ECAT Chemistry Chapter 9 Solutions

Sr	Questions	Answers Choice
1	How many g of dibasic acid (mol. wt. 200) should be present in 100 ml of the aqueous solution to give 0.1 Normality?	A. 1 g B. 2 g C. 10 g D. 20 g
2	A solution is a homogeneous mixture of two or more kinds different :	A. Molecular. B. Covalent substance C. Ionic Substances D. Both (a) and (c)
3	Camphor is often used in molecular mass determination because	A. It is solvent for organic substances B. It is readily available C. It has a very high cryoscopic constant D. It is volatile
4	Isotonic solutions have same	A. Molar concentration B. Molality C. Normality D. None of these
5	The molarity of solution containing 14.5 grams urea (N_2H_4CO) dissolved in 100 $cm^3$of the solution is	A. 1 molar B. 0.1 molar C. 0.2 molar D. 0.25 molar
6	Hydrolysis of potassium acetate produces	A. Acidic solution B. Neutral solution C. Basic solution D. None of these
7	The temperature at which the vapour pressure of a liquid becomes equal to external pressure is	A. Melting point B. Sublimation point C. Inversion point D. Boiling point
8	The weight of pure NaOH required to prepare 250 $cm^3$of 0.1 N solution is	A. 4 g B. 1 g C. 2 g D. 5 g
9	What mass of NaOH is required to prepare 2.5 $dm^3$of 1.5 M NaOH solution	A. 130 g B. 140 g C. 150 g D. 160 g
10	Aqueous solution of glucose $C_6H_{12}O_6$, boils at 100.052°C. The solution contains	A. 180 grams glucose in 1 kg water B. 18 grams glucose in 1 kg water C. 1.8 grams glucose in 1 kg water D. 3.6 grams glucose in 1 kg
11	The molal depression constant depends upon	A. Nature of solute B. Nature of solvent C. Δ H_{solution} D. Vapour pressure of solution
12	When the solute is present in trace quantities the following expression is used	A. Gram per million B. Milligram percent C. Microgram percent D. Parts per million
13	Which of the following liquid pairs shown a positive deviation from Raoult's law	A. CH₃COOH₃+ CH₃Cl B. C₆H₆+ CH₄OH C. H₂O + HCl D. H₂O + HNO₃
		A. 1/5

14	18 g glucose is dissolved in 90 g of water. The relative lowering vapor pressure is equal to :	B. 5.1 C. 1/51 D. 6
15	10% aqueous solution of NaCl has molarity	A. 1.7 M B. 2.7 M C. 0.17 M D. 3.7 M
16	The sum of mole percent of all the components of solution is always equal to :	A. Less than 100 B. One C. 100 D. 10
17	The osmotic pressure of solution increases if	A. Temperature is decreased B. Solution constant is increased C. Number of solute molecules are increased D. Volume is increased
18	A solution has 92 g of ethyl alcohol, 96 g of methyl alcohol and 90 g of water. Mole percentage of ethyl alcohol in the solution is	A. 10 B. 20 C. 25 D. 50
19	Which one of the following is an ideal solution that obeys Rault's law	A. Ethanol + water B. Benzene + toluene C. HCl + water D. Acetone + chloroform
20	Which of the following is a colligative property?	A. Melting point B. Osmotic pressure C. Freezing point D. Sublimation temperature