ECAT Chemistry Chapter 10 Electrochemistry Online Test

Sr	Questions	Answers Choice
1	Question Image	A. A strong reducing agent B. A strong oxidising agent C. Better oxidising agent than hydrogen D. Less reducing agent than hydrogen
2	Purification of an impure copper is made by electrolytic cell, in which impure copper is anode and pure copper is cathode, and the electrolyte used is	A. H₂SO₄ B. CuSO₄ C. ZnSO₄ D. Na₂SO₄
3	What will be the weight of deposited silver on passing 965 coulombs of electricity in solution of $AgNO_3$?	A. 1.08 g B. 2.16 g C. 0.54 g D. 0.27 g
4	In a compound an atom has negative oxidation state because	A. Atom is negatively charged B. Atom acts as cathode C. Atom is more electronegative D. Atom has lowest ionization energy
5	Alkaline battery has a voltage of	A. 4.5 V B. 3.5 V C. 2.5 V D. 1.5 V
6	An electrolyte	A. Forms complex ions solution B. Gives ions only when electricity is passed C. Possesses ions even in solid state D. Gives ions only when dissolved in water
7	During electrolysis of KNO_3, $H_2$is evolved	A. Anode B. Cathode C. Both a and b D. None
8	Which of the following will be good conductor of electricity	A. Pure distilled water B. Molten NaCl C. Dilute solution of glucose D. Chloroform
9	In $KO_2$the oxidation state of oxygen is	A. -2 B. -1 C. +1/2 D. -1/2
10	Which statement is correct for the fuel cells	A. These cells operate at low temperature B. These cells operate at low temperature C. No catalyst used for the formation of water D. MnO₂is used as electrolyte
11	The equivalent conductivity of 0.1 M week acid is 100 times less than at infinite dilution. The degree of dissociation is	A. 100 B. 10 C. 0.01 D. 0.001
12	In the presence of copper electrodes the electrolysis of aqueous $CuSO_4$produces at cathode	A. H₂gas B. O₂gas C. SO₂gas D. Cu metal
13	In lead accumulator the electrolyte is H_2SO_4solution is	A. 30% H₂SO₄ B. 60% H₂SO₄ C. 80% H₂SO₄ D. 90% H₂SO₄

14	Most metals are conductors of electricity because of the :	A. Light weight. B. Immobility of the electrons. C. Lustrous surfaces D. Relatively free movement of their electrons
15	In a Galvanic cell	A. Chemical energy is converted into electricity B. Chemical energy is converted into heat C. Electrical energy is converted into heat D. Electrical energy is converted into chemical energy
16	2.5 faradays of electricity is passed through solution of $CuSO_4$. The number of gram equivalents of copper deposited on the cathode would be	A. 1 B. 2 C. 2.5 D. 1.25
17	Prevention of corrosion of iron by Zn coating is called	A. Galvanization B. Cathodic protection C. Electrolysis D. Photoelectrolysis
18	Pure water does not conduct electricity because it	A. Has low boiling point B. Is almost unionized C. Is neutral D. Is readily decomposed
19	The electrode through which the electrons enter the electrolytic solution is electrolytic solution is	A. Anode B. Cathode C. May be anode or cathode D. None of these
20	The number of coulombs required for the deposition of 107.870 g of silver is	A. 96500 B. 48250 C. 193000 D. 10000