Atomic Structure

Sr	Questions	Answers Choice
1	The shape of d-orbitals are like	A. Complicated B. Polar C. Spherical D. Cloverleaf
2	Formula for calculating the value of (Azimuthal quantum number)	A. n+1 B. n-1 C. n+2 D. 2n2
3	s-subshell has.....orbital	A. 1 B. 2 C. 3 D. 4
4	Two electrons in the same orbital have.	A. Same Spin B. Differece spin C. May same or different D. Parallel spin
5	Atom having same atomic numebr but different mass numebr are called.	A. Isotopes B. Isotones C. Isobars D. Isoelecrons
6	How many electron presetn in Cl are.	A. 17 B. 18 C. 19 D. 35
7	If we remove electrons from an atom until the nucleus is left, then it is called.	A. Ionization energy B. Ionization potential C. Successive ionization energy D. All
8	Number of neutrons present in Cl are	A. 18 B. 17 C. 19 D. 25
9	An oribtal cna accommodate how many electrons.	A. 1 B. 2 C. 3 D. 4
10	The space orientation for f-subshell is.	A. 3 B. 7 C. 1 D. 5
11	In which subshell electrons are filled first according to n+1 rule.	A. 2p B. 3p C. 4p D. 4s
12	Which of the following atoms show anomalous configuration.	A. Cu B. CO C. Fe D. Ni
13	Orbital having same energy are called.	A. Hybrid orbials B. Valance orbitals C. Degenerate orbitals D. d-orbital
14	Atoms having same atomic mass but different atomic numebr are called.	A. Isobars B. Isotopes C. Isotones D. Isoelectrons
15	In n = 3, then elecron sin this shell are.	A. 2 B. 8 C. 18

		D. 32
16	No two electrons in an atom have same all four quantum numbers, this is called.	A. Aufbau Principle B. Hund's rule C. Pauli's exclusing principle D. None
17	1st ionization energy is.	A. Exothermic B. Endothermic C. Both D. Depend upon atom
18	Electronic cloud around teh nucleus is uniformly distributed in all directions in.....subshell	A. f B. s C. d D. p
19	An electron having the set of values n = 4, 1 = 0 ,m = 0 and s =+1/2 lies in	A. 2 s B. 4 s C. 3s D. 5s
20	How many d orbitals are there in a givne energy level.	A. 1 B. 5 C. 3 D. 7