ECAT Chemistry Chapter 3 Gases

Sr	Questions	Answers Choice
1	Pressure remaining constant, at which temperature volume of gas will become twice of what it is at 0°C ?	A. 546<o:p></o:p> B. 200 C. 546K D. 273K
2	The deviation of a gas from ideal behavior is maximum at :	A. -10<span style="font-size:11.0pt;line-height:107%; font-family:"Calibri",sans-serif;mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:"Times New Roman";mso-bidi-theme-font:minor-bidi; mso-ansi-language:EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA">°C and 5.0 atm B. -10°C and 2.0 atm C. 100°C and 2.0 atm D. 0°C and 2.0 atm
3	The intramolecular forces in gases are	A. Weak B. Normal C. Very weak D. Strong
4	The relation ships b/w volume of given amount of gas and prevailing conditions of temperature and pressure are :	A. Charles's law B. Graham's law C. Boyle's law D. Gas law
5	The rate of diffusion of a gas in	A. Inversely proportional to its density B. Inversely proportional to square root of its molecular mass C. Directly proportional to molecular mass D. Directly proportional to its density
6	How should condition be changed to prevent the volume of a given gas from expanding whine its mass is increased ?	A. Temperature is lowered and pressure is increased B. Temperature is increasedand pressure is lowered C. Temperature and pressure both are lowered D. Temperature and pressure both are increased
7	Absolute temperature of a gas is proportional to	A. Rotational kinetic energy B. Translational kinetic energy C. Vibrational kinetic energy D. Potential energy
8	Equal masses of methane and oxygen are mixed in an empty container at 25°CThe fraction of total pressure exerted by oxygen is :	A. 1/2 B. 8/9 C. 1/9 D. 16/17
9	Gases shows uniform behavior toward their :	A. Internal conditions B. External conditions C. Internal and external conditions D. None of above
10	The product of pressure and volume remains constant when temperature and quantity of gas is	A. zero B. variable C. kept constant

		D. None of these
11	Which of the following will have the same number of molecules at STP ?	A. 280 CM³ of CO² and 280 CM³ of N₂O<o:p></o:p> B. 11.2 dm3 of O₂ and 32 g of O₂<o:p></o:p> C. 44 g of CO₂ and 11.2 dm³ of CO<o:p></o:p> D. 28 g of N2 and 5.6 dm³ of oxygen<o:p></o:p>
12	The movement of molecules from a region of high pressure to vacuum is called :	A. Evaporation B. Effusion C. Conduction D. Difusion
13	Boyle's law does not fall even	A. Temperature is extremely high B. Pressure is extremely high C. Mixture of gases is taken D. all of above
14	Gases of air always remain in random motion and do not settle due to :	A. Difference in molecular masses of air gases. B. Difference in partial pressure of gas molecules. C. Unequal number of different gas molecules. D. Elastic collision of gas molecules.
15	The rate of diffusion of a gas is :	A. Inversely proportional to its density B. Inversely proportional to square root of its molecular mass C. Directly proportional to molecular mass D. Directly proportional to its density
16	Keeping the temperature constant, if the gas is expanded	A. kinetic energy of molecules will increase B. Number of gas molecules increases C. Temperature will increases D. Pressure will decrease
17	For a gas obeying Boyle's law if pressure is doubled, the volume becomes :	A. Remain constant B. Double C. One half D. None of above
18	Cooling happens under the Joule Thomson Effect due to sudden :	A. Contraction B. Absortption C. Expansion D. All of above
19	Boyle's law doesn't fail even :	A. Temperature is extremely high B. Pressure is extremely high C. Mixture of gas is taken D. All of above
20	Gases of air, always remain in random motion and do not settle due to	A. Difference in molecular masses of air gases B. Difference in partial pressure of gas molecules C. Unequal number of different gas molecules D. Elastic collision of gas molecules