

Fsc Part 1 Chemistry MCQ's Test For Full Book

Sr	Questions	Answers Choice
1	Acid rain is mainly due to	A. SO ₃ and NO ₃ B. CO ₂ C. H ₂ D. CH ₄
2	Which law states that enthalpy change is independent of the path taken.	A. Hess's law B. Boyle's law C. Avogadro's law D. Dalton's law
3	The rate determining step in a multistep reaction is	A. The slowest step B. Always the first step C. Always the last step D. The fastest step
4	Major human sources of CO ₂ increase in atmosphere is	A. Fossil fuel combustion and deforestation B. Photosynthesis C. Ocean uptake D. Nitrogen fixation
5	1st ionization energy is.	A. Exothermic B. Endothermic C. Both D. Depend upon atom
6	The anode is electrolysis of molten NaCl is	A. Na ⁺ B. Cl ⁻ C. Na D. H ⁻
7	Liquid crystals are used in thermometers because.	A. They are highly reflective B. They glow at high temperatures C. Their color changes with temperature D. They conduct electricity
8	The chromyl chloride test is a specific confirmatory test for	A. Chloride ions B. Bromide ion C. Iodide ions D. Sulfate ions
9	NO ₂ in the atmosphere contributes to	A. Acid rain formation and photochemical smog B. Ozone layer repair C. Decreasing greenhouse effect only D. Building of clouds only
10	In contact process arsenic purifier uses.....for absorption of arsenic oxide.	A. Gelatinous ferric hydroxide B. Vanadium pentoxide C. Iron pyrite D. None of above
11	Classical smog is	A. Oxidation B. Reducing C. Both oxidizing and reducing D. None of these
12	Which of the following is a characteristic of crystalline solids.	A. They do not have a definite geometrical shape B. They melt over a wide temperature range C. They have a sharp melting point D. Their properties do not depend on the direction of measurement
13	If the electronegativity difference between bonded atoms is more than 0.5 but less than 1.7, the bond will be.	A. Ionic B. Polar Covalent C. Non polar covalent D. Coordinate covalent

14	The number of atoms present in 6 g of Mg metal is.	<p>A. 12.4×10^{23}</p> <p>B. 6.02×10^{23}</p> <p>C. 1.5×10^{23}</p> <p>D. 3.01×10^{23}</p>
15	The number of moles of CH ₄ in 4 g of gas is	<p>A. 0.2 moles</p> <p>B. 0.25 mole</p> <p>C. 0.3 mol</p> <p>D. 0.4 mole</p>
16	The more positive the standard reduction potential	<p>A. The stronger the reducing agent</p> <p>B. The weaker the oxidizing agent</p> <p>C. The stronger the oxidizing agent</p> <p>D. No effect</p>
17	The chemical species responsible for ozone depletion by chlorofluorocarbons is.	<p>A. Chlorine free radical</p> <p>B. CFC molecule itself</p> <p>C. UV rays</p> <p>D. Organic compounds</p>
18	A reversible reaction is one which	<p>A. Proceeds to completion</p> <p>B. Occurs only in one direction</p> <p>C. Proceeds in both directions</p> <p>D. Has no products</p>
19	27 g of Al will react completely with how much mass of O ₂ to produce Al ₂ O ₃ .	<p>A. 8 g</p> <p>B. 24 g</p> <p>C. 16 g</p> <p>D. 32 g</p>
20	Standard electrode potential is measured under	<p>A. 100 °C and 1 atm</p> <p>B. 0 °C and 1 M concentration</p> <p>C. 25 °C and 1 M concentration</p> <p>D. 100 °C and 1 M concentration</p>