

Electrochemistry

Sr	Questions	Answers Choice
1	In an electrochemial cell, which ion migrates to cathode through the salt bridge.	A. Anions B. Cations C. Eletrons D. Protons
2	The anode is electrolysis of molten NaCl is	A. Na+ B. Cl- C. Na D. H-
3	Electroysis requires.	A. Redox reaction B. Spontaneous cell C. Non spontaneous reaction D. Combustion
4	The salt brige allows tranfer ofin Zn -Cu voltaic cell	A. SO4 ions B. Zn2 ions C. Both D. None of thes
5	Electroplating involves.	A. Electrolysis for coating B. Spontaneous oxidation C. Redox without electordes D. Thermolysis
6	Which component maintains elecrical neutrality in an electronchemical cell	A. Salt bridge B. Voltmeter C. Electrolyte D. Electrodes
7	Which of the following metals would most raadily displace hydrogen gas from dilute acids.	A. Copper B. Silver C. Magnesium D. Platinum
8	The more positive the standard reduction potential	A. The stronger the reducing agent B. The weaker the oxidizing agent C. The stronger the oxidizing agent D. No effect
9	The metal deposited at cathode during electolysis of NaCl is.	A. Cl2 B. Na C. H2 D. Cu
10	Which species is educed in the followig rection ? $Zn + Cu \rightarrow Zn + Cu$	A. Cu2+ B. Zn C. Cu D. Zn3+
11	Which element has $E_o = 0.00$ V?	A. H+ B. H2 C. SHE D. All of the above
12	If E_o cell is +1.10 V , the reactio is	A. Spontaneous B. Non spontaneous C. At equilibrium D. Endothermic only
13	What is the unit of cell potanetial.	A. ampere B. Volt C. Ohm D. Farad
14	A more genative E_o value means the electrode	A. Is a stong oxidizing agent B. Is a strong reducing agent C. Is neurtal D. Has high electronegativity
15	Electrons in a galvanic cell flow from	A. Anode to cathode B. Cathode to anode C. Salt bridge to cathode D. Elecrolyte to anode

16	Which of the following will react with HCl to liberate H ₂ gas	A. Ag B. Mg C. Au D. Cu
17	Down cell is used for extraction of.	A. Zinc B. Iron C. Sodium D. Aluminum
18	Electrochemical equivalent is.	A. Mass per mole B. Mass per coulomb C. Charge per second D. Current per mass
19	Electrolysis of CuSO ₄ using copper electrodes results in	A. Increase in electrolyte concentration B. Decrease in Cu ²⁺ concentration C. No change in electrolyte composition D. Formation of new compound
20	Oxidation number of oxygen in H ₂ O	A. 1 B. -2 C. 0 D. +1