

## Chemical Bonding

Sr	Questions	Answers Choice
1	When molten copper and molten zinc are mixed together, they give rise to a new substance called brass. Predict what type of bond is formed between copper and zinc.	A. Ionic bond B. Coordinate Covalent bond <b>C. Metallic bond</b> D. Covalent Bond
2	How many electrons are there in the valence shell of sodium atom.	<b>A. One</b> B. Two C. Three D. Four
3	Triple covalent bond involves how many electrons.	<b>A. Six</b> B. Four C. Eight D. Three
4	How many covalent bonds do N <sub>2</sub> molecules have	<b>A. 3</b> B. 4 C. 2 D. 5
5	Water has a high boiling point as compared to alcohol due to	A. Low density B. High surface tension <b>C. Hydrogen bonding</b> D. High vapour pressure
6	During the formation of an ionic bond, heat is	A. Remains same B. Absorbed <b>C. Released</b> D. Both a and b
7	When an electronegative element combines with an electropositive element, the type of bonding is.	A. Covalent B. Polar Covalent <b>C. Ionic</b> D. Coordinate Covalent
8	Dative covalent bonds are also known as	A. Covalent bond B. Ionic Bond C. Metallic Bond <b>D. Coordinate covalent bond</b>
9	How many lone pairs are present on nitrogen in an ammonia molecule.	<b>A. One</b> B. Two C. Three D. Four
10	Which metal has the lowest melting point?	A. Li B. Na <b>C. Rb</b> D. K
11	Metals have the tendency to lose electrons due to	A. High ionization energies <b>B. Low ionization energies</b> C. Low electron affinity D. None of the above
12	The boiling point of alcohol is	A. $44^{\circ}\text{C}$ <b>B. <math>78^{\circ}\text{C}</math></b> C. $53^{\circ}\text{C}$ D. $19^{\circ}\text{C}$
13	A bond pair in covalent molecules usually has	A. One electron <b>B. Two electrons</b> C. Three electrons D. Four electrons
14	Nitrogen molecule contains	A. Polar covalent bond <b>B. Triple Covalent bond</b> C. Double covalent bond D. Single covalent bond
15	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 <b>B. Group -17</b> C. Group -2 D. Group -18

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16	Which molecule contains a single covalent bond.	A. CH <sub>4</sub> B. C <sub>2</sub> H <sub>4</sub> C. C <sub>2</sub> H <sub>2</sub> D. O <sub>2</sub>
17	Noble gases are non-reactive, because they do not.	A. Gain electrons B. Lose electrons C. Share electrons D. All of these
18	Ionic compounds are good conductors of electricity in	A. Solution B. Molten state C. Solid state D. both a and b
19	Non-polar compounds are insoluble in	A. Alcohol B. Benzene C. Ether D. Water
20	All the noble gases have their valence electrons.	A. Incomplete B. Partially filled C. Completely filled D. None of the above

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