ECAT Chemistry Chapter 20 Aromatic Hydrocarbons Online Test

Sr	Questions	Answers Choice
1	Among the following, poly cyclic compound is:	A. styrene B. cumene C. napthalene D. xylene
2	Which of the following is explosive:	A. Trinitropheonal B. Nitropheonal C. Nitromethane D. Nitrobenzene
3	The electrophile in aromatic sulphonation Is	A. H₂SO₄ B. HSO₄⁻ C. SO₃ D. SO₃⁺
4	The hybirdization in benzene is:	A. sp³ B. sp² C. sp² D. dsp²
5	Which term was derived from atomos?	A. atom B. hydrocarbon C. aromatic D. aliphatic
6	Arenes are also called	A. atom B. hydrocarbons C. aromatic D. benzene
7	Benzene cannot undergo	A. substitution reactions B. addition reactions C. oxidation reactions D. elimination reactions
8	Aromatic hydrocarbons are derivatives of	A. normal series of paraffins B. alkene C. benzene D. cyclohexane
9	Characterstics of aromaic are:	A. how hydrogen carbon ratio than alkanes B. characterstics oudour C. a &B D. Characterstics properties
10	Simplest aromatic compound is	A. bezene B. toluene C. aniline D. phenol
11	Simple aromatic compound is:	A. benzene B. toluene C. aniline D. phenol
12	Among the following, poly cyclic compound is	A. styrene B. cumene C. naphthalene D. xylene
13	Aroma meas:	A. Fragrant B. Invisible C. latest affinityt D. benzene
14	Benzene was discovered by first of all	A. Michael Faraday B. Hofmann C. Ainderson D. Sorenbon
15	Monosubstituted benzene can have disubstitution at position:	A. Ortho B. meta C. para D. a, b, c

16	The conversion of n-hexane into benzene by heating in the presence of CO, is called	A. Isomerization B. Aromatization C. Dealkylation D. Rearrangement
17	The benzene molecule contains:	A. Three double bond B. Two double bond C. One double bond D. Delocalized<span style="font-size:12.0pt;line-height:107%; font-family:"Plantagenet Cherokee","serif";mso-fareast-font-family:Calibri; mso-fareast-theme-font:minor-latin;mso-bidi-font-family:"Times New Roman"; mso-ansi-language:EN-US;mso-fareast-language:EN-US;mso-bidi-language:AR-SA"> electron charge
18	Resonance energy of benzene is (in KJ mol^{-1}):	A. 120 B. 150 C. 170 D. 180
19	During nitration of benzene. the active nitrating agent is	A. NO₃⁻¹ B. NO₂⁺ C. NO₂⁻¹ D. HNO₃
20	During nitration of benzene, the active nitrating agent is:	A. NO₃⁻¹ B. NO₂₊ C. NO₂⁻¹ D. HNO₃
21	Aromatic hydrocarbons are the derivatives of:	A. Normal series of parrafins B. Alkene C. Benzene D. Cyclohexane
22	Bond angle in benzene is :	A. 109..5^o B. 180^o C. 120^o D. 107.2^o
23	Benzene gives reactions generally:	A. Electropholic subsitution B. addition C. synthesis D. addition and electropholic subsitution
24	Which of the following is explosive?	A. Trinitrophenol B. Nitrophenol C. Nitromethane D. Nitrobenzene
25	Which of the following acid can be used as a catalyst in Friedel Craft's reaction?	A. AlCl₃ B. HNO₃ C. BeCl₂ D. NaCl
26	Benzene acid can be prepared from the oxidation of:	A. benzene B. ethyl benzene C. benzonic acid D. toluene
27	Benzoic acid can be prepared from the oxidation of	A. benzene B. ethyl benzene C. benzoic acid D. toluene
28	Kekule structures contributed towards actual structure of benzene	A. 60% B. 70% C. 80% D. 90%
29	Aromatic compounds burn with sooty flame cause:	A. They have high percentage of hydrogan B. They have ring structure C. They have high percentage of carbon D. They resist reaction with air
30	Acyiation of benzene to produce aliphatic aromatic ketones is known as	A. Friedel Craft's reaction B. benzenecondensation C. hydroformylation D. Clemmensen reduction
31	Bond angle in benzene is	A. 109.5° B. 180° C. 120° D. 107.20
32	Monosubstituted benzene can have disubstitution at position	A. ortho B. meta C. para D. a, b, c

