

## ECAT Chemistry Chapter 10 Electrochemistry

Sr	Questions	Answers Choice
1	The unit of specific conductivity is	A. Ohm cm <sup>-1</sup> B. Ohm cm <sup>-2</sup> C. Ohm <sup>-1</sup> cm D. Ohm <sup>-1</sup> cm <sup>-1</sup>
2	Which statement is incorrect for NICAD battery	A. The electrolyte is alkali B. Cd acts as anode C. MnO <sub>2</sub> acts as electrolyte D. NiO <sub>2</sub> acts as cathode
3	Question Image	A. Cu B. H C. N D. O
4	Coinage metals like Au, Pt, Ag and Cu are the least reactive metals and don't liberate H <sub>2</sub> gas when treated with acids because	A. These have very high positive values of reduction potentials B. These have very high negative values of reduction potentials C. Their ionization potentials are lowest D. Their reduction potentials are close to SHE
5	A cell constant is generally found by measuring the conductivity of aqueous solution of	A. BaCl <sub>2</sub> B. KCl C. NaCl D. MgCl <sub>2</sub>
6	Sodium can be obtained by :	A. Electrolysis of acidified water. B. By heating NaCl and water at 100°C C. Electrolysis of molten sodium chloride. D. Electrolysis of aqueous sodium chloride.
7	In a Galvanic cell, the electrons flow from	A. Anode to cathode through the solution B. Cathode to anode through the external circuit C. Cathode to anode through the external circuit D. Anode to cathode through the external circuit
8	Corrosion reaction are	A. Spontaneous redox reaction B. Non-spontaneous acid-base reactions C. Spontaneous acid-base reactions D. None of these
9	In a compound an atom has negative oxidation state because	A. Atom is negatively charged B. Atom acts as cathode C. Atom is more electronegative D. Atom has lowest ionization energy
10	Best way to prevent rusting of iron is by	A. Making iron cathode B. Putting it in saline water C. Both of these D. None of these
11	Question Image	A. -1.10 V B. +1.10 V C. -0.42 V D. +0.42 V
12	Silver oxide battery has a voltage of	A. 2.0 V B. 1.5 V C. 2.5 V

C. 2.5 V  
D. 1.0 V

13	The conductivity of strong electrolyte	A. Increases on dilution slightly B. Does not change on dilution C. Decreases on dilution D. Depends on density of electrolyte it self
14	When fused $\text{PbBr}_2$ is electrolyzed	A. Bromine appears at the cathode B. Lead is deposited at the cathode C. Lead appears at the anode D. None of these happens
15	Metallic conduction is due to the	A. Movement of electrons B. Movement of ions C. Both a and b D. None of these
16	An electrochemical cell is based upon	A. Acid-base reaction B. Redox reaction C. Nuclear reaction D. None of the above
17	A cell in which electric current is produced as a result of spontaneous redox reaction is called :	A. Dry cell B. Electrolytic cell C. Galvanic cell D. Standard cell
18	In electrolytic cells, the chemical changes may be :	A. Either spontaneous or non-spontaneous B. Always spontaneous C. Always non-spontaneous D. More spontaneous and less non-spontaneous.
19	In a solution of $\text{CuSO}_4$ how much time will be required to precipitate 2g copper by 0.5 ampere current?	A. 12157.48 sec B. 102 sec C. 510 sec D. 642 sec
20	The reference electrode is made by using	A. $\text{ZnCl}_2$ B. $\text{CuSO}_4$ C. $\text{HgCl}_2$ D. $\text{Hg}_2\text{Cl}_2$