Biochemistry

Sr	Questions	Answers Choice
1	Which force is responsible for bonding the oppositely charged ions together in a salt.	A. Nuclear force B. Electrostaitic force of attraction C. Gravitational force D. Megnetic force
2	Ionic compunds serve as	A. Weak electroytes B. Bad electroytes C. good electrolytes D. Strong electrolytes
3	Higher charges and smaller ion size lead to stronger attrractions and hence	A. Higher boiling points B. Lower melting points C. Lower boiling points D. Higher melting points
4	In which state of matter ions are held together by storng forces of attraction	A. Solid state B. Liquid state C. Gaseous state D. All of these
5	What are the primary products formd when an acid neutralizes a base or an alkali.	A. Salt and water B. Salt and hydrogen gas C. Salt and a metal D. Salt and an oxide.
6	What happens to the ions in an ionic compounds when it is heated to its melting point.	A. they become fixed B. They disappear C. They become free to move D. None of these
7	According to the solubility rules, which of these groups of salt are always soluble in water.	A. All sodium, potassium, and ammonium salts. B. Barium and lead sulfates C. Silver and lead chlorides D. All of these
8	Which of the following salts is only partially soulube in water.	A. Potassium hydroxide B. Calcium hydroxide C. Ammonium hydroxide D. Sodkum hydroxide
9	Ionic compunds typically have high melting points of.....electrostatic of attraction are present between the ions.	A. Strong B. Very Strong C. Weak D. Very Weak
10	Which base is not soluble in water.	A. KOH B. Mg(OH)2 C. NaOH D. Na2CO3
11	How was common salt used in the past	A. As currency B. As a fuel C. As a source of light D. As source of radiation
12	Reaction of water solubel acid with water soluble base is also known as.	A. Filtration B. Crystallization C. Titration method D. Evaporation
13	If a salt being prepared is stabel towards heat, how should it be recovered from the solution.	A. filtraiton B. Evapooration to dryness C. Cooling D. Heating
14	in the reaction of NaOH +HCl.....NaCl = H2O What does NaCl represent.	A. Hydrochlroic acid B. Sodium hydroxide C. Sodium chloride D. Hydrogen oxide
15	All carbonates are insoluble in water with the exception of carbonates of.	A. Sodium B. Potassium C. Ammonium

		D. All of these
16	In which state doionic compunds NOT conduct electricity.	A. Molten state B. Solid state C. Liquid state D. Aqueous solution
17	In reaction of an acid with an insoluble base, what is the step taken to remvoe the excess insoluble base from the mixture.	A. Filtration B. Evaporation C. Crystallization D. Stirring
18	Ions are held together by storng forces of attractions in	A. Liquid state B. Solid state C. Gaseous state D. All of these
19	Which of the following salt is used as afungicide.	A. Magnesium hydroxide B. Potassium nitrate C. Sodium carbonate D. Copper sulphate
20	Where are the smaller sodium ions located in the sodium crystal cube.	A. At the corners of teh cube B. At the center of each face C. In the edges of teh cube D. Floating outside the cube.