

ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Amorphous solids:	<p>A. Have a sharp melting point</p> <p>B. Undergo clean cleavage when cut with knife</p> <p>C. Have a perfect arrangement of atoms</p> <p>D. Can possess small regions of orderly arrangement of atoms</p>
2	Diamond is a bad conductor because	<p>A. It has a tight structure</p> <p>B. It has a high density</p> <p>C. there is no free electron present in the crystal of diamond of conduct electricity</p> <p>D. Is transparent to light</p>
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4	H ₂ S is a gas while H ₂ O is a liquid at room temperature. It is due to:	<p>A. Less inter-molecular forces in water.</p> <p>B. Covalent bonding H-O in water molecule.</p> <p>C. Hydrogen bonding in water molecules</p> <p>D. Ionic character in water molecules</p>
5	In a group on going downward, polarizability generally:	<p>A. Decreases</p> <p>B. Increases</p> <p>C. Remain constant</p> <p>D. Negligible</p>
6	The density of water decreases, When it is freeze at 0°C	<p>A. Change of bond length</p> <p>B. Change of bond angles</p> <p>C. Cubic structure of ice</p> <p>D. Empty spaces present in the structure of ice</p>
7	Debye forces are present in one of the following pairs	<p>A. Na⁺ ion and water</p> <p>B. Argon and water</p> <p>C. Argon and Na⁺ ion</p> <p>D. Ne and Water</p>
8	The intermolecular forces in liquids are	<p>A. Negligible</p> <p>B. Very weak</p> <p>C. Very strong</p> <p>D. Reasonably strong</p>
9	At sea level and at 100°C the vapour pressure of water in an open system is	<p>A. 1000 mm Hg</p> <p>B. 760 mm Hg</p> <p>C. 730 mm Hg</p> <p>D. 670 mm Hg</p>
10	Evaporation of water is possible at:	<p>A. Above 100°C</p> <p>B. 0°C</p>

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D. At all temperature

11	Force of attraction b/w atoms of He is :	A. London dispersion forces B. Hydrogen bonding C. Coordinate covalent bond D. Covalent bond
12	Diamond is a bad conductor because:	A. It <i>has</i> tight structure B. It has a high density C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light
13	Which of the following liquids has low vapor pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
14	Hydrogen bonding is present in one of the following pairs:	A. NH ₃ B. H ₂ O C. HF D. All of above
15	Which one of the following is the weakest intermolecular force	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatics forces between ions D. Dipole-dipole forces
16	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.
17	Ionic solids are characterized by :	A. Low meeting points B. Good conductivity in solid state C. High vapor pressure D. Solubility in polar solvents
18	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
19	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.
20	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds