Stoichiometry

Sr	Questions	Answers Choice
1	Formula of Ozone.	A. O_2 B. S_8 C. CO_2 D. O_3
2	Formula of common salt is	A. NaCl B. AgCl C. LiCl D. KCl
3	Whof the following represent sand?	A. NaCl B. CaCO3 C. H2O D. CH2O
4	What mass of 95% CaCO3 will be requred to neutralize 50 cm 3 of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
5	How many atoms are present in one gram of H2O?	A. 1002×10^{23} atom B. 6.022×10^{23} atom C. 0.334×10^{23} atom D. 2.004×10^{23} atom
6	How many number of moles are equivalent to 8 gram of CO2.	A. 0.18 B. 0.15 C. 0.24 D. 0.21
7	Empirical formula of acetic acid (CH3COOH) is	A. CHO B. CH C. CH2O D. None of these
8	Which one of the following compounds will hae the highest percentage of the mass of nitrogen?	A. N2H4 B. CO (NH2)2 C. NH3 D. NH2OH
9	Which of the following is insoluble salt	A. KCl B. AgCl C. NaCl D. CaCl2
10	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
11	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
12	A nceklaee has 6 g of diamonds in it . What are the numberof carbon atoms in it?	A. 3.01×10^{23} B. 1.003×10^{23} C. 12.04×10^{23} D. 6.02×10^{23}
13	Avogadrao was a scientist	A. Italian B. Greek C. Garman D. African
14	How many moels of molecules are there in 16 g oxygen.	A. 0.05 B. 0.1 C. 0.5 D. 1
15	The mass number of sodium is.	A. 19 B. 31 C. 27 D. 23

16	A compound with chemicla formula Na_2CX_3 has formula mass 106 ami. Atomicmass of the elemetn X is.	A. 16 B. 23 C. 12 D. 106
17	How many moles are there in 25 g of H_2SO_4?	A. 0.765 Moles B. 0.255 moles C. 0.4 moles D. 0.51 moles
18	Numbr of hydrogen atoms present in 18 g of water.	A. 2 x NA B. N_A C. 3 x N_A D. 1/2 N_A
19	Empirical formula of hydrogen peroxide.	A. HO B. CO C. CHO D. CH
20	Which is the correct formula of calcium phosphide.	A. CaP B. Ca_3P_2 C. CaP_2 D. Ca_2P_3