

ECAT Chemistry Chapter 4 Liquids & Solids Online Test

Sr	Questions	Answers Choice
1	In a group on going downward, polarizability generally:	A. Decreases B. Increases C. Remain constant D. Negligible
2	Ionic solids are characterized by :	A. Low meeting points B. Good conductivity in solid state C. High vapor pressure D. Solubility in polar solvents
3	The density of water decreases, when it is freezed at 0°C because of	A. Change of bond lengths B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of Ice
4	H ₂ S is a gas while H ₂ O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
5	The gases can be converted into liquids by	A. increasing the pressure only B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
6	Hydrogen bonding is present in one of the following pairs:	A. NH ₃ B. H ₂ O C. HF D. All of above
7	The intermolecular forces in liquids are	A. Negligible B. Very weak C. Very strong D. Reasonably strong
8	Which of the following is a pseudo solid?	A. CaF_2 B. Glass C. NaCl D. All
9	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
10	Chloroform and acetone are soluble in each other due to:	A. Instantaneous dipole interactions. B. Dipole-dipole interactions. C. Inter molecular hydrogen bonding. D. All of above
11	Polypeptide chains are coiled about one another into a spiral by:	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
12	The boiling point of NH ₃ is maximum among the hydrides of group V elements due to:	A. Enhanced electronegative character of Nitrogen B. Pyramidal structure of NH ₃ C. Very small size of Nitrogen. D. Enhanced electropositive character of Nitrogen
13	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of water in solid state D. Molecules of water in gaseous state

		<p>C. Molecules of solid I_2</p> <p>D. Molecular of hydrochloric acid gas</p>
14	The boiling point of NH_3 is maximum among the hydrides of group V elements due to	<p>A. Enhanced electronegative character of Nitrogen</p> <p>B. Pyramidal structure of NH_3</p> <p>C. Very small size of Nitrogen</p> <p>D. Enhanced electropositive character of Nitrogen</p>
15	At sea level and at $100^\circ C$ the vapour pressure of water in an open system is	<p>A. 1000 mm Hg</p> <p>B. 760 mm Hg</p> <p>C. 730 mm Hg</p> <p>D. 670 mm Hg</p>
16	Hydrogen bonding is present between the molecules of	<p>A. NH_3</p> <p>B. H_2O</p> <p>C. HF</p> <p>D. All of above</p>
17	Which of the following liquids has low vapor pressure at $25^\circ C$:	<p>A. Diethyl ether</p> <p>B. Acetone.</p> <p>C. Water.</p> <p>D. Ethyl alcohol.</p>
18	Diamond is a bad conductor because:	<p>A. It has tight structure</p> <p>B. It has a high density</p> <p>C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light</p>
19	The strongest forces are:	<p>A. Debye forces</p> <p>B. London dispersion</p> <p>C. Dipole-dipole attraction</p> <p>D. Hydrogen bonding</p>
20	At sea level and at $100^\circ C$ the vapor pressure of water in an open system is:	<p>A. 1000 mm Hg</p> <p>B. 760 mm Hg</p> <p>C. 730 mm Hg</p> <p>D. 670 mm Hg</p>
21	Which one of the following is the weakest intermolecular force	<p>A. Dipole induced dipole forces</p> <p>B. Ionic dipole forces</p> <p>C. Electrostatics forces between ions</p> <p>D. Dipole-dipole forces</p>
22	Evaporation of water is possible at:	<p>A. Above $100^\circ C$</p> <p>B. $0^\circ C$</p> <p>C. $100^\circ C$</p> <p>D. At all temperature</p>
23	Escape of high energy molecules from the surface of a liquid is called	<p>A. Sublimation</p> <p>B. Distillation</p> <p>C. Condensation</p> <p>D. Evaporation</p>
24	In a group on going downward, polarizability generally	<p>A. Decreases</p> <p>B. Increases</p> <p>C. Remains constant</p> <p>D. Negligible</p>
25	Debye forces are present in on of the following pairs:	<p>A. Na^+ and water</p> <p>B. Argon and water</p> <p>C. Argon and Na^+</p> <p>D. Na and water</p>

26	Heat of vapourization for liquids with strong dipole-dipole forces will have	A. Negligible Values B. Reasonably high values C. Very high values D. very low values
27	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
28]Which of the following is a pseudo solid?	A. Atoms of He is gaseous stat at high temperate. B. Molecules of water in liquid state. C. Molecules of solid I ₂ D. Molecular of hydrochloric acid gas
29	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
30	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
31	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
32	HF has exceptionally low acidic strength due to:	A. Smaller size of fluorine. B. Stronger polar bond between H and F. C. Electronegativity of fluorine D. Strong hydrogen bonding
33	H ₂ S is a gas which H ₂ O is liquid at room temperature. it is due to	A. Less intermolecular forces in water B. Covalent bond in H-O in water molecule C. Hydrogen bonding in water molecules D. Ionic characters in water molecules
34	The inter-molecular forces in liquids are:	A. Negligible B. Very weak C. Very strong D. Reasonably strong
35	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
36	London dispersion forces are alos called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
37	Amorphous solids:	A. Have a sharp melting point B. Undergo clean cleavage when cut with knife C. Have a perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
38	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
39	Which of the following is a pseudo solid	A. CaF ₂ B. Glass C. NaCl D. All
40	The molecules of CO ₂ in dryice from the :	A. Ionic crystals B. Coverlet crystals C. Molecular crystals D. Any type of crystals
41		A. Dipole induced dipole forces B. Ionic diiole forces

41	Which one of the following is weakest inter molecular force?	C. Electrostatic forces b/w ions D. Dipole dipole forces
42	Which of the following liquids has low vapour pressure at 25 ⁰ C	A. Diethyl ether B. Acetone C. Water D. Ethyl alochol
43	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
44	Force of attraction b/w atoms of He is :	A. London dispersion forces B. Hydrogen bonding C. Coordinate covalent bond D. Covalent bond
45	Chloroform and acetone are soluble in each other due to	A. Instantaneous dipole interactions B. Dipole-dipole interactions C. Intermolecular hydrogen bonding D. All of above
46	A liquid on evaporation causes	A. Heating effect B. Cooling effect C. Suffication D. All of above
47	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electicity D. Is transparent to light
48	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
49	The molecules of CO ₂ in dry ice form the	A. Ionic crystals B. Covalent crystals C. Molecular crystals D. Any type of crystal
50	Hydrocarbon molecules with large chain lengths experience	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
51	The density of water decreases, When it is freezed at 0°C	A. Change of bond length B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of ice
52	The gases can be converted into liquids by:	A. Increasing the pressure only. B. Lowering temperature and increasing pressure C. Increasing pressure and bringing temperature below critical points D. Lowering temperature only
53	Force of attraction between atoms of He is	A. London dispersion forces B. Hydrogen bondign C. Coordinate covalent bond D. Covalent bond
54	Trend of boiling point of halogens from fluorine to Iodine is that it	A. Decreases B. Is negligible C. Increases D. Remains constant
55	Which of the following liquids has low vapour pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
56	Debye forces are present in one of the following pairs	A. Na ⁺ ion and water B. Argon and water C. Argon and Na ⁺ ion D. Ne and Water
57	Rate of evaporation and rate of condensation at equilibrium:	A. Become very low B. Become very high C. Become equal D. Can never be equal.

58	Amorphous solids	A. Have sharp melting point B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
59	Evaporation of water is possible at	A. Above 100 ⁰ B. 0 ⁰ C. 100 ⁰ D. At all temperature
60	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.