

Energetics

Sr	Questions	Answers Choice
1	The enthalpy of reaction $\text{H}_2 + \text{I}_2 \rightarrow 2\text{HI}$	A. -571.6 kJ B. +53.8 kJ C. 11 kJ D. -393.5 kJ
2	Which is released in anaerobic respiration.	A. Stearic acid B. Citric acid C. Lactic acid D. Amino Acid
3	The word energy is used in physics for the first time.	A. 1902 B. 1858 C. 1805 D. 1802
4	The enthalpy of reaction $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$	A. -571.6 kJ B. -393.5 kJ C. +53.8 kJ D. -110.5 kJ
5	When new bonds are formed, the energy is	A. Consume B. Remain same C. Release D. None of these
6	Bond formation energy of one O-H bond is.....	A. 488 kJ/mol B. 484 kJ/mol C. 486 kJ/mol D. 489 kJ/mol
7	Bond dissociation for O_2 is	A. 505 kJ/mol B. 705 kJ/mol C. 605 kJ/mol D. 498 kJ/mol
8	The enthalpy of reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	A. -571.6 kJ B. -110.5 kJ C. -393.5 kJ D. +53.8 kJ
9	Aerobic respiration releases.....energy than anaerobic respiration.	A. Equal B. Less C. More D. None of these
10	All chemical reaction involves.	A. Enzymes B. Catalyst C. Energy changes D. All of these
11	Formation of NO is	A. Exothermic B. Endothermic C. No Heat Change D. None of these
12	During the glycolysis net ATP produced are.	A. 2 B. 4 C. 6 D. 8
13	No reaction occurs if the energy of reacting particles.....activation energy.	A. Lower than B. Greater than C. Nearest to D. Equal to
14	Who used the word energy for the first time	A. Rutherford B. Bohr C. Thomas Young D. None of these
15	----- of the energy used by traditional electric bulb is wasted in producing heat.	A. 60% B. 50% C. 70% D. 90%

16	If the Delta H value is negative then reaction will be	A. Endothermic B. Exothermic C. May or may not be exothermic or endothermic D. None of these
17	Which is not produced in an aerobic respiration.	A. Carbon dioxide B. Lactic acid C. Water D. Energy
18	----- acts a catalyst promoting the breakdown of ozone.	A. I ₂ B. Br ₂ C. Cl ₂ D. None
19	When NaOH and HCl are mixed the temperature increases. The reaction	A. Exothermic with a negative enthalpy change. B. Endothermic with a positive enthalpy change. C. Endothermic with a negative enthalpy change D. Exothermic with a positive enthalpy change
20	When old bonds are broken, the energy is.	A. Release B. Remain same C. Consume D. None of these