Acid bases and Salts

Sr	Questions	Answers Choice
1	It is esential for optimizing chemical reactions in many industries.	A. Stoichimetry B. Biochemistry C. Organic chemistry D. Electrochemistry
2	What is teh room temperature for standard conditions.	A. O oC B. 25 oC C. 75 oC D. 100 oC
3	How many moles of natural gas CH4 will be present in a cylinder if its volume is 18 dm3 at RTP?	A. 0.75 Mol B. 0.67 mol C. 0.85 mol D. 0.56 mol
4	How many moles of HCl are needed to react with 1 mole of Ca CO3 in the given reaction.	A. 1 mole B. 2 moles C. 3 moles D. 4 moles
5	Empirical formula of water is.	A. H2O B. 2H2O C. H2O2 D. HO
6	Emperical data required for titration calculations includes.	A. Volumes of two solutions used B. The atmospheric pressure only C. The weight of the burette D. None of these
7	What is teh standard unit for the mass/volume concentration.	A. mol/kg B. g/cm3 C. mg/l D. g/dm3
8	Benzene and acetyle both have the same emperical formula which is.	A. CH2 B. C2H C. CH D. CH4
9	In a titratin, the solution is added unitll the ..... point is reached.	A. starting B. Boiling C. End D. Freezing
10	Which factor does NOT chagne teh volume occupied by one mole of a gas a RTP.	A. Temperature B. Pressrue C. Size of the molecules D. Numebr of moles
11	The quantity of product actually obtained when a chemical raetion is performed is called.	A. theoreticla yield B. Actual yield C. Percent purity D. Stoichiometric yield
12	Molar concentration is expressed in which of the following units.	A. g/dm3 B. mol/g C. mol/dm3 D. dm3/mol
13	Glucose has an empirical formula of CH2O and a molecular mass of 180 .What is its molecular formula.	A. C2H4O3 B. C4H8O4 C. C6H12O6 D. C4H8O4
14	For titration purposes, you need to know the concentration of.	A. Both solutions beforehand B. The other solution C. Indicator D. Both a and b
15	The formula that represents the simpleast whole numebr ratio of atoms in compund is the	A. Empricial formula B. Molecular formula C. Molar volume D. Percentage composition

16	In the formula for concentration, what is the denominator.	A. Mass of solute in grams B. Mass of solvent in kg C. Volume of solution in dm3 D. Temperature of the solution.
17	When natural gas burns in air, which component is the limiting reactant.	A. Both components B. Neither component is a limiting react C. Air D. natural gas
18	The reactant that is consumed earlier in a reaction and controls the amount product formed is the.	A. Excess reactant B. Catalyst C. Limiting reactant D. Theoretical reactant
19	the pressure value used for standard RTP conditons is.	A. 100 kpa B. 1 atmosphere (760 torr) C. 273 torr D. 24 atmosphere
20	Concentratio of solution is a measure of the amount of solute dissolved in given amount of.	A. solution or solvent B. Only Solute C. Solid crystals D. Catalyst