

## Equilibria

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Sr	Questions	Answers Choice
1	Formation of ammonia from Nitrogen and hydrogen is an.	A. Exothermic reaction     B. Endothermic reaction     C. Both a and b     D. No heat change
2	What conditon Should be met for the reversible reactio to achieve the state of equilibrium.	A. The concentratio of all the reactants and the prodduct should become constant B. all the reactants should be converted into the product C. 50% of the reactant shuld be converted into prodcuts. D. One of the product should be removed from the reaction mixture.
3	A reverse reaction is one that	A. Speeds up gradually B. Proceeds from left to right C. In which reactants react to form products D. Slow down gradually
4	If reversibe reaction useful for preparing compounds on large scale.	A. Yes B. No C. They are useful only when equilirbium lies far to the left side D. They are useful only when equilibrium lies far to the right side
5	When system is at quilibrium state.	A. The rate of the forward and reverse rections become equal B. The concentrationof reatants and product becoes equal C. The oposing reactions stop D. The rate of the reverse reactio becomes very low
6	The forward reaction takes place from	A. Right to left B. Left to right C. Both a and b D. None of these
7	The new substance formed in a chemcial reaction is.	A. Reverse B. Reactant C. Forward D. Product
8	Which compound is used a thinner is paint industry?	A. H2O B. C2H3OH C. CH3COOC2H5 D. CH3COOH
9	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
10	The characteristics of reversible reactions are the following except.	A. Product over recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. The proceed in both ways
11	When the rate of the forward reaction takes place at the rate of reverse reaction the composition of the reaction mixgure remains consant. It is called.	A. Chemical Equilibrium B. Static equilirbirum C. Both a and b D. None of the above
12	In an irreversibel reaction equiirbrium is.	A. The forward reaction will be fovoured B. No effect on forward or backward reaction C. No effect on bakcwars reaction

		D. The backward reaction will be favoured
13	Concentratin of reactants and product at equilibrium remains unchanged if	A. Concentration of any reactant or product is not changed B. Temperaure of the reation is not changed C. Pressure or volume of the system is not changed D. All of the above are observed
14	At what temperature, rate of ammonia formation and decomposition is the highest.	A. 200 <sup>o</sup> C B. 300 <sup> o</sup> C C. 400 <sup>o</sup> C D. 500 <sup>o</sup> C
15	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
16	The colour of anhydrous copper (II)sulphate solid is	A. Pink B. Black C. White D. Blue
17	CaO or lime is used extensive in steel, glass and paper industries. It is produced inan exothermic reversitble reaction by the decompositon of lie . Choose the conditions to produce maximum amount of lime.	A. Heating at high temperatur ein an open vessel B. Heating at high temperatur ein a closed vessel C. Cooling it in a closed vessel D. Colling it in an open vessel
18	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium
19	The reaction in which the products do not reombine to form reatants are called	A. Addition reactions B. Decomposition reactions C. Irreversitble reactions D. reversible reactions
20	Which type of reactions speed up gradully?	A. Decomposition reaction     B. Forward reaction     C. Reverse reactions     D. Irreversibel reactions