

Electrochemistry

Sr	Questions	Answers Choice
1	What is the unit of cell potential.	A. ampere B. Volt C. Ohm D. Farad
2	The activity series of metals arranges metals in order of their.	A. Atomic mass B. Density C. Ease of oxidation D. Ease of reduction
3	Which element has $E^\circ = 0.00$ V?	A. H^+ B. H_2 C. SHE D. All of the above
4	The reducing agent in a redox reaction	A. Gains electrons B. Loss electrons C. Accepts protons D. Donates protons
5	Redox reactions always involve.	A. Gain of protons B. Transfer of electrons C. Loss of neutrons D. Nuclear change
6	If 1 Faraday of electricity is passed the mass deposited equals.	A. 1 gram equivalent B. 1 gram C. 1 mole D. 1 atm
7	The more positive the standard reduction potential	A. The stronger the reducing agent B. The weaker the oxidizing agent C. The stronger the oxidizing agent D. No effect
8	Electrolysis is the process of	A. Chemical energy into electrical B. Electrical energy into chemical C. Heat energy into mechanical D. Mechanical energy into chemical
9	Which species is reduced in the following reaction? $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$	A. Cu^{2+} B. Zn C. Cu D. Zn^{2+}
10	Electroplating involves.	A. Electrolysis for coating B. Spontaneous oxidation C. Redox without electrodes D. Thermolysis
11	Standard electrode potential is measured under	A. 100 oC and 1 atm B. 0 oC and 1 M concentration C. 25 oC and 1 M concentration D. 100 oC and 1 M concentration
12	Which of the following changes would typically lead to an increase in the rate of electrolysis.	A. Decreasing the concentration of the electrolyte B. Increasing the distance between the electrodes C. Decreasing the surface area of the electrolytic cell D. Increasing the current passed through the electrolytic cell
13	Faraday's First Law states that mass deposited is directly proportional to that in	A. Time B. Voltage C. Resistance D. Quantity of electricity passed
14	Cell potential is the difference between	A. Temperature and pressure B. Concentration of ions C. Electrode potentials of cathode and anode D. Mass of electrodes

15 The anode is electrolysis of molten NaCl is
A. Na^+
B. Cl^-
C. Na
D. H

16 Which metal is the best reducing agent.
A. Cu
B. K
C. Zn
D. Fe

17 An Increase in temperature generally
A. No Effect
B. Increases cell voltage
C. Decreases cell voltage
D. Depends on pressure

18 Down cell is used for extractio of.
A. Zinc
B. Iron
C. Sodium
D. Aluminum

19 A Daniell cell produces electricity through
A. Electrolysis
B. Radioactivity
C. Spontaneous redox reaction
D. Endothermic reaction

20 Which of the followig will react with HCl to liberate H₂ gas
A. Ag
B. Mg
C. Au
D. Cu
