Halogens

Sr	Questions	Answers Choice
1	Which force governs the volatility trend in halogens.	A. Dipole dipole B. Hydrogen bonding C. Instantaneous dipole induced dipole D. Ionic bonding
2	What ause diffrent colors in halogens.	A. Diffrente atomic radii B. Presence of d-orbitals C. Electron transititons D. Ionization energy
3	The weakest rdducing halide ion is	A. F- B. I- C. Cl- D. Br-
4	What trend in volatility is observed in halognes from chlorine to iodine.	A. Decreases B. Increases C. Fluctuates D. Remains constatant
5	The volatility of the halogens generallly.....as you move down the group .	A. Increawe B. Decreases C. Remains teh same D. Fluctuates unpredicatbly
6	The weakest halogen -alogen bond is in	A. F2 B. Br2 C. Cl2 D. I2
7	What is the bleacig action of chlorine due to.	A. O2 B. Cl2 C. CO2 D. HCl
8	Hydrogen halides in water form	A. Salts B. Hydrohalic acids C. Bases D. Amides
9	Why is F-F bond weaker than Cl - Cl	A. Large atomci radius of F B. Stronger id-id forces in Cl C. Lone pair repulsion in F D. Lower electronegativity of Cl
10	Down the group, the oxidizing power of halogens.	A. Increases B. Remain Constant C. Decreases D. Show no trend
11	Reaction of NaCl and conc. H2SO4	A. Yellow ppt B. Steamy white HCl Fumes C. Violet vapor D. Br2 Fumes
12	The term halogen means.	A. Salt dostryoyer B. Salt producer C. Water purifier D. Acid former
13	the dercreasing volatility doen the group is due to.	A. Decreasing molecular mass B. Increasing London Forces C. Decreasing atomic radius D. Increasing electronegativity
14	Which statemetn about the reaction between halogesn and hydrogen is correct.	A. Iodine reacts most vigorously with hydrogen B. Chlorine and hydrogen explode in darkness C. Fluorine comines explosively with hydrogen even n cod and dark conditions D. Bromine and hydroen do not

		react at all
15	Why are halogens highly reactive.	A. They have full outer shells B. They gain one electron easily C. They lose electrons easily D. They have low electronegativity
16	Why is fluorine the ost reactive halogen.	A. Bond length int he halogen olecule B. Bond strength int he halogen molecule C. Electronegativity of the halogen D. Number of electrons in the halogen molecule.
17	Which halogen exists as a volatile liquid at room temperature.	A. Iodine B. Bromine C. Chlorine D. Fluorine
18	Oxidizing power of halogesn decreases due to.	A. Increasing atomix size B. Increasing bond strength C. Increasing ionization energy D. Increasing hydration energy
19	Chlorofluorocarbon (CFCs) contribute to	A. Ozone formation B. Air purification C. ozone depletion D. Global warming directly
20	Which one of the following halogen molecules has strongest oxideizing power.	A. Br2 B. F2 C. Cl2 D. I2