Equilibria

Sr	Questions	Answers Choice
1	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
2	When system is at quilibrium state.	A. The rate of the forward and reverse rections become equal B. The concentrationof reatants and product becoes equal C. The oposing reactions stop D. The rate of the reverse reactio becomes very low
3	What conditon Should be met for the reversible reactio to achieve the state of equilibrium.	A. The concentratio of all the reactants and the prodduct should become constant B. all the reactants should be converted into the product C. 50% of the reactant shuld be converted into prodcuts. D. One of the product should be removed from the reaction mixture.
4	Concentratin of reactants and product at equilibrium remains unchanged if	A. Concentration of any reactant or product is not changed B. Temperaure of the reation is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
5	Which of the following does not happpen, when a system is at equilibrium state.	A. Reaction continues to occur in both the directions B. Concentration of reactants and products stop changig C. Forward and reverse reactions stop D. Forward and reverse rates become equal
6	What will hapen to the concentrations of the prodcut if a reversible reaction at eqiulibrium is not distrubed.	A. They will keepon increasing B. They will keep on decreasing C. They will remain constant D. They willremain constant for some time and then start decreasing
7	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium
8	Formation of ammonia from Nitrogen and hydrogen is an.	A. Exothermic reaction B. Endothermic reaction C. Both a and b D. No heat change
9	Such reaction which continue in both directins are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
10	The reaction in which the products can recombine to formreactants are called.	A. Reversible Reaction B. Irreversible reactions C. Decomposition reactions D. Addition reactions
11	In an irrversibel reaction equilibrium	A. Never eastablised B. Established quickly C. Established slowly D. Established when reaction stops

12	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
13	The colour of anhydrous cobalt(II) cholride solid	A. White B. Black C. Pink D. Blue
14	The new substance formed in a chemcial reaction is.	A. Reverse B. Reactant C. Forward D. Product
15	At what temperature , rate of ammonia formation and decomposition is the highest.	A. 200 ^oC B. 300^oC C. 400 ^oC D. 500 ^oC
16	Why the gas starts coming out when you open a can of fizzy drink.	A. Because gteh solubiliyt of the gas increases B. Because the gas is dissolved under pressure hence it comes out when pressur eis decreased C. Becaue the gas is insoluble in water D. Because the solubility of hte gas decreases at high pressure.
17	In an irreversibel reaction equiirbrium is.	A. The forward reaction will be fovoured B. No effect on forward or backward reaction C. No effect on bakcwars reaction D. The backward reaction will be favoured
18	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
19	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
20	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually