Chemistry Fsc Part 1 Chapter 10 Online Test

Sr	Questions	Answers Choice
1	In given equation underlined element is. P+HNO3---------- H2PO4+NO+H2O	A. Oxidized B. Reduced C. Neither oxidized nor reduced D. Botha a and b
2	The cathodic reaction in the electrolysis of dill. $H_2SO_4$with Pt electrodes is	A. Reduction B. Oxidation C. Both oxidation and reduction D. Neither oxidation nor reduction
3	In electrolysis of aqueous NaCl, Cl-ions are.	A. Oxidized at anode B. Oxidized at cathode C. Reduced at cathode D. Neither oxidized nor reduced
4	Cell potential depends upon	A. Temperature B. Concentration of ions C. Nature of electrolyte D. All of above
5	During the electrolysis of molten NaCl, the ion which is reduce is	
6	The electrode reaction of a voltaic cell can be reversed when	A. Concentrations of solutions are changed B. Temperature is increased C. Electrodes are interchanged D. Electric circuit is employed to supply the source of electricity
7	In which of the following changes, nitrogen in reduced.	A. NH3 to NO B. NH3 to NO3 C. N2 to NH3 D. N-3 to N2
8	The cathodic reaction in the electrolysis of dil H2SO4 with Pt electrodes is.	A. Reduction B. Oxidation C. Both oxidation or reduction D. Neither oxidation nor raduction
9	Which has greater reduction potential	A. Na B. H2 C. Zn D. F2
10	The difference of potential of two electrodes when concentration of solution is 1 M each at 25 ^{O}C and 1 atmosphere is called.	A. Electrode potential B. Standard cell potential C. Cell reaction D. Cell voltage
11	When a non-spontaneous redox reaction is carried out by using the electrical current, then the process is called	A. Decomposition of the substances B. Cracking C. Hydrolysis D. Electroiysis
12	In H2SO4the oxidation number of 'S' is	A. +2 B. +6 C. +8 D. +4
13	Oxidation number of Cr in a C_2CrO_4 is	A. +2 B. +4 C. +6 D. +8
14	Standard hydrogen electrode has an arbitrarily fixed potential	A. 0.00 volts B. 1.00 volt C. 0.10 volt D. None of these
15	The cell in which a non spontaneous redox reaction takes place as a result of electricity is known as.	A. Voltaic cell B. Denial cell C. dry Cell D. Electrolytic cell

16	Fuel cells are the means by which chemical energy may be converted into	A. Heat energy B. Magnetic energy C. Sound energy D. Electric energy
17	The reduction potential of Zn is.	A. +0.76 V B. -0.34 B C. +0.34 V D. -0.76 V
18	Electrolysis is used for	A. Electroplating B. Refining of copper C. Manufacture of caustic soda D. All of the above
19	Stronger the oxidizing agent greater is the	A. Oxidation potential B. Reduction potential C. Redox potential D. E.M.F of cell
20	In Daniel cell, if salt bridge is removed between the two half cells, the voltage.	A. Drops to zero B. Does not changes C. Increases gradually D. Increases rapidly
21	In the reaction 2Fe+3Cl2 ---------- FeCl2	A. Fe is reduced B. Fe is oxidized C. Cl2 is oxidized D. None of these happens
22	A cell in which electric current is produced as a result of spontaneous redox reaction is called.	A. Electrolytic cell B. Galvanic cell C. Half cell reaction D. Down's cell
23	Oxidation number of phosphorus in the compound is.	A. +3 B. +4 C. +5 D. +6
24	In lead accumulator cathode is made up of.	A. Pb B. Pb coated with PbO2 C. PbSO4 D. Mixture of Pb and PbO2
25	If the salt bridge is not used between two half cells, then the voltage.	A. Decrease rapidly B. Decrease slowly C. Drops to zero D. Does not change
26	Electrochemical series is the arrangement of the electrodes in	A. Increasing order of reduction potentials B. Decreasing order of reduction potentials C. Increasing order of oxidation reduction potential D. There is no fixed arrangement
27	The oxidation number of C in C12H22O11 is	A. Zero B. - 6 C. + 6 D. 12
28	The cell in which a non spontaneous redox reaction takes place as a result of electricity is known as.	A. Voltaic cell B. Denial cell C. dry Cell D. Electrolytic cell
29	Alkali and alkaline earth metal are usually obtained by	A. Decomposition of their carbonates B. By heating their hydroxide C. electrolysis of molten metal oxides D. Electrolysis of molten metal halides
30	In NICAD dry cell, the cathode and anode is made up of.	A. Ca and Ag B. Ni and CdO2 C. NiO2 and Cd D. Ag and Ag2O
31	Stronger the oxidizing agent, greater is the	A. oxidation potential B. Reduction potential C. Redox potential D. E.M.F of cell
32	Electrode of the lead storage battery are immersed in dilute H2SO4 which has strength by mass	A. 100% B. 98% C. 30% D. 10%

