Equilibria

Sr	Questions	Answers Choice
1	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually
2	In an irreversibel reaction equiirbrium is.	A. The forward reaction will be fovoured B. No effect on forward or backward reaction C. No effect on bakcwars reaction D. The backward reaction will be favoured
3	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
4	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
5	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
6	When a reaction will be come a reversible one?	A. If the actvation eneryg of the forward reaction is comparrble to that of backward reaction B. If the activation energy of the forward reaction is higher than that of backward reaction C. If the activationenergy of the forward reaction is lower than that of backward reaction D. If the enthalpy change of both the reactions is zero.
7	Such reaction which continue in both directins are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
8	The characteristics of reversible reactions are the following except.	A. Product nver recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. The proceed in both ways
9	Predict which components of the amosphere react in the presence of lightening.	A. N2 and H2O B. O2 and H2O C. N2 and O2 D. CO2 and O2
10	The colour of anhydrous cobalt(II) cholride solid	A. White B. Black C. Pink D. Blue
11	An Inorganic chemistry places one mle of PCl5 in container A and one mole of each Cl2 and PCl3 in container B. Both the containers were sealed and heated to the same temperture to reach the stte of equilibrium Guess about the composition of mixtures in both the containers.	A. Both the containers wil have zero concentration of its reactants. B. Both the containers wil hae the same composition of mixtures C. Container A will have more concentraion of PCl3 than B. D. Container A will have less concentraion of PCl3 than B.
		A. Concentration of any reactant or

12	Concentratin of reactants and product at equilibrium remains unchanged if	product is not changed B. Temperaure of the reation is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
13	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
14	Why the gas starts coming out when you open a can of fizzy drink.	A. Because gteh solubiliyt of the gas increases B. Because the gas is dissolved under pressure hence it comes out when pressur eis decreased C. Becaue the gas is insoluble in water D. Because the solubility of hte gas decreases at high pressure.
15	Which compound is used a thinner is paint industry?	A. H2O B. C2H3OH C. CH3COOC2H5 D. CH3COOH
16	What will hapen to the concentrations of the prodcut if a reversible reaction at eqiulibrium is not distrubed.	A. They will keepon increasing B. They will keep on decreasing C. They will remain constant D. They willremain constant for some time and then start decreasing
17	The new substance formed in a chemcial reaction is.	A. Reverse B. Reactant C. Forward D. Product
18	Industrialy, ammonia is produced by which process.	A. Halogenation B. Solvay process C. Haber Process D. Hydrogenation
19	In an irrversibel reaction equilibrium	A. Never establised B. Established quickly C. Established slowly D. Established when reaction stops
20	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium