Atomic Structure

Sr	Questions	Answers Choice
1	What information does the pricnipal quantum numerb (n) gives us about orbitals	A. size B. Shape C. Spin D. Size and shape
2	Subshells in an atom are filled with electrons in an increasing order of their energy values called.	A. aufbau pricniple B. Hund's rule C. Pauli's Exclusion Principle D. None
3	No two electrons in an atom have same all four quantum numbers, this is called.	A. Aufbau Principle B. Hund's rule C. Pauli's exclusing principle D. None
4	The value for (n) Principal quantum numebr can not be.	A. 0 B. 1 C. 2 D. 3
5	If we remove electrons from an atom until the nucleus is left, then it is called.	A. Ionization energy B. Ionization potential C. Successive ionization energy D. All
6	How many d-orbitals are showing their lobes along axis.	A. 1 B. 3 C. 2 D. All
7	An electron having the set of values n = 4, 1 = 0 ,m = 0 and s =+1/2 lies in	A. 2 s B. 4 s C. 3s D. 5s
8	An oribtal cna accommodate how many electrons.	A. 1 B. 2 C. 3 D. 4
9	Atom having same atomic numebr but different mass numebr are called.	A. Isotopes B. Isotones C. Isobars D. Isoelecrons
10	Quantum numebr values for 2 p orbitals are.	A. n = 2, , l = 1 B. n = 1, 1= 2 C. n = 1, 1= 0 D. n= 2,1 = 0
11	The value of 'm" cannot be	A. Fractional B. Zero C. Positive D. Negative
12	How many d orbitals are there in a givne energy level.	A. 1 B. 5 C. 3 D. 7
13	1st ionization energy is.	A. Exothermic B. Endothermic C. Both D. Depend upon atom
14	The shape of d-orbitals are like	A. Complicated B. Polar C. Spherical D. Cloverleaf
		A. 18

15	Number of neutrons present in Cl are	B. 17 C. 19 D. 25
16	Two electrons in the same orbital have.	A. Same Spin B. Differece spin C. May same or different D. Parallel spin
17	In which subshell electrons are filled first according to n+1 rule.	A. 2p B. 3p C. 4p D. 4s
18	Formula for calculating the value of (Azimuthal quantum number)	A. n+1 B. n-1 C. n+2 D. 2n2
19	When 3d subshell is completely filed, the next entering electron goes into.	A. 4f B. 4 p C. 4s D. 4 d
20	d-block elements are also called.	A. Inner transition B. Outer transition C. Typical transition D. None