

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
2	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
3	Which of the following liquids has low vapor pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
4	Force of attraction b/w atoms of He is :	A. London dispersion forces B. Hydrogen bonding C. Coordinate covalent bond D. Covalent bond
5	A liquid on evaporation causes	A. Heating effect B. Cooling effect C. Suffocation D. All of above
6	In a group on going downward, polarizability generally:	A. Decreases B. Increases C. Remain constant D. Negligible
7	The molecules of CO ₂ in dry ice form the	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal
8	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid ₂ D. Molecular of hydrochloric acid gas
9	The density of water decreases, When it is freezed at 0°C	A. Change of bond length B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of ice
10	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
11	Hydrogen bonding is present in one of the following pairs:	A. NH ₃ B. H ₂ O C. HF D. All of above
12	The strongest forces are:	A. Debye forces B. London dispersion C. Dipole-dipole attraction D. Hydrogen bonding
13	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg
14	The inter-molecular forces in liquids are:	A. Negligible B. Very weak C. Very strong D. Reasonably strong
15	Chloroform and acetone are soluble in each other due to	A. Instantaneous dipole interactions B. Dipole-dipole interactions C. Intermolecular hydrogen bonding

D. All of above

16 London dispersion forces are also called:

A. Hydrogen bonding.
B. Debye forces
C. Van der Waal's forces
D. Instantaneous dipole-induced dipole forces.

17 H_2S is a gas which H_2O is liquid at room temperature. it is due to

A. Less intermolecular forces in water
B. Covalent bond in $\text{H}-\text{O}$ in water molecule
C. Hydrogen bonding in water molecules
D. Ionic characters in water molecules

18 Ionic Solids are characterized by

A. Low melting points
B. Good conductivity in solid state
C. High vapour pressure
D. Solubility in polar solvents

19 Debye forces are present in one of the following pairs

A. Na^{+} and water
B. Argon and water
C. Argon and Na^{+}
D. Ne and Water

20 Evaporation of water is possible at:

A. Above 100°C
B. 0°C
C. 100°C
D. At all temperature