Chemistry Fsc Part 1 Chapter 11 Online Test

Sr	Questions	Answers Choice
1	Dilatometer method is useful for the reactions that involve.	A. Ionic species B. Where reactant absorb U.V. visible or infrared radiations C. Small volume changes in solutions D. Change in refractive indices
2	Chemical reactivity of different substance is controlled by	A. Atomic number B. Electronic arrangement C. Mass number D. Number of isotope of reactant elements
3	When a chemical reaction is completed.	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
4	If reactants have very low activation energy it means that reaction is.	A. Slow B. Fast C. Endothermic D. Exothermic
5	The mathematical relation between the rate of reaction and the concentrations of the reactants is known as the	A. Rate equation B. Rate law C. Arrhenius equation D. Both a and b
6	All radio active disintegration nuclear reaction are of.	A. First order B. Zero order C. 2nd order D. Third order
7	a zero order reaction is one is which	A. Reactants do not react B. One reactant is in large excess C. Concentration of reactant do not change with passage of time D. Rate is of affected by changing concentration of reactants
8	A substance which itself is not a catalyst but increases the activity of a catalyst is called.	A. Promoter B. Poisoner C. Inhibitor D. Enzyme
9	The quantitative relationship between rate and concentration is given by.	A. Law of mass action B. Rate law C. Both of these D. Le Chatelier's principle
10	Which properties of liquid is measured by polarimeter	A. Conductance B. Refractive index C. Optical activity D. Change in volume
11	The power of which the concentration of a substance appears in the rate expression is known as	A. order of reaction with respect to that substance B. Rate of reaction C. Order of reaction D. Molecularity of reaction
12	When a reaction proceeds in a sequence of steps, the overall rate is determined by	A. Fastest step B. Slowest step C. Order of different steps D. Molecularity of all steps
13	In zero order reaction, the rate is independent of	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these
14	The rate of reaction determined at any given time is called.	A. Average rate B. Instantaneous rate C. Spontaneous rate

		D. Over all rate
15	The unit of the rate constant is same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
16	The rate of reaction	A. Increases B. Decreases C. Remains the same D. May decrease of increase
17	Question Image	A. 1 B. 2 C. 3 D. None of these
18	Which statement is incorrect about order of reaction	A. It cannot be determined experimentally B. It is determined experimentally C. Sum of exponents in rate equation D. It can have fraction value
19	The rate of reaction	A. Increases as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reaction proceeds D. May decrease or increase as the reaction proceeds
20	When a reaction occurs in many steps than the slowest step is.	A. Mechanism step B. Rate determining step C. enthalpy determining step D. None of the above
21	The rate of reaction	A. Increase as the reaction proceeds B. Decreases as the reaction proceeds C. Remains the same as the reactions proceeds D. May decrease or increase as the reaction proceeds
22	Rate of a chemical reaction generally increase rapidly even for small increase in temperature because of rapid increase in the	A. Collisions frequency B. Activation energy C. Average KE of molecules D. Fraction of molecules with energies more than activation energy
23	The unit of the rate constant is the same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Zero order reaction D. Third order reaction
24	Which technique is used ot determine the absorption of radiations.	A. Spectrometry B. dilatometer method C. Refractometric method D. Optical rotation method
25	The unit of rate constant depends upon	A. Number of reactants B. Concentration terms C. Order of reaction D. Molecularity of reaction
26	The unit of rate constant is same as that of rate of reaction in	A. first order reaction B. Second order reaction C. Third order reaction D. Zero order reaction
27	Half -Life period of a first order reacting is independent of.	A. Initial concentration of the compound B. Conditions of temperature C. Presence of catalyst D. All the above
28	After 3 half life of a chemical reaction, amount of reactant un reactive will be.	A. 50% B. 25% C. 12.5% D. 6.25%
29	Which statement is incorrect about activated complex	A. Short lived B. Maximum energy C. Unstable combination of atoms D. Less energy than Ea
30	The unit of rate constant is the same as that of the rate of reaction is.	A. First order reaction B. Second order reaction C. Zero order reaction

31	In zero order reaction the rate is independent of.	A. Temperature of reaction B. Concentration of reactants C. Concentration of products D. None of these
32	With increases of 10 oC temperature the rate of reaction doubles. This increase in rate of reactions is due to.	A. Decrease in activation energy of reaction B. Decrease in the number of collisions between reactant molecules C. Increase in activation energy of reactants D. Increase in number of effective collisions
33	When a chemical reaction is completed the	A. Instantaneous rate > average rate B. Instantaneous rate = average rate C. Instantaneous rate is zero D. Both average and instantaneous rates become zero
34	In zero order reaction, the rate is independent of.	A. Temperature of reaction B. Concentration of reactants C. concentration of products D. None of these
35	With increase in 10°C temperature, the rate of reaction double. This increase in rate of reaction is due to	A. Decrease in activation energy of reaction B. Decrease in the number of collisions between reactant molecules C. Increases in activation energy of reactants D. Increase in number of effect collisions
36	Velocity constant is the rate of reaction when the concentrations of reactants are	A. Zero B. Unity C. Two D. Three
37	A substance which rerareds the rate of a reaction is called	A. Inhibitor B. Activator C. Auto-catalyst D. None of these
38	The influence of temperature on reaction rate is predicated by	A. Free energy change of reaction B. Arrhenius equation C. Van der weal's equation D. Kinetic equation
39	Photosynthesis a photochemical reaction has order of reaction	A. 0 B. 1 C. 2 D. Fractional order
40	Glucose can be converted into ethanol by an enzyme.	A. Lipase B. Zymase C. Sucrose D. Urease
41	Which one of the following statements is incorrect.	A. Enzymes are protein in nature B. Enzymes are catalyst C. Enzymes can catalyze any reaction D. Urease is an enzyme
42	If the energy of the activated complex lies close to energy of reactants, it means that reaction is	A. Slow B. Fast C. Exothermic D. Endothermic
43	Half -Life for a given reaction is doubled if initial concentration is doubled. The order of reaction is.	A. 0 B. 1 C. 2 D. 3
44	The substance which decrease the activity of a catalyst is called.	A. Promoter B. Activator C. Inhibitor D. Positive catalyst
45	The true representation for the units of rate constant K for the first order reaction	A. sec^{-1} B. $mole\ dm^{-3}s$ C. $mole\ dm^{-3}s^{-1}$ D. $mole^{-1}dm^{+3}s^{-}$

46	A type of meals are usually used as catalyst.	A. Coinage metal B. Alkali metals C. Transition metals D. alkaline earth metals
47	the rate of reaction when concentration of reactants are taken unity is called.	A. Average rate B. Instantaneous rate C. Specific rate D. Rate equation
48	Half life period for 235 U is 92	A. 710 million years B. 810 milion years C. 720 million years D. 820 million years
49	After 3 half lives of a chemical reaction, the % fraction of the amount left is	A. 6.25 B. 75 C. 12.5 D. 50
50	Unit of rate constant is the same as that of the rate of reaction in	A. Zero order reaction B. 1st order reaction C. 2nd order reaction D. 3rd order reaction
51	The reaction that involves gases, its rate does not depend upon	A. Catalyst B. Temperature C. Moles dm^{-3} D. Partial pressure