Biochemistry

Sr	Questions	Answers Choice
1	According to the solubility rules, which of these groups of salt are always soluble in water.	A. All sodium, potassium, and ammonium salts. B. Barium and lead sulfates C. Silver and lead chlorides D. All of these
2	When an ionic compound dissolves in water the forces of attraction between its ions.	A. Become stronger B. Break down C. Remain the same D. Completely finish
3	Which force is responsible for bonding the oppositely charged ions together in a salt.	A. Nuclear force B. Electrostaitic force of attraction C. Gravitational force D. Megnetic force
4	In which state doionic compunds NOT conduct electricity.	A. Molten state B. Solid state C. Liquid state D. Aqueous solution
5	Which indicator is used to neutrialize soium hydroxide with sulphuric acid.	A. Methyl orange B. Phenophthalein C. Litmus D. a and b both
6	Ions are held together by storng forces of attractions in	A. Liquid state B. Solid state C. Gaseous state D. All of these
7	Which of these salts is typically insoluble in water.	A. Sodium Carbonate B. Ammonium Nitrate C. Lead Sulphate D. Potassium chloride
8	What is the geometric shape of the structure of NaCl lattice.	A. Hexagonal prism B. Pyramid C. Face centered cube D. Octa hedron
9	What is the three dimensional network structure formed by ions in a salt called.	A. Atomic cloud B. Crystal lattice C. Liquid matric D. Molecular chain
10	in the reaction of NaOH +HCl.....NaCl = H2O What does NaCl represent.	A. Hydrochlroic acid B. Sodium hydroxide C. Sodium chloride D. Hydrogen oxide
11	Ionic compunds serve as	A. Weak electroytes B. Bad electroytes C. good electrolytes D. Strong electrolytes
12	The forces of attraction between ions break down and the ions are again to move independently, when ionic compounds are.	A. In solid state B. In molten state C. Dissolved in water D. a and b
13	All carbonates are insoluble in water with the exception of carbonates of.	A. Sodium B. Potassium C. Ammonium D. All of these
14	Which of the following salts is only partially soulube in water.	A. Potassium hydroxide B. Calcium hydroxide C. Ammonium hydroxide D. Sodkum hydroxide
15	If a salt being prepared is stabel towards heat, how should it be recovered from the solution.	A. filtraiton B. Evapooration to dryness C. Cooling

		D. Heating
16	In which state of matter ions are held together by storng forces of attraction	A. Solid state B. Liquid state C. Gaseous state D. All of these
17	How should pure salt crystals be dried once they are filtered.	A. By heating B. By leaving in a beaker C. Between the folds of filter paper D. None of these
18	Ionic compunds typically have high melting points of.....electrostatic of attraction are present between the ions.	A. Strong B. Very Strong C. Weak D. Very Weak
19	What happens to the ions in an ionic compounds when it is heated to its melting point.	A. they become fixed B. They disappear C. They become free to move D. None of these
20	To prepare soluble potassium Nitrate which acid and base should be used in a titration method.	A. Hydrochloric acid and Potassium hydroxide B. Sulpuric acid and potassium carbonate C. Nitric acid and potassium hydroxide D. Nitric acid and sodium hydroxide.