ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	]Which of the following is a pseudo solid?	A. Atoms of He is gaseous stat at high temperate. B. Molecules of water in liquid state. C. Molecules of solid I₂ D. Molecular of hydrochloric acid gas
2	Which of the following is a pseudo solid?	A. CaF₂<o:p></o:p> B. Glass<o:p></o:p> C. NaCL<o:p></o:p> D. All<o:p></o:p>
3	Hydrogen bonding is present in one of the following pairs:	A. NH₃ B. H₂O C. HF D. All of above
4	Ionic Solids are characterized by	A. Low melting points B. Good conductivity in solid state C. High vapour pressure D. Solubility in polar solvents
5	HF has exceptionally low acidic strengths due to	A. Smaller size of fluorine B. Strong polar bond between H and F C. Electronegativity of fluorine D. Strong hydrogen bonding
6	Which of the following liquids has low vapour pressure at 25^0C	A. Diethyl ether B. Acetone C. Water D. Ethyl alochol
7	Evaporation of water is possible at:	A. Above 100<span style="font-size: 11pt; line-height: 107%; font-family: Arial, sans-serif; color: rgb(34, 34, 34); background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">°C B. 0<span style="font-size: 11pt; line-height: 107%; font-family: Arial, sans-serif; color: rgb(34, 34, 34); background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">°C C. 100<span style="font-size: 11pt; line-height: 107%; font-family: Arial, sans-serif; color: rgb(34, 34, 34); background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">°C D. At all temperature
8	Which of the following is a pseudo solid	A. CaF₂ B. Glass C. NaCl D. All
9	At sea level and at 100°C the vapor pressure of water in an open system is:	A. 1000 mm Hg B. 760 mm Hg C. 730 mm Hg D. 670 mm Hg

10	Hydrocarbon molecule with large chain lengths experience:	A. Weaker attractive forces B. Stronger attractive forces C. Repulsive forces D. No attractive forces
11	The molecules of CO_2 in dry ice from the	A. Ionic crystalls B. Covalent crystals C. Molecular crystals D. Any type of crystal
12	The strongest forces are	A. Debye forces B. London dispersion forces C. Dipole-dipole attraction D. Hydrogen
13	Escape of high energy molecules from the surface of liquid is called:	A. Sublimation. B. Distillation. C. Condensation. D. Evaporation.
14	Debye forces are present in on of the following pairs:	A. Na+ion and water B. Argon and water C. Argon and Na+ion D. Ne and water
15	The strongest forces are:	A. Debye froces B. London dispersion C. Dipole-dipole attraction D. Hydrogen bonding
16	Ionic solids are characterized by :	A. Low meeting points B. Good conductivity in solid state C. High vapor pressure D. Solubility in polar solvents
17	Which one of the following is weakest inter molecular force?	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatic forces b/w ions D. Dipole dipole forces
18	Rate of evaporation and rate of condensation at equilibrium:	A. Become very low B. Become very high C. Become equal D. Can never be equal.
19	Debye forces are present in one of the following pairs	A. Na⁺ion and water B. Argon and water C. Argon and Na⁺ion D. Ne and Water
20	London dispersion forces are also called:	A. Hydrogen bonding. B. Debye forces C. Van der Waal's forces D. Instantaneous dipole-induced dipole forces.