		D. a, b, c
33	Amongst the following, the compound that can be most readily sulphonated is:	A. Toluene B. Benzene C. nitrobenzene D. chlorobenzene
34	Benzene gives reactions generally:	A. Electropholic subsitution B. addition C. synthesis D. addition and electropholic subsitution
35	Which compound is the most reactvie one:	A. benzene B. ethene C. ethane D. ethyne
36	Benzene was discovered by first of all:	A. Micheal Faraday B. Hofmann C. Ainderson D. Sorenbon
37	Resonance energy of benzene is (in KJ mol^{-1})	A. 120 B. 150 C. 170 D. 180
		A. Friedel Craft's reaction B. Benzene condensation C. Hydroformylation D. Cellemense <!--[if gte msEquation 12]><m:oMathPara><m:oMath><m:rad><m:radPr><m:degHide m:val="on"/><span style='font-size:12.0pt;mso-ansi-font-size:12.0pt; mso-bidi-font-size:12.0pt;font-family:"Cambria Math","serif";mso-ascii-font-family: "Cambria Math";mso-hansi-font-family:"Cambria Math";mso-bidi-font-family: "Times New Roman";font-style:italic;mso-bidi-font-style:normal'><m:ctrlPr></m:ctrlPr></m:radPr><m:deg></m:deg><m:e></m:e></m:rad></m:oMath></m:oMathPara><![endif]--><!--[if !msEquation]--><span style="font-size:11.0pt;line-height:107%;font-family:"Calibri","sans-serif"; mso-ascii-theme-font:minor-latin;mso-fareast-font-family:Calibri;mso-fareast-theme-font: minor-latin;mso-hansi-theme-font:minor-latin;mso-bidi-font-family:"Times New Roman"; mso-bidi-theme-font:minor-bidi;mso-ansi-language:EN-US;mso-fareast-language: EN-US;mso-bidi-language:AR-SA"><v:shapetype id="_x0000_t75" coordsize="21600,21600" o:spt="75" o:preferrelative="t" path="m@4@5l@4@11@9@11@9@5xe" filled="f" stroked="f"> <v:stroke joinstyle="miter"> <v:formulas> <v:f eqn="if lineDrawn pixelLineWidth 0"> <v:f eqn="sum @0 1 0"> <v:f eqn="sum 0 0 @1"> <v:f eqn="prod @2 1 2"> <v:f eqn="prod @3 21600 pixelWidth"> <v:f eqn="prod @3 21600 pixelHeight"> <v:f eqn="sum @0 0 1"> <v:f eqn="prod @6 1 2"> <v:f eqn="prod @7 21600 pixelWidth"> <v:f eqn="sum @8 21600 0"> <v:f eqn="prod @7 21600 pixelHeight"> <v:f eqn="sum @10 21600 0"> </v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:formulas> <v:path o:extrusionok="f" gradientshapeok="t" o:connecttype="rect"> <o:lock v:ext="edit" aspectratio="t"> </o:lock></v:path></v:stroke></v:shapetype><v:shape id="_x0000_i1025" type="#_x0000_t75" style="width:20.25pt; height:19.5pt"> <v:imagedata src="file:///C:\Users\SHOPER.PK\AppData\Local\Temp\msohtmlclip1\01\clip_image001.png" o:title="" chromakey="white"> </v:imagedata></v:shape><!--[endif]--><o:p></o:p> <!--[if gte msEquation 12]><m:oMathPara><m:oMath><m:rad><m:radPr><m:degHide m:val="on"/><span style='font-size:12.0pt;mso-ansi-font-size:12.0pt; mso-bidi-font-size:12.0pt;font-family:"Cambria Math","serif";mso-ascii-font-family: "Cambria Math";mso-hansi-font-family:"Cambria Math";mso-bidi-font-family: "Plantagenet Cherokee";font-style:italic;mso-bidi-font-style:normal'><m:ctrlPr></m:ctrlPr></m:radPr><m:deg></m:deg><m:e><span style='font-size:12.0pt; line-height:107%;font-family:"Cambria Math","serif";mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-font-family:"Plantagenet Cherokee"; mso-bidi-font-family:"Plantagenet Cherokee";mso-ansi-language:EN-US; mso-fareast-language:EN-US;mso-bidi-language:AR-SA'><m:r>23</m:r></m:e></m:rad></m:oMath></m:oMathPara><![endif]--><!--[if !msEquation]--><span style="font-size:11.0pt;line-height:107%;font-family:"Calibri","sans-serif"; mso-ascii-theme-font:minor-latin;mso-fareast-font-family:Calibri;mso-fareast-theme-font: minor-latin;mso-hansi-theme-font:minor-latin;mso-bidi-font-family:"Times New Roman"; mso-bidi-theme-font:minor-bidi;mso-ansi-language:EN-US;mso-fareast-language: EN-US;mso-bidi-language:AR-SA"><v:shapetype id="_x0000_t75" coordsize="21600,21600" o:spt="75" o:preferrelative="t" path="m@4@5l@4@11@9@11@9@5xe" filled="f" stroked="f"> <v:stroke joinstyle="miter"> <v:formulas> <v:f eqn="if lineDrawn pixelLineWidth 0"> <v:f eqn="sum @0 1 0"> <v:f eqn="sum 0 0 @1"> <v:f eqn="prod @2 1 2"> <v:f eqn="prod @3 21600 pixelWidth"> <v:f eqn="prod @3 21600 pixelHeight"> <v:f eqn="sum @0 0 1"> <v:f eqn="prod @6 1 2"> <v:f eqn="prod @7 21600 pixelWidth"> <v:f eqn="sum @8 21600 0"> <v:f eqn="prod @7 21600 pixelHeight"> <v:f eqn="sum @10 21600 0"> </v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:f></v:formulas> <v:path o:extrusionok="f" gradientshapeok="t" o:connecttype="rect"> <o:lock v:ext="edit" aspectratio="t"> </o:lock></v:path></v:stroke></v:shapetype><v:shape id="_x0000_i1025" type="#_x0000_t75" style="width:21pt; height:17.25pt"> <v:imagedata src="file:///C:\Users\SHOPER.PK\AppData\Local\Temp\msohtmlclip1\01\clip_image001.png" o:title="" chromakey="white"> </v:imagedata></v:shape><!--[endif]--><!--[if gte msEquation 12]><m:oMathPara><m:oMath><m:rad><m:radPr><m:degHide m:val="on"/><span style='font-size:12.0pt;mso-ansi-font-size:12.0pt; mso-bidi-font-size:12.0pt;font-family:"Cambria Math","serif";mso-ascii-font-family: "Cambria Math";mso-hansi-font-family:"Cambria Math";mso-bidi-font-family: "Plantagenet Cherokee";font-style:italic;mso-bidi-font-style:normal'><m:ctrlPr></m:ctrlPr></m:radPr><m:deg></m:deg><m:e><span style='font-size:12.0pt; line-height:107%;font-

