Chemistry Fsc Part 1 Chapter 10 Online Test

Sr	Questions	Answers Choice
1	The difference of potential of two electrodes when concentration of solution is 1 M each at 25 ^{O}C and 1 atmosphere is called.	A. Electrode potential B. Standard cell potential C. Cell reaction D. Cell voltage
2	The over all positive value for the reaction potential predicts that process is energetically.	A. Not feasible B. Feasible C. Impossible D. No indication
3	When aqueous NaCl is electrolyzed, which of the following ions gas discharged at anode.	A. Cl- B. OH- C. Na+ D. H+
4	If a strips of Cu metal is placed in a solution of FeSO4	A. Cu will be precipitated down B. Fe is precipitated out C. Cu and Fe both dissolve D. No reaction takes palce
5	In given equation underlined element is. P+HNO3---------- H2PO4+NO+H2O	A. Oxidized B. Reduced C. Neither oxidized nor reduced D. Botha a and b
6	The oxidation number of C in C12H22O11 is	A. Zero B. - 6 C. + 6 D. 12
7	In the reaction 2Fe+3Cl2 ---------- FeCl2	A. Fe is reduced B. Fe is oxidized C. Cl2 is oxidized D. None of these happens
8	Fuel cells are the means by which chemical energy may be converted into	A. Heat energy B. Magnetic energy C. Sound energy D. Electric energy
9	The cathodic reaction in the electrolysis of dill. $H_2SO_4$with Pt electrodes is	A. Reduction B. Oxidation C. Both oxidation and reduction D. Neither oxidation nor reduction
10	Question Image	A. Fe is reduced B. Fe is oxidized C. Cl₂is oxidized D. None of these
11	If a strip of Cu metal is placed in a solution of $FeSO_4$	A. Cu will be precipitated out B. Fe is precipitated out C. Cu and Fe both dissolve D. No reaction takes place
12	If the salt bridge is not used between two half cells, then the voltage.	A. Decrease rapidly B. Decrease slowly C. Drops to zero D. Does not change
13	In lead accumulator cathode is made up of.	A. Pb B. Pb coated with PbO2 C. PbSO4 D. Mixture of Pb and PbO2
14	The voltage Nickel Cadmium cell is	A. 1 V B. 1.2 V C. 1.4 V D. 1.6 V
15	Oxidation number of phosphorus in the compound is.	A. +3 B. +4 C. +5 D. +6

16	Which of the following statements is correct about galvanic cell	A. Anode is negative charged B. Reduction occurs at anode C. Cathode is positively charged D. Reduction occurs at cathode
17	In a electrolytic cell the electrons flow from	A. Cathode to anode B. Anode to cathode C. From cathode to anode or opposite, depending upon the nature of electrolyte D. All of the above
18	The cell in which a non spontaneous redox reaction takes place as a result of electricity is known as.	A. Voltaic cell B. Denial cell C. dry Cell D. Electrolytic cell
19	If strip of Cu metal is placed in the solution of FeSO4	A. Cu will be precipitated out B. Fe is precipitated out C. Cu and Fe both dissolves D. No reaction takes place
20	What is the oxidation state of sulphur in $SO3^{2-}$	A. -4 B. -2 C. +2 D. +4
21	The best reducing agent is	A. F⁻¹ B. Cl⁻¹ C. Br⁻¹ D. I⁻¹
22	When a non-spontaneous redox reaction is carried out by using the electrical current, then the process is called	A. Decomposition of the substances B. Cracking C. Hydrolysis D. Electroiysis
23	Cell potential depends upon	A. Temperature B. Concentration of ions C. Nature of electrolyte D. All of above
24	Standard hydrogen electrode has an arbitrarily fixed potential	A. 0.00 volts B. 1.00 volt C. 0.10 volt D. None of these
25	In Daniel cell, if salt bridge is removed between the two half cells, the voltage.	A. Drops to zero B. Does not changes C. Increases gradually D. Increases rapidly
26	The cathodic reaction in the electrolysis of dil H2SO4, with pt electrode sis.	A. Reduction B. Oxidation C. Both oxidation and reduction D. Neither oxidation nor reduction
27	A single lead cell provides volts	A. 2 B. 4 C. 6 D. 8
28	The cathodic reaction in the electrolysis of dil H2SO4 with Pt electrodes is.	A. Reduction B. Oxidation C. Both oxidation or reduction D. Neither oxidation nor raduction
29	In NICAD dry cell, the cathode and anode is made up of.	A. Ca and Ag B. Ni and CdO2 C. NiO2 and Cd D. Ag and Ag2O
30	Electrolysis is used for	A. Electroplating B. Refining of copper C. Manufacture of caustic soda D. All of the above
31	In silver oxide battery, the cathode is mad up of.	A. AgO B. Ag2O C. Ag2O3 D. Ag
32	Stronger the oxidizing agent, greater is the	A. oxidation potential B. Reduction potential C. Redox potential D. E.M.F of cell
		A. KMnO4