33	Which of the following statements is correct about galvanic cell	A. Anode is negative charged B. Reduction occurs at anode C. Cathode is positively charged D. Reduction occurs at cathode
34	In a electrolytic cell the electrons flow from	A. Cathode to anode B. Anode to cathode C. From cathode to anode or opposite, depending upon the nature of electrolyte D. All of the above
35	According to classical concept, oxidation involves	A. Addition of oxygen B. Addition of electron C. Removal of hydrogen D. All are correct
36	The oxidation of O -atom in OF3 is.	A. -2 B. +2 C. -1 D. +1
37	Electrolysis is a process in which a chemical reaction takes place at the expense of	A. Chemical energy B. Electrical energy C. Heat energy D. None of these
38	The voltage Nickel Cadmium cell is	A. 1 V B. 1.2 V C. 1.4 V D. 1.6 V
39	If strip of Cu metal is placed in the solution of FeSO4	A. Cu will be precipitated out B. Fe is precipitated out C. Cu and Fe both dissolves D. No reaction takes place
40	The best reducing agent is	A. F⁻¹ B. Cl⁻¹ C. Br⁻¹ D. I⁻¹
41	Oxidation number of carbon in $NaHCO_3$	A. +4 B. -6 C. +6 D. +2
42	If a strips of Cu metal is placed in a solution of FeSO4	A. Cu will be precipitated down B. Fe is precipitated out C. Cu and Fe both dissolve D. No reaction takes palce
43	That cell in which electrical energy is converted into chemical energy is called	A. Galvanic cell B. Electrolytic cell C. Fuel cell D. Daniel cell
44	Fuel cells are mostly used in space air crafts as the source of.	A. Power only B. Drinking water C. Drinking water and power D. Fuel and drinking water
45	What is oxidation state of chlorine in Ca(ClO3)2	A. +1 B. +3 C. +5 D. +7
46	The over all positive value for the reaction potential predicts that process is energetically.	A. Not feasible B. Feasible C. Impossible D. No indication
47	In silver oxide battery, anode is made of.	A. Zinc B. Copper C. Lead D. Graphite
48	In which compound oxidation state of chlorine is +5	A. NaCl B. HOCl C. NaClO3 D. NaClO2
49	In which compound the oxidation number of Mw is +6	A. KMnO4 B. K2MnO4 C. MnO2 D. MnO
50	A single lead cell provides volts	A. 2 B. 4

50	A single lead cell provides volts	C. 6 D. 8
51	If the salt bridge is not used between two half cells, then the voltage	A. Decreases rapidly B. Decreases slowly C. Does not change D. Drops to zero
52	Which is not chargeable cell	A. Lead accumulator B. NiCAD cell C. Fuel cell D. Alkaline battery
53	In silver oxide battery, the cathode is mad up of.	A. AgO B. Ag2O C. Ag2O3 D. Ag
54	Which is not use of electrochemical series.	A. Feasibility of reaction B. Measurement of EMF of cell C. Comparison of reactivity with water or acids D. Determination of atomic and ionic radii
55	What is the oxidation state of sulphur in $SO3^{2-}$	A. -4 B. -2 C. +2 D. +4
56	If a strip of Cu metal is placed in a solution of $FeSO_4$	A. Cu will be precipitated out B. Fe is precipitated out C. Cu and Fe both dissolve D. No reaction takes place
57	Which statements not correct about Galvanic cell.	A. Anode in negatively charge B. Reduction occur at anode C. Cathode is positively charged D. Reduction occur at cathode
58	If salt bridge is not used between two half cells, than the voltage.	A. Decreases rapidly B. Decreases slowly C. Does not change D. Drops of zero
59	Question Image	A. Fe is reduced B. Fe is oxidized C. Cl₂is oxidized D. None of these
60	When an atom reacts chemically and loses one or more electrons it is.	A. Decomposed B. Reduced C. Oxidized D. Catalyzed
61	When aqueous NaCl is electrolyzed, which of the following ions gas discharged at anode.	A. Cl- B. OH- C. Na+ D. H+
62	Electromotive force of the cell is the	A. Difference of two electrode potentials B. May be sum or the difference of two electrode potentials C. Sum of two electrode potential D. Depends upon the nature of the cell
63	The gain of electron is known as.	A. Oxidation B. Reduction C. Dehydration D. Dehydrogenation
64	The cathodic reaction in the electrolysis of dil H2SO4, with pt electrode sis.	A. Reduction B. Oxidation C. Both oxidation and reduction D. Neither oxidation nor reduction