

Chemistry Fsc Part 1 Chapter 4 Online Test

Sr	Questions	Answers Choice
1	In which system all the three axes are of equal length and all angles are at right angle.	A. Cubic B. Tetragonal C. Orthorhombic D. Hexagonal
2	London dispersion force are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecule of solid iodine D. Molecules of hydrogen chloride gas
3	The repulsion of electronic clouds of the molecules are responsible for the attractive forces among the molecules. These forces are	A. Dipole-induced dipole forces B. Ion-dipole forces C. Instantaneous dipole-induced dipole forces D. Dipole-dipole forces
4	How many allotropic forms are present in carbon	A. Two B. Three C. Four D. Five
5	In order to mention the boiling point of water at 110 °C, the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure
6	Acetone and chloroform are soluble in each other due to	A. Intermolecular hydrogen bonding B. Dipole-dipole interaction C. Instantaneous dipoles D. All of the above
7	Acetone and chloroform are soluble in each other due to.	A. Intermolecular hydrogen bonding B. Dipole dipole interaction C. Instantaneous dipoles D. All of the above
8	Which of the following elements in its crystalline form will have the lowest enthalpy change of vaporization	A. Chlorine B. Argon C. Phosphorus D. Silicon
9	When liquid water changes to ice its volume expands. The expansion in volume is.	A. 5% B. 9% C. 10% D. 18%
10	The lightest value of lattice energy is for which one of these ionic compounds.	A. NaI B. NaF C. NaBr D. NaCl
11	Isomorphous crystals show	A. Same chemical properties B. Same physical properties C. Same crystalline form D. Same melting point
12	NH ₃ shows a maximum boiling point among the hydrides of the group V elements due to.	A. Very small size of nitrogen B. Long pair of electrons present on nitrogen C. Enhanced electronegative character of nitrogen D. Pyramidal structure of NH ₃
13	Which has greater enthalpy of vaporization	A. F ₂ B. Cl ₂ C. Br ₂ D. I ₂
14	A pressure cooker reduces cooking time because.	A. Heat is uniformly distributed B. Boiling point of water rises C. A large flame is used D. Vapour pressure of liquid reduces

15	NH ₃ shows a maximum boiling point among the hydrides of V-A group elements due to.	A. Very small size of nitrogen B. Lone pair of electron present on nitrogen C. enhanced electronegative character of nitrogen D. Pyramidal structure of NH ₃
16	Which one of the following substances is not amorphous	A. Polymer B. Rubber C. Glass D. AgNO ₃
17	The molecules of CO ₂ in dry ice form the.	A. Ionic crystals B. Molecular crystals C. Amorphous D. Covalent crystals
18	Amorphous solids.	A. Have sharp melting points. B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possess small regions of orderly arrangement of atoms.
19	Crystal to diamond is.	A. Ionic B. Molecular C. Covalent D. Metallic
20	When water freezes at 0°C, its density decreases due to	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
21	Diamond is a bad conductor because.	A. It has tight structure. B. It has a high density C. There is no free electron present in the crystal of diamond to conduct electricity D. None of the above
22	In cubic and hexagonal closest packing which layer has different arrangement.	A. First B. Second C. Third D. Fourth
23	London dispersion forces are the only forces present among the	A. Molecules of water in liquid state B. Atoms of helium in gaseous state at high temperature C. Molecules of solid iodine D. Molecules of hydrogen chloride gas
24	On which factor boiling point of a liquid depends.	A. Amount of the liquid B. Shape of the container of the liquid C. Type of burner used for boiling D. External pressure
25	Long chains of amino acids are coiled about one another into a spiral by	A. Covalent bond B. Ionic bond C. Hydrogen bond D. Van Der Waal's forces
26	Boiling point of a liquid is high when	A. There is no hydrogen bonding B. Dipole moment is zero C. Inter molecular forces are weak D. Hydrogen bonding is present
27	Which is not use of liquid crystals.	A. Temperature sensor B. Liquid crystal display C. Skin thermography D. Energy supply in electrical devices.
28	Which is pseudo solid	A. CaF ₂ B. Glass C. NaCl D. CaCl ₂
29	NH ₃ shows a maximum boiling point among the hydrides of V-A group elements due to	A. Very small size of nitrogen B. Lone pair electrons present on Nitrogen C. Enhanced electronegative character of Nitrogen D. Pyramidal structure of NH ₃
30	The shape of diamond crystal is.	A. cubic B. Hexagonal C. Tetrahedral

