Stoichiometry

Sr	Questions	Answers Choice
1	Empirical formula of acetic acid (CH3COOH) is	A. CHO B. CH C. CH2O D. None of these
2	What mass of 95% CaCO3 will be requred to neutralize 50 cm 3 of 0.5 M HCl solution.	A. 9.5 g B. 1.45 g C. 1.32 g D. 1.25 g
3	Avogadrao was a scientist	A. Italian B. Greek C. Garman D. African
4	The molecular mass of H2SO4 is.	A. 9.8 g B. 98 amu C. 98 g D. 9.8 amu
5	Whcih law is obeyed in chemicla calculations?	A. Law of mass action B. Law of conservationof mass C. Law of definite proportion D. Both a and b
6	Mass of 3 moles of oxygen atoms is.	A. 64 g B. 16 g C. 32 g D. 48 g
7	Molecular mass of Acetic Acid	A. 70 amu B. 43 amu C. 60 amu D. 80 amu
8	What is the mass of 4 moles of hydrogen gas.	A. 1 g B. 1.008 g C. 8.064 g D. 4.032 g
9	The mass of one molecule of water is.	A. 18 g B. 18 mg C. 18 kg D. 18 amu
10	40 g of H_3PO_4 contains numebr of moles	A. 0.58 g B. 4.8 g C. 5.8 g D. 0.408 g
11	If one mole of carbon contains x atoms what is the number of atoms contained in 12 g of Mg.	A. 1.5 x B. 0.5 x C. x D. 2x
12	Without stoichiometry which industry cannot exist.	A. Meta B. Petroleium C. Leather D. Chemical
13	1 gram formula of NaCl contan is grams.	A. 100 g B. 58.5 g C. 32 g D. 49 g
14	Whof the following represent sand?	A. NaCl B. CaCO3 C. H2O D. CH2O
15	Limes is another name of	A. Sodium hydroxide B. Sodium Carbonate C. Calcium Carbonate D. Silicon dioxide

16	A compound with chemicla formula Na_2CX_3 has formula mass 106 ami. Atomicmass of the elemetn X is.	A. 16 B. 23 C. 12 D. 106
17	1- gram atom of carbon contain hwo many moles.	A. 1 mole B. 2 mole C. 6 moles D. 12 moles
18	Formula of Ozone.	A. O₂ B. S₈ C. CO₂ D. O₃
19	Which one of the following compounds will hae the highest percentage of the mass of nitrogen?	A. N2H4 B. CO (NH2)2 C. NH3 D. NH2OH
20	Which compund has same molecular and emiprical formula.	A. C6H12 O6 B. H2O2 C. H2O D. C6H6