ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	An electrolyte	A. Forms complex ions solution B. Gives ions only when electricity is passed C. Possesses ions even in solid state D. Gives ions only when dissolved in water
2	Cell potential depends upon :	A. Concentration of ions B. Nature of electrolyte C. Temperature D. All of above
3	For spontaneity of a cell, which is correct?	A. ΔG = 0,ΔE = 0 B. ΔG = -ve ,ΔE = 0 C. ΔG = +ve ,ΔE = 0 D. ΔG = -ve
4	The conductivity of strong electrolyte	A. Increases on dilution slightly B. Does not change on dilution C. Decreases on dilution D. Depends on density of electrolyte it self
5	2.5 faradays of electricity is passed through solution of $CuSO_4$. The number of gram equivalents of copper deposited on the cathode would be	A. 1 B. 2 C. 2.5 D. 1.25
6	The oxidation number of Mn is $KMnO_4$is	A. +2 B. +4 C. +6 D. +7
7	A standard hydrogen electrode (S.H.E) consists of a platinized platinum electrode dipped in 1 molar solution of H^+ions and hydrogen gas is passed at a pressure of	A. One pascal B. One kilo pascal C. One atmoshpere D. Then atmoshpere
8	In a solution of $CuSO_4$how much time will be required to precipitate 2g copper by 0.5 ampere current?	A. 12157.48 sec B. 102 sec C. 510 sec D. 642 sec
9	In the reaction $K_2Cr_2O_7$+ HCl + $CrCl_3$+ Cl_2+ H_2O the element which is reduced is	A. K B. Cl C. Cr D. H
10	Which of the following statements is true about Galvanic cell	A. Anode is negatively charged B. Reduction occurs at anode C. Cathode is positively charge D. Reduction occurs at cathode
11	Reduction or oxidation potential of standard hydrogen electrode is :	A. 0.0 Volt B. 0.8Volt C. 1.0Volt D. 1.8Volt
12	In the presence of copper electrodes the electrolysis of aqueous $CuSO_4$produces at cathode	A. H₂gas B. O₂gas C. SO₂gas D. Cu metal
		A. Greater positive value of standard

13	Strong reducing agents gave	reduction potential B. Greater negative value of standard reduction potential C. Lesser positive value of standard reaction potential D. None of these
14	During electrolysis of KNO_3, $H_2$is evolved	A. Anode B. Cathode C. Both a and b D. None
15	Question Image	A. Iodine can oxidise iron B. Bromine can oxidise iron C. Iodine can oxidise bromine D. Iron can oxidise bromine
16	96500 C of electricity liberates from $CuSO_4$solution	A. 63.5 g of Cu B. 31.75 g of Cu C. 96500 g of Cu D. 100 g of Cu
17	The substance having highest conductivity at room temperature among the following is	A. 0.1 N HCl B. 0.1 N NaCl C. Graphite D. Glass
18	An electric current is passed through silver voltameter connected to a water voltmeter. The cathode of the silver voltameter is 0.108 g more at the end of the electrolysis. The volume of oxygen evolved at STP is	A. 56 cm³ B. 550 cm³ C. 5.6 cm³ D. 11.2 cm³
19	The oxidation number of free element is always taken to be	A. 0 B. 1 C. 2 D. -1
20	In lead accumulator the electrolyte is H_2SO_4solution is	A. 30% H₂SO₄ B. 60% H₂SO₄ C. 80% H₂SO₄ D. 90% H₂SO₄