 4°C C. It is less dense than water D. It has no regular arrangement of molecules.
53	The nature of copper crystals is	A. Metallic B. Ionic C. Covalent D. Molecular
54	In which , case particles are separated from each other.	A. Fusion B. Condensation C. Neutralizations D. Vaporization
55	London dispersion forces are the only forces present among the	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecules of solid iodine D. Molecules of hydrogen chloride gas
56	In cubic and hexagonal closest packing which layer has different arrangement.	A. First B. Second C. Third D. Fourth
57	On which factor boiling point of a liquid depends.	A. Amount of the liquid B. Shape of the container of the liquid C. Type of burner used for boiling D. External pressure
58	The distillation of a solution under reduced pressure is called	A. Fractional distillation B. Destructive distillation C. Distillation D. Vacuum distillation
59	When liquid water changes to ice its volume expands. The expansion in volume is.	A. 5% B. 9% C. 10% D. 18%
60	Which of the following can form H-bonds	A. NH_3 B. C_2H_6 C. NaCl D. CHCl_3
61	In triclinic unit cell	A. All axial lengths are equal B. All internal lengths and angles are equal C. Both axial lengths and angles are equal D. Both axial lengths and angles are unequal
		A. Very small size of nitrogen B. Lone pair electrons present on Nitrogen

62	NH ₃ shows a maximum boiling point among the hydrides of V-A group elements due to	C. Enhanced electronegative character of Nitrogen D. Pyramidal structure of NH ₃
63	The molecules of CO ₂ in dry ice form the	A. Ionic crystal B. Covalent crystals C. Molecular crystals D. Any type of crystals
64	One of the following liquids has lowest vapour pressure at 32°C. Indicate that liquid	A. Ether B. Chloroform C. Ethanol D. Water
65	Which is pseudo solid	A. CaF ₂ B. Glass C. NaCl D. CaCl ₂
66	Which one of the following inter molecular forces are present in neon gas molecules.	A. Hydrogen bond B. dipole -Dipole attraction C. London dispersion force D. Hydrogen bonding and London dispersion force
67	Conductivity of metal decreases by increasing temperature because.	A. Atoms are converted to ions B. Atoms oscillates and hinder the movement of free electrons. C. Ions are converted into atoms D. Velocity of mobile electrons increases
68	The molecule of CO ₂ in dry ice form the.	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystals
69	The process in which liquid can be made to boil at low temperature is known asdistillation	A. Simple B. Thermal C. Steam D. Vacuum
70	Down the VII -A group, polarizability generally.	A. Increases B. Decreases C. Remain constant D. Negligible
71	The boiling point of water at the top mount Everest is.	A. 59°C B. 69°C C. 83°C D. 75°C
72	Which one of the following substances is not amorphous	A. Polymer B. Rubber C. Glass D. AgNO ₃
73	Existence of an element in more than one crystalline form is known as.	A. Anisotropy B. Allotropy C. Isomorphism D. Unit cell
74	Which pair of molecule have Debye forces in them	A. Ar and Ar B. Argon and water C. Na ⁺ ions and water D. water and water
75	When external pressure is 23.7 torr, boiling point of water is	A. 100 °C B. 200 °C C. 98 °C D. 25 °C