

Chemical Energetics

C-	Quantiana	Anguaga Chair-
Sr	Questions	Answers Choice
1	The sun of all the energies of atoms, molecule, ion, within system is called	A. Enthalpy B. K.E. of the system C. Internal energy D. None
2	The amount of heat evolved or absorbed in a chemical reaction, when the molar quantities of products and reactants being the same as represented in chemical equation. is called	A. Heat of reaction B. Free energy of reaction C. Entropy of reaction D. None of these
3	In the chemical combination of hydrogen and fluorine to form HF:	A. Sodium atom donates major share of its electrons. B. Hydrogen atom donates the major share of its electrons. C. Both the atoms share the electrons equally. D. None of above.
4	Heat absorbed by a system when its volume does not change is equal to	A. Internal energy of system B. Enthalpy of system C. Increase in internal energy of system D. Increase in enthalpy of system
5	A system undergoes a change to attain the state of	A. High energy B. Low energy C. Moderate energy D. None of these
6	Elements combine together due to inherent tendency to stabilize themselves by:	A. Losing electron.B. Sharing electrons.C. Gain in electrons.D. All of above.
7	Which is not state function	A. Temperature B. Enthaply C. Entropy D. Internal energy
8	Molecular orbitals are filled with the available electrons according to:	A. Hund's of rule.B. Pauli's exclusion principle.C. Aufbau principle.D. All of above.
9	Which statement is true for Na and Na ⁺	A. Both have equal sizes. B. Both have same properties. C. Size of Na is smaller than Na ⁺ D. Size of Na is greater than Na ⁺
10	Internal energy of a system is equal to	A. Kinetic energy of particles B. Potential energy due to binding forces between particles C. Sum of kinetic energy plus potential energy of the particles D. Heat contents
11	Question Image	A. Positive sign B. Negative sign C. Without any sign D. None
12	The change in heat energy at constant temperature is called	A. Enthalpy change B. Heat of vaporisation C. Bond energy D. Internal energy change
13	When the degree of freedom increase the entropy	A. Decreases B. Increases C. Remains same D. All
14	Question Image	A76 KJ B57 KJ
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15	If heat absorbed in the reaction, the process is said to be	A. Exothermic B. Isothermal C. Adiabatic D. Endothermic
16	Heat,. work and internal energy of the system and surroundings are related into an equation which is called	A. First law of thermodynamics B. Hess's law C. Henry's law D. Born-haber cycle
17	Which of the following molecules has zero dipole moment?	A. NH ₃ B. CHCl ₃ C. H ₂ O D. BF ₃
18	A state function of the system which describes together the internal energy and the work done is called	A. Enthaply B. Internal energy C. Work D. Free energy
19	One calorie is equal to	A. 5.184 J B. 3.184 J C. 4.184 J D. 7.184 J
20	The standard enthalpy change in the formation of 1 mole of a compound from its elements in their standard physical states is	A. Enthalpy of formation B. Enthalpy of atomization C. Enthalpy of neutralization D. Internal energy change