		C. Tetragonal D. Orthorhombic
31	Polymorphic substances have	A. Same physical and chemical properties B. Different physical and chemical properties. C. Same physical but different chemical properties D. Different physical and same chemical properties.
32	When water freezes, its volume increase.	A. 12% B. 9% C. 15% D. 18%
33	The boiling point of water at Murree Hills.	A. 90 ^o C B. 98 ^o C C. 100 ^o C D. 120 ^o C
34	Select the correct answer out of the following alternative suggestions London dispersion forces are the only forces present among the.	A. Molecules of water in liquid state B. Atoms of helium is gaseous state at high temperature. C. Molecules of solid I ₂ D. Molecule of H-Cl gas
35	Which substances has diffused melting point.	A. Crystalline solids B. Amorphous solids C. Metallic solids D. Covalent solids
36	Exceptionally low acidic strength of HF is due to.	A. Strong polar bond between H and F B. Smaller size of fluorine C. Strong hydrogen bonding D. electronegativity of fluorine
37	The boiling point of pure water at 1 atm pressure is.	A. 98 ^o C B. 100 ^o C C. 69 ^o C D. 120 ^o C
38	In triclinic unit cell	A. All axial lengths are equal B. All internals lengths and angles are equal C. Both axial lengths and angles are equal D. Both axial lengths and angles are unequal
39	The boiling point of water at the top mount Everest is.	A. 59 ^o C B. 69 ^o C C. 83 ^o C D. 75 ^o C
40	The nature of copper crystals is	A. Metallic B. Ionic C. Covalent D. Molecular
41	Dipole-dipole forces are present among.	A. Molecules of Iodine B. Atoms of Neon i gaseous state C. Chloroforms' molecules D. CCl ₄ molecules
42	The long chains of amino acids are coiled about one another onto a spiral by	A. Ionic bond B. Van der walls forces C. Hydrogen bonding D. Overlapping of orbitals
43	The boiling point of glycerin at one atm is.	A. 280 ^o C B. 290 ^o C C. 100 ^o C D. 110 ^o C
44	When external pressure is 23.7 torr, boiling point of water is	A. 100 ^o C B. 200 ^o C C. 98 ^o C D. 25 ^o C
45	Which of the following can form H-bonds	A. NH ₃ B. C ₂ H ₆ C. NaCl D. CHCl ₃
46	The molecule of CO ₂ in dry ice form the.	A. Ionic crystals B. Covalent crystals C. Molecular crystals

		D. Any type of crystals
47	Conductivity of metal decreases by increasing temperature because.	A. Atoms are converted to ions B. Atoms oscillates and hinder the movement of free electrons. C. Ions are converted into atoms D. Velocity of mobile electrons increases
48	Diamond is bad conductor because.	A. It has a tight structure B. It has a high density C. It is transparent to light D. There are no free electrons present in the crystal of diamond to conduct electricity.
49	Molecular crystals are generally	A. Hard B. Relatively soft C. Unstable D. do not exist
50	Ice float over water because.	A. Its structure is diamond like B. Its density is maximum at 4 °C C. It is less dense than water D. It has no regular arrangement of molecules.
51	NaCl is face centered cubic structure. The Na ion at the face of the unit cell is shared by	A. 2-unit cells B. 4-unit cells C. Only one unit cell D. 8-unit cells
52	In order to mention the B.P. of water at 110°C, the external pressure should be	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 760 torr D. Any value of pressure
53	The molecules of CO ₂ in dry ice form the	A. Ionic crystal B. Covalent crystals C. Molecular crystals D. Any type of crystals
54	Existence of an element in more than one crystalline form is known as.	A. Anisotropy B. Allotropy C. Isomorphism D. Unit cell
55	The distillation of liquid under reduced pressure is called.	A. Destructive distillation B. Vacuum distillation C. Simple distillation D. Fractional distillation
56	One of the following liquids has lowest vapour pressure at 32°C. Indicate that liquid	A. Ether B. Chloroform C. Ethanol D. Water
57	Which one of the following inter molecular forces are present in neon gas molecules.	A. Hydrogen bond B. dipole -Dipole attraction C. London dispersion force D. Hydrogen bonding and London dispersion force
58	When water freezes at 0 °C, its density decrease due to.	A. Cubic structure of ice B. Empty spaces present in the structure of ice C. Change of bond lengths D. Change of bond angles
59	The polarizabilities of elements mostly increase down the group due to the reason that	A. The atomic numbers increase B. Number of protons increase C. Number of shells increase along with increase of shielding effect D. The behavior of the elements remain the same
60	Which liquid has low boiling point with.	A. Less intermolecular force and higher V.P B. Greater intermolecular forces and low V.P C. Bigger size and greater polarizability D. High hydrogen bonding in it
61	In order to mention the B.P of water at 110 °C the external pressure should be.	A. Between 760 torr and 1200 torr B. Between 200 torr and 760 torr C. 765 torr D. Any value of pressure

62	Diamond is a bad conductor of electricity because	A. It has a tight structure B. It has a high density C. There are no free electrons present in the crystal of diamond to conduct electronics D. None of these
63	The distillation of a solution under reduced pressure is called	A. Fractional distillation B. Destructive distillation C. Distillation D. Vacuum distillation
64	Ionic solid are characterized by.	A. Low melting point B. Good conductivity in solid state C. High vapours pressure D. solubility in polar solvent
65	Which pair of compound are isomorphous in nature.	A. NaCl and KNO ₃ B. KNO ₃ and MgO C. MgO and NaF D. CaF and CaCO ₃
66	Vapour pressure of a substance does not depend upon.	A. Temperature B. Intermolecular forces C. Surface area D. Physical state of water
67	The number of Na ⁺ ions which surround each Cl ⁻ ion in the NaCl crystal lattice is	A. 8 B. 12 C. 6 D. 4
68	Allotropy is the property of.	A. Compound B. Element C. Atom D. Mixture
69	The process in which liquid can be made to boil at low temperature is known asdistillation	A. Simple B. Thermal C. Steam D. Vacuum
70	Which of the following is a pseudo solid	A. CaF ₂ B. Glass C. NaCl D. All
71	Which of the given has hydrogen bonding.	A. CH ₄ B. CCl ₄ C. NH ₃ D. NaCl
72	Ionic solids are characterized by.	A. Low melting points B. High vapour pressures C. Good conductivity in solid state D. Solubility in polar solvents
73	In which case particles are separated from each other.	A. Fusion B. Condensation C. Neutralizations D. Vaporization
74	Down the VII -A group, polarizability generally.	A. Increases B. Decreases C. Remain constant D. Negligible
75	Which pair of molecules have Debye forces in them	A. Ar and Ar B. Argon and water C. Na ⁺ ions and water D. water and water