

## ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	A liquid on evaporation causes:	<p>A. Heating effect.  <b>B. Cooling effect.</b>            C. Suffocation .            D. All of above</p>
2	Chloroform and acetone are soluble in each other due to	<p>A. Instantaneous dipole interactions            B. Dipole-dipole interactions  <b>C. Intermolecular hydrogen bonding</b>            D. All of above</p>
3	Which of the following is a pseudo solid?	<p>A. <math>\text{CaF}_2</math>  <b>B. Glass</b>            C. <math>\text{NaCl}</math>            D. All</p>
4	HF has exceptionally low acidic strength due to:	<p>A. Smaller size of fluorine.            B. Stronger polar bond between H and F.            C. Electronegativity of fluorine  <b>D. Strong hydrogen bonding</b></p>
5	Hydrocarbon molecules with large chain lengths experience	<p>A. Weaker attractive forces  <b>B. Stronger attractive forces</b>            C. Repulsive forces            D. No attractive forces</p>
6	Which of the following liquids has low vapor pressure at 25°C:	<p>A. Diethyl ether            B. Acetone.  <b>C. Water.</b>            D. Ethyl alcohol.</p>
7	A liquid on evaporation causes	<p>A. Heating effect  <b>B. Cooling effect</b>            C. Suffication            D. All of above</p>
8	The boiling point of $\text{NH}_3$ is maximum among the hydrides of group V elements due to	<p><b>A. Enhanced electronegative character of Nitrogen</b>            B. Pyramidal structure of <math>\text{NH}_3</math>            C. Very small size of Nitrogen            D. Enhanced electropositive character of Nitrogen</p>
9	The boiling point of $\text{NH}_3$ is maximum among the hydrides of group V elements due to:	<p>A. Enhanced electronegative character of Nitrogen  <b>B. Pyramidal structure of <math>\text{NH}_3</math></b>            C. Very small size of Nitrogen.            D. Enhanced electropositive character of Nitrogen</p>
10	Evaporation of water is possible at:	<p>A. Above 100°C            B. 0°C  <b>C. 100°C</b></p>

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D. At all temperature

11 Ionic solids are characterized by :

- A. Low melting points
- B. Good conductivity in solid state
- C. High vapor pressure
- D. Solubility in polar solvents

12 London dispersion forces are also called

- A. Hydrogen bonding
- B. Debye forces
- C. Van der Waal's forces
- D. Instantaneous dipole-induced dipole forces

13 Polypeptide chains are coiled about one another into a spiral by

- A. Ionic bonds
- B. Covalent bonds
- C. Van der Waal's forces
- D. Hydrogen bonds

14 Ionic Solids are characterized by

- A. Low melting points
- B. Good conductivity in solid state
- C. High vapour pressure
- D. Solubility in polar solvents

15 The strongest forces are

- A. Debye forces
- B. London dispersion forces
- C. Dipole-dipole attraction
- D. Hydrogen

16 Heat of vapourization for liquids with strong dipole-dipole forces will have

- A. Negligible Values
- B. Reasonably high values
- C. Very high values
- D. very low values

17 Which of the following liquids has low vapour pressure at 25°C:

- A. Diethyl ether
- B. Acetone.
- C. Water.
- D. Ethyl alcohol.

18 Which one of the following is the weakest intermolecular force

- A. Dipole induced dipole forces
- B. Ionic dipole forces
- C. Electrostatics forces between ions
- D. Dipole-dipole forces

19 Trend of boiling point of halogens from fluorine to iodine is that it

- A. Decreases
- B. Is negligible
- C. Increases
- D. Remains constant

20 Debye forces are present in on of the following pairs:

- A. Na<sup>+</sup>ion and water
- B. Argon and water
- C. Argon and Na<sup>+</sup>ion
- D. Ne and water