

Electrochemistry

C.	Overtions	Anguaya Chair-
Sr	Questions	Answers Choice
1	Which of the following yield both hydrogen and chlorine on electrolysis:	A. Electrolysis of acidified water. B. Electrolysis of molten NaCl C. Electrolysis of aqueous NaCl D. Electrolysis of saline water
2	F ₂ , Cl ₂ , Br ₂ and l ₂ lie below SHE is the Electro chemical series that is why these	A. Undergo reduction B. Undergo oxidation C. Liberate H ₂ gas with steam D. None of these
3	In a Galvanic cell	A. Chemical energy is converted into electricity B. Chemical energy is converted into heat C. Electrical energy is converted into heat D. Electrical energy is converted into chemical energy
4	If a salt bridge is removed between the two half cells, the voltage	A. Drops to zero B. Does not change C. Increases gradually D. Increases rapidly
5	The equivalent conductivity of 0.1 M week acid is 100 times less than at infinite dilution. The degree of dissociation is	A. 100 B. 10 C. 0.01 D. 0.001
6	The unit of specific conductivity is	A. Ohm cm ⁻¹ B. Ohm cm ⁻² C. Ohm ⁻¹ cm D. Ohm ⁻¹ cm ⁻¹
7	96500 C of electricity liberates from CuSO ₄ solution	A. 63.5 g of Cu B. 31.75 g of Cu C. 96500 g of Cu D. 100 g of Cu
8	While balancing an equation by ion electron method, the number of oxygen atoms are balanced by	A. OH ⁻ B. H ₂ O C. O ₂ D. H ⁺
9	Which of the following will be good conductor of electricity	A. Pure distilled water B. Molten NaCl C. Dilute solution of glucose D. Chloroform
10	Question Image	A. lodine can oxidise iron B. Bromine can oxidise iron C. lodine can oxidise bromine D. Iron can oxidise bromine
11	The oxidation number of Mn is KMnO4is	A. +2 B. +4 C. +6 D. +7
12	Pure water does not conduct electricity because it	A. Has low boiling point B. Is almost unionized C. Is neutral D. Is readily decomposed
13	Which statement is correct for the fuel cells	A. These cells operate at low temperature B. These cells operate at low temperature C. No catalyst used for the formation of water D. MnO ₂ is used as electrolyte

14	Same amount of electric current is passed through solutions of AgNO3and HCl. If 1.08 g of silver is obtained in the first case, the amount of hydrogen liberated as S.T.P in the second case is	A. 112 cm ³ B. 22400 cm ³ C. 224 cm ³ D. 1.008 g
15	In the presence of copper electrodes the electrolysis of aqueous CuSO ₄ produces at cathode	A. H ₂ gas B. O ₂ gas C. SO ₂ gas D. Cu metal
16	96500 C electricity is passed through CuSO ₄ . The amount of copper precipitated is	A. 0.25 mole B. 0.5 mole C. 1.0 mole D. 2.00 mole
17	The galvanic or voltaic cells which are rechargeable called as	A. Primary cells B. Secondary cells C. Dry cells D. Infinite cells
18	Metallic conduction is also called as :	A. lonic conduction. B. Protonic conduction. C. Electronic conduction D. Super conduction
19	What are the products electrolysis of aqueous sodium chloride at two electrodes	A. Chlorine at anode and oxygen at cathode. B. Hydrogen at anode and chlorine acathode. C. Chlorine at anode and hydrogen at cathode. D. Chlorine at anode and sodium at cathode.
20	The reduction potential to copper electrode is +0.34 V and that of Zn electrode is -0.76 V. when these two are coupled the e.m.f. of the cell is	A0.42 V B. +0.42 C1.10 V D. +1.10 V