ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	Question Image	A. -1.10 V B. +1.10 V C. -0.42 V D. +0.42 V
2	Question Image	A. Fe is reduced B. Fe is oxidized C. Cl₂is oxidized D. None of these
3	Which of the following statements is true about Galvanic cell	A. Anode is negatively charged B. Reduction occurs at anode C. Cathode is positively charge D. Reduction occurs at cathode
4	In electronic watches or electronic calculators the tiny batteries used are	A. Alkaline battery B. NICAD battery C. Fuel cell D. Silver oxide battery
5	A smuggler could not carry gold by chemically depositing iron on the gold surface since	A. Gold is denser B. Iron rusts C. Gold has higher reduction potential then iron D. Gold has lower reduction potential then iron
6	Cell potential depends upon :	A. Concentration of ions B. Nature of electrolyte C. Temperature D. All of above
7	The equivalent conductivity of 0.1 M week acid is 100 times less than at infinite dilution. The degree of dissociation is	A. 100 B. 10 C. 0.01 D. 0.001
8	2.5 faradays of electricity is passed through solution of $CuSO_4$. The number of gram equivalents of copper deposited on the cathode would be	A. 1 B. 2 C. 2.5 D. 1.25
9	Standard reduction electrode potential of three metals A, B and C are erespectively + 0.05 V, -3.0 and -1.2V. The reducting power of	A. B > C > A B. A > B > C C. C > B > A D. A > C > B
10	While balancing an equation by ion electron method, the number of oxygen atoms are balanced by	A. OH⁻ B. H₂O C. O₂ D. H⁺
11	During electrolysis, the reaction that takes place at cathode is :	A. Reduction B. Both (a) and (c) C. Oxidation D. No reaction occurs
		A. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">H⁺<o:p></o:p> B. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial;

12	When aqueous NaCl is electrolyzed, which of the following ions get discharged at anode :	background-clip: initial;">Na⁺<o:p></o:p> C. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">OH⁻<o:p></o:p> D. <span style="font-size: 10.5pt; line-height: 107%; font-family: Arial, sans-serif; background-image: initial; background-position: initial; background-size: initial; background-repeat: initial; background-attachment: initial; background-origin: initial; background-clip: initial;">Cl⁻<o:p></o:p>
13	The unit of specific conductivity is	A. Ohm cm⁻¹ B. Ohm cm⁻² C. Ohm⁻¹cm D. Ohm⁻¹cm⁻¹
14	The electrolysis of $CuSO_4$aqueous solution using copper as cathode as well an anode the substance which deposits at cathode is	A. H₂SO₄ B. Oxygen C. Copper D. Hydrogen
15	When electrically is passed through molten Al_2O_3+ $Na_3AlF_6$and 13.5 gms of Al are deposited, the number of farady must be	A. 0.5 B. 1.0 C. 1.5 D. 2.0
16	In an electrolytic cell, the electrons flow from :	A. Cathode to anode or opposite B. Cathode to anode C. Anode to cathode D. Random flow
17	The substance having highest conductivity at room temperature among the following is	A. 0.1 N HCl B. 0.1 N NaCl C. Graphite D. Glass
18	The reaction is galvanic cell is	A. Spontaneous B. Non-spontaneous C. Acid-base D. None of these
19	Which of the following will be good conductor of electricity	A. Pure distilled water B. Molten NaCl C. Dilute solution of glucose D. Chloroform
20	The voltaic or galvanic cells which cannot be recharged are called	A. Primary cells B. Secondary cells C. Infinite cells D. Fuel cells