

Lab Safety and Practical Skills

Sr	Questions	Answers Choice
1	What is the the typical volume of solution a pipette delivers in titration	A. 25 ml B. 10 ml C. 50 ml D. 100 ml
2	Which of the followign pairs would require methly orinage as a a indicator	A. Strong acid and strong base B. Weak acid and sstrong base C. Strong acid and weak base D. Weak acia and weak base
3	Which cation gives a white gelatinous precipitate upon the addition of aueoeous ammonia.	A. Cr ³⁺ B. Cr ²⁺ C. Zn ²⁺ D. Al ³⁺
4	Which apparatus is used ot deliver a variiable volume of soluton in titration .	A. Pipette B. Burette C. Beaker D. Measuring cylinder
5	Before starting a titration , the burette should be	A. Headed B. Drived completely C. Washed with distilled water only D. Rinsed with the solutin it will contain
6	Which of the followign is a primary standard solution.	A. NaOH B. Na ₂ CO ₃ C. KOH D. HCl
7	While takig able a reading with a burette, why it is always advisibel to read the lower meniscus for the colorless liquids and the upper meniscus for the coloured liquids.	A. Because it is more convenient B. Becuse colorless liquids have more surface tension than colored liquids C. Because of the parallax effect D. Beacuse lower meniscus does not exist for colored liquids
8	A white precipitate that dissolves in excess ammonia is likely to be	A. Zinc hydroxide B. Calcium carbonat C. Lead iodide D. Barium sulphate
9	The end point of a titration is indicated by	A. Formation of a gas B. Change in temperatutere C. Colour change of the indicator D. Change in pressure
10	The brown ring tests a confirmatory test for which acid radical	A. Nitrate B. Chloride C. Sulfate D. Carbonate
11	On dry heating test for salt analyshat turens limewater milky suggests the presence of.s, the evolution of colorieess, odourieess gas ta	A. Carbonate ion B. Chloridei on C. Sulfate ion D. Nitrate ion
12	Why is it important to usea pippte in titration	A. To hold the acid B. To measure a fixed volume of the solution accurately C. To detect the endpoint D. To mix the acid and alkali
13	Which cation forms a white precipitate with NaOH that dissolves in excess NaOH	A. Mg ²⁺ B. Zn ²⁺ C. Ca ²⁺ D. Ba ²⁺
14	Which of the followig is not a component of titration	A. Pipette B. Burette C. Test Tube D. Conicla flask

15	addition of NH_4OH to an aqueous solution of cation gives a green precipitate which turns brown upon standing. Which basic radical is indicated.	<p>A. Cr^{3+} B. Cu^{2+} C. Fe^{2+} D. Fe^{3+}</p>
16	Why phenolphthalein indicator is more appropriate to use during the titrations which involve a strong acid and a strong base.	<p>A. Because it is itself weakly acidic B. Because the pH at the equivalence point as well as the pH over where the colour of phenolphthalein changes match each other C. Because the solution at the end of titration is acidic D. Because the solution at the end of titration is basic</p>
17	Which is the proper way to smell a chemical	<p>A. Inhale deeply B. Waft the vapour toward your nose C. Sniff directly D. Open the bottle under our nose</p>
18	Lab coats are important because they.	<p>A. Are fashionable B. Prevent contamination and protect clothing C. Make you feel like a scientist D. Keep you warm</p>
19	The chromyl chloride test is a specific confirmatory test for	<p>A. Chloride ions B. Bromide ion C. Iodide ions D. sulfate ions</p>
20	Eating and drinking in the lab is	<p>A. Allowed during breaks B. Allowed if you're careful C. Strictly prohibited D. Allowed in the chemistry lab but not physics</p>