Electrochemistry

Sr	Questions	Answers Choice
1	A Daniell cell produces electricity through	A. Electrolysis B. Radioactivity C. Spontaneous redox reaction D. Endothermic reaction
2	Electrochemical equivalent is.	A. Mass per mole B. Mass per coulomb C. Charge per second D. Current per mass
3	In electolysis, reduction always occurs at	A. Anode B. Cathode C. Salt bridge D. Electroyte
4	Which element has Eo = 0.00 V?	A. H+ B. H2 C. SHE D. All of the above
5	The process of purifying copper using electricity is called.	A. Electrolysis B. Electrplating C. Electrorefinging D. Electrogavimetry
6	The anode is electrolysis of molten NaCl is	A. Na+ B. Cl- C. Na D. H-
7	In an electrochecmial cell, which ion migrates to cathode through the salt bridge.	A. Anions B. Cations C. Eletrons D. Protons
8	If salt bridge in not used, between twohalf cells in a Galvanic cell, than the voltage.	A. Decrease slowly B. Decrease repidly C. Does not change D. Drope to zero
9	The metal depostied at cathode during electolysis of NaCl is.	A. Cl2 B. Na C. H2 D. Cu
10	Which of the following is not a use of electrochemistry	A. Manufacture of explosives B. Electroplating C. Extractio of metals D. Electrorefining
11	The reaction at the cathode is alwyas	A. Oxidation B. Neutral C. Reduction D. Ionization
12	Electrons in a galvanic cell flow from	A. Anode to cathode B. Cathode to anode C. Salt bridge to cathode D. Elecrolyte to anode
13	Which of the following metals would most raadily displace hydrogen gas from dilute acids.	A. Copper B. Silver C. Magnesium D. Platinum
14	The salt brige allows tranfer of .....in Zn -Cu voltaic cell	A. SO4 Ions B. Zn2 ions C. Both D. None of thes
15	The process of coating a metal with zinc is called.	A. Alloying B. Galvanization C. Electroefining D. Electrplating

16	Electroysis of brine produces.	A. Hydrogen, oxygen , NaOH B. Sodium, Chlorine , H2O C. Hydrogen, Chlorine, NaOH D. NaOH , Cl2, CO2
17	Which component maintains elecrical neutrality in an electronchemical cell	A. Salt bridge B. Voltmeter C. Electrolyte D. Electrodes
18	The principle of measuring DO by Wrinkler's Method is based on.	A. Iodimetry B. Iodometry C. Acid Bse titration D. Complexometry
19	Electrolysis of CuSO4 using copper eletrodes results in	A. Increase in eletolyte concentration B. Decrease in Cu2+ concentration C. No change in electrolyte composition D. Formation of new compound
20	The experimental determination of Avogadro's numebr through electrolysis typically involves measuring	A. The current and voltage applied B. The mass of the substance deposited or liberated by a konown charge C. The conductivity of the electrolytic solution D. The temperature changes duirng electrolysis