ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	Caustic soda is obtained by electrolysis of conc. aqueous solution of NaCl in a cell called	A. Daniell's cell B. Nelson's cell C. Down's cell D. Voltaic cell
2	The oxidation number of free element is always taken to be	A. 0 B. 1 C. 2 D. -1
3	The electrode through which the electrons enter the electrolytic solution is electrolytic solution is	A. Anode B. Cathode C. May be anode or cathode D. None of these
4	Out of Cu, Ag, Fe and Zn the metal which can displace all others from theri salt solution is	A. Ag B. Cu C. Zn D. Fe
5	Strong oxidizing agents have	A. Greater positive value of standard reduction potential B. Lesser positive value of standard C. Greater negative value of standard D. None of these
6	When electrically is passed through molten Al_2O_3+ $Na_3AlF_6$and 13.5 gms of Al are deposited, the number of farady must be	A. 0.5 B. 1.0 C. 1.5 D. 2.0
7	In a solution of $CuSO_4$how much time will be required to precipitate 2g copper by 0.5 ampere current?	A. 12157.48 sec B. 102 sec C. 510 sec D. 642 sec
8	The function of salt bridge in the galvanic or voltaic cell is to	A. Carry out oxidation at anode B. To carry out reduction at cathode C. Carry out electrolysis D. To prevent the net charge accumulation in either of the half cells
9	The reference electrode is made by using	A. ZnCl₂ B. CuSO₄ C. HgCl₂ D. Hg₂Cl₂
10	Question Image	A. Fe is reduced B. Fe is oxidized C. Cl₂is oxidized D. None of these
11	The cell in which a non-spontaneous redox reaction takes place as a result electricity is known as :	A. Electrolytic cell. B. Voltaic cell. C. Daniel cell. D. Dry cell.
12	In a Galvanic cell, the electrons flow from	A. A node to cathode through the solution B. Cathode to anode through the external circuit C. Cathode to anode through the external circuit D. Anode to cathode through the external circuit
13	Metallic conduction is due to the	A. Movement of electrons B. Movement of ions C. Both a and b D. None of these
		A. The reaction is splitted into two half reactions B. H₂O and H<sup>+

14	Which statement is incorrect for balancing of redox reactions by ion-electron method	</sup>ions are added for acidic or neutral reaction to balance O and H atoms C. To balance H, HCi, is added D. To balance O and H in the alkaline reaction OH⁻added
15	Question Image	A. Cu B. H C. N D. O
16	When aqueous solution of NaOH is electrolysed useing graphite electrodes, the product obtained at anode is	A. O₂gas B. H₂gas C. Na metal D. Na₂O
17	In the reaction $K_2Cr_2O_7$+ HCl + $CrCl_3$+ Cl_2+ H_2O the element which is reduced is	A. K B. Cl C. Cr D. H
18	The standard e.m.f. of a galvanic cell involving cell reaction with n = 2 is found to be 0.2965 V at 25°C. The equilibrium constant of the reaction would be	A. 1.0 x 10¹⁰ B. 2.0 x 10¹¹ C. 4.0 x 10¹² D. 1.0 x 10²
19	Purification of an impure copper is made by electrolytic cell, in which impure copper is anode and pure copper is cathode, and the electrolyte used is	A. H₂SO₄ B. CuSO₄ C. ZnSO₄ D. Na₂SO₄
20	The equivalent conductivity of 0.1 M week acid is 100 times less than at infinite dilution. The degree of dissociation is	A. 100 B. 10 C. 0.01 D. 0.001