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39	The benzene molecule contains	A. Three double bond B. Two double bond C. One double bond D. Delocalized pie electron charge
40	Benzene gives reactions generally	A. electrophilic B. addition substitution C. synthesis D. addition and electrophilic substitution
41	empirical formula mass of benzene is time lesser than moleculer formula mass:	A. four B. five C. six D. seven
42	Amongst the following, the compound that can be most readily sulphonated is	A. toluene B. benzene C. nitrobenzene D. chlorobenzene
43	Which is the property of benzene?	A. decolourizes KMnO₄ B. straight chain structure C. only double bond is present D. triple and double bond
44	The hydridization in benzene is	A. sp³ B. sp² C. sp⁴ D. dsp²
45	Benzene cannot undergo:	A. Substitutions reactions B. Addition reactions C. Oxidation reactions D. Elimination reactions
46	Aromatic compounds burn with sooty flame cause	A. They have high percentage of hydrogen B. They have a ring structure C. They have high percentage of carbon D. They resist reaction with air
47	empirical formula mass of benzene is time lesser than moleculer formula mass:	A. four B. five C. six

		D. seven
48	Which terms was derived from " aroma"?	A. atom B. hydrocarbons C. aromatic D. aliphatic
49	Ozonolysis of benzene gives	A. nitration B. sulphonation C. ozonide D. glyoxal
50	Ozonolysis of benzene gives:	A. Nitration B. sulphonation C. ozonide D. glyoxal
51	Which compound is the most reactive one?	A. benzene B. ethene C. ethane D. ethyne
52	Empirical formula mass of benzene is times lesser than molecular formula mass	A. four B. five C. six D. seven
53	Ratio of carbon to hydrogen in aromatic compounds is	A. Low than alkanes B. High than alkanes C. Low than alkenes not high than alkanes D. High than high than alkenes
54	Kekule structures contributed towards actual structure of benzene	A. 60% B. 70% C. 80% D. 90%
55	Which term was derived from "aroma":	A. Atom B. Hydrocarbons C. aromatic D. aliphatic
56	Which of the following acid can be used as a catalyst in Friedel craft's reaction	A. AlCl³ B. HNO₃ C. BeCl₂ D. NaCl
57	Ratio of carbon to hydrogen in aromatic compounds is:	A. Low than alkanes B. High than alkanes C. Low than alkanes not high than alkanes D. High than alkanes
58	Which is the property of benzene:	A. Decolourizes KMnO₄ B. straight chain structure C. only double bond is present D. triple and double bond
59	Benzene reacts with Cl, in sunlight to give the end product	A. C₆H₆Cl₆ B. C₆H₆Cl C. O - C₆H₄Cl₂ D. P-C₆H₄Cl₂
60	The electrophile in aromatic sulphontion is:	A. H₂SO₄ B. HSO₄ C. SO₃ D. SO₃₊
61	The conversion of n-hexane into benzene by heating in the presence of Cr_2O_3 is called:	A. Isomerization B. Aromatization C. Dealkylation D. Rearrangment