33	In which compound the oxidation number of Mw is +6	A. KMnO4 B. K2MnO4 C. MnO2 D. MnO
34	The electrode reaction of a voltaic cell can be reversed when	A. Concentrations of solutions are changed B. Temperature is increased C. Electrodes are interchanged D. Electric circuit is employed to supply the source of electricity
35	In H2SO4the oxidation number of 'S' is	A. +2 B. +6 C. +8 D. +4
36	Electrode of the lead storage battery are immersed in dilute H2SO4 which has strength by mass	A. 100% B. 98% C. 30% D. 10%
37	A cell in which electric current is produced as a result of spontaneous redox reaction is called.	A. Electrolytic cell B. Galvanic cell C. Half cell reaction D. Down's cell
38	Which statements not correct about Galvanic cell.	A. Anode in negatively charge B. Reduction occur at anode C. Cathode is positively charged D. Reduction occur at cathode
39	If salt bridge is not used between two half cells, than the voltage.	A. Decreases rapidly B. Decreases slowly C. Does not change D. Drops of zero
40	Electrolysis is a process in which a chemical reaction takes place at the expense of	A. Chemical energy B. Electrical energy C. Heat energy D. None of these
41	The reduction potential of Zn is.	A. +0.76 V B. -0.34 B C. +0.34 V D. -0.76 V
42	If the salt bridge is not used between two half cells, then the voltage	A. Decreases rapidly B. Decreases slowly C. Does not change D. Drops to zero
43	In silver oxide battery, anode is made of.	A. Zinc B. Copper C. Lead D. Graphite
44	Which is not chargeable cell	A. Lead accumulator B. NiCAD cell C. Fuel cell D. Alkaline battery
45	Fuel cells are mostly used in space air crafts as the source of.	A. Power only B. Drinking water C. Drinking water and power D. Fuel and drinking water
46	The gain of electron is known as.	A. Oxidation B. Reduction C. Dehydration D. Dehydrogenation
47	In which of the following changes, nitrogen in reduced.	A. NH3 to NO B. NH3 to NO3 C. N2 to NH3 D. N-3 to N2
48	Oxidation number of Cr in a C_2CrO_4 is	A. +2 B. +4 C. +6 D. +8
49	According to classical concept, oxidation involves	A. Addition of oxygen B. Addition of electron C. Removal of hydrogen D. All are correct
50	What is oxidation state of chlorine in Ca(ClO3)2	A. +1 B. +3 C. +5

		D. +7
51	Electrochemical series is the arrangement of the electrodes in	A. Increasing order of reduction potentials B. Decreasing order of reduction potentials C. Increasing order of oxidation reduction potential D. There is no fixed arrangement
52	Electromotive force of the cell is the	A. Difference of two electrode potentials B. May be sum or the difference of two electrode potentials C. Sum of two electrode potential D. Depends upon the nature of the cell
53	In electrolysis of aqueous NaCl, Cl-ions are.	A. Oxidized at anode B. Oxidized at cathode C. Reduced at cathode D. Neither oxidized nor reduced
54	Which has greater reduction potential	A. Na B. H2 C. Zn D. F2
55	The cell in which a non spontaneous redox reaction takes place as a result of electricity is known as.	A. Voltaic cell B. Denial cell C. dry Cell D. Electrolytic cell
56	During the electrolysis of molten NaCl, the ion which is reduce is	
57	Oxidation number of carbon in $NaHCO_3$	A. +4 B. -6 C. +6 D. +2
58	That cell in which electrical energy is converted into chemical energy is called	A. Galvanic cell B. Electrolytic cell C. Fuel cell D. Daniel cell
59	Stronger the oxidizing agent greater is the	A. Oxidation potential B. Reduction potential C. Redox potential D. E.M.F of cell
60	When an atom reacts chemically and loses one or more electrons it is.	A. Decomposed B. Reduced C. Oxidized D. Catalyzed
61	Which is not use of electrochemical series.	A. Feasibility of reaction B. Measurement of EMF of cell C. Comparison of reactivity with water or acids D. Determination of atomic and ionic radii
62	Alkali and alkaline earth metal are usually obtained by	A. Decomposition of their carbonates B. By heating their hydroxide C. electrolysis of molten metal oxides D. Electrolysis of molten metal halides
63	The oxidation of O -atom in OF3 is.	A. -2 B. +2 C. -1 D. +1
64	In which compound oxidation state of chlorine is +5	A. NaCl B. HOCl C. NaClO3 D. NaClO2