Chemical Equilibrium

Sr	Questions	Answers Choice
1	When the magnitude of Kc is very small in indicates.	A. Equilibrium will never establish B. All reactants will converted to products. C. Reaction will go to completion D. The amount of products is negligible
2	In chemical reaction , the substances that combine are called.	A. Reactants B. Products C. Equlibrium D. Numerator
3	In irreversible reaction, dynamic equilibrium:	A. Never establishes B. Establishes before the completion of reaction C. Establishes after the completion of reaction D. Establishes readily
4	In the beginning the rate of reverse reaction is	A. Negligible B. Moderate C. Very fast D. Slow
5	Which type of reactions do not go to completion?	A. Irreversible reaction B. Reversible reactions C. Addition reactions D. Decomposition reactions
6	In a Reversible Reaction if Qc = Kc then.	A. Reaction is occuring in forward direction B. Reaction is occuring in Reverse direction C. Equilibrium has been allained D. Reaction is not at equilibrium
7	The unit of molar concentration is:	A. moldm⁻³ B. moldm⁺³ C. molcm⁻³ D. molcm⁺³
8	For which reaction Kf is the rate constant?	A. Forward reaction B. Backward reaction C. Upward reaction D. Downward reaction
9	The color of hydrogen Iodide is:	A. Colourless B. Black C. Red D. Pink
10	Active mass is expressed as:	A. { } B. [ ] C. ( ) D. I I
11	Who proposed "Law of mass action"?	A. Newton B. Boyle C. Guldberg and waage D. Lavoisier
12	The characteristics of reversible reactions are the following except:	A. Products never recombine to form reactants. B. They never complete C. The proced in both ways D. They have a double arrow between reactants and products.
13	When the numbers of moles of both sides are equal in a reaction, then the unit of K_c will be:	A. No unit B. mol⁻² dm⁶ C. mol dm³ D. mol⁻²dm
		A. Reversible reaction

14	The reaction in which the products can recombine to form reactants are called.	B. Irreversible reaction C. Decomposition D. Addition
15	At dynamic equilibrium state:	A. Rate of forward reaction≠ Rate of reverse reaction B. Rate of forward reaction > Rate of reverse reaction C. Rate of forward reaction = Rate of reverse reaction D. Rate of forward reaction < rate of reverse reaction
16	An equilibrium is achievable only in a:	A. big system B. small system C. open system D. closed system
17	The color of Iodine is:	A. Purple B. Black C. Red D. Pink
18	Law of Mass Action was put forward by:	A. G.N.Lewis B. Lowry C. Arrhenius D. Gulbderg and Waage
19	When the magnitude of Kc is very large it indicates.	A. Reaction mixture consist of almost all products. B. Reaction mixture consist of almost all reactants. C. Reaction has not gone to completion D. Reaction mixture has negligible products.
20	When a system is at equilibrium states?	A. The concentration of reactants and products becomes equal B. The opposing reactions C. The rate of the reverse reaction becomes very low D. The rates of the forward and reverse reactions becomes equal.