ECAT Chemistry Chapter 11 Reaction Kinetics

Sr	Questions	Answers Choice
1	Decreases om concentration of reactant is denoted by	A. dc/dt B. -dc/dt C. +dc/dt D. None
2	With the progressive of the reaction the slope of the curve between concentration of product and time	A. Gradually becomes more steep B. Gradually becomes less steep C. No change occurs in slope D. None of these occurs
3	Which of the following may affect the rate constant (k) fro a reaction :	A. Change in concentration. B. Change in pressure. C. Change in pH. D. Change in temperature.
4	If half life period of a reaction is independent of the concentration of the reactants, then the reaction is	A. Zero order B. First order C. Second order D. Order is in fraction
5	Which of the following is not affected by light	
6	Factor which slows down the rate of reaction is	A. Small size of the particles of the reactant B. High temperature of reaction C. More concentration of reactant D. Lowering the temperature
7	In the rate equation when the concentration of reactants are unity, then rate is equal to	A. Instantaneous rate B. Average rate C. Active mass of products D. Specific rate constant
8	It is common observation that rates of chemical reactions differ :	A. Greatly. B. A little bit. C. Moderately.
9	The unit of rate constant k is the same as that of the rate of reaction in	A. First order reaction B. Second order reaction C. Third order reaction D. Zero order reaction
10	The rate of reaction :	A. Decreases as the reaction proceeds B. Increases as the reaction proceeds C. May decrease or increase reaction proceeds D. Remains same as the reaction proceeds
11	For a chemical reaction to take place the particles must have sufficient energy for the effective collisions, the energy is called	A. Average energy B. Activation energy C. Potential energy D. Collision energy
12	If the rate of reaction is independent of the concentration of the reactant, the reaction is of	A. Zero order B. First order C. Second order D. Third order
13	which one of the following is a heteroheneous catalysis	
14	With increases of 10°C temperature the rate of reactiondoubles. This increase in rate of reactionis due to :	A. Decrease in activation energy or reaction. B. Decrease in number of collisions between reactant molecules. C. Increase in activation energy of reactants. D. Increase in number of effective collisions
		A. Increase as the reaction proceeds.

| 15 | The rate of reaction : | C. Remains the same as the reaction proceed.
D. May decrease or increase as the reaction proceeds. |

| 16 | Question Image | A. Three times
B. Six times
C. Nine times
D. Two times |

| 17 | A white precipitate of silver chloride immediately formed on addition of : | A. Silver nitrate solution to sodium chloride solution.
B. Silver chloride solution to sodium nitrate solution.
C. Silver nitrate solution to potassium chloride solution
D. Silver nitrate solution to hydrogen chloride solution. |

| 18 | Hydrolysis of ethyl-acetate (ester) has order of reaction : | A. 3
B. 2
C. 1
D. 1 |

| 19 | If the rate equation of a reaction 2A+B-------->Products is , Rate = K[A] [B], and A is present in large excess, then order of reaction is : | A. 1
B. 2
C. 3
D. Above |

| 20 | When initial concentration of reactants an order of reaction is given, then its half life period can be calculated by the